

Supporting Information

In situ transformation of a Bi-based MOF to highly active catalyst for CO₂ reduction

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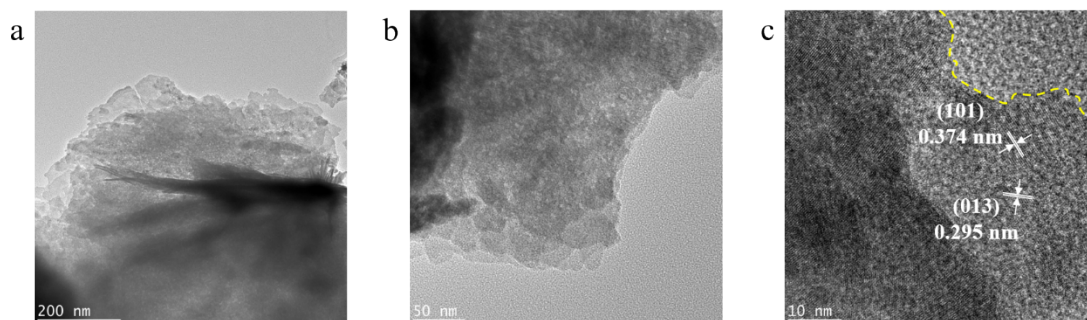


Fig. S1 a, b) TEM images and c) HR-TEM of Bi-TATB soaked in CO₂-saturated KHCO₃ electrolyte.

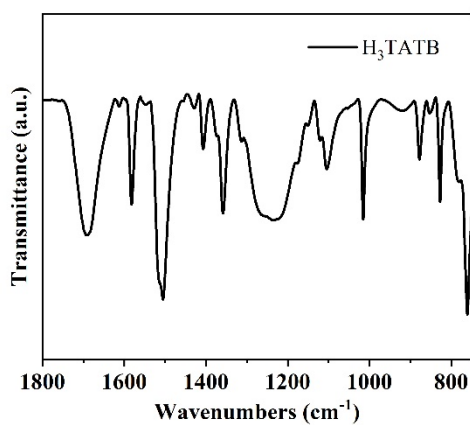


Fig. S2 FTIR spectra of H₃TATB.

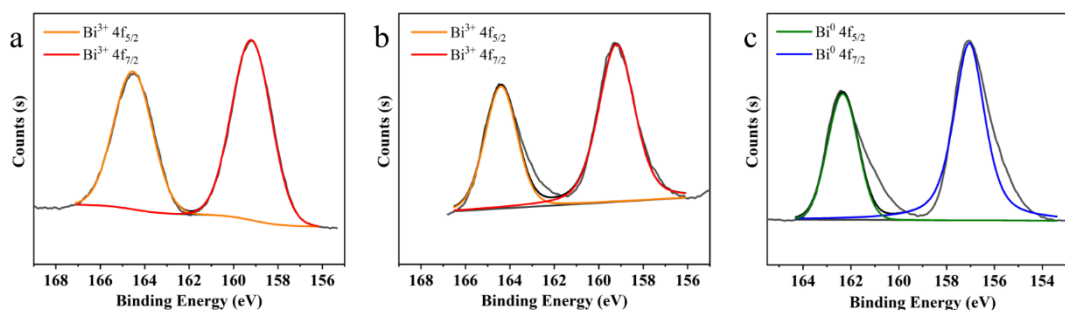


Fig. S3 a-c) High-resolution XPS spectra centered in the binding energy regions of Bi 4f for Bi-TATB, Bi-TATB soaked in KHCO₃ electrolyte and Bi-TATB after electrochemical activation.

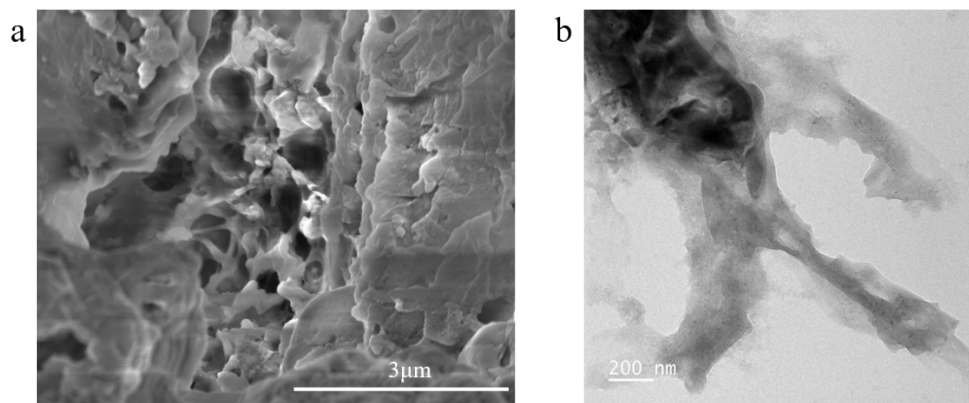


Fig. S4 a) SEM image of Bi-TATB on a carbon paper using Nafion as binder after activation. b) TEM image of Bi-TATB after activation.

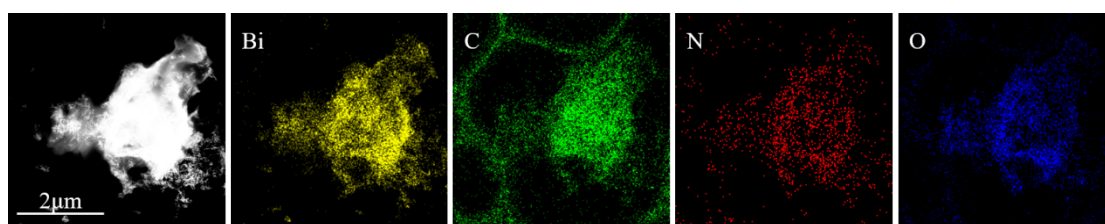


Fig. S5 HAADF-STEM and EDX mapping images of Bi-TATB after activation.

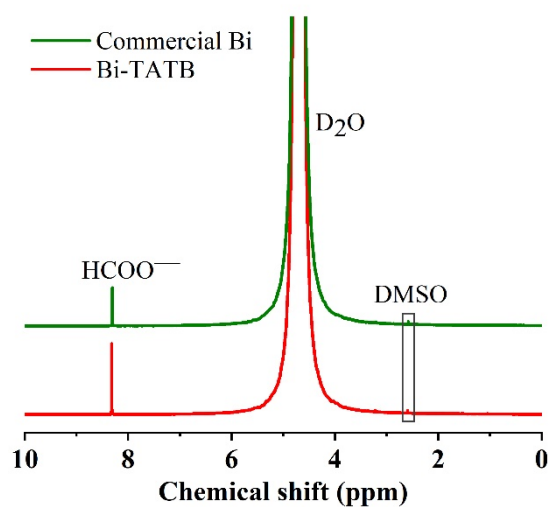


Fig. S6 NMR hydrogen spectrum of liquid products of commercial Bi and Bi-TATB after electrolysis at -1.08 V vs RHE for 2 hours.

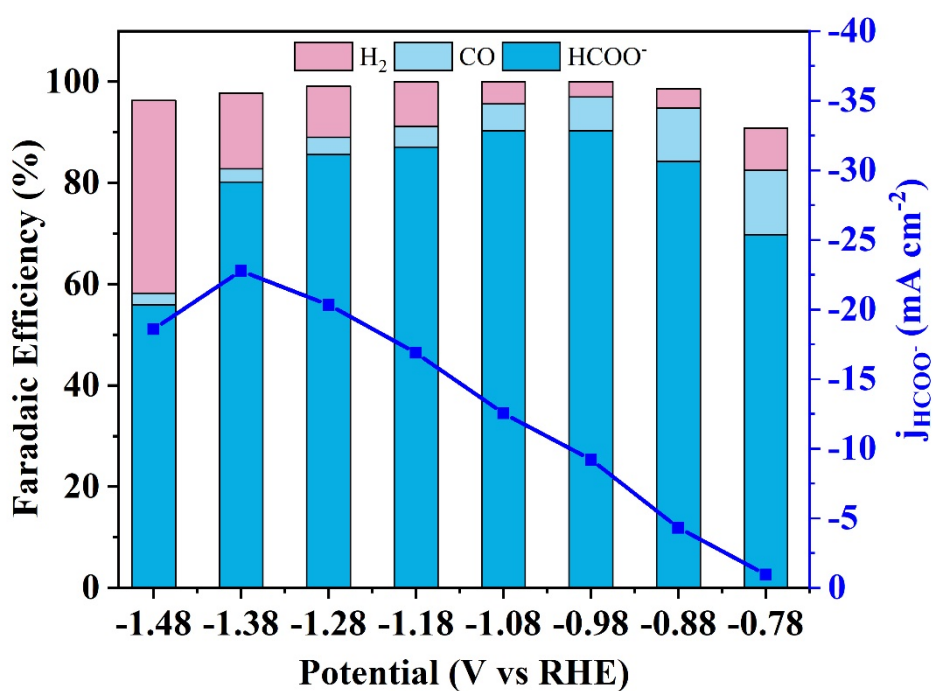


Fig. S7 Potential-dependent Faradaic efficiencies and partial current densities of HCOO⁻ for commercial Bi in CO₂-saturated 0.5 M KHCO₃.

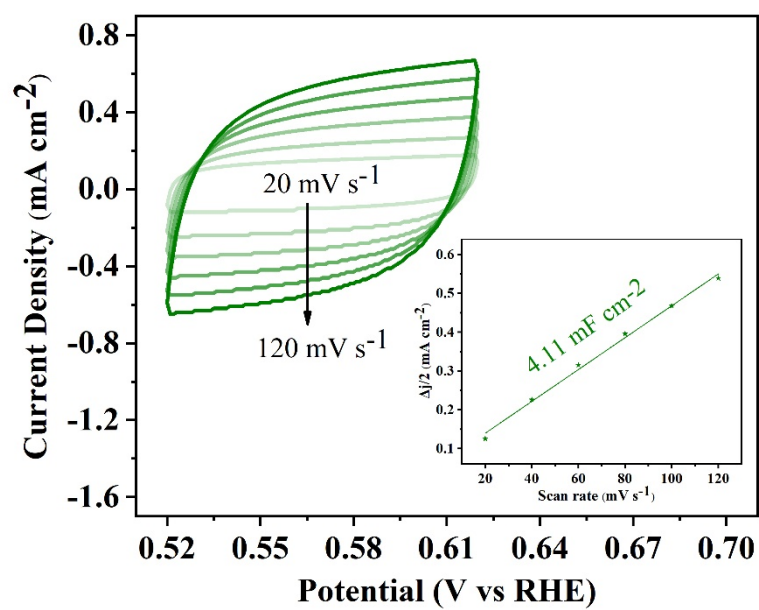


Fig. S8 C_{dl} measurement of commercial Bi.

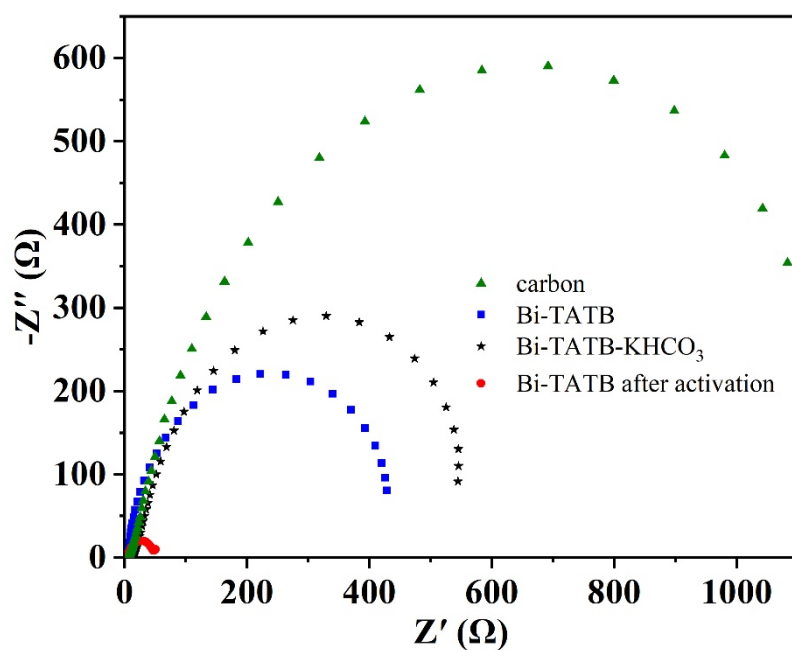


Fig. S9 EIS of different samples.

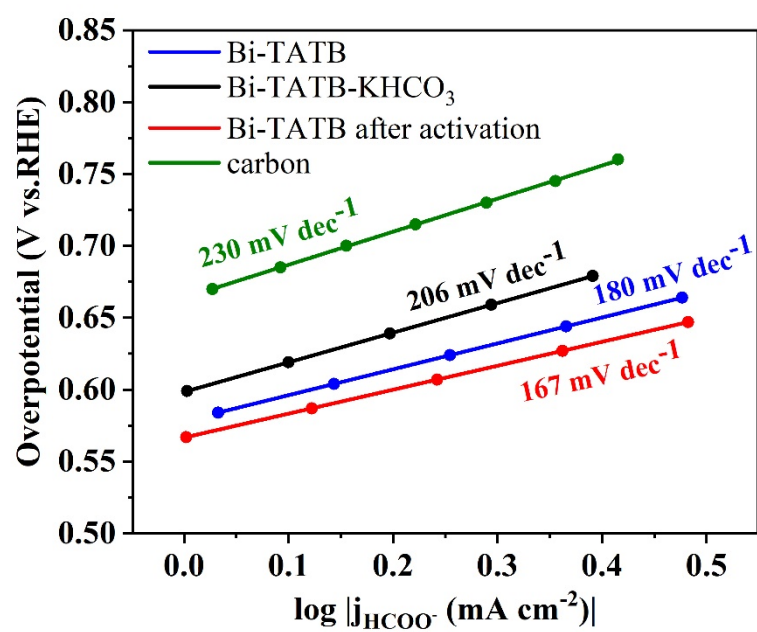


Fig. S10 Tafel slope of different samples.

Table S1. Electrocatalytic performance comparisons of the Bi-TATB with previously reported Bi-based electrocatalysts for CO₂-to-formate.

Catalyst	Electrolyte	Potential (V vs.RHE)	FE _{formate}	j _{formate} (mA cm ⁻²)	Stability (h)
this work	0.5 M KHCO ₃	-1.08	94.3%	20.0	14.5
CAU-17 ¹	0.1 M KHCO ₃	-0.9	92.2%	15.0	30
Bi-ZMOF ²	0.1 M KHCO ₃	-1.1	91%	-	12
MFM-220 ³	0.1 M KHCO ₃	-1.1	90.4%	20.8	-
Bi ₂ O ₂ CO ₃ derived Bi ⁴	0.1 M KHCO ₃	-1.0	90%	4.8	-
BiPO ₄ derived Bi ⁵	0.5 M KHCO ₃	-0.9	91.4%	7.2	15
Bi ₂ S ₃ derived Bi ⁶	0.5 M NaHCO ₃	-0.75	84%	4.2	24
Bi-NFs ⁷	0.1 M KHCO ₃	-0.9	92.3%	11.4	12
Bi@C nanotubes ⁸	0.1 M KHCO ₃	-0.9	90.9%	6.0	24
Bi nanosheets ⁹	0.1 M KHCO ₃	-1.1	86%	16.5	10
α-Bi ₂ O ₃ ¹⁰	0.5 M KHCO ₃	-1.58	93%	-	-
f-Bi ₂ O ₃ ¹¹	0.1 M KHCO ₃	-1.2	87%	20.9	-

Reference

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