Study of two benzophenone-based difluoroboron compounds containing triphenylamine units mechanofluorochromic behaviors and latent fingerprints imaging

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Table S1 Photophysical data of two compounds insolvents19**Table S2** The emission wavelengths of two compounds in mixtures of THF/water19

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Figure S7 HRMS of compound TPA-BP-BF₂-1.



Figure S8 HRMS of compound TPA-BP-BF₂-2.



Figure S9 The Φ_f and τ of compound **TPA-BP-BF₂-1** in THF/water mixture of $f_w = 0\%$.



Figure S10 The Φ_f and τ of compound **TPA-BP-BF₂-1** in THF/water mixture of f_w = 80%.









Figure S14 The Φ_f and τ of compound **TPA-BP-BF₂-2** in THF/water mixture of f_w = 100%.



Figure S15 SEM images of compounds (a) **TPA-BP-BF₂-1** and (b) **TPA-BP-BF₂-2** in pure water (c = 1×10^{-5} mol/L), respectively.



Figure S16 The Φ_f of solid of compound **TPA-BP-BF**₂-1 before (up) and after grinding (down).





Figure S18 The Φ_f of solid of compound **TPA-BP-BF**₂-2 before (up) and after grinding (down).



Figure S19 The τ of solid of compound TPA-BP-BF₂-2 before (left) and after grinding (right).



Figure S20 The emission wavelength of TPA-BP-BF₂-2 upon treated by grinding and fuming with CH_2Cl_2 repeatedly.



Figure S21 The Φ_f (up) and τ (down) of the developer with compound **TPA-BP-BF**₂-1.



Figure S22 The Φ_f (up) and τ (down) of the developer with compound **TPA-BP-BF**₂-2.



Figure S23 SEM images of developers with (a) TPA-BP-BF₂-1 and (b) TPA-BP-BF₂-2, respectively.



Figure S24 Photograph of the LFPs imaging on mobile phone tempered film and tinfoil substrates after being processed with compounds **F-1** and **F-2** under 365 nm UV light.

			TPA-BP-BF ₂ -1	TPA-BP-BF ₂ -2
		<i>n</i> -hexane	303, 381	305, 382
		CH ₂ Cl ₂	304, 384	305,389
	UV (λ _{abs} , nm)	CH₃CN	301, 376	279, 380
		DMF	304, 379	305, 384
colvente		CH₃OH	301, 380	304, 384
solvents	PL (λ _{nm} , nm)	<i>n</i> -hexane	481	499
		CH ₂ Cl ₂	518	517
		CH₃CN	529	519
		DMF	523	510
		CH₃OH	486	512

Table S1 Photophysical data of compounds TPA-BP-BF₂-1 and TPA-BP-BF₂-2 in solvents.

Table S2 The emission wavelengths of compounds TPA-BP-BF₂-1 and TPA-BP-BF₂-2 in mixtures of THF/water.

<i>f</i> _w (%)	0	10	20	30	40	50	60	70	80	90	100
TPA-BP-BF ₂ -	471	493	488	481	480	485	490	472	488	600	603
1											
TPA-BP-BF ₂ -	488	502	506	504	504	509	512	499	494	497	586
2											

Computational details

Kohn-Sham density functional theory (DFT) has been employed to optimize the ground state geometries of the investigated complexes at the B3LYP/6-31G(d, p) level. All the optimized geometries were tested to be local minima by frequency calculations at the same level. To get insight into the photophysical properties of the investigated complexes, time-dependent density functional theory (TD-DFT) calculations at the CAM-B3LYP¹/6-31G(d, p) have been performed. The effect of the solvent was considered in all DFT and TD-DFT calculations utilizing the integral equation formalism polarized continuum model (IEF-PCM) with the dichloromethane as solvent which has been employed in the experiment. All the DFT and TD-DFT calculations were performed using the Gaussian 16 software suit.²

According to the hole-electron analysis method, the "hole" and "electron" denote where the excited electron leaves and goes, respectively. In many cases, any excitation can be identified as a definitive distribution of hole and electron. The theory proved to be a useful and powerful method in unraveling nature of electron excitations.³ The wavefunction analysis was calculated by means of the Multiwfn version 3.8(dev) code⁴ and plotted using VMD⁵ and GaussView 6.0⁶ software.

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Ν	8.290435	0.807232	0.004609	н	9.760440	2.654708	-1.222174
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С	9.384067	0.048573	-0.489059	н	9.130437	5.775246	1.658625
С	10.611722	0.066406	0.177447	Н	7.693708	4.242003	2.982690
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Н	-6.623028	2.101296	-3.233724	С	10.095265	4.266507	1.410840
Н	-11.369449	0.777830	2.273688	С	9.556446	4.276928	2.693285
Н	-10.376465	-0.320095	1.306576	С	8.835784	3.173255	3.138268
Н	-9.608303	0.933638	2.281743	С	8.661649	2.066877	2.317174
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Н	-9.906099	3.404180	1.666824	С	-7.517456	0.149470	0.361322
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Н	-11.750786	0.750008	2.314297	Н	-7.704755	5.088608	-0.326252
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