

A green and efficient strategy to utilize spent SCR catalyst carrier: in-situ remediation of Cu@TiO₂ for photocatalytic hydrogen evolution

*Zhuo Wang^a, Ling Ma^a, Bingzhang Chen^a, Yubo Zhang^a, Kai Hong Wong^b, Wei Zhao^c, Chunxia Wang^{a, *}, Guoyong Huang^{a, *}, Shengming Xu^d*

^a State Key Laboratory of Heavy Oil Processing, College of New Energy and Materials, China University of Petroleum (Beijing), Beijing 102249, China.

^b State Key Laboratory of Urban Water Resource and Environment, Harbin Institute of Technology, Harbin 150090, P. R. China.

^c Green Living and Innovation Division, Hong Kong Productivity Council, HKPC Building, 78, Tat Chee Avenue, Kowloon 999077, Hong Kong.

^d Beijing Key Lab of Fine Ceramics, Institute of Nuclear and New Energy Technology, Tsinghua University, Beijing, 100084, China.

*Corresponding authors: E-mail: cxwang@iccas.ac.cn; huanggy@cup.edu.cn.

Table S1 Content of each element in the original SCR catalyst before recycling treatment

ICP-OES								
element	Ti	Al	Si	W	V	Fe	Ca	-
content	42.49%	2.15%	5.78%	2.21%	0.10%	0.08%	0.77%	-
Corresponding oxides	TiO ₂	Al ₂ O ₃	SiO ₂	WO ₃	V ₂ O ₅	Fe ₂ O ₃	CaO	Other
content	70.92%	8.12%	12.40%	2.79%	0.36%	0.23%	1.08%	4.10%

Table S2 Elemental content of spent SCR catalysts after recycling treatment

	ICP-OES							
element	Ti	Al	Si	W	V	Fe	Ca	-
content	57.55%	0.96%	0.14%	0.25%	0.05%	0.06%	0.06%	-
Corresponding oxides	TiO ₂	Al ₂ O ₃	SiO ₂	WO ₃	V ₂ O ₅	Fe ₂ O ₃	CaO	Other
content	96.07%	3.61%	0.30%	0.31%	0.19%	0.18%	0.09%	-

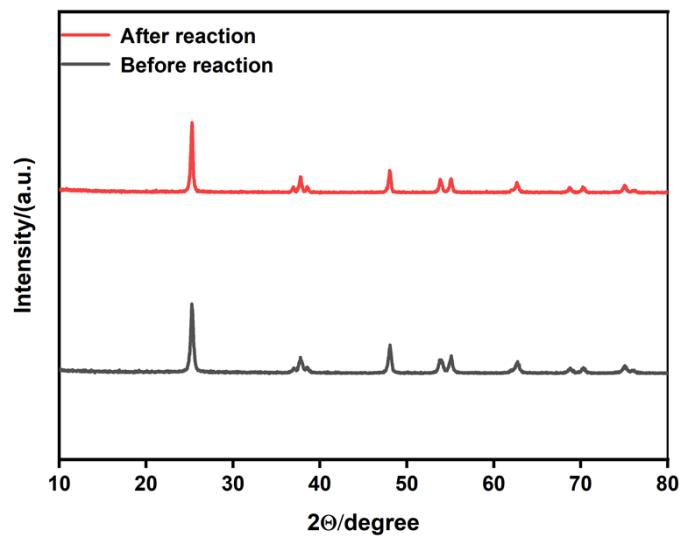


Fig. S1. XRD comparison of samples before and after photocatalytic hydrogen production reaction

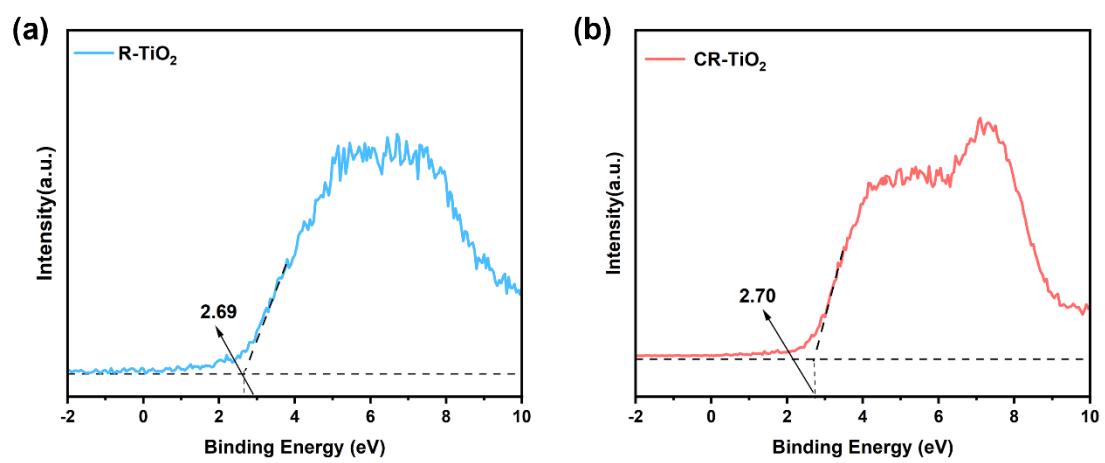


Fig. S2. XPS valence-band spectra of TiO₂, (b) and Cu–TiO₂

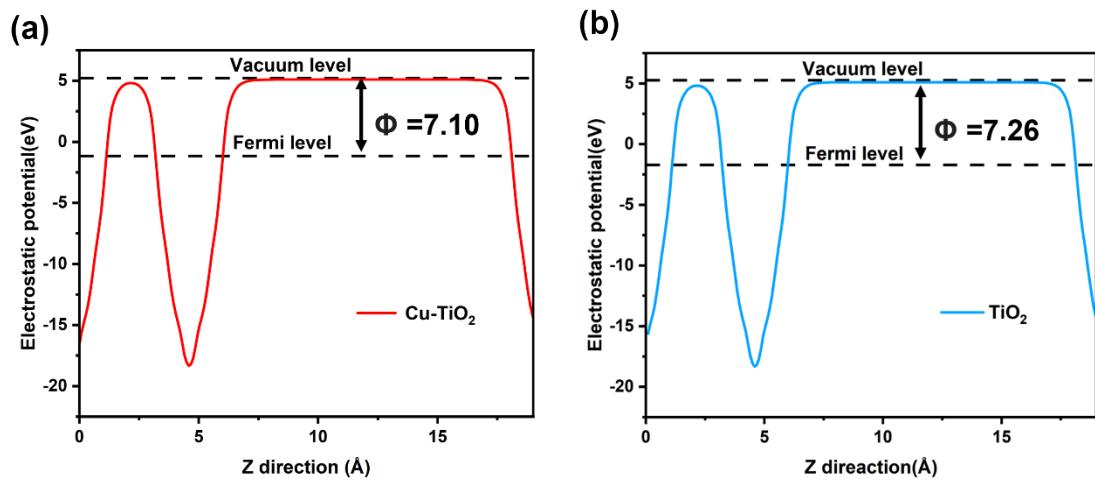


Fig. S3. Workfunction of Cu-TiO₂ and TiO₂

Table S3 Content of different types of oxygen in XPS analysis.

catalyst	O1	O2	O3
Cu-TiO₂	79.98%	14.26%	5.76 %
TiO₂	81.40%	8.29 %	10.31%

Table S4. A comparison of g-C₃N₄-based photocatalysts for hydrogen production.

Photocatalysts	Light sources	Sacrificial agents	HER rates (μmol h ⁻¹ g ⁻¹)	Refs.
Cu-TiO ₂ (Recovery)	300 W Xe lamp	20% (vol.) TEOA solution	368	This work
Ni-7.5@TCN	300 W Xe lamp $\lambda > 365$ nm	10% (vol.) TEOA	134	[1]
Pt@N-g-C ₃ N ₄	300 W Xe lamp $\lambda > 420$ nm	10% (vol.) TEOA	64	[2]
Pt/AgNbO ₃ /g-C ₃ N ₄	300 W Xe lamp $\lambda > 420$ nm	20 % (vol.) CH ₃ OH	88	[3]
TiO ₂ /g-C ₃ N ₄	420 W Xe lamp $\lambda > 300$ nm	10 % (vol.) CH ₃ OH	36.4	[4]
NiO/g-C ₃ N ₄	420 W Xe lamp $\lambda > 365$ nm	10% (vol.) TEOA	68.8	[5]
Pt/t-ZrO ₂ /g-C ₃ N ₄	300 W Xe lamp $\lambda > 420$ nm	10% (vol.) TEOA	261	[6]
Cd _x Zn _{1-x} S/Au/g-C ₃ N ₄	300 W Xe lamp $\lambda > 420$ nm	0.1 mol L ⁻¹ glucose	123.2	[7]
WS ₂ /g-C ₃ N ₄	300 W Xe lamp	20% (vol.) TEOA	154	[8]

Reference

[1] *Journal of Materials Science & Technology*. **2024**, *175*, 3, 104–114.

[2] *Carbon*. **2016**, *99*, 111–117.

[3] *J. Colloid Interface Sci.* **2019**, *534*, 163–171.

[4] *Int. J. Hydrom. Energy*. **2017**, *42*, 6738–6745.

- [5] *Appl. Catal. B.* **2018**, *222*, 35–43.
- [6] *Appl. Catal. B.* **2019**, *251*, 305–312.
- [7] *Sci. Bull.* **2017**, *62*, 602–609.
- [8] *Int. J. Hydrog. Energy.* **2019**, *44*, 14927–14937.