

Carbonylative Cyclization of Biaryl Enones with Aldehydes and Oxamic Acids

Chada Raji Reddy,^{*,a,b} Dattahari H. Kolgave,^{a,b,‡} Sana Fatima^{a,b,‡} and Remya Ramesh^{a,b}

^aDepartment of Organic Synthesis & Process Chemistry, CSIR-Indian Institute of Chemical Technology, Hyderabad 500007, India.

^bAcademy of Scientific and Innovative Research (AcSIR), Ghaziabad 201 002, India

[‡]Both these authors contributed equally

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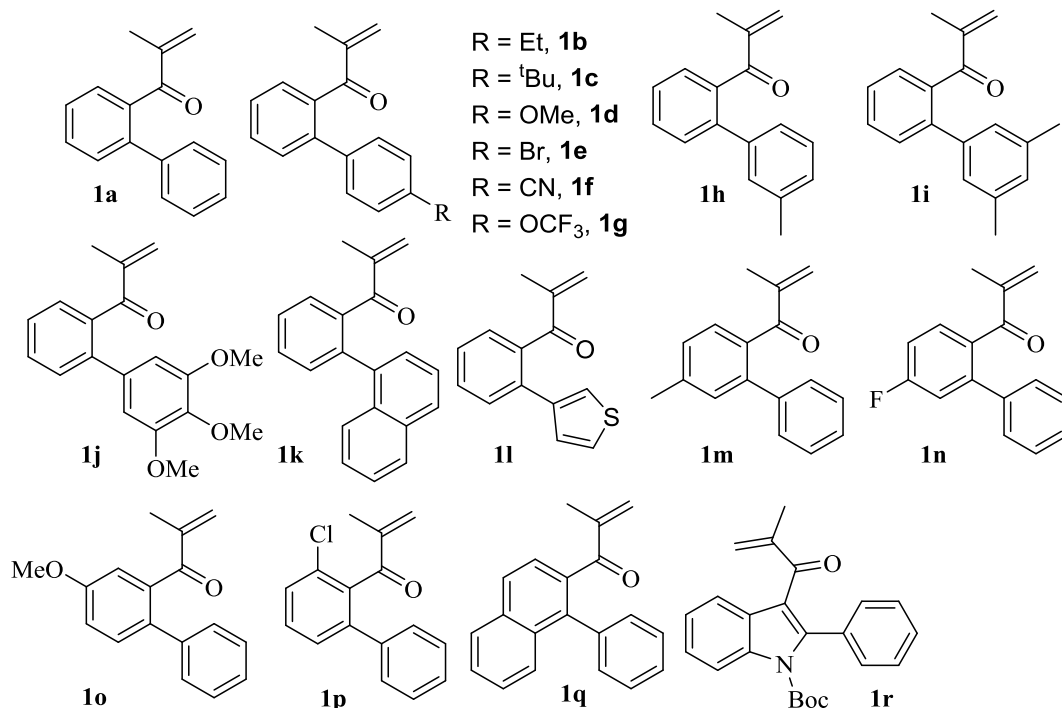
1. General information

All the reactions were performed in oven-dried glass apparatus, the air and moisture sensitive reactions were carried out under inert atmosphere (nitrogen) using freshly distilled anhydrous solvents. Commercially available reagents were used as such without further purification. All reactions were monitored by thin-layer chromatography carried out on silica plates using UV-light and anisaldehyde for visualization. Column chromatography was performed on silica gel (100-200 mesh) using hexanes and ethyl acetate as eluent. ¹H NMR was recorded in CDCl₃, DMSO on 500 MHz and 400 MHz and ¹³C NMR was recorded on 151 MHz, 126 MHz and 101 MHz. δ 7.26 and δ 7.77 and δ 2.5, δ 39.5 are corresponding to CDCl₃ and DMSO-*d*₆ in ¹H NMR and ¹³C NMR respectively. Chemical shifts were reported in δ (ppm) relative to TMS as an internal standard and *J* values were given in Hz (hertz). Multiplicity is indicated as, s (singlet); d (doublet); t (triplet); m (multiplet); dd (doublet of doublets), etc. FTIR spectra were recorded on Alpha (Bruker) Infrared Spectrophotometer. High resolution mass spectra (HRMS) [ESI+] were obtained using either a TOF or a double focusing spectrometer.

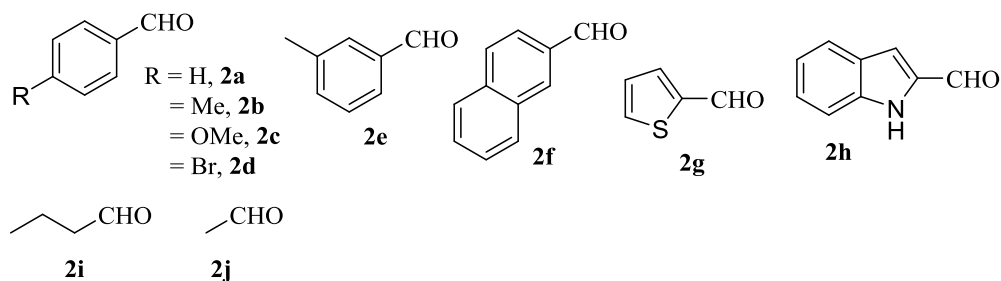
Structures of Biaryl Enones:

All the biaryl enones were prepared following the literature procedures.¹⁻³

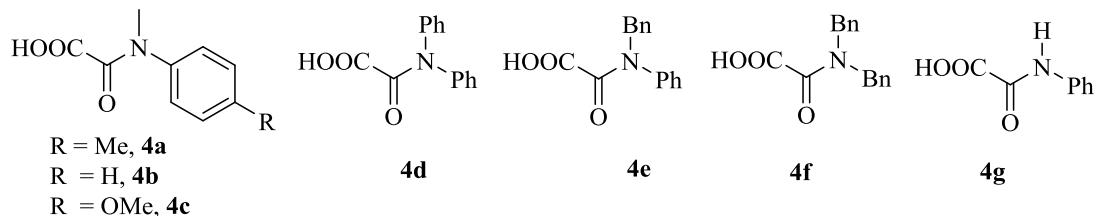
Structures of Biaryl Enones 1a to 1q:



Structures of Aldehydes 2a to 2j:



Structures of Oxamic Acids 4a to 4f:

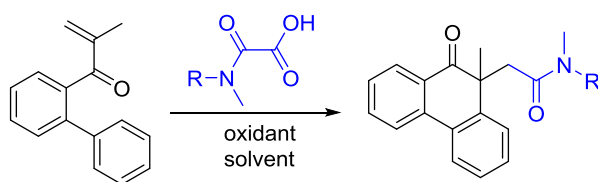


2. Optimization table, Control experiments and Plausible mechanism:

A. Optimization of reaction conditions:

Optimization of the reaction conditions, the cyclic reaction of 1-([1,1'-biphenyl]-2-yl)-2-methylprop-2-en-1-one (1a) with 2 equiv. of oxoacetic acid as the carbamoylating agent was chosen as a model reaction.

Table S1: Optimization of carbamoyl radical-promoted cyclization of biaryl enones:

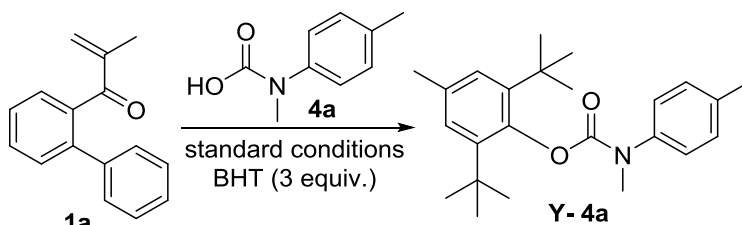


| Entry | Oxidant | Catalyst | Solvent | Temp (°C) | Yield (%) ^b |
|----------|---|-------------------------------------|-------------------------|-----------|------------------------|
| 1 | K ₂ S ₂ O ₈ | AgOAc | CH ₃ CN | 80 | 33 |
| 2 | K ₂ S ₂ O ₈ | AgOTf | CH ₃ CN | 80 | 28 |
| 3 | K ₂ S ₂ O ₈ | AgNO ₃ | CH ₃ CN | 80 | 60 |
| 4 | K₂S₂O₈ | Ag₂CO₃ | CH₃CN | 80 | 86 |
| 5 | Na ₂ S ₂ O ₈ | Ag ₂ CO ₃ | CH ₃ CN | 80 | 12 |
| 6 | (NH ₄) ₂ S ₂ O ₈ | Ag ₂ CO ₃ | CH ₃ CN | 80 | 43 |
| 7 | TBHP | Ag ₂ CO ₃ | CH ₃ CN | 80 | 18 |
| 8 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | DMF | 80 | 48 |
| 9 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | toluene | 80 | - |
| 10 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | DMSO | 80 | 42 |
| 11 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | 1,4-dioxane | 80 | 50 |
| 12 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | THF | 60 | 30 |
| 13 | K ₂ S ₂ O ₈ | Ag ₂ CO ₃ | CH ₃ CN | rt | - |
| 14 | K ₂ S ₂ O ₈ | - | CH ₃ CN | 80 | - |
| 15 | - | Ag ₂ CO ₃ | CH ₃ CN | 80 | - |

Unless otherwise stated all the reactions were performed using **1a** (0.3 mmol) with **4a** (0.6 mmol), Ag₂CO₃ (20mol%) oxidant (1.2 mmol) in 3 mL of solvent, ^bIsolated yield

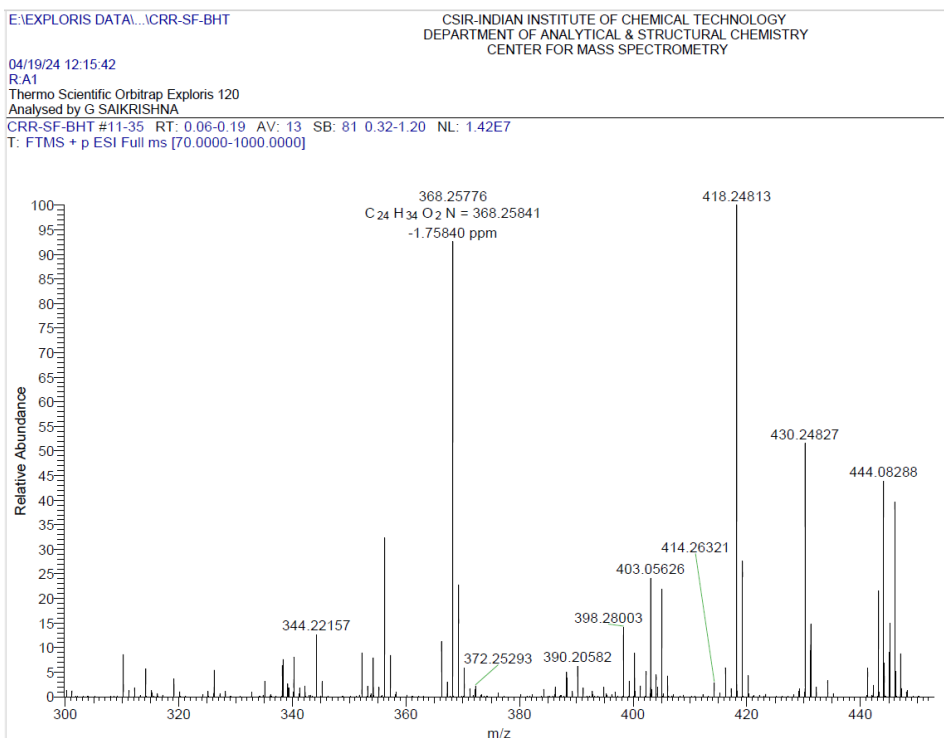
B. Radical trapping experiment:

Control experiment was conducted by adding methyl(p-tolyl) carbamic acid to biaryl enone in presence of 3.0 equiv of butylated hydroxytoluene (BHT; radical scavenger), and was found that the reaction was completely inhibited and no product (**5a**) formation, instead BHT-adduct **Y-4a** was isolated (confirmed by the reaction mass HRMS).



BHT-carbamoyl adduct
confirmed by HRMS found for
C₂₄H₃₄NO₂: 368.2577; Calculated : 368.2511

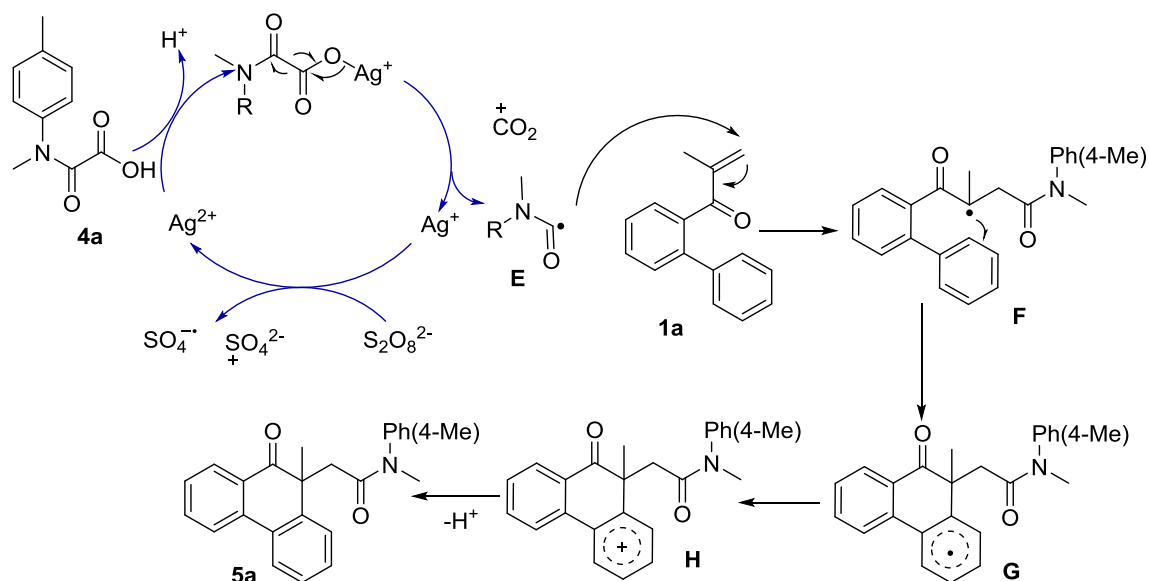
Scheme S1



HRMS for BHT-carbamoyl adduct Y-4a

Plausible mechanism:

A plausible mechanism is proposed based on the result of control experiment and literature survey (scheme 5). First, Ag(I) is oxidized to an Ag (II) species by S₂O₈²⁻. Then, the Ag (II) species oxidizes oxamic acid to form carbamoyl radical **E** *via* decarboxylation. Regioselective addition of **E** on to olefin produces the tertiary radical **F**, which undergoes an intramolecular cyclization with the adjacent phenyl ring to give radical intermediate **G**. Then, a single electron transfer (SET) from **G** to oxidant, generates the cation **H**, which then rearomatizes through the loss of a proton producing the phenanthrenone **5a**.



Scheme S2

3. X-ray Crystallography:

X-ray data for the compound was collected at room temperature on a Bruker D8 QUEST instrument with an I μ S Mo microsource ($\lambda = 0.7107$ Å) and a PHOTON-III detector.⁴ The raw data frames were reduced and corrected for absorption effects using the Bruker Apex 3 software suite programs [1]. The structure was solved using intrinsic phasing method [2] and further refined with the SHELXL [2] program and expanded using Fourier techniques.⁵ Anisotropic displacement parameters were included for all non-hydrogen atoms. All C bound H atoms were positioned geometrically and treated as riding on their parent C atoms [C-H = 0.93-0.97 Å, and Uiso(H) = 1.5Ueq(C) for methyl H or 1.2Ueq(C) for other H atoms].

A. Crystal structure determination of 3a:

Crystal Data for C₂₃H₁₈O₂ ($M = 326.37$ g/mol): monoclinic, space group P2₁/c (no. 14), $a = 16.843(8)$ Å, $b = 13.341(7)$ Å, $c = 16.884(7)$ Å, $\beta = 116.018(17)^\circ$, $V = 3409(3)$ Å³, $Z = 8$, $T = 294.15$ K, $\mu(\text{MoK}\alpha) = 0.080$ mm⁻¹, $D_{\text{calc}} = 1.272$ g/cm³, 29333 reflections measured ($2.69^\circ \leq 2\theta \leq 50^\circ$), 5915 unique ($R_{\text{int}} = 0.1299$, $R_{\text{sigma}} = 0.1005$) which were used in all calculations. The final R_1 was 0.0599 ($I > 2\sigma(I)$) and wR_2 was 0.1517 (all data). **CCDC 2344358** deposition numbers contains the supplementary crystallographic data for this paper which can be obtained free of charge at <https://www.ccdc.cam.ac.uk/structures/>

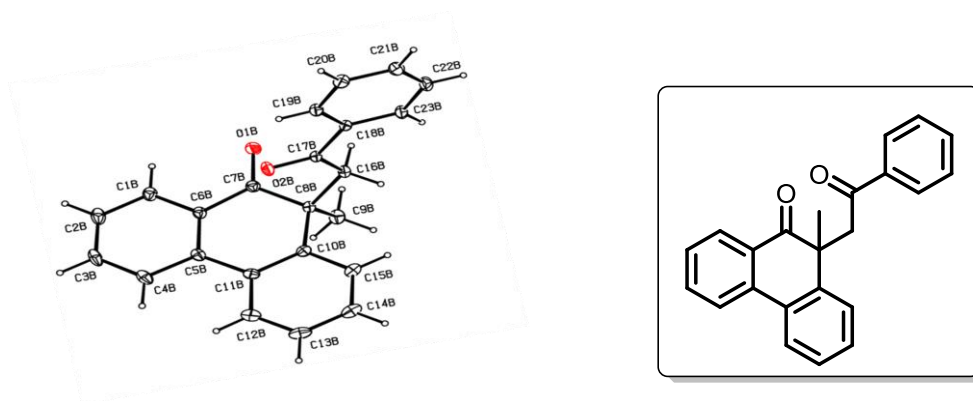


Figure S1: ORTEP diagram of **3a** compound with the atom-numbering. Displacement ellipsoids are drawn at the 30% probability level and H atoms are shown as small spheres of arbitrary radius.

B. Crystal structure determination of **7**:

Crystal Data for $C_{17}H_{14}O_2$ ($M = 250.28$ g/mol): orthorhombic, space group $Pbca$ (no. 61), $a = 18.905(6)$ Å, $b = 7.073(2)$ Å, $c = 18.994(7)$ Å, $V = 2539.8(14)$ Å³, $Z = 8$, $T = 294.15$ K, $\mu(\text{MoK}\alpha) = 0.085$ mm⁻¹, $D_{\text{calc}} = 1.309$ g/cm³, 14223 reflections measured ($4.288^\circ \leq 2\theta \leq 57.834^\circ$), 3225 unique ($R_{\text{int}} = 0.0342$, $R_{\text{sigma}} = 0.0372$) which were used in all calculations. The final R_1 was 0.0471 ($I > 2\sigma(I)$) and wR_2 was 0.1500 (all data). **CCDC 2344359** deposition numbers contains the supplementary crystallographic data for this paper which can be obtained free of charge at <https://www.ccdc.cam.ac.uk/structures/>

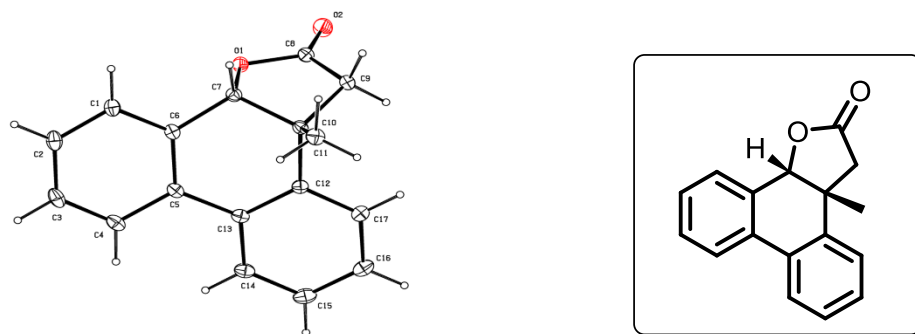
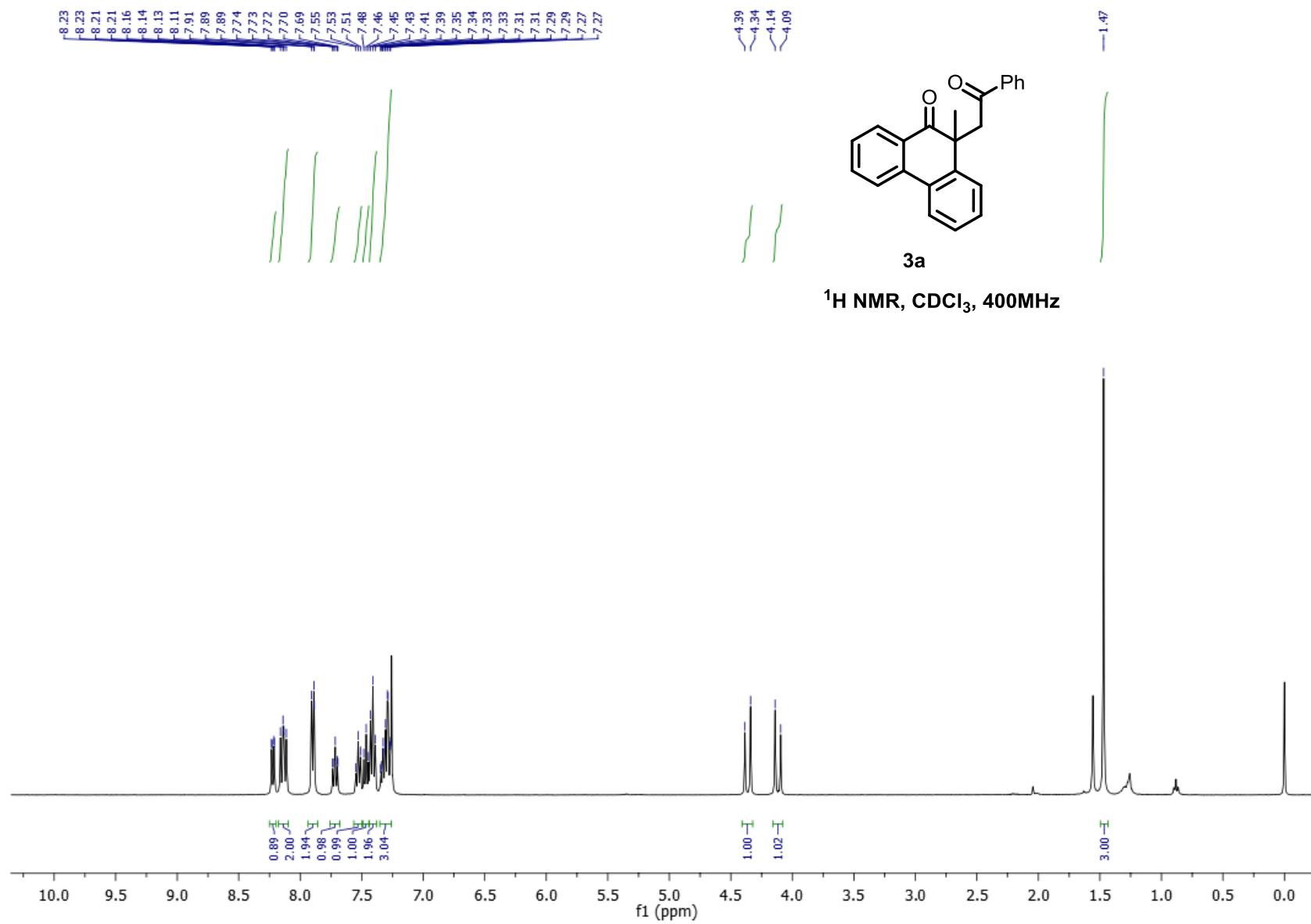
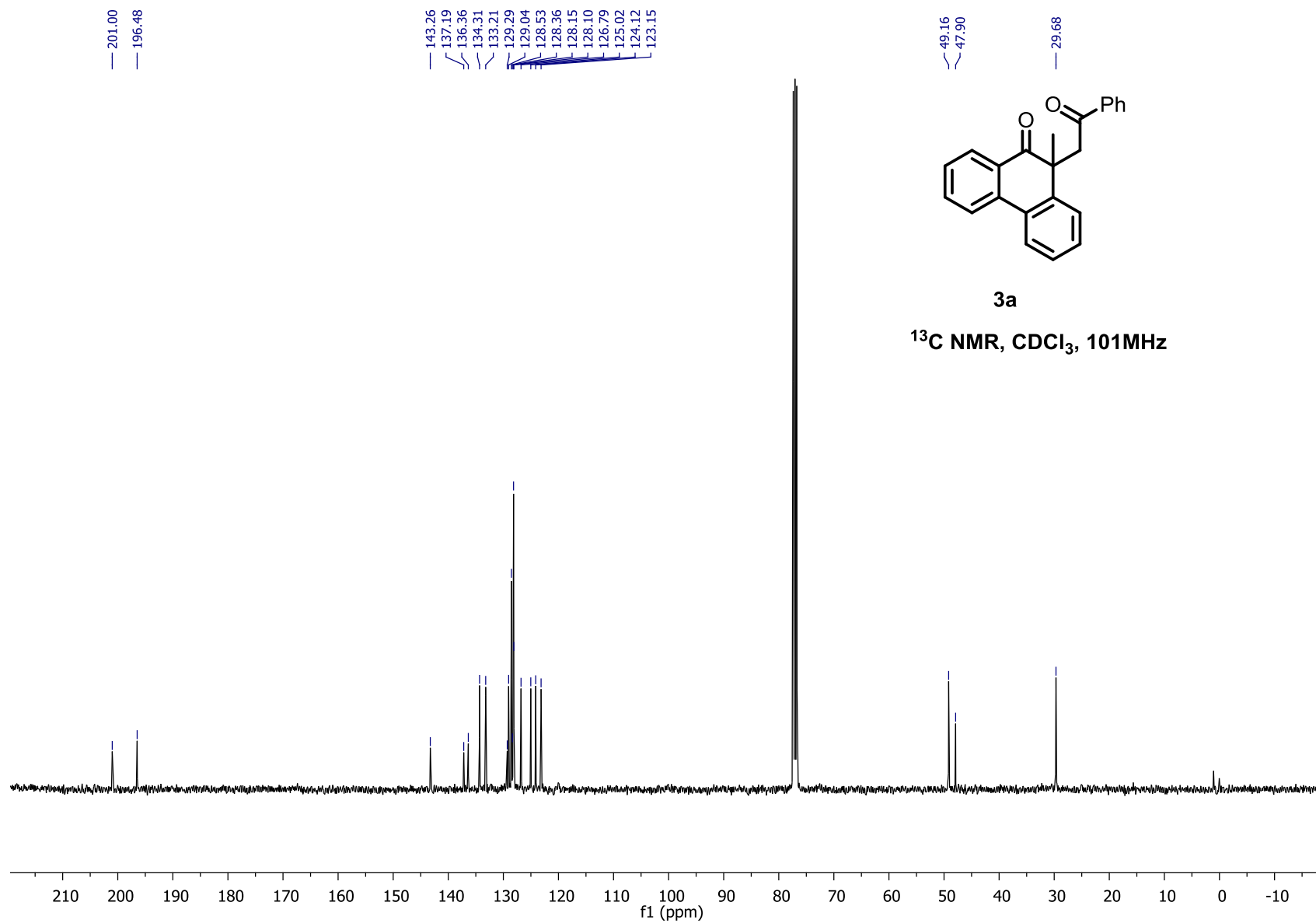


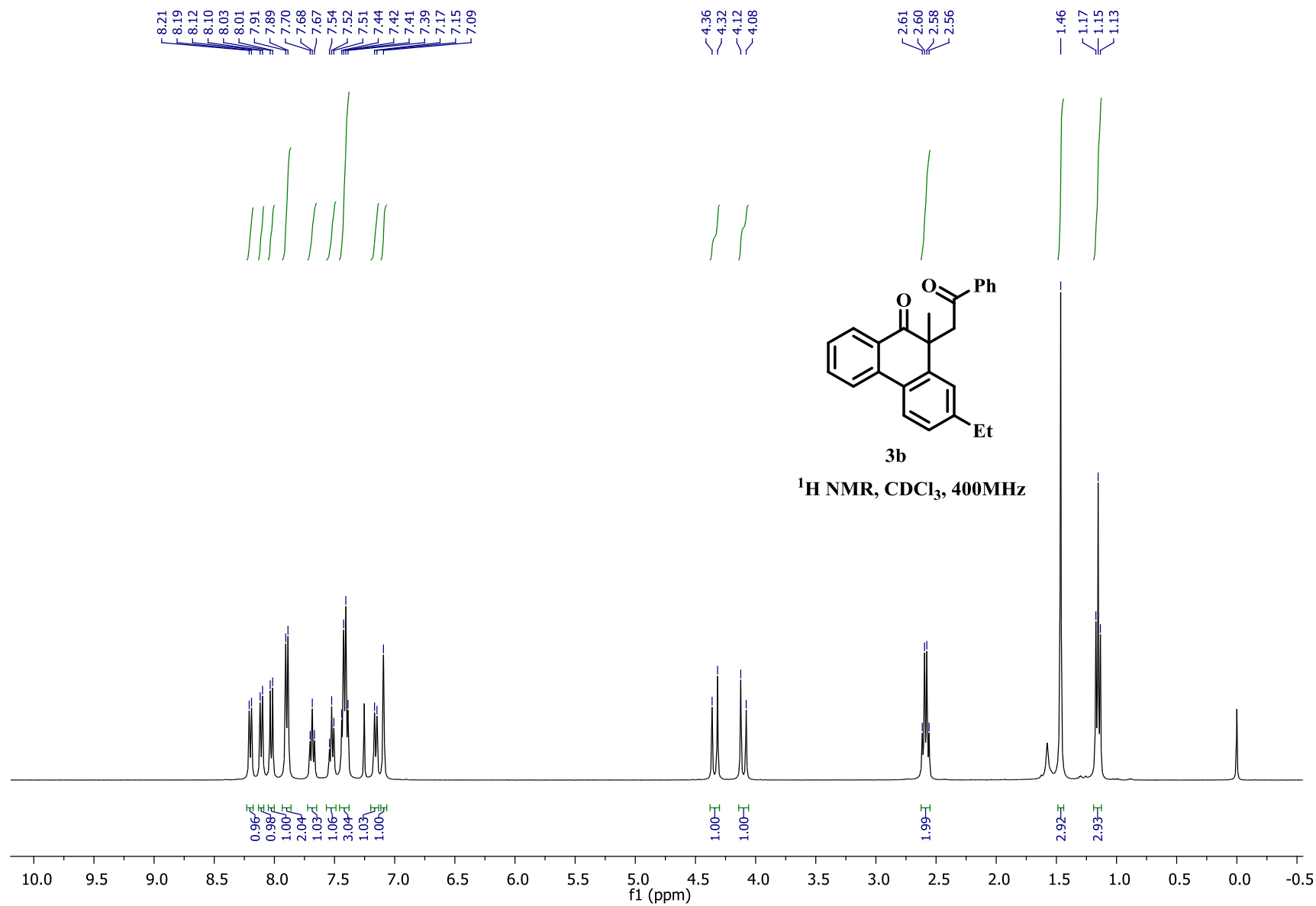
Figure S2: ORTEP diagram of **7** compound with the atom-numbering. Displacement ellipsoids are drawn at the 30% probability level and H atoms are shown as small spheres of arbitrary radius.

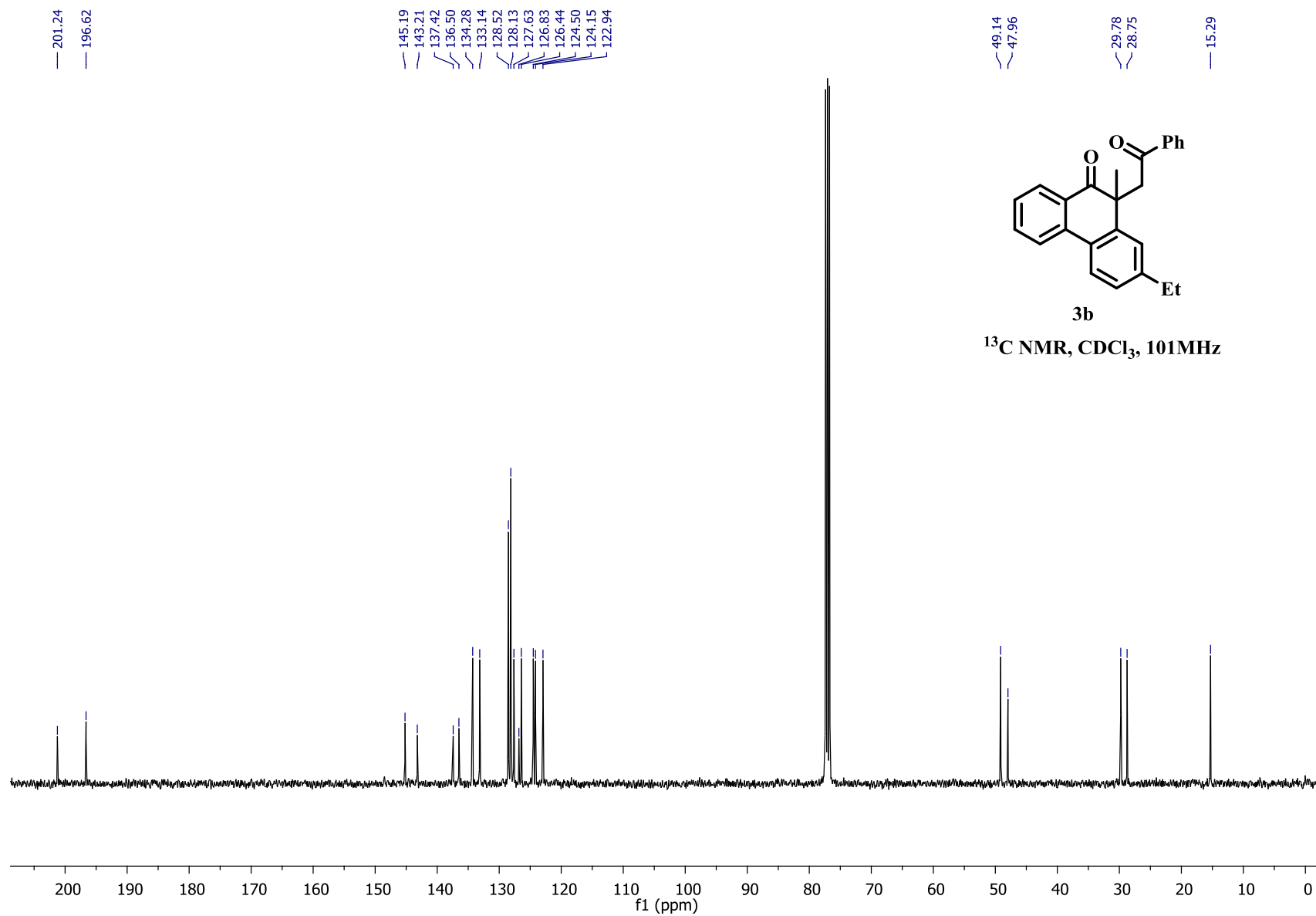
4. References:

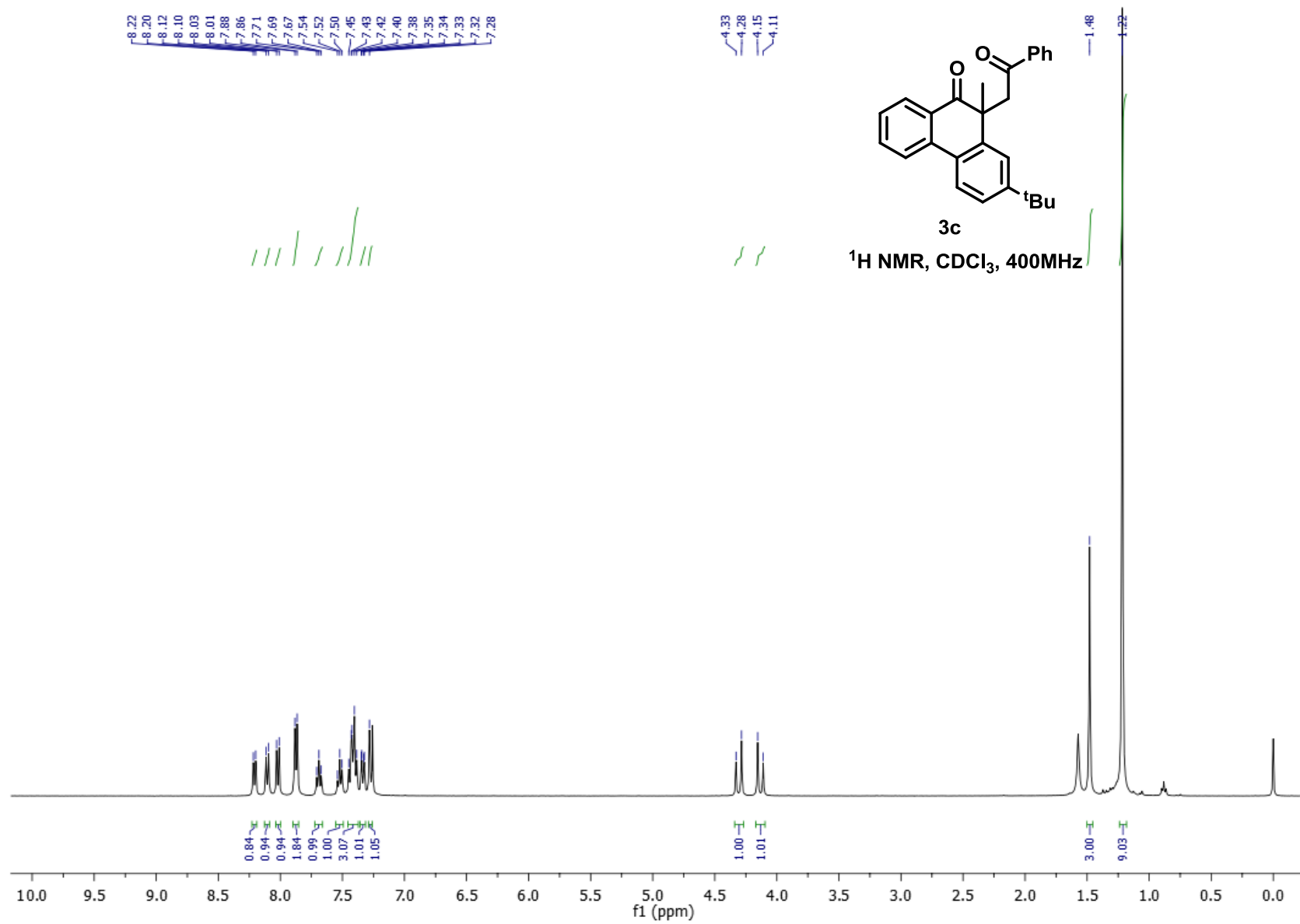
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2. X. Yang, J. Liu, Y. Gao, L. Wang, Y. Zhang, and P. Li, *Asian J. Org. Chem.*, 2022, **11**, e202200269.
3. X. Yang, G. Zhang, J. Zhou, C. Zhou, L. Wang and P. Li, *Org. Biomol. Chem.*, 2023, **21**, 4018–4021.
4. Bruker (2016). APEX3, SAINT and SADABS. Bruker AXS, Inc., Madison, Wisconsin, USA.
5. Sheldrick G. M. (2015). *Acta Crystallogr C* **71**: 3–8.

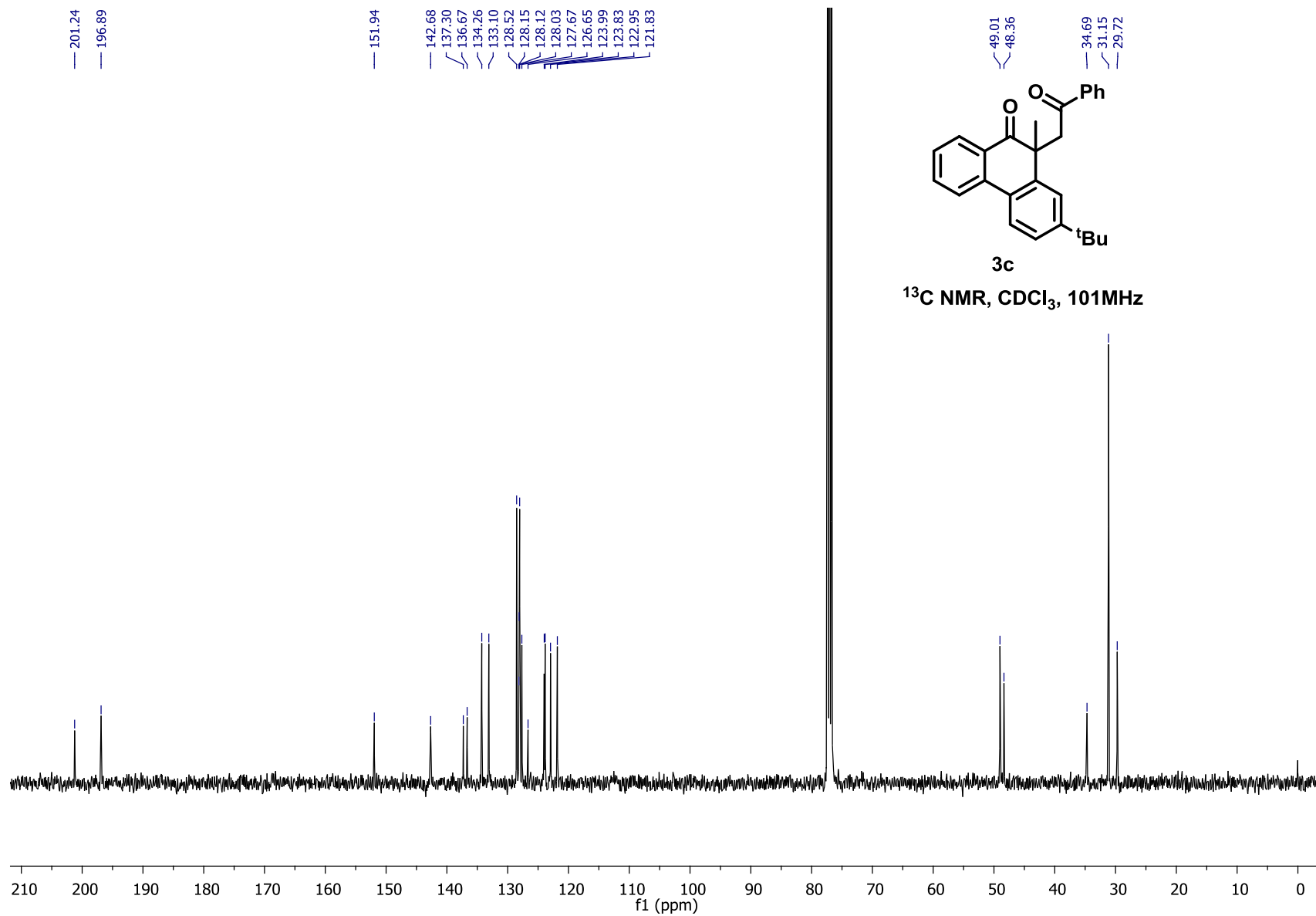


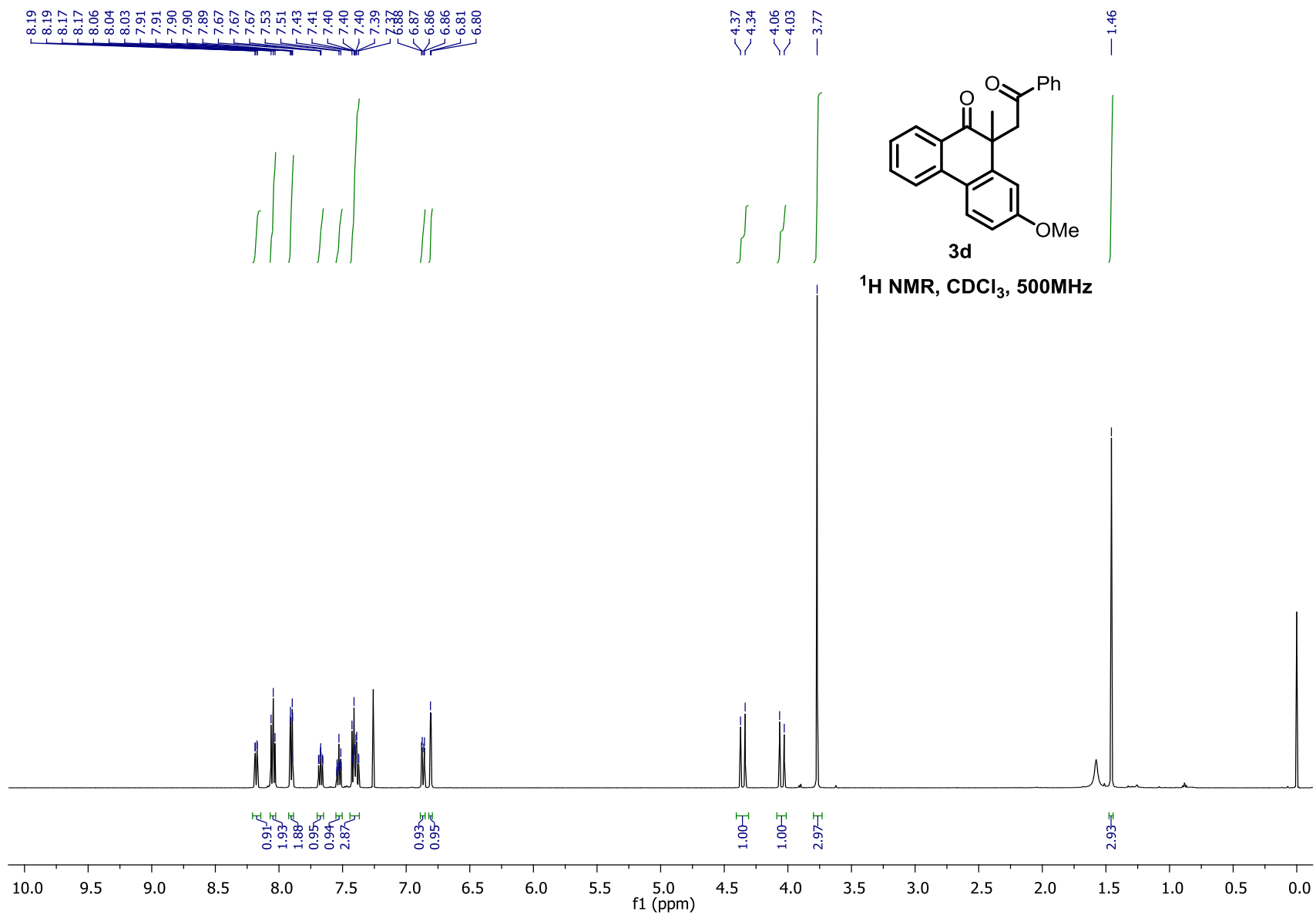


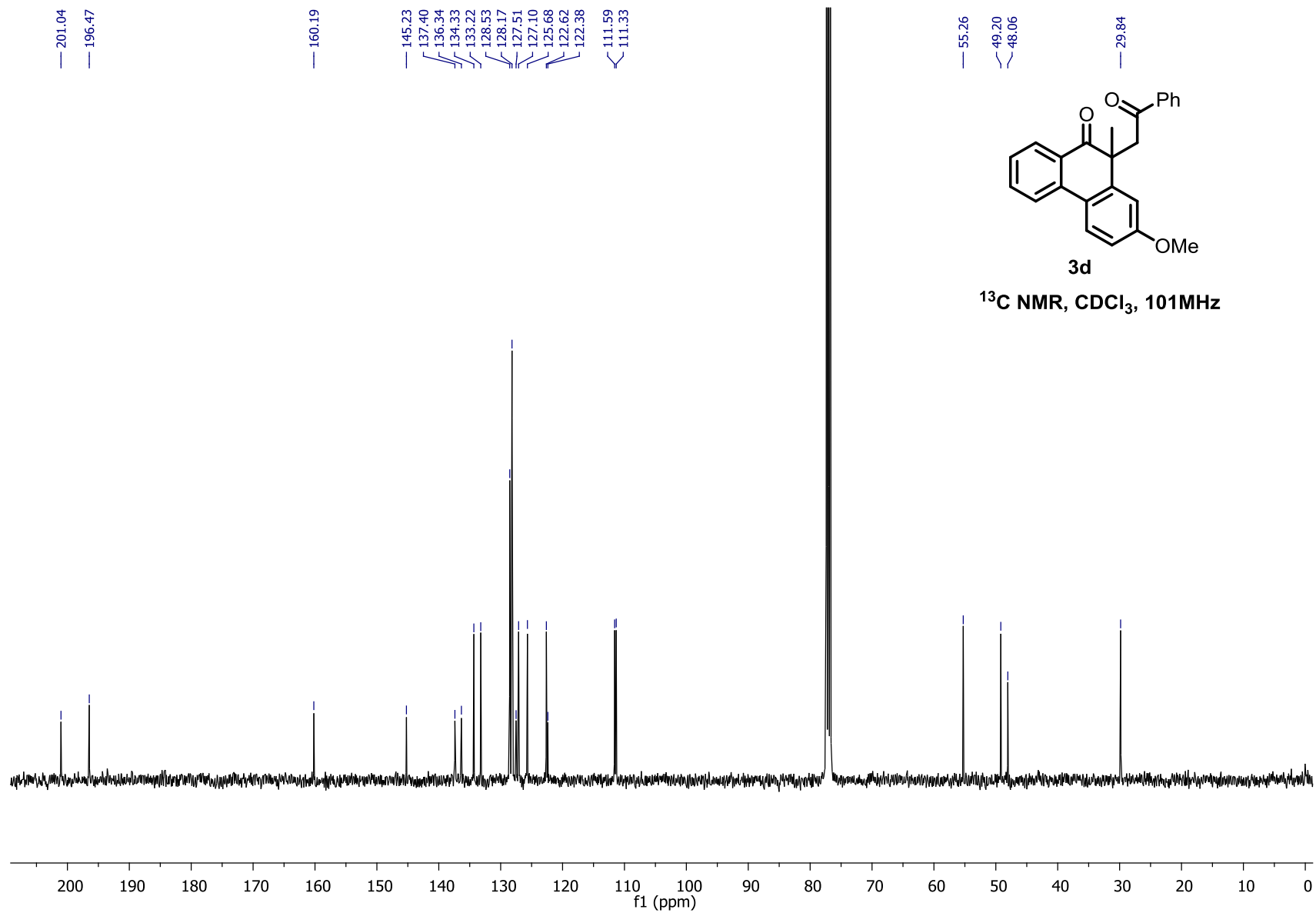


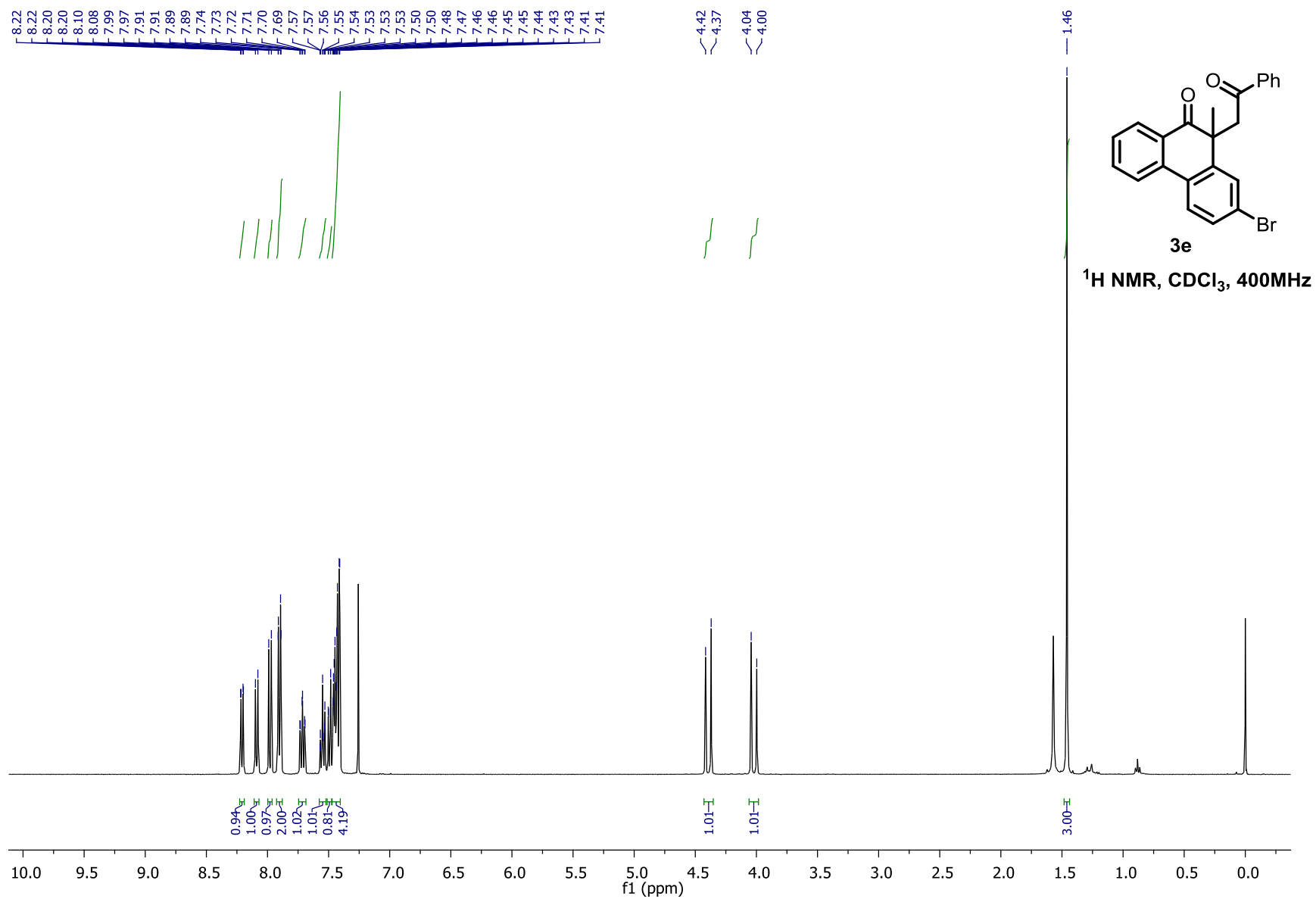


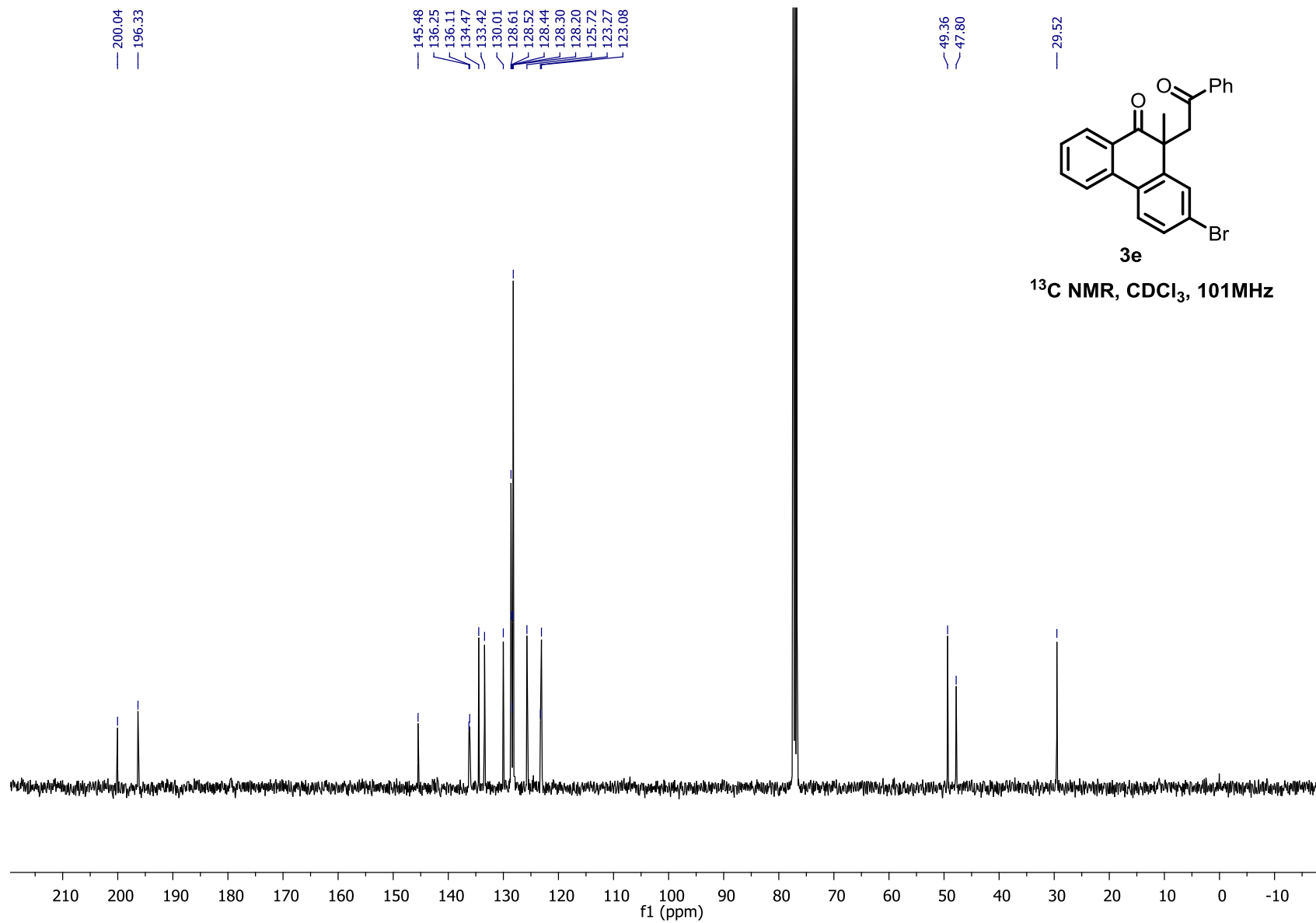


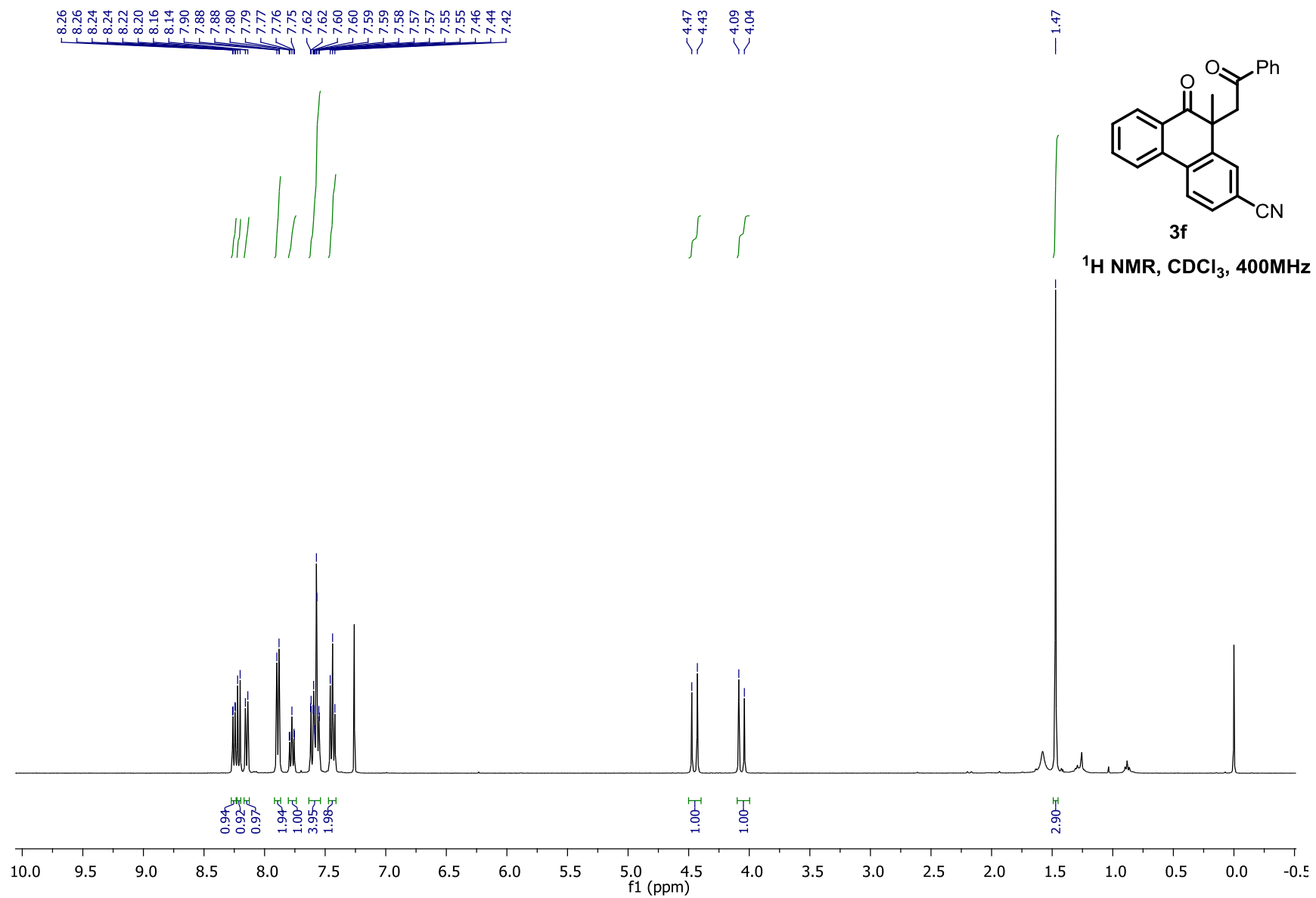


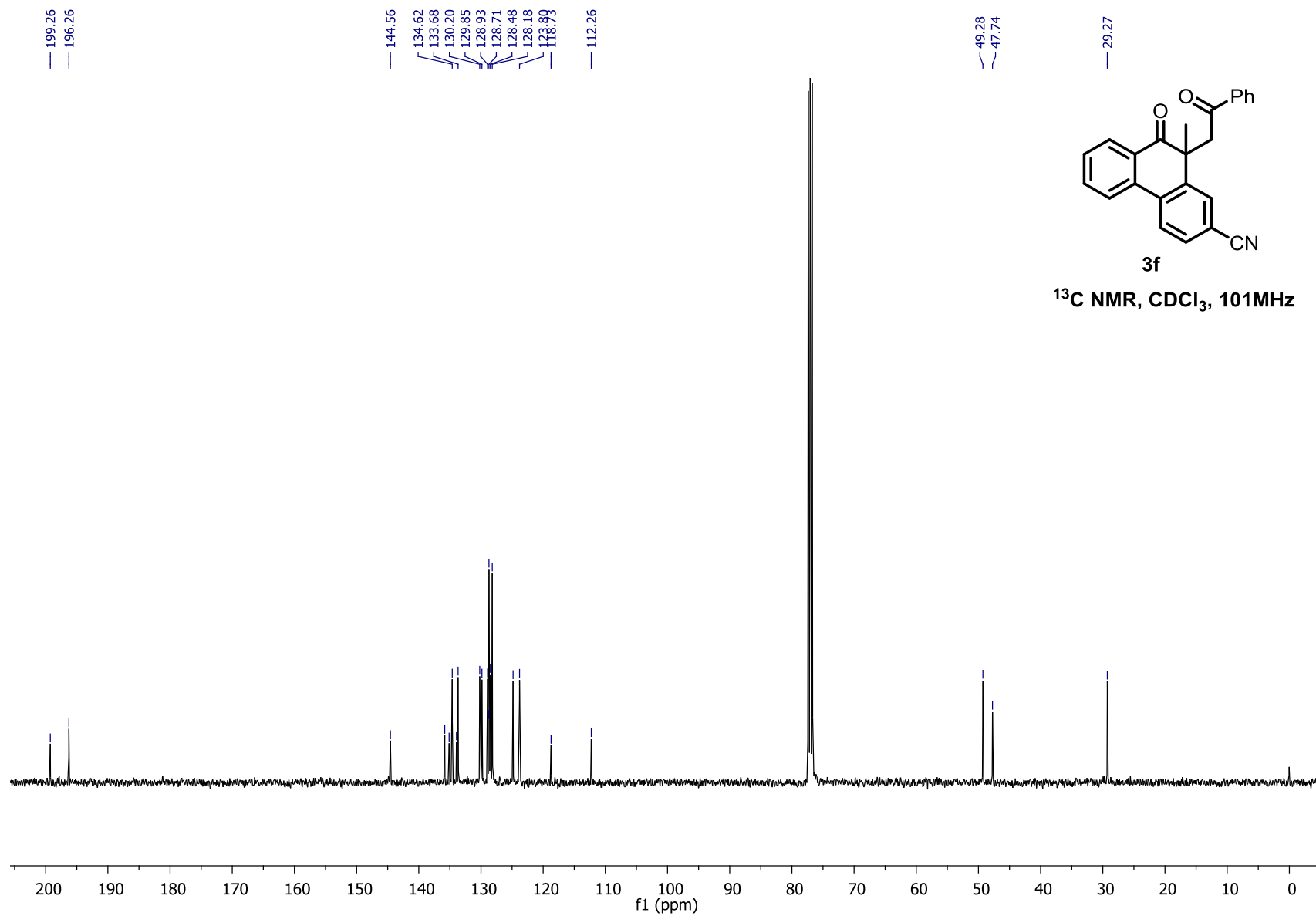


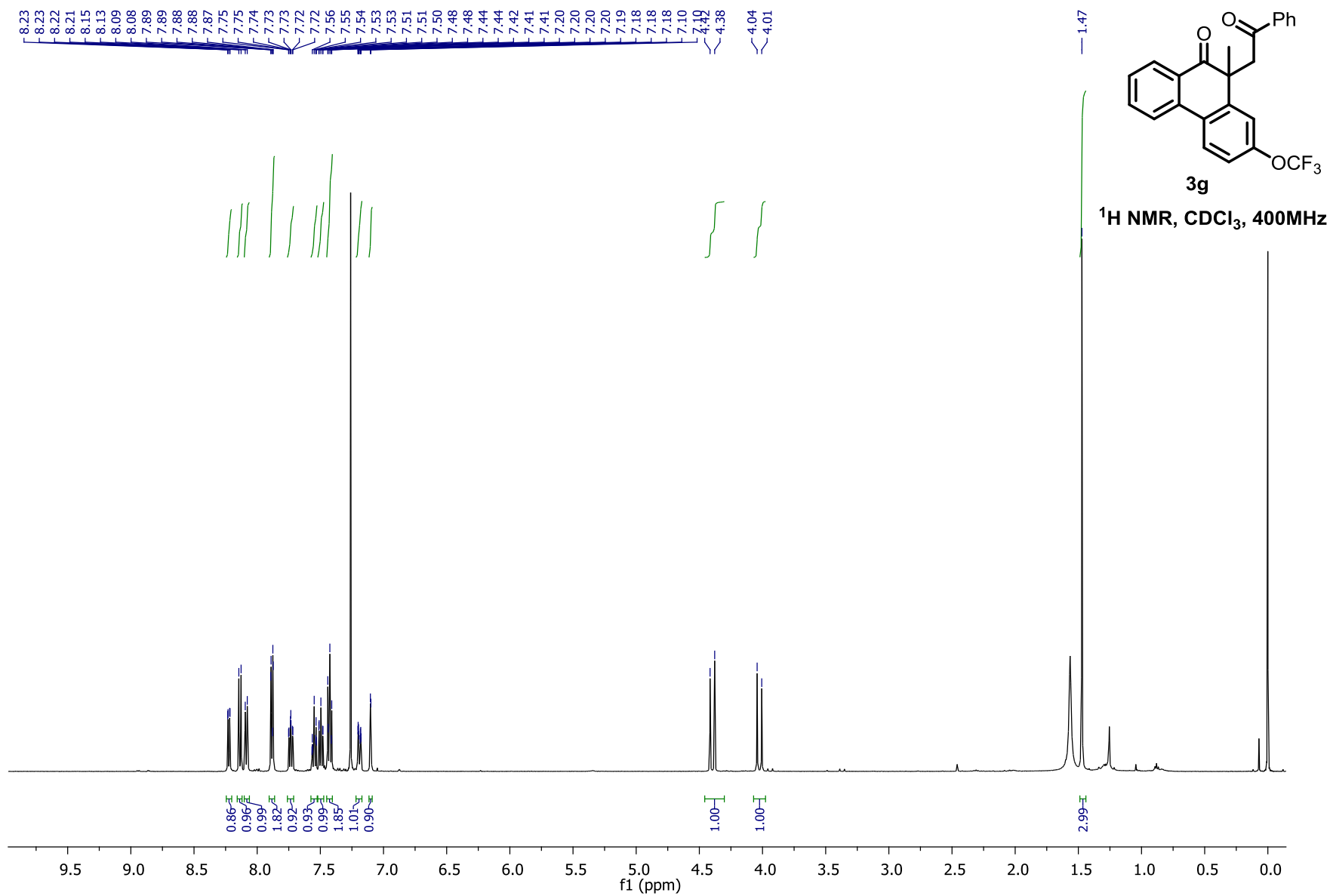


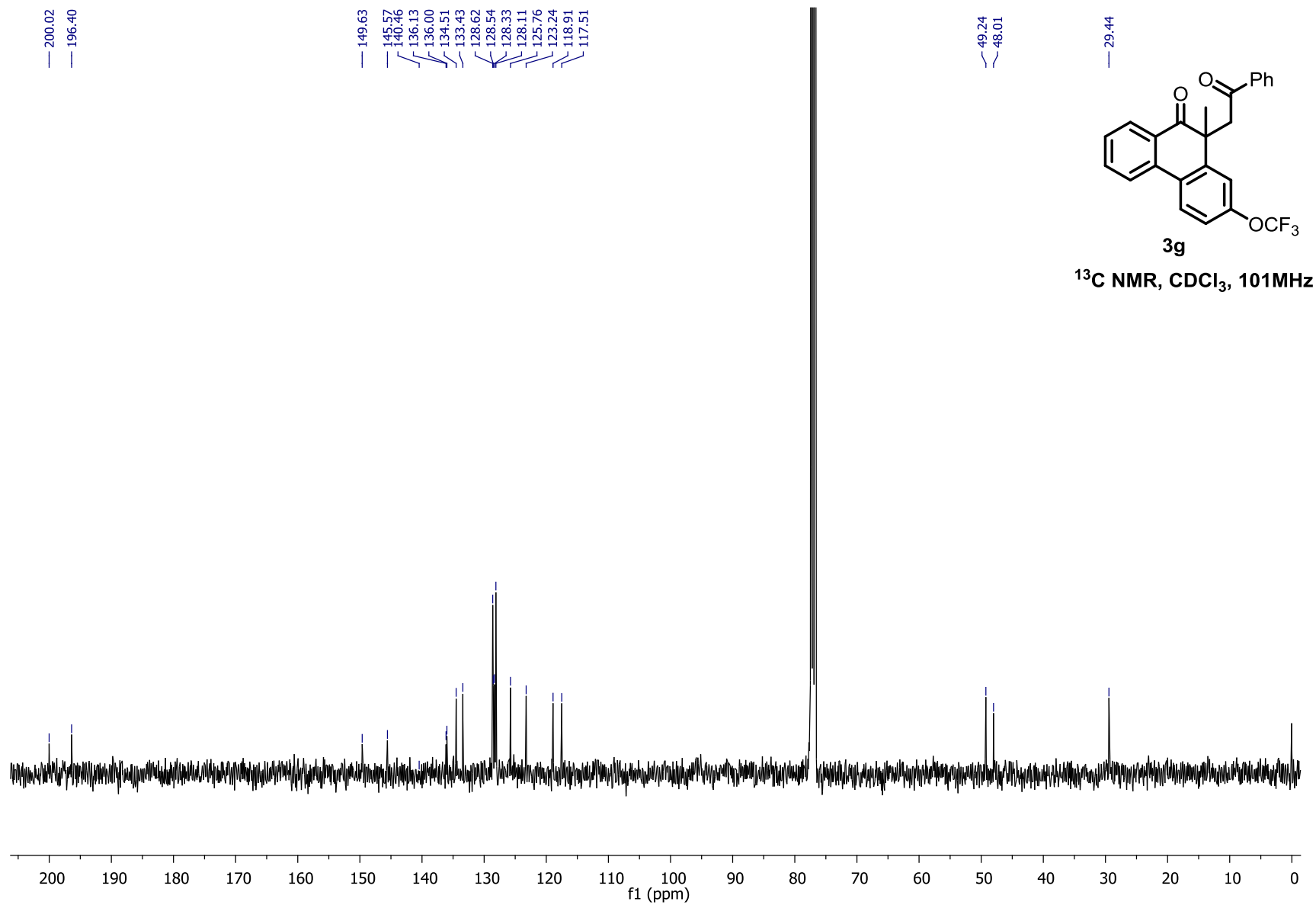


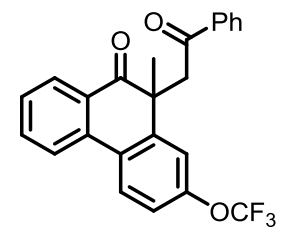






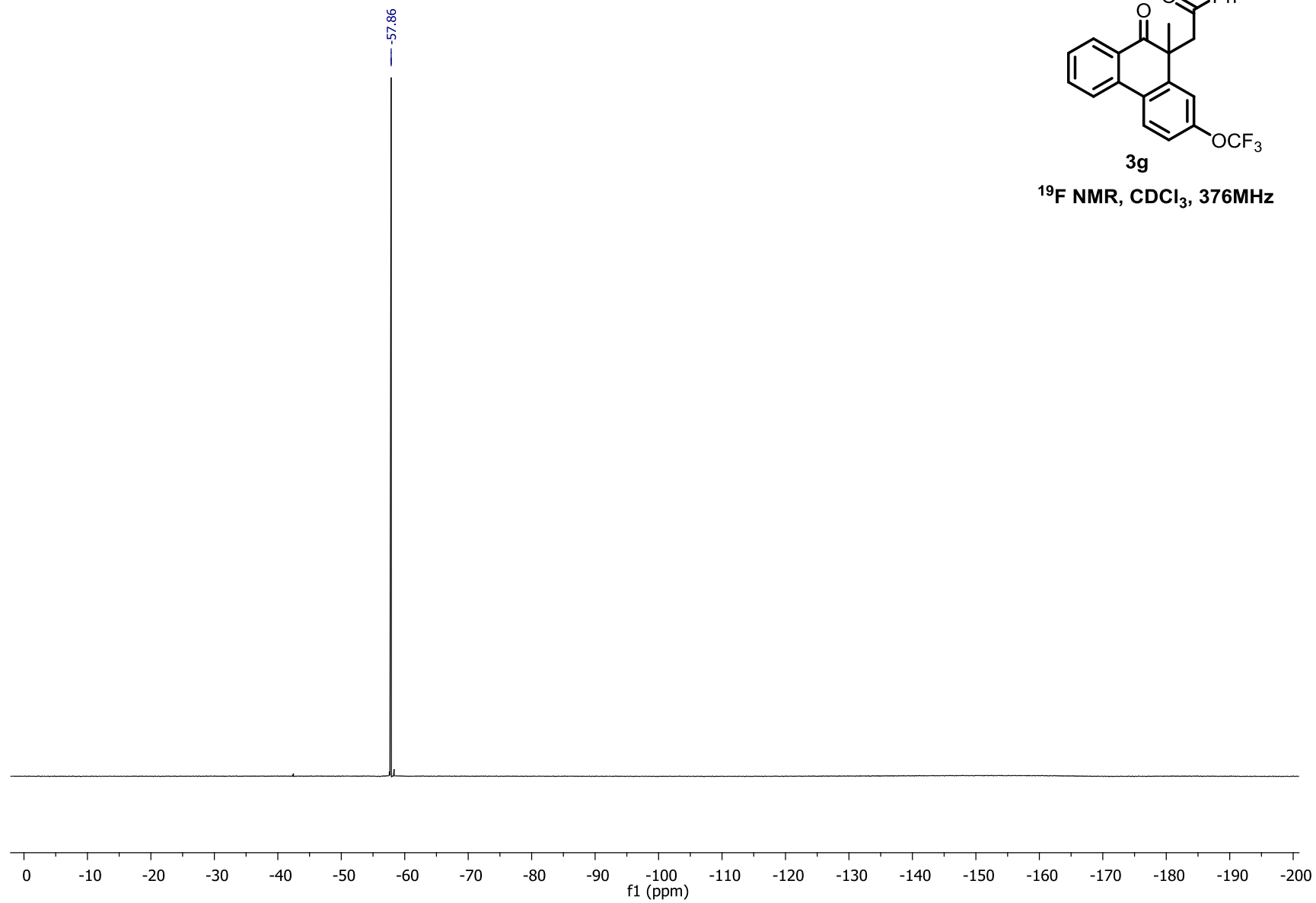


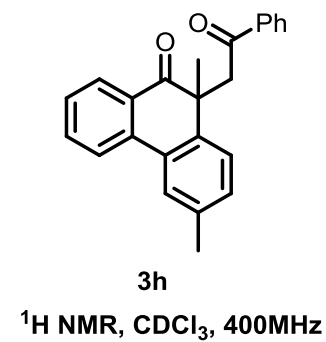
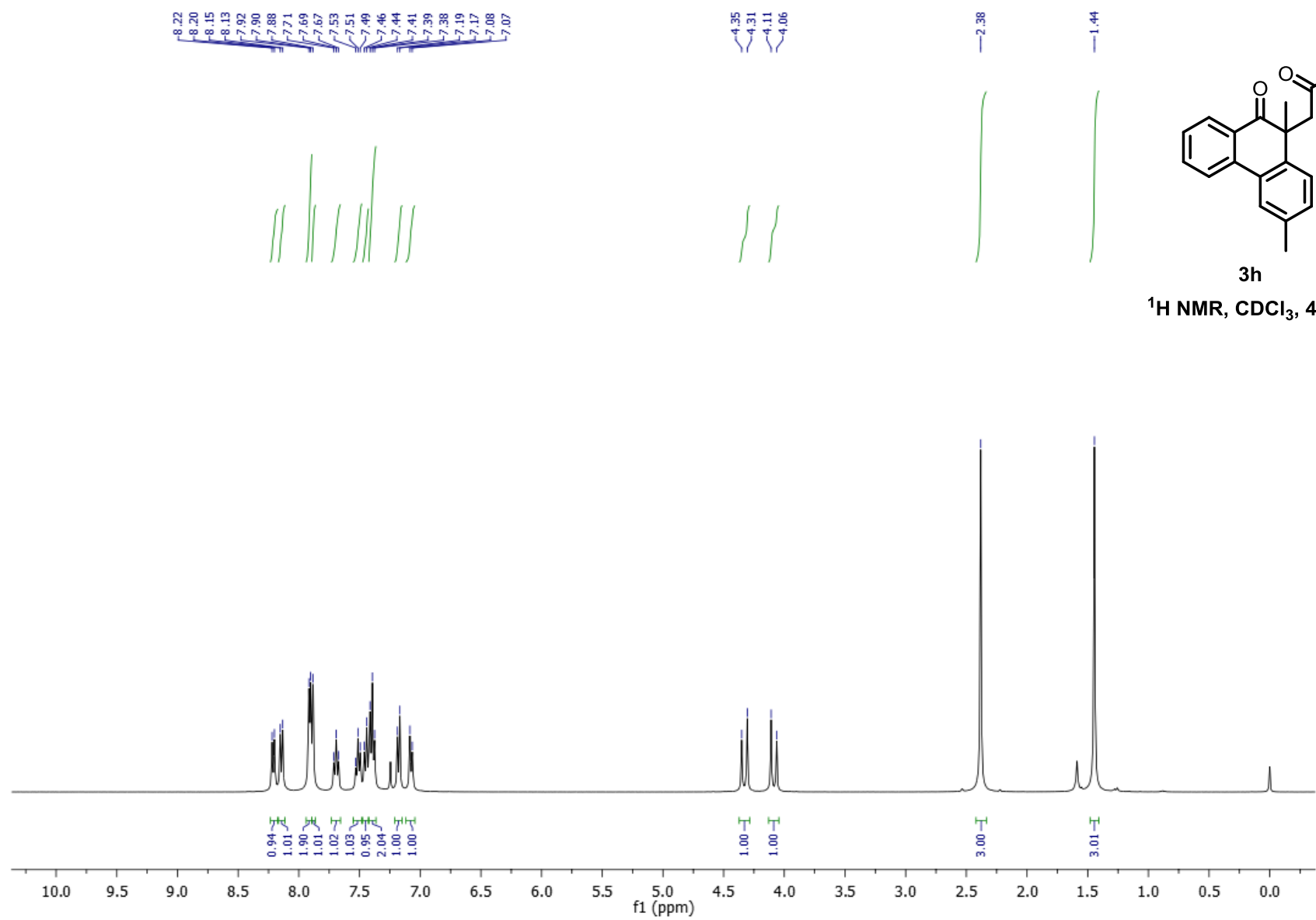


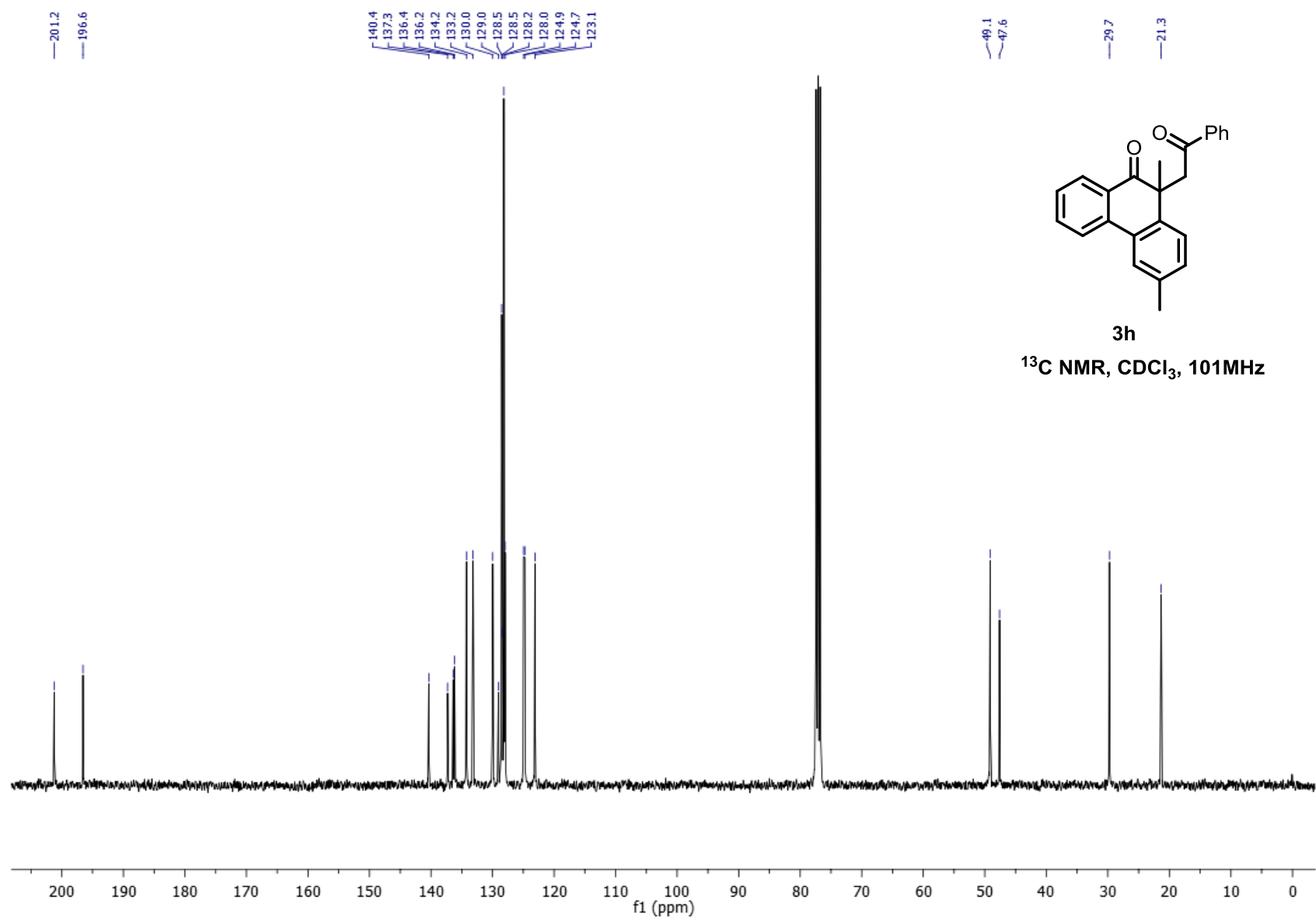


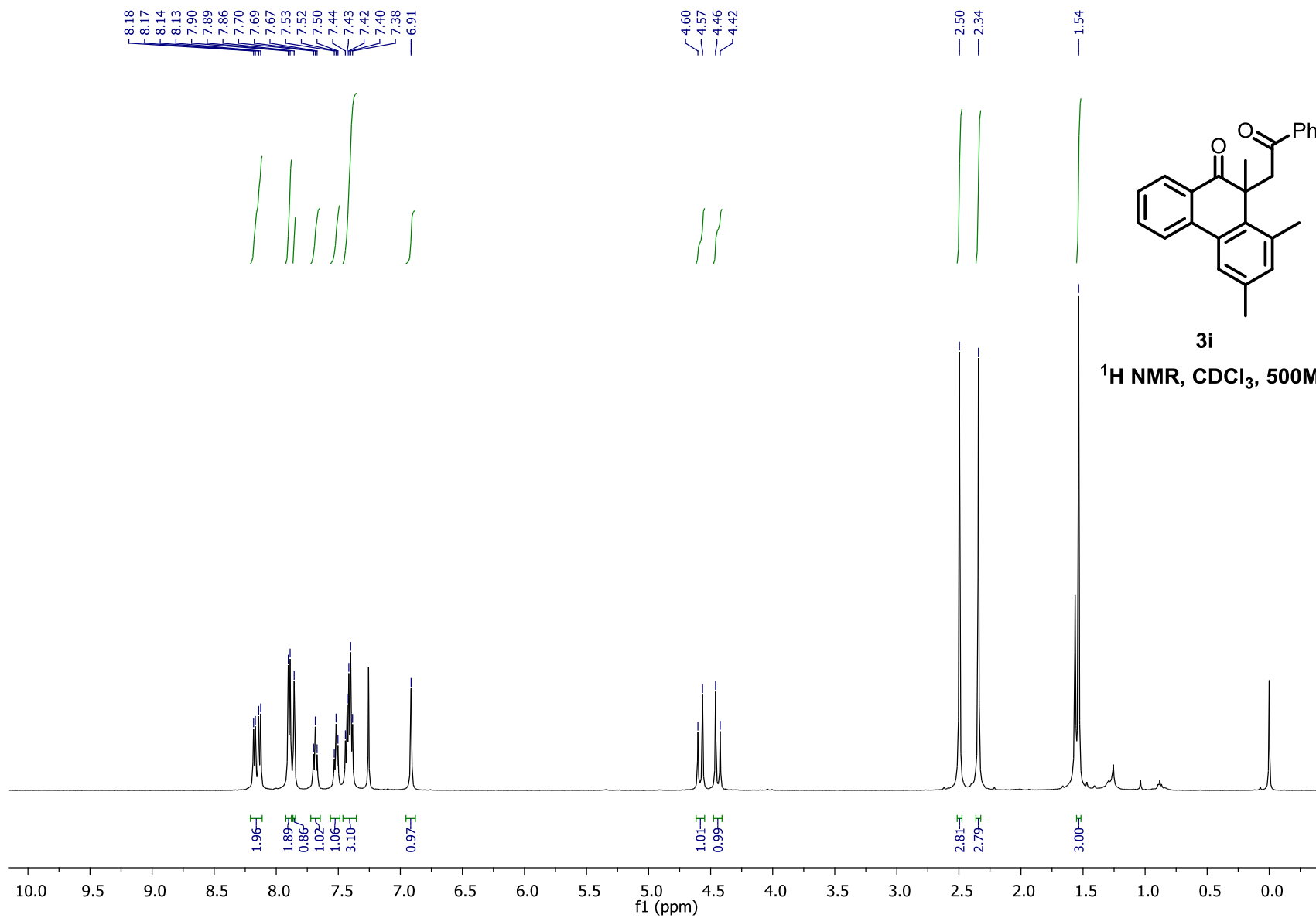
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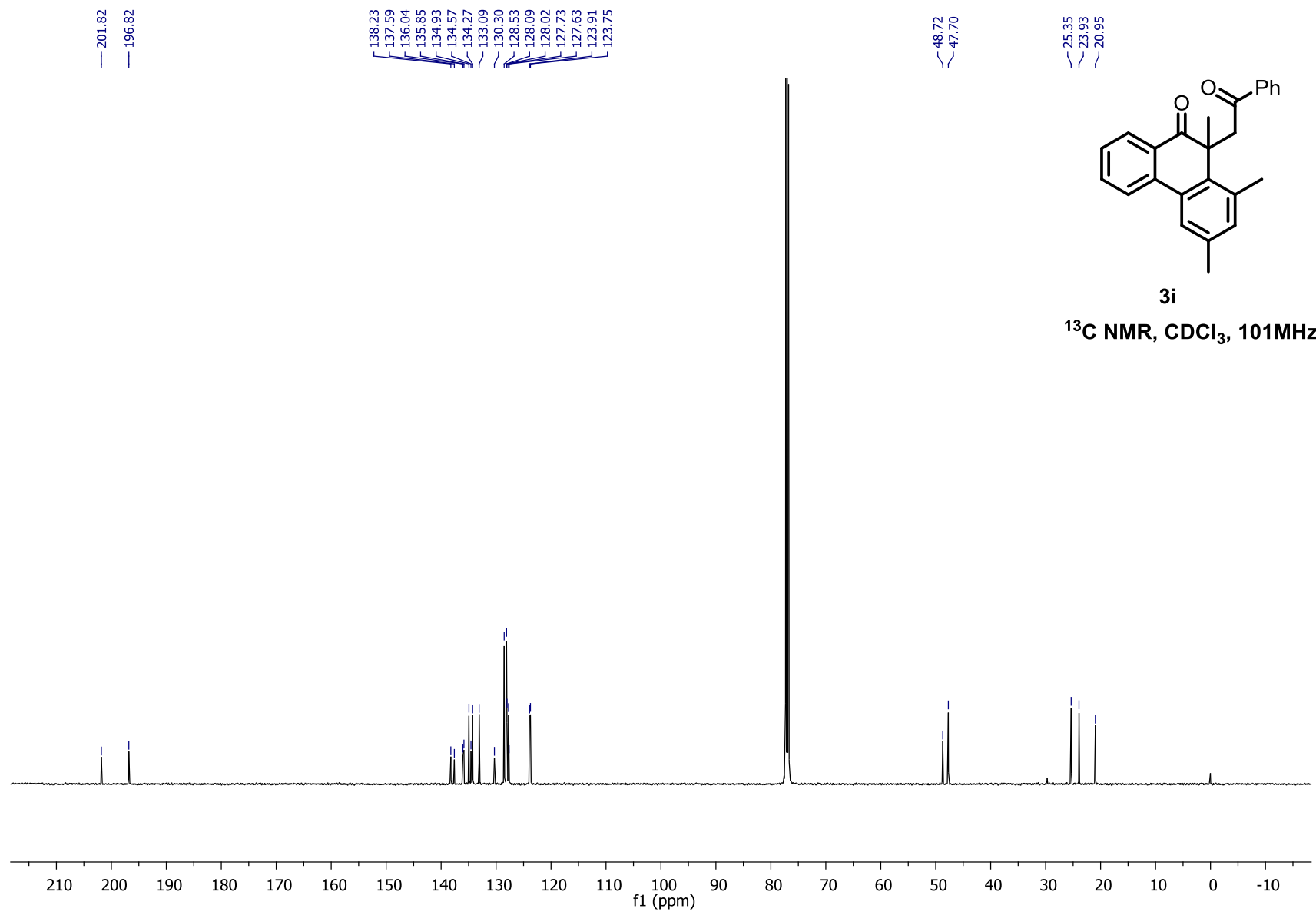
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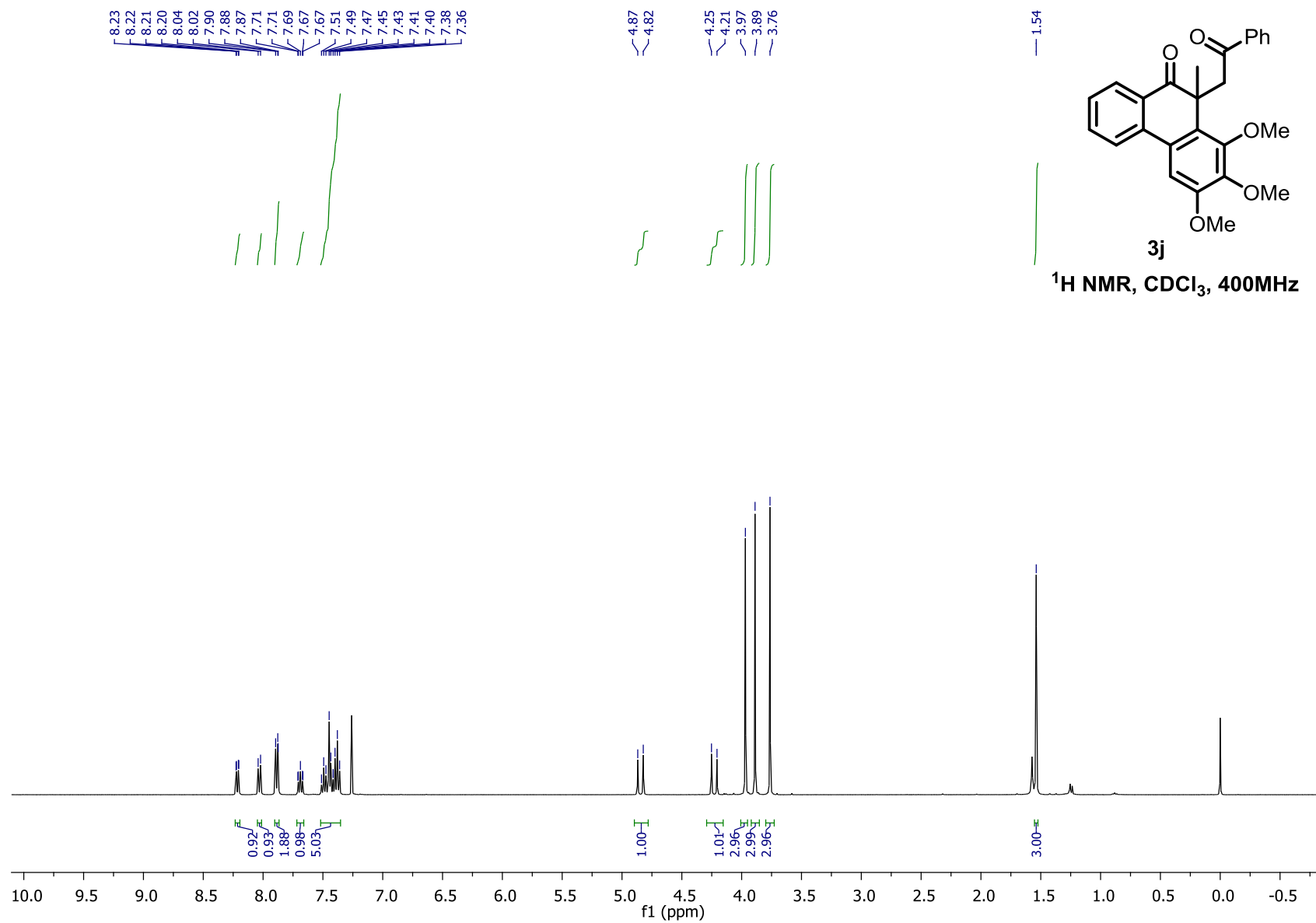


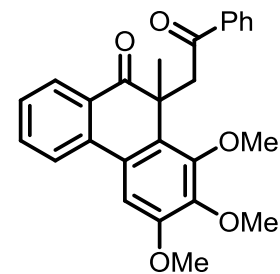
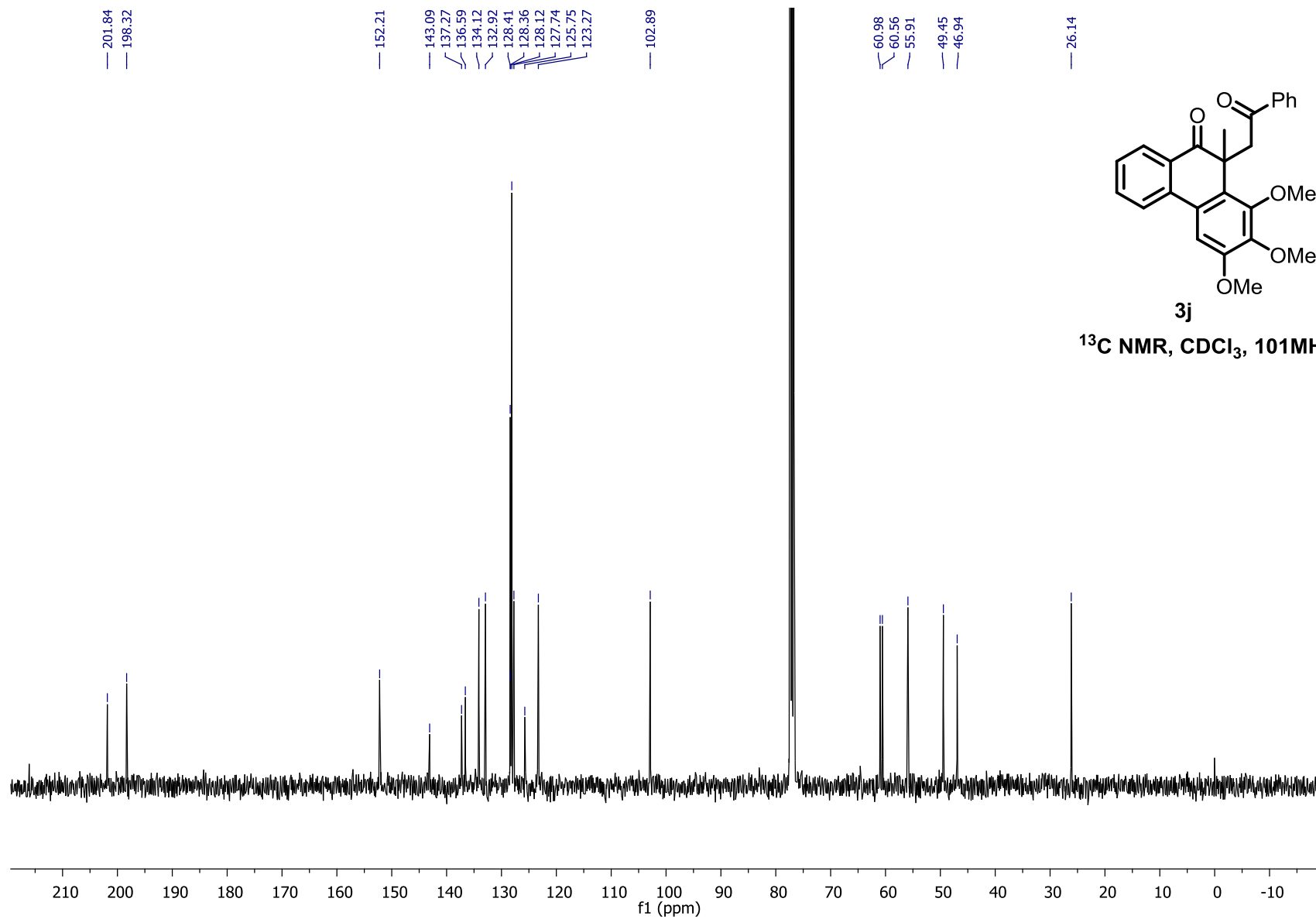






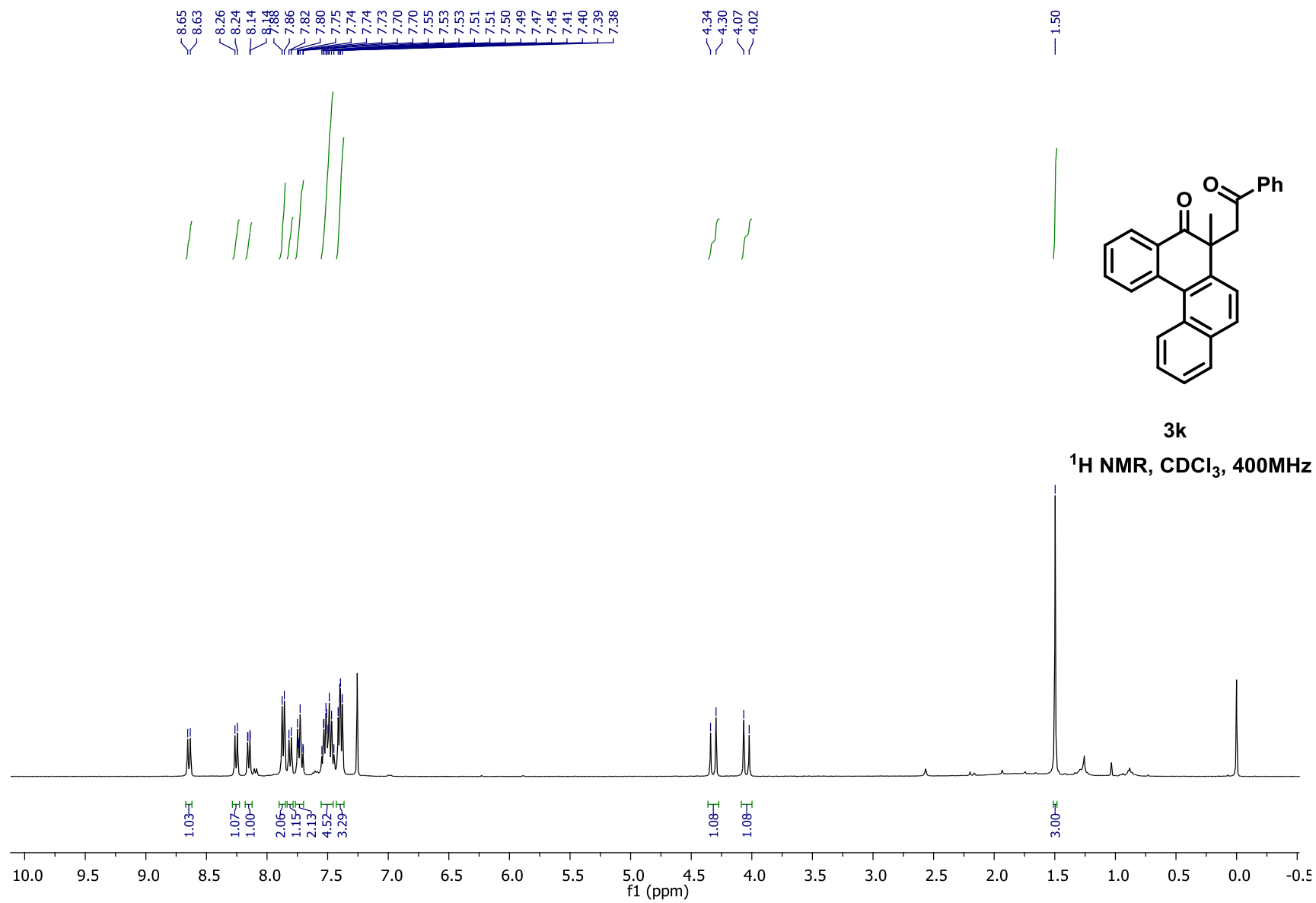


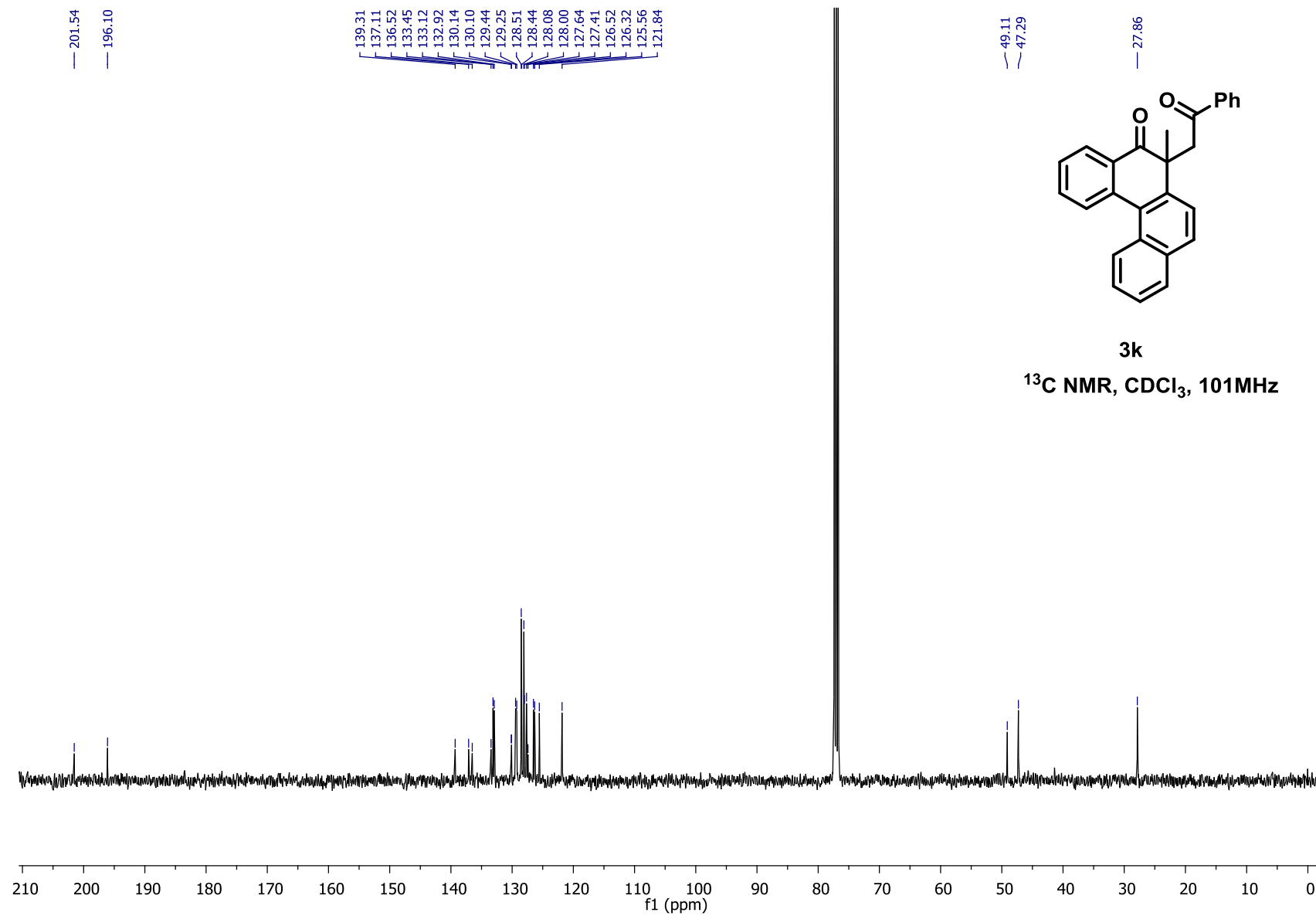


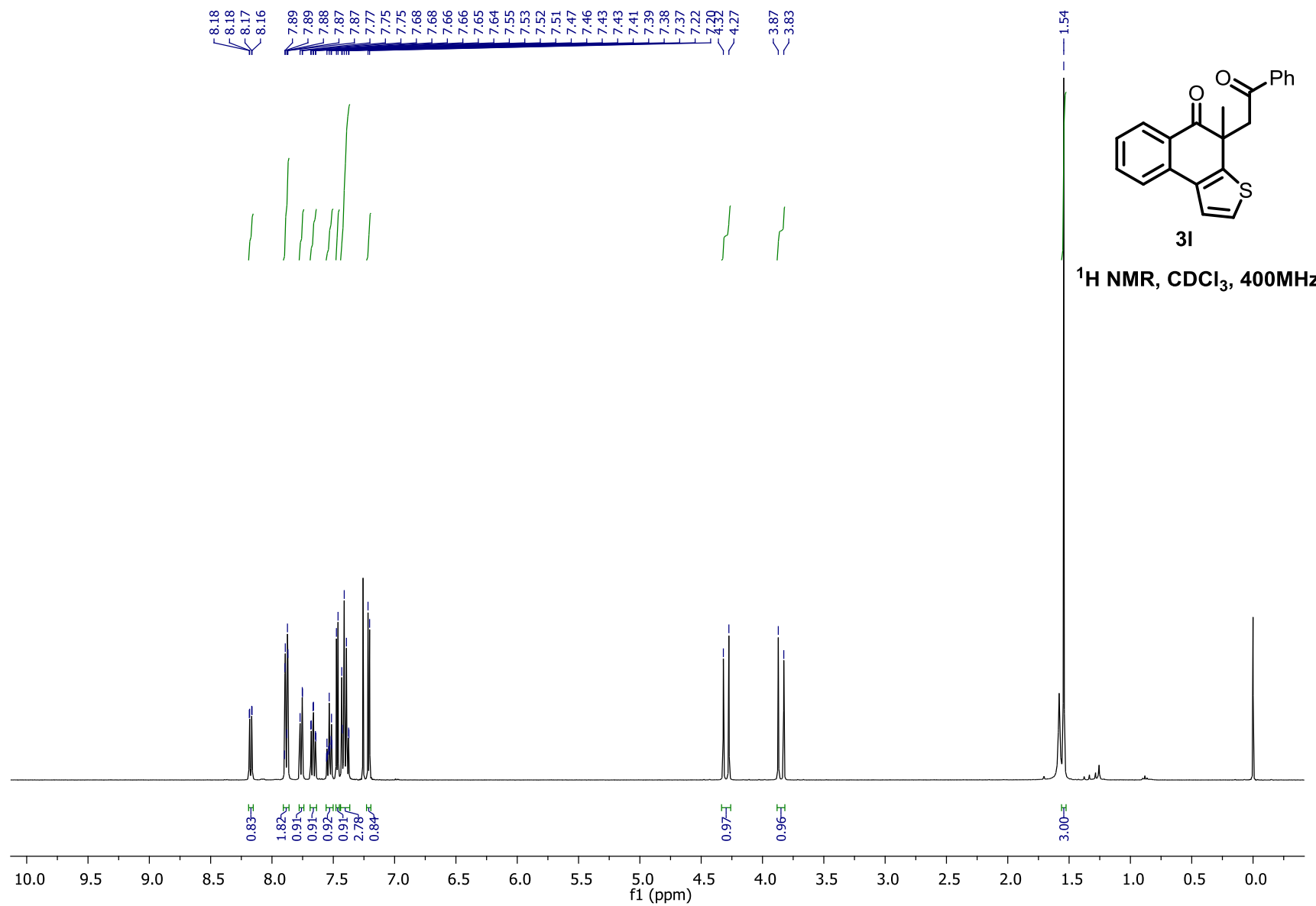


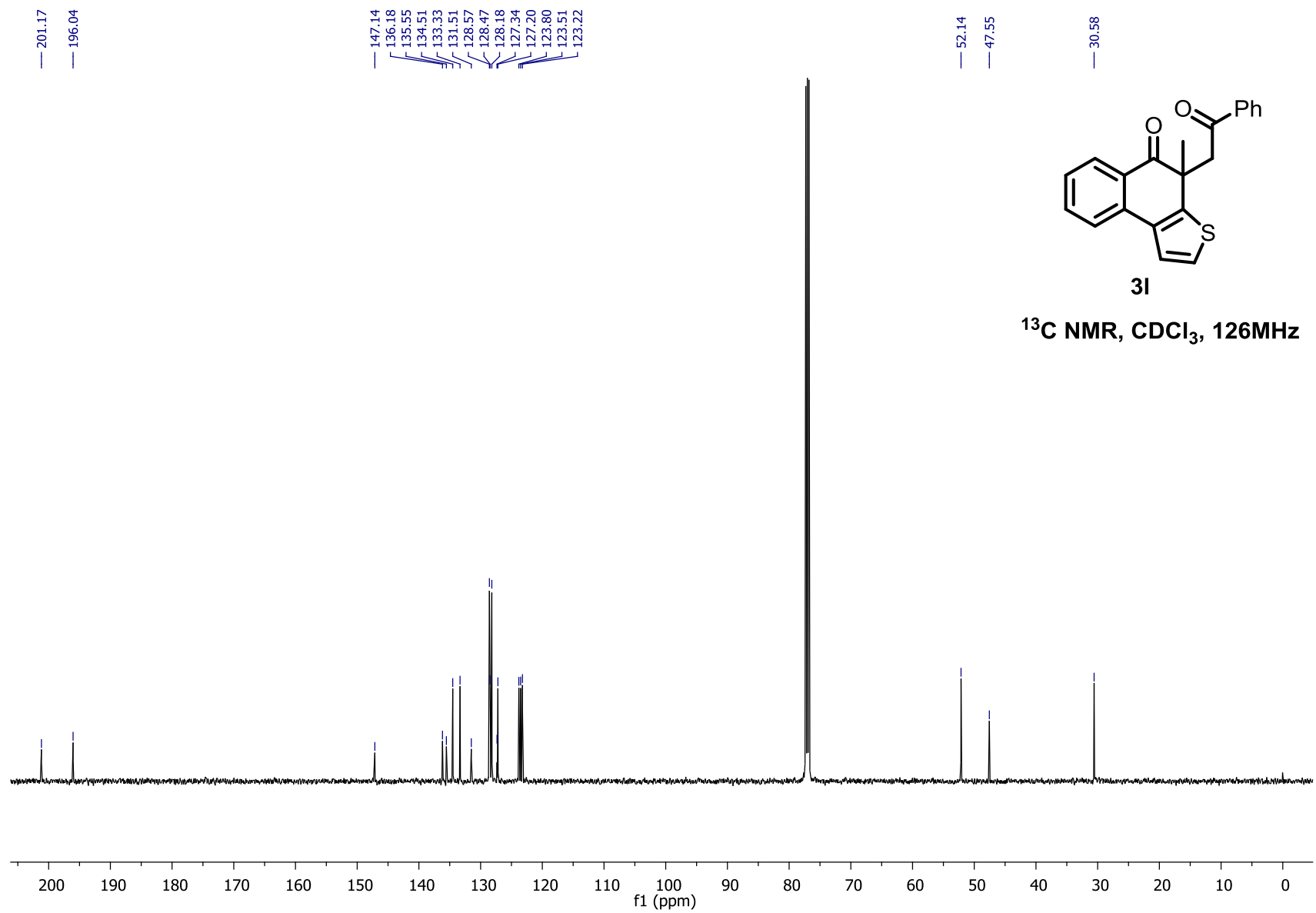
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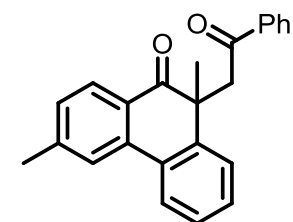
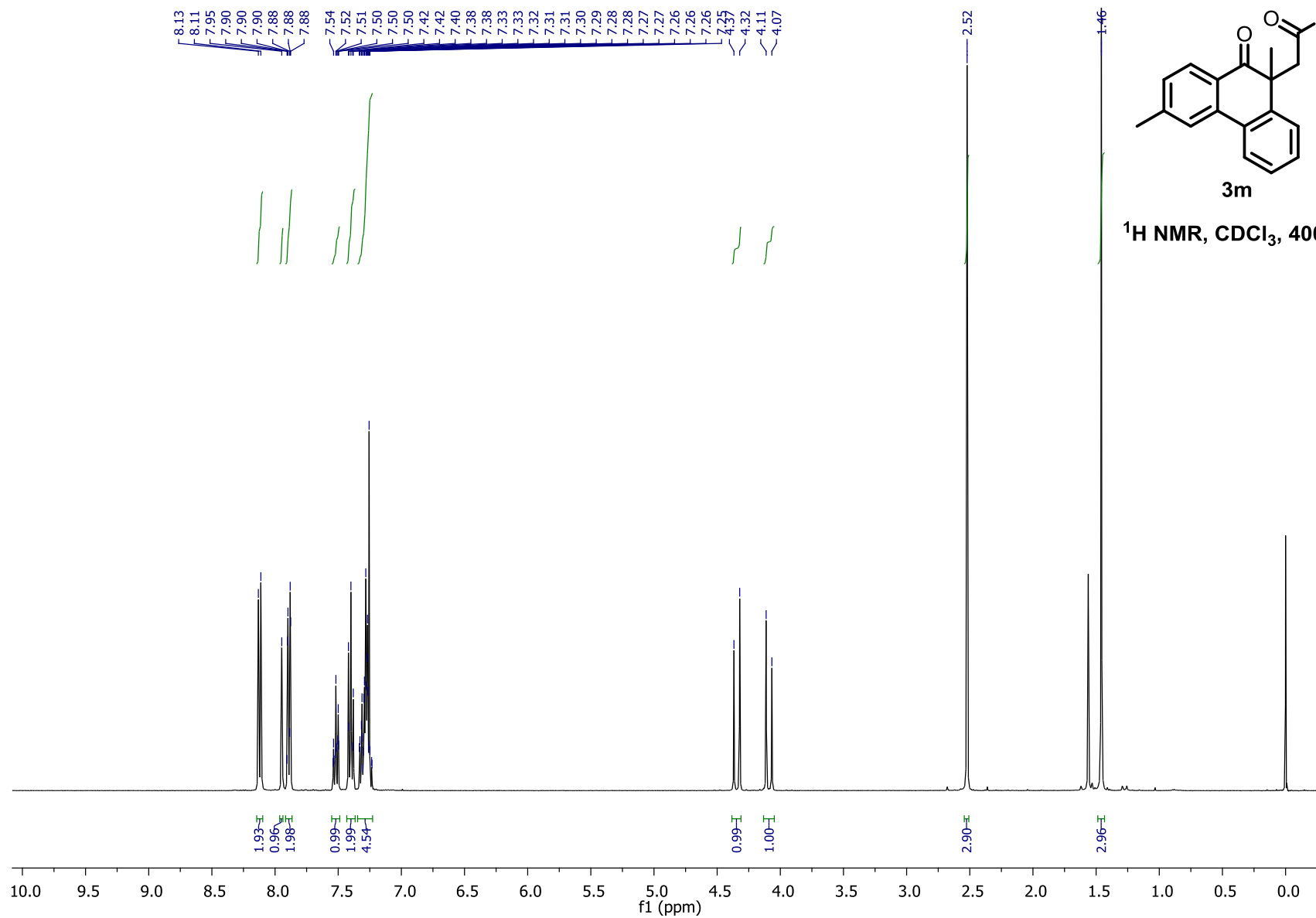
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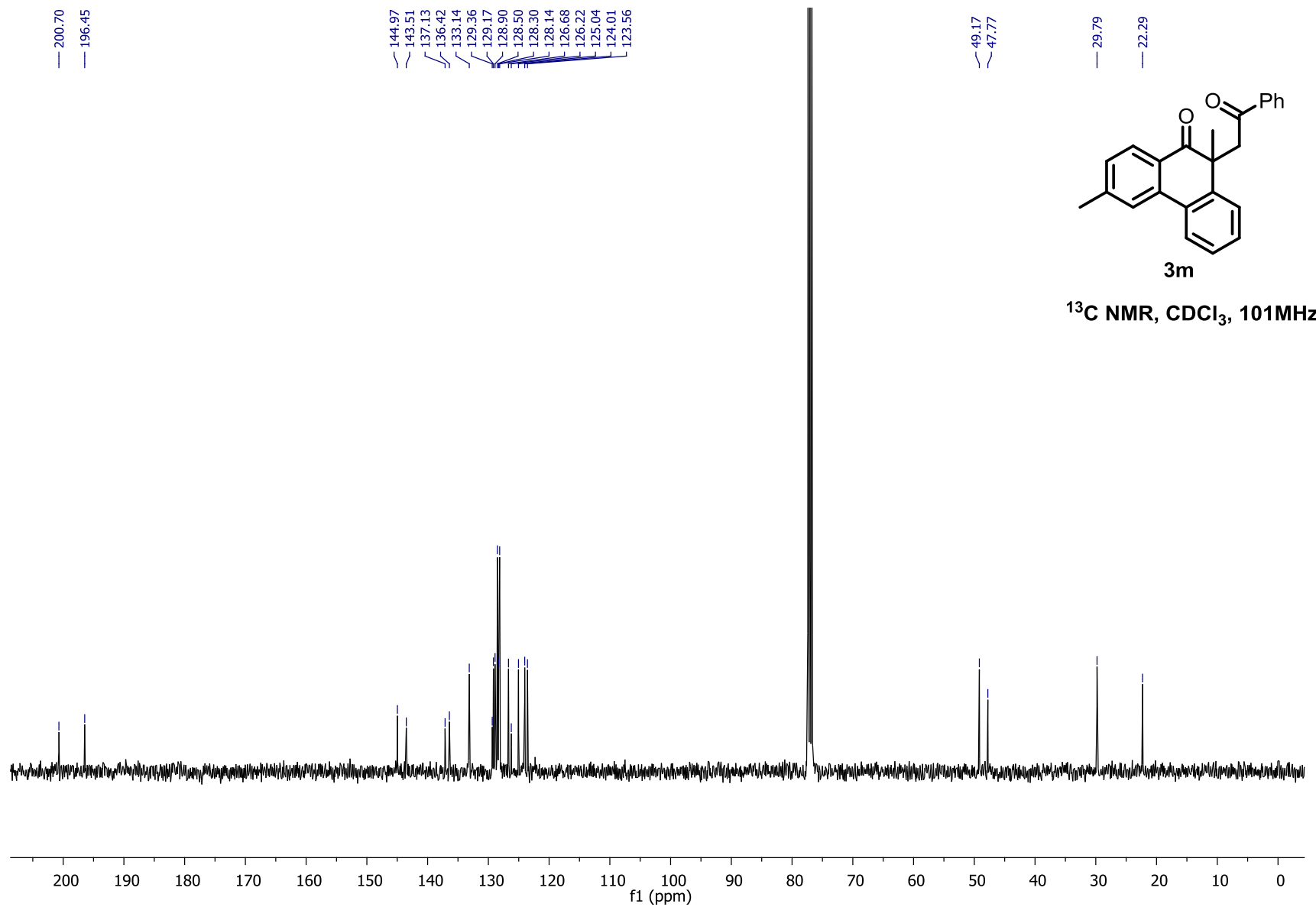


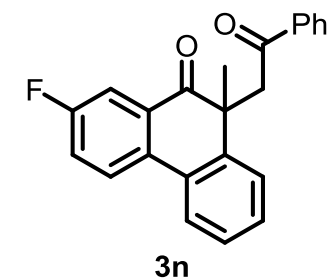
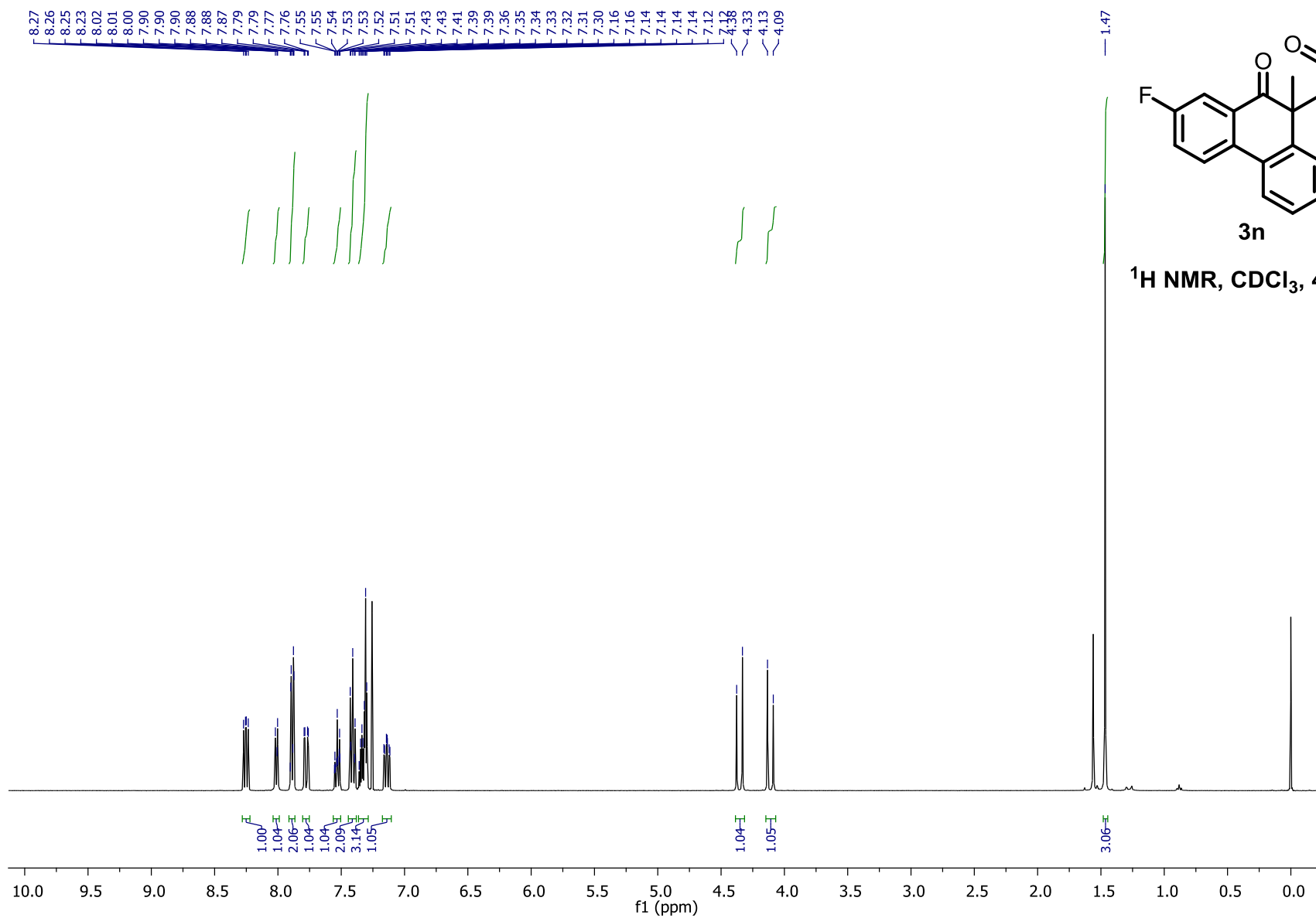




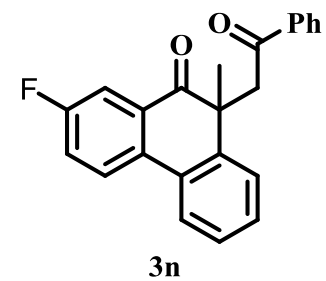
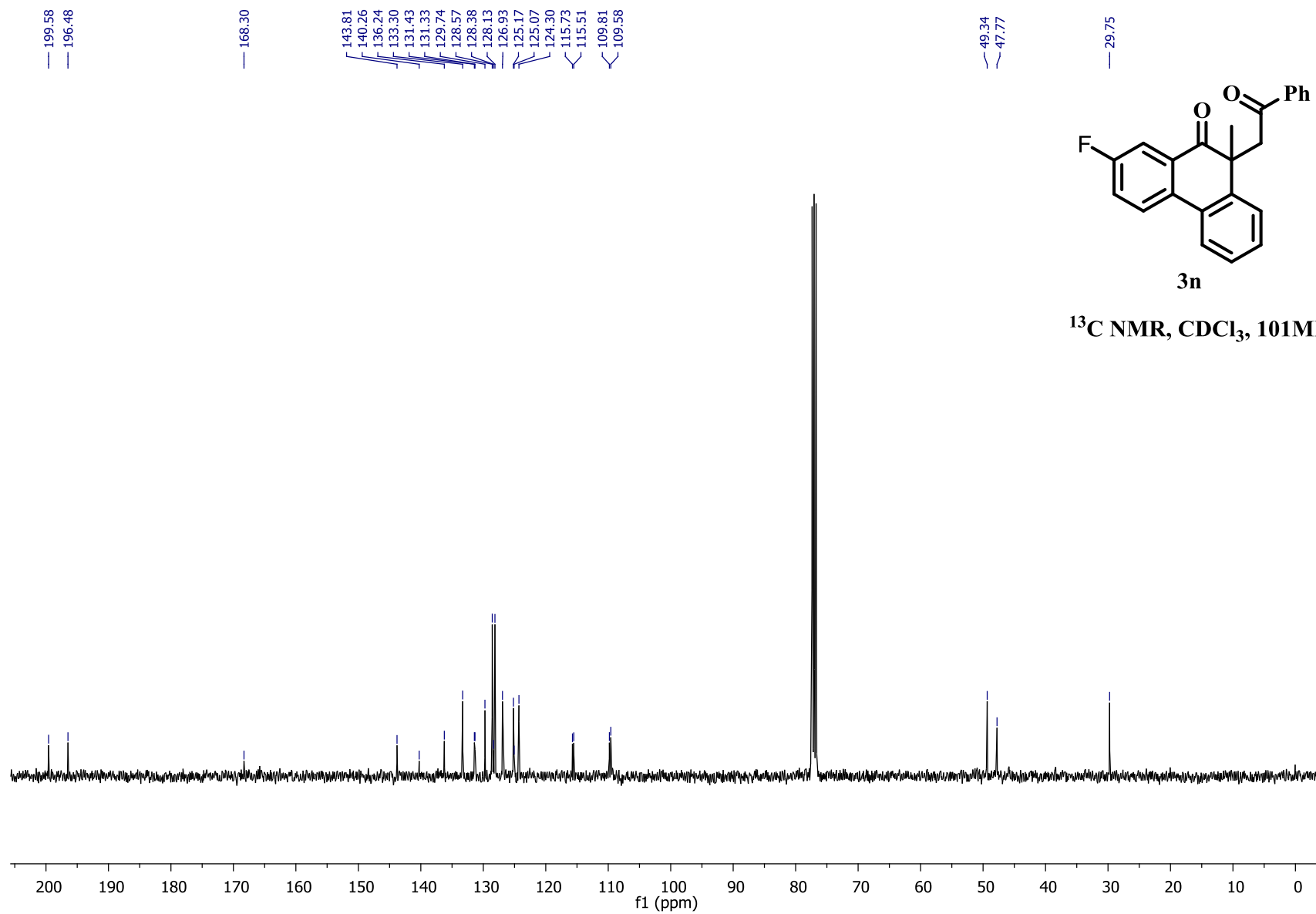
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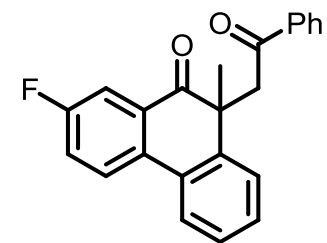




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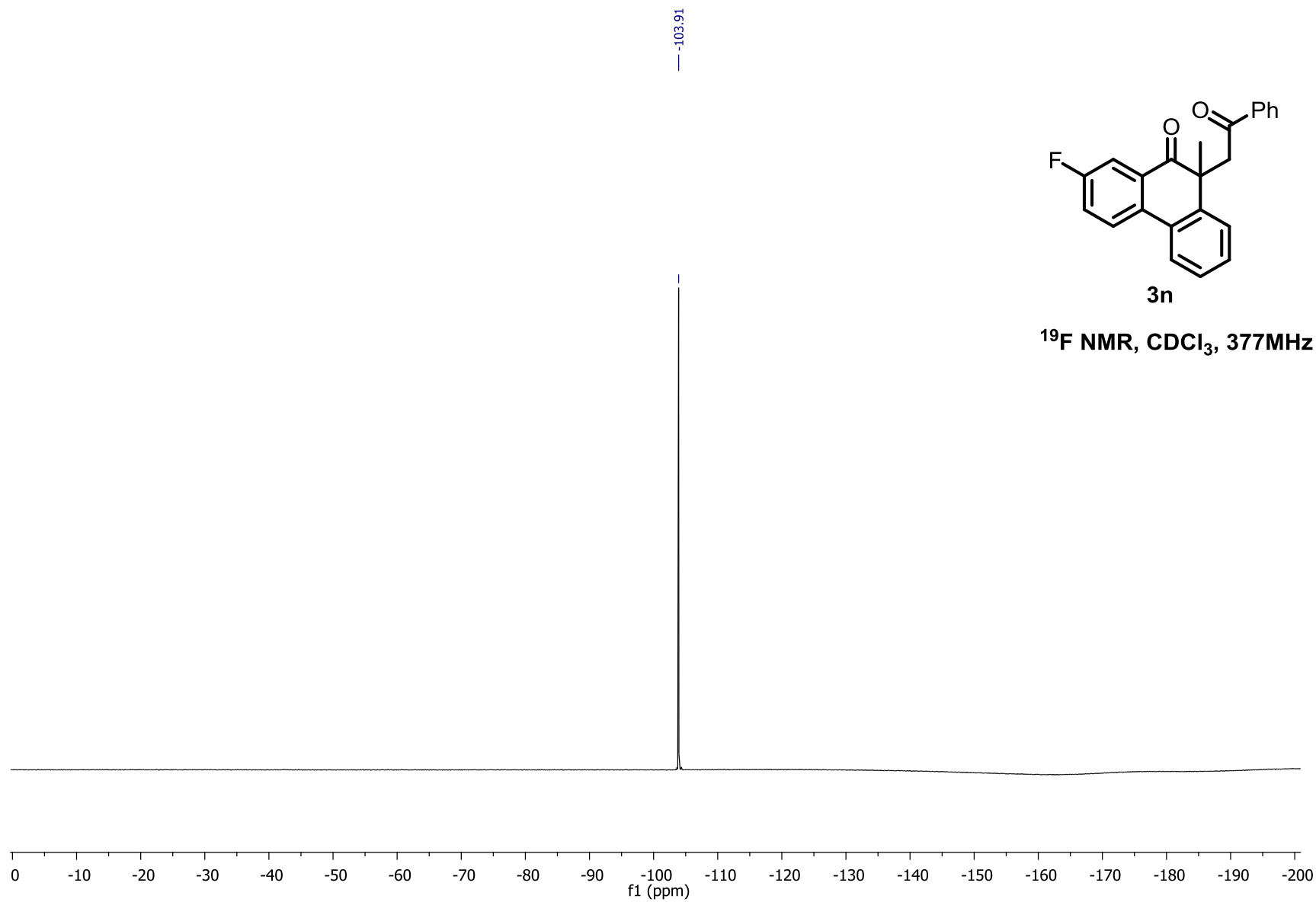


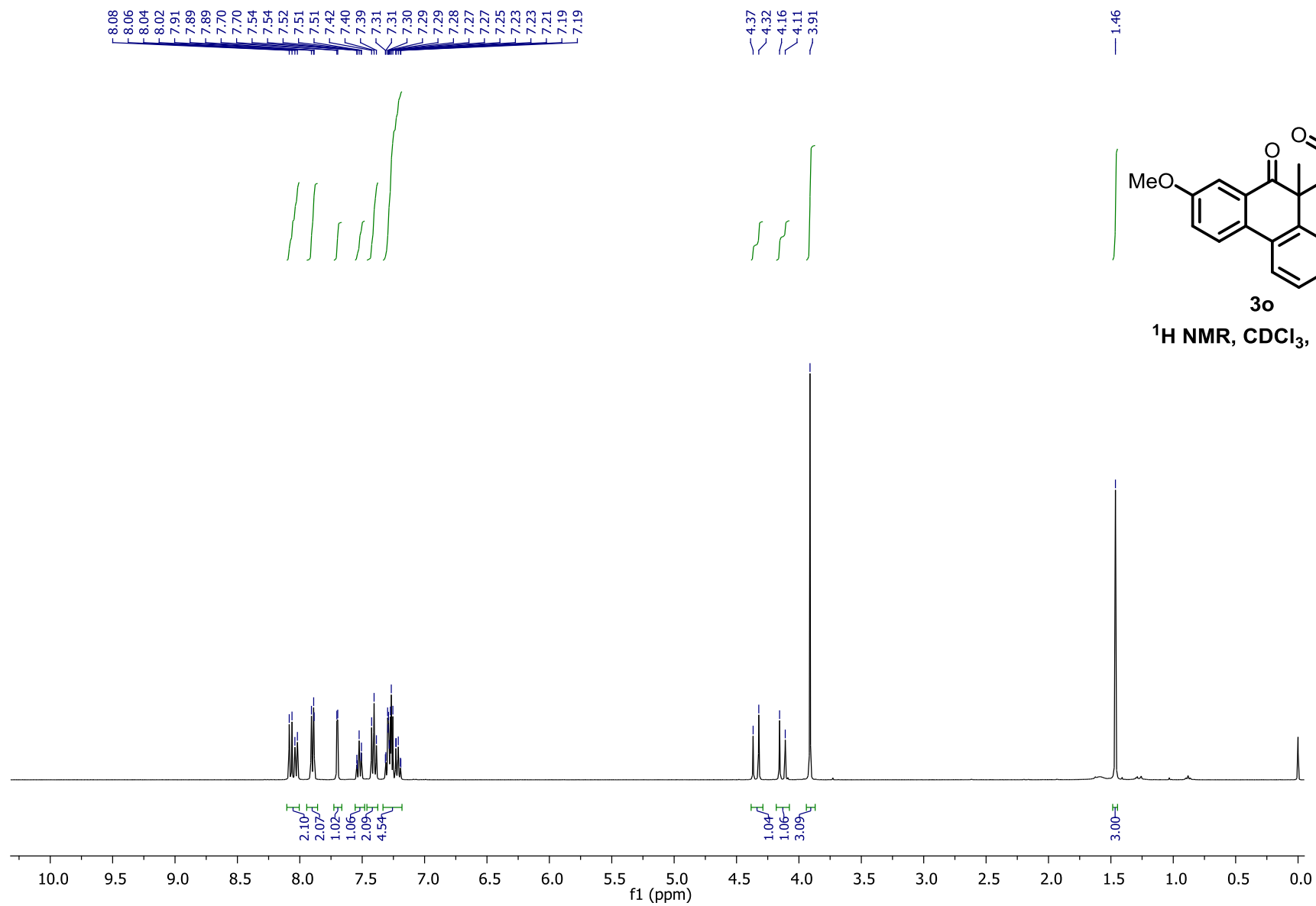
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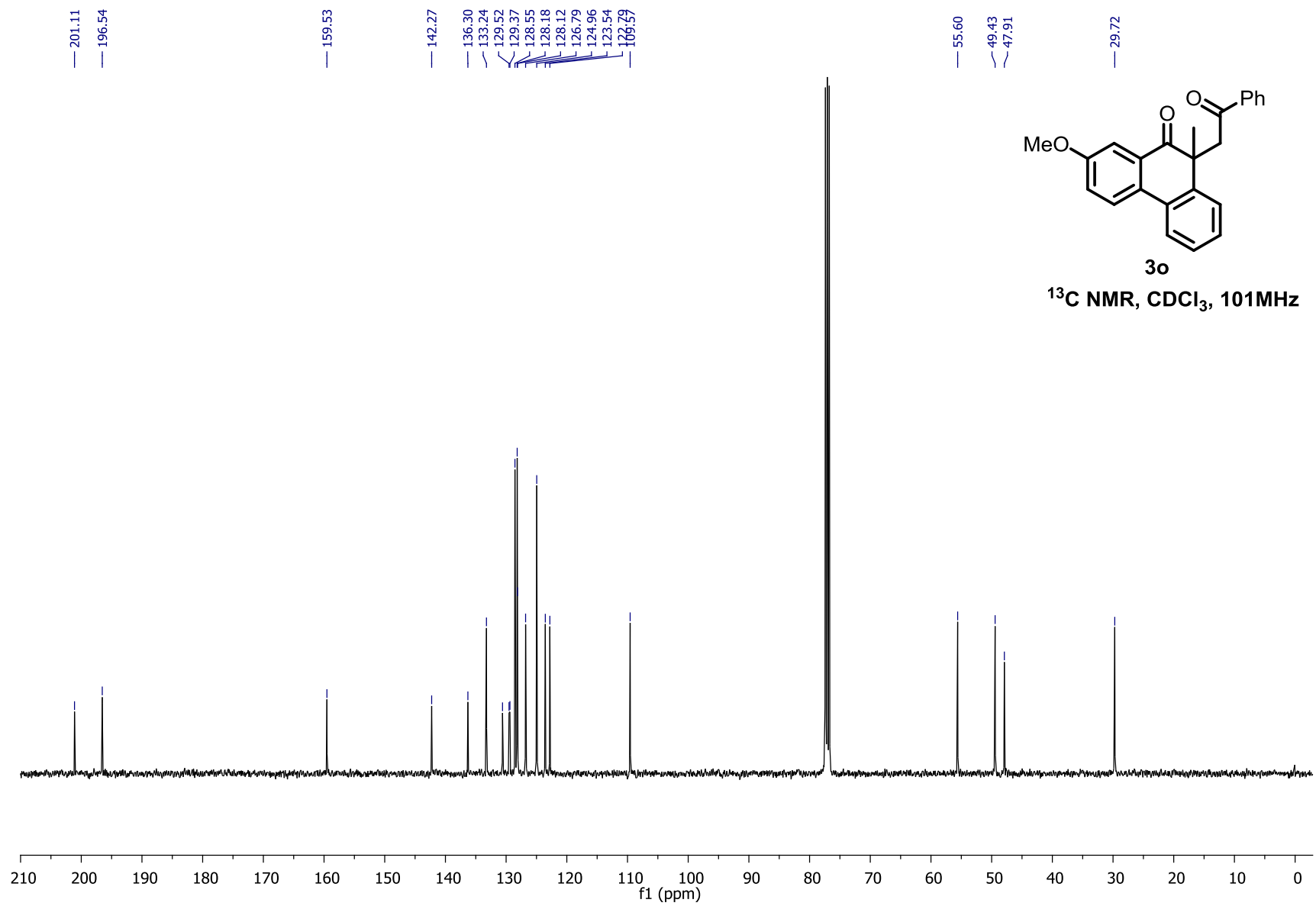


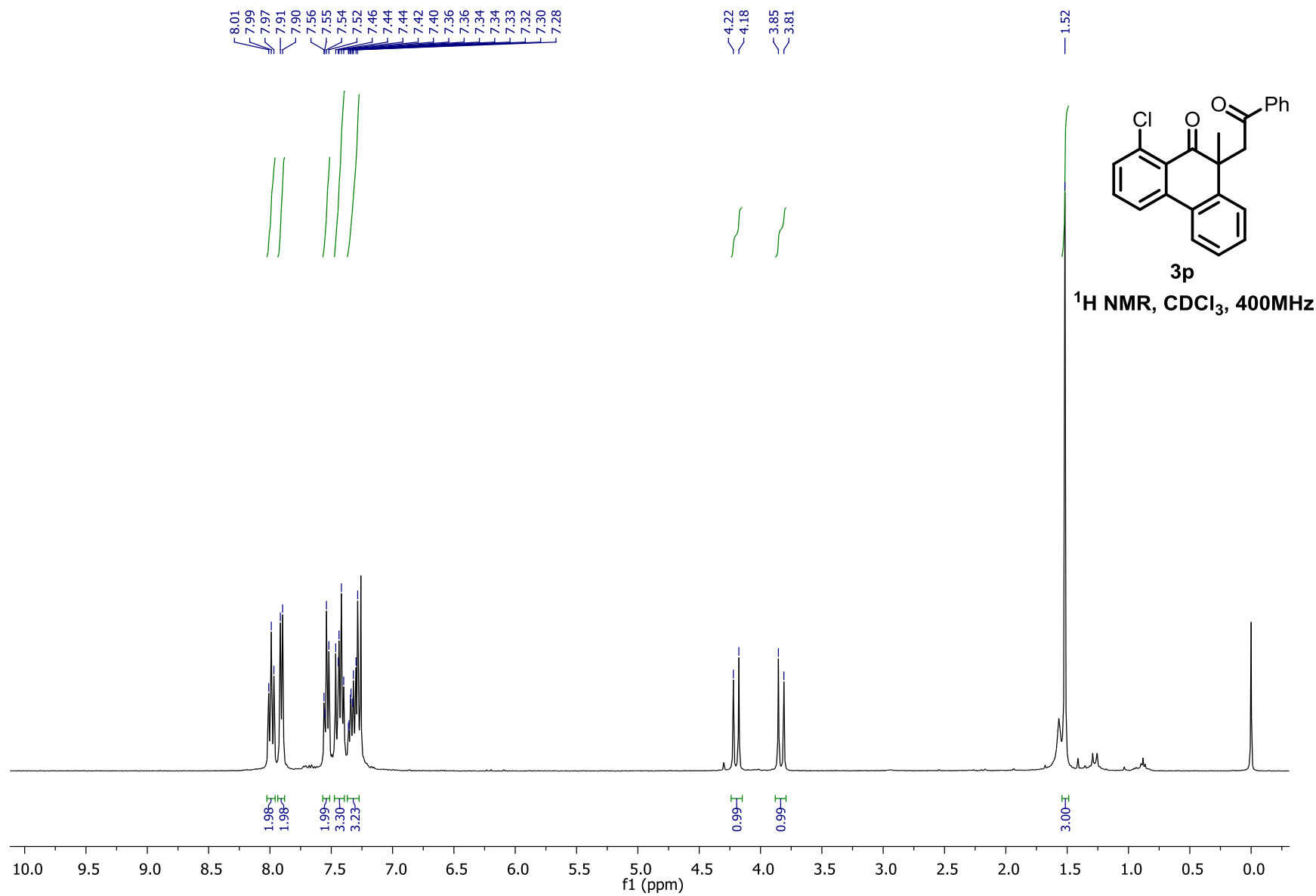
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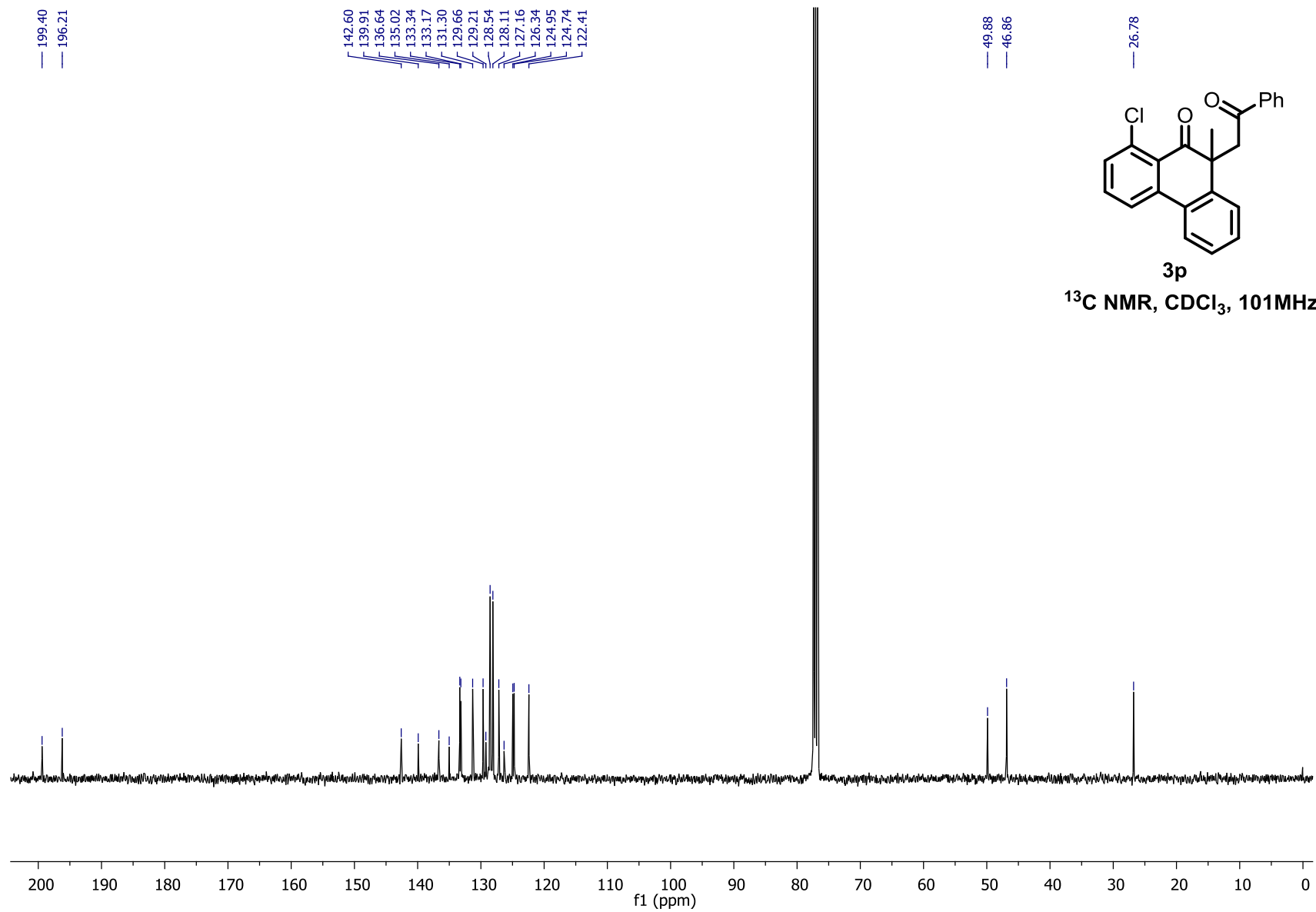
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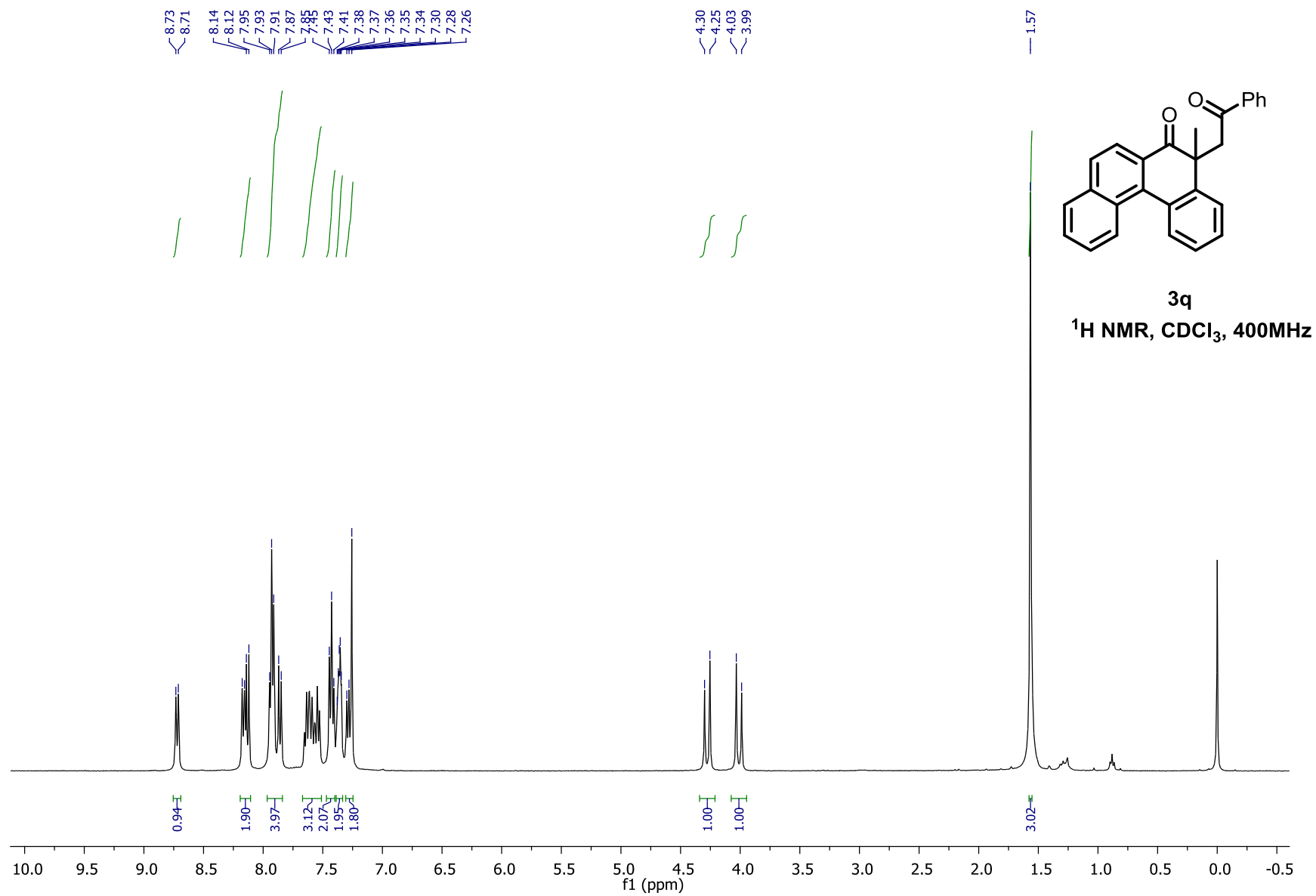


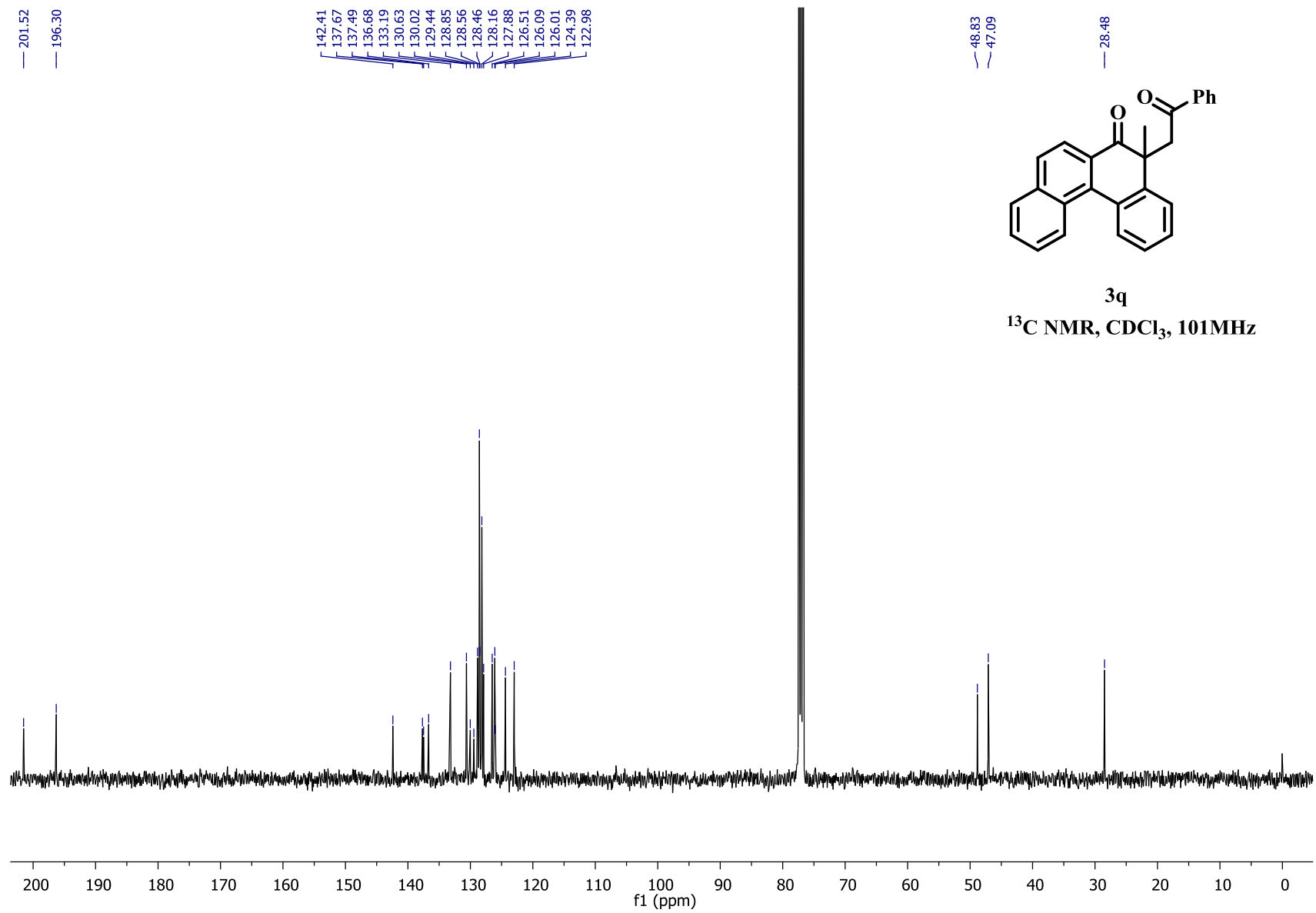


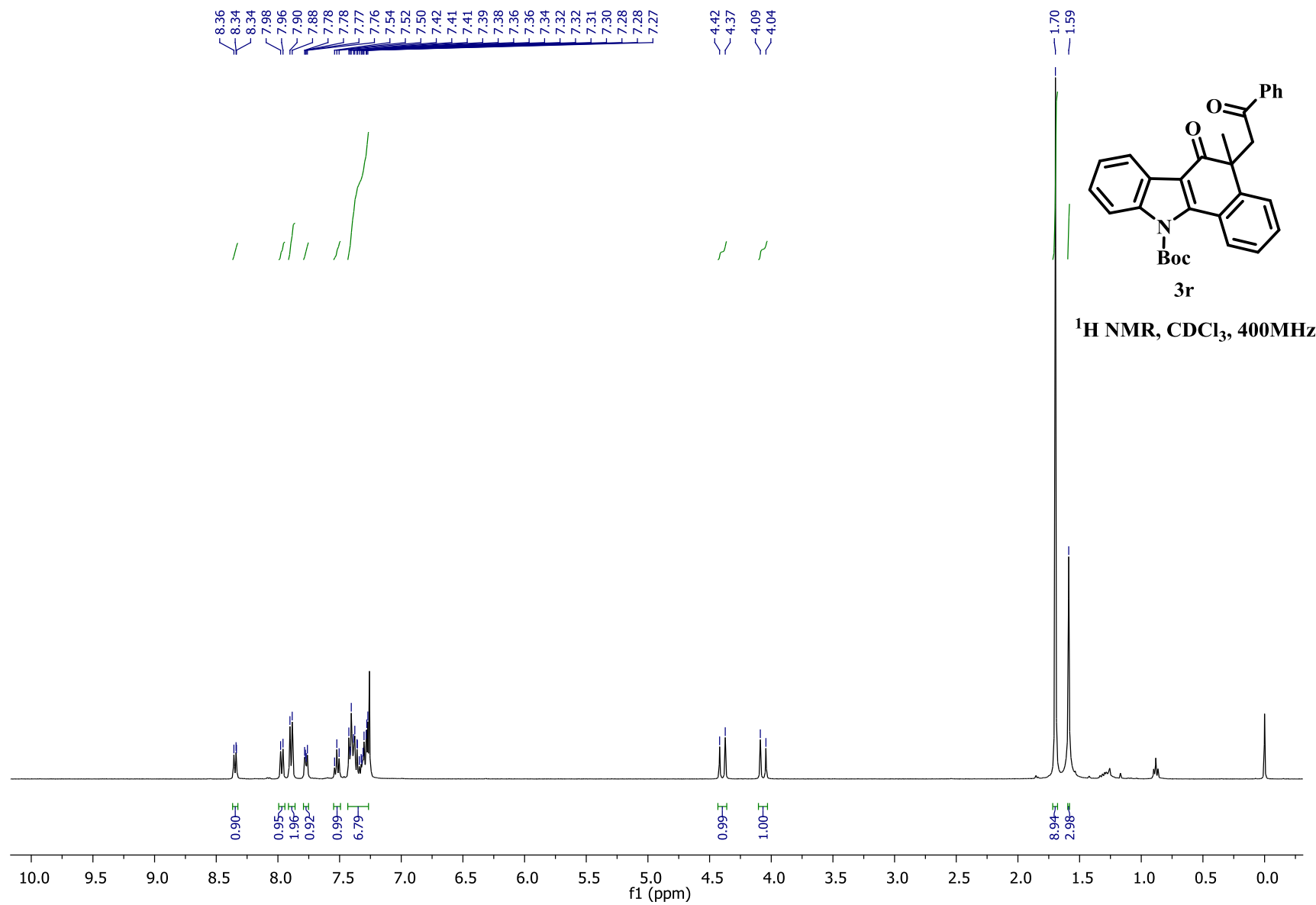


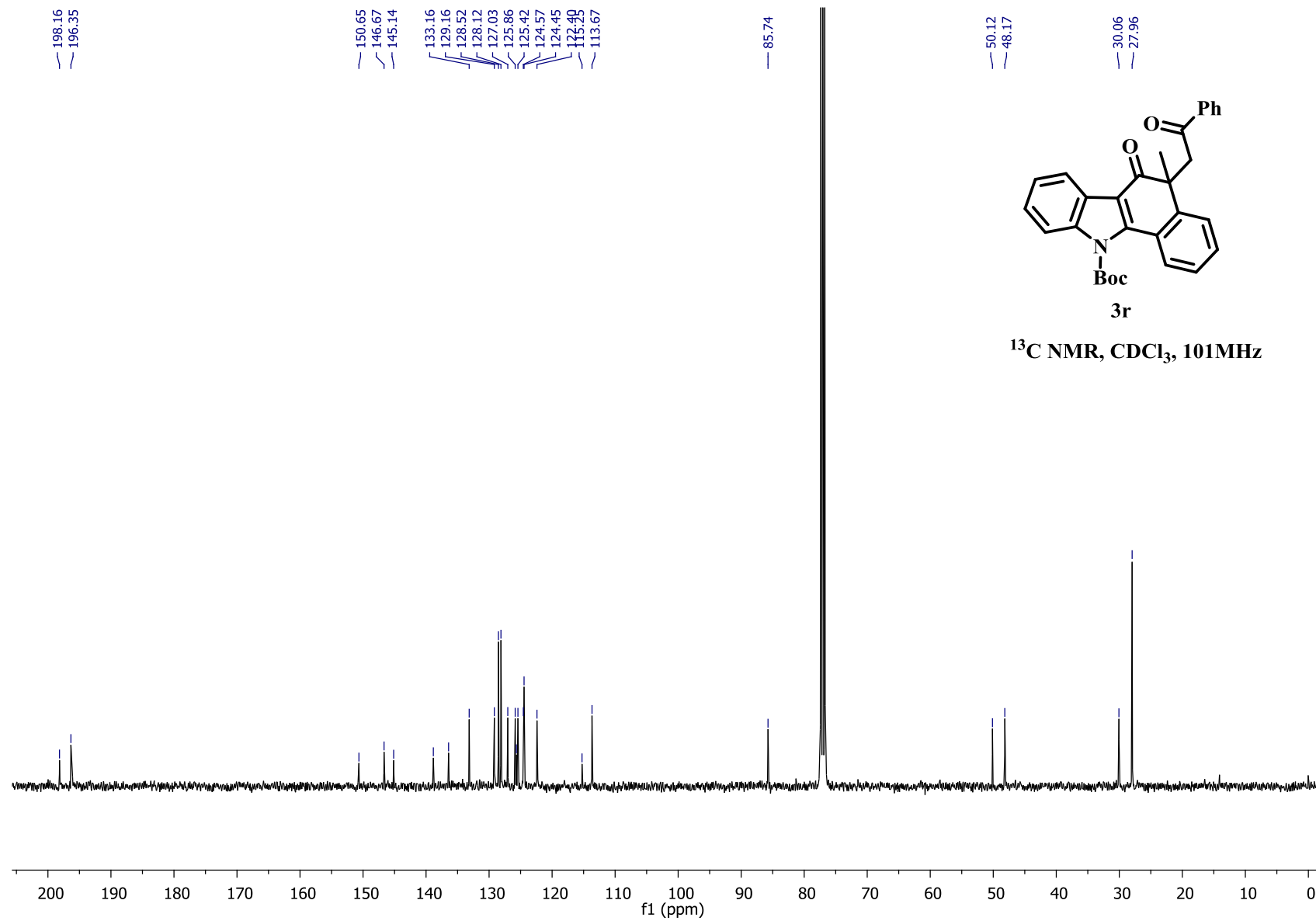


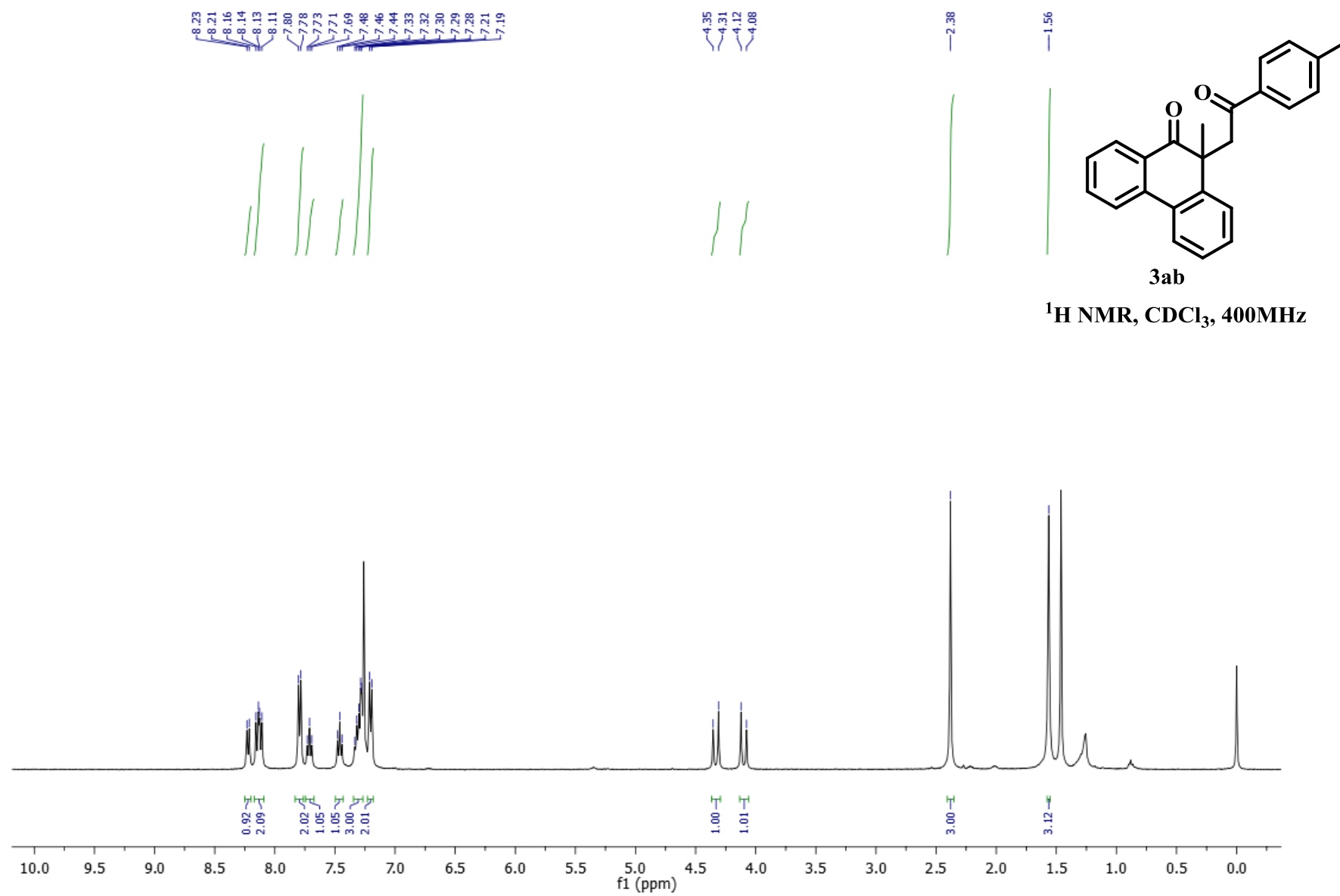


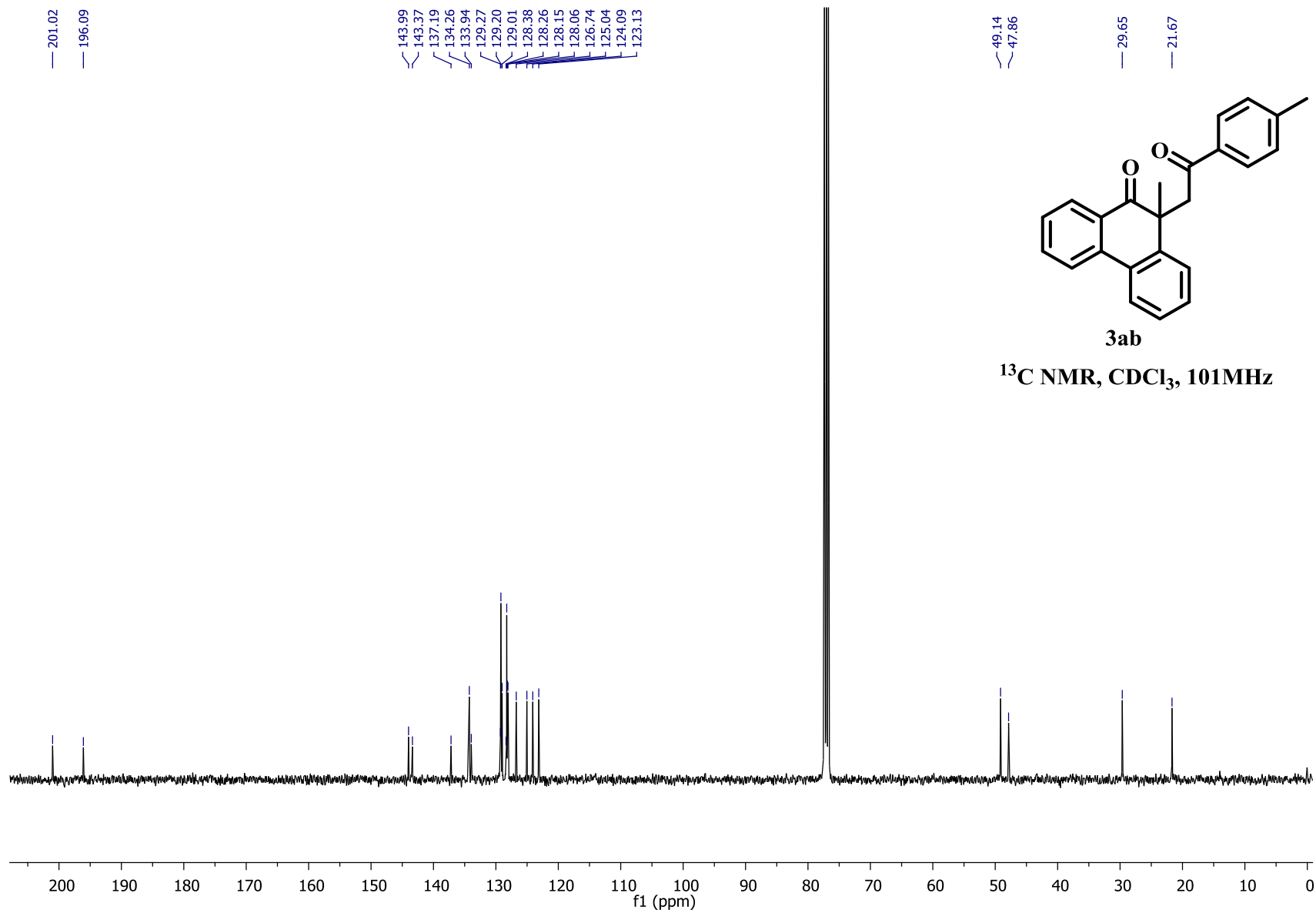


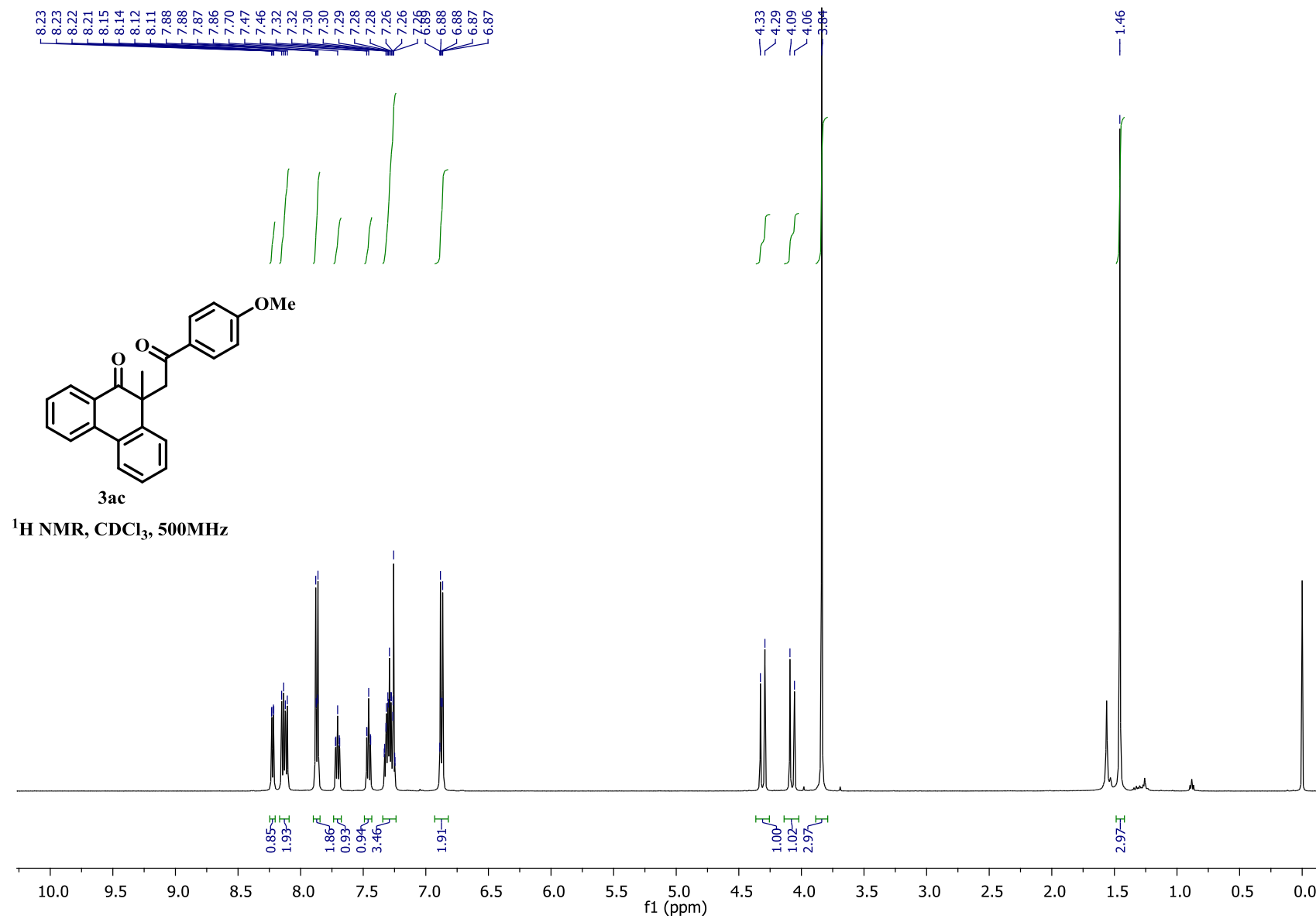


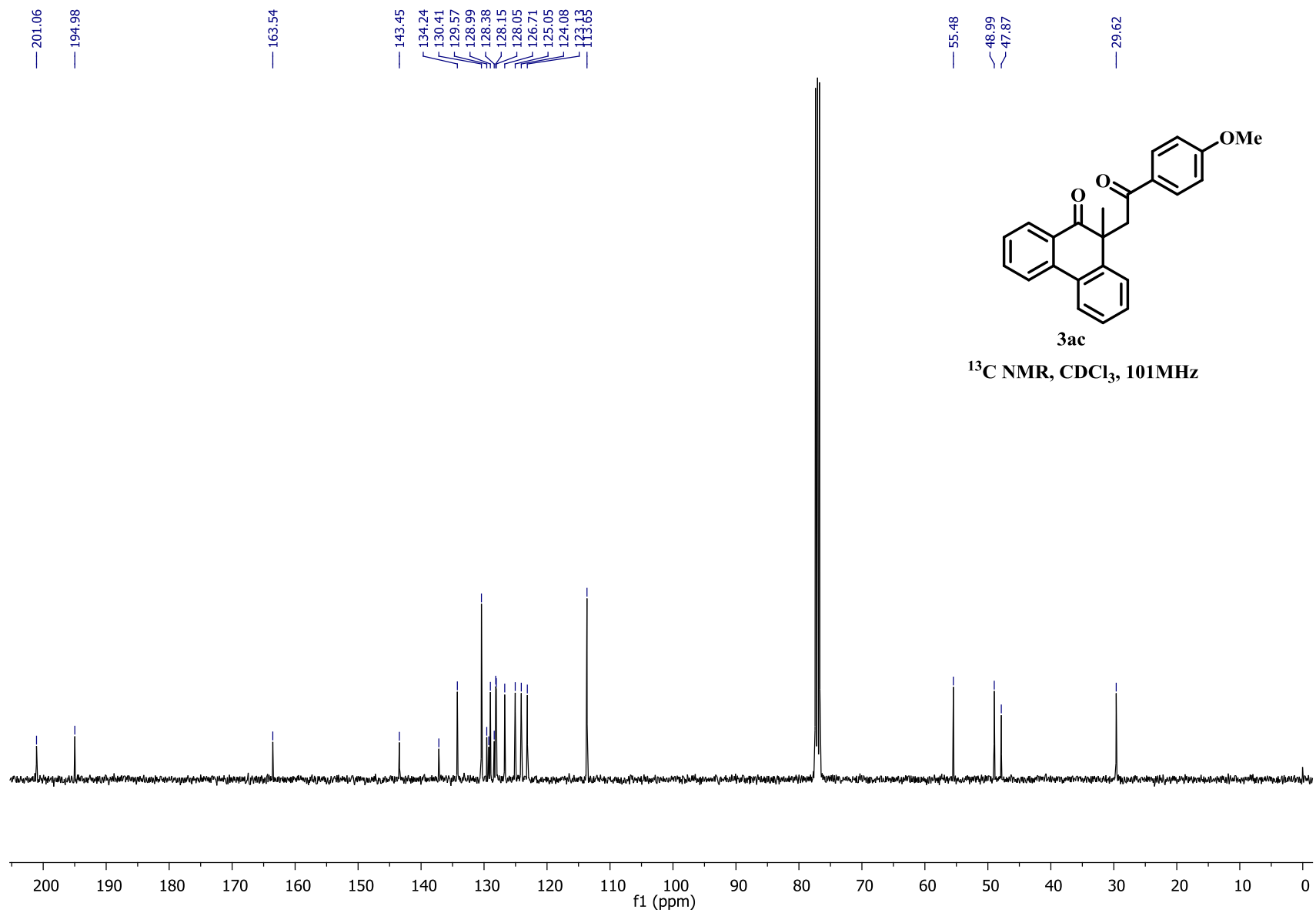


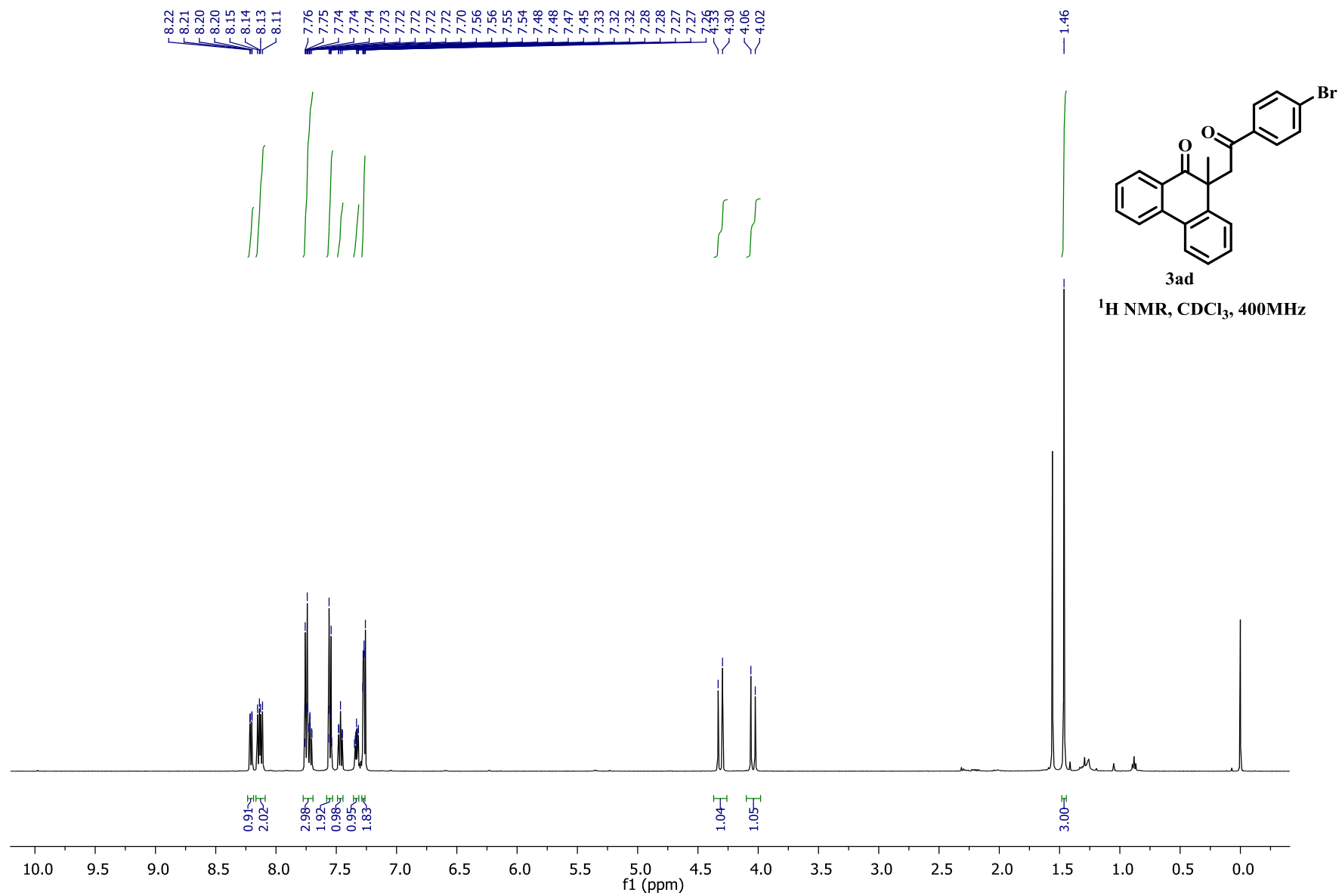


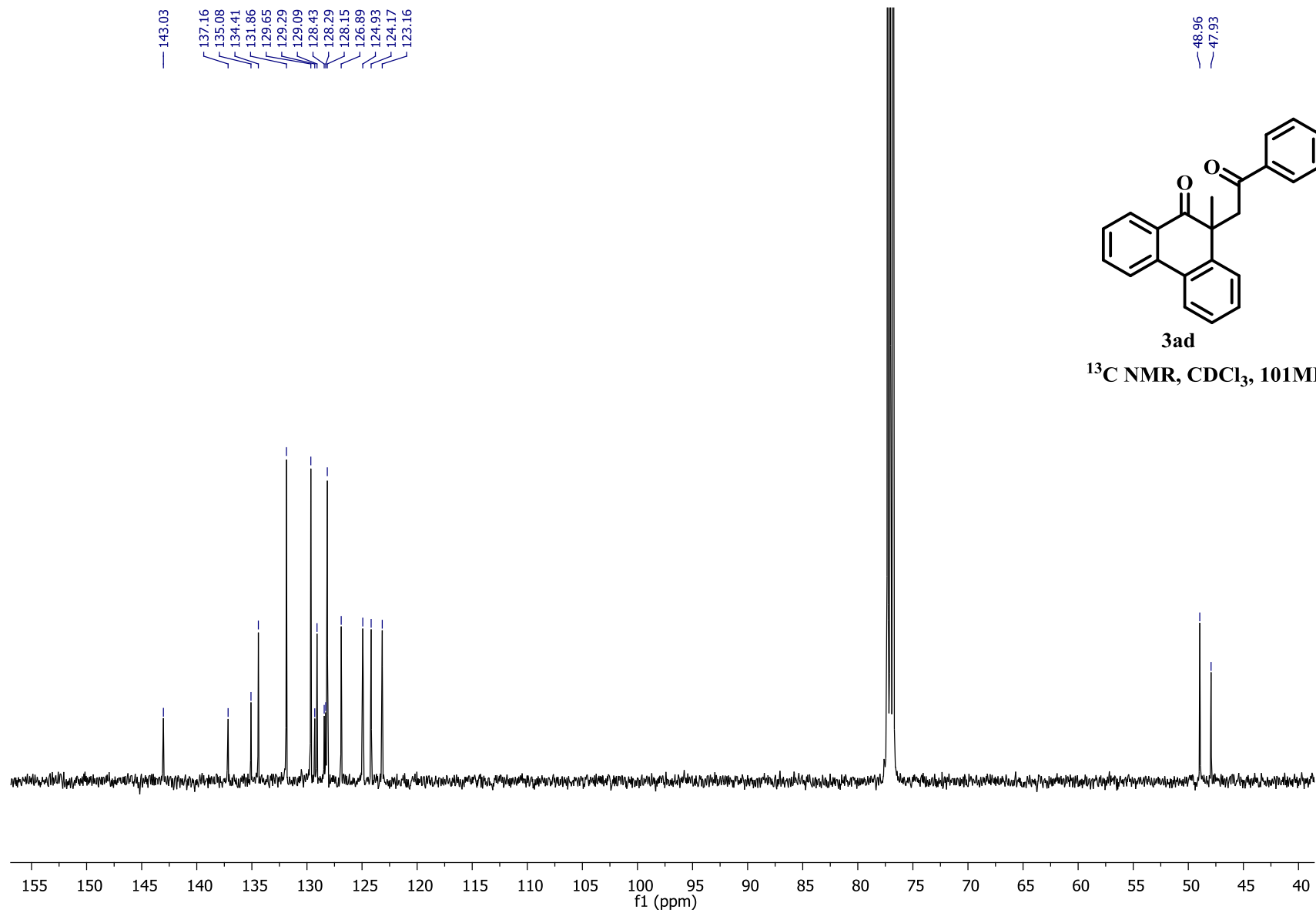


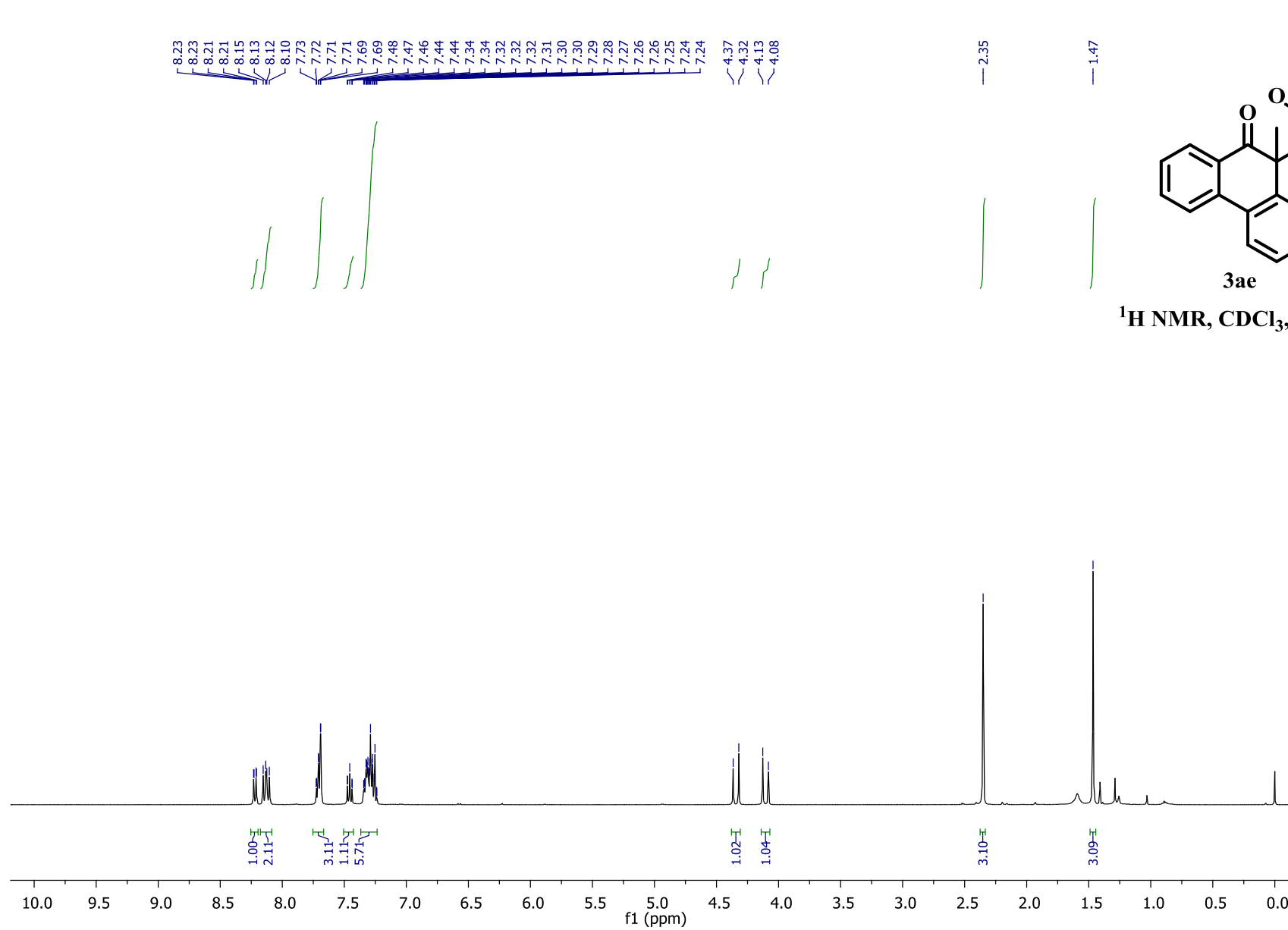


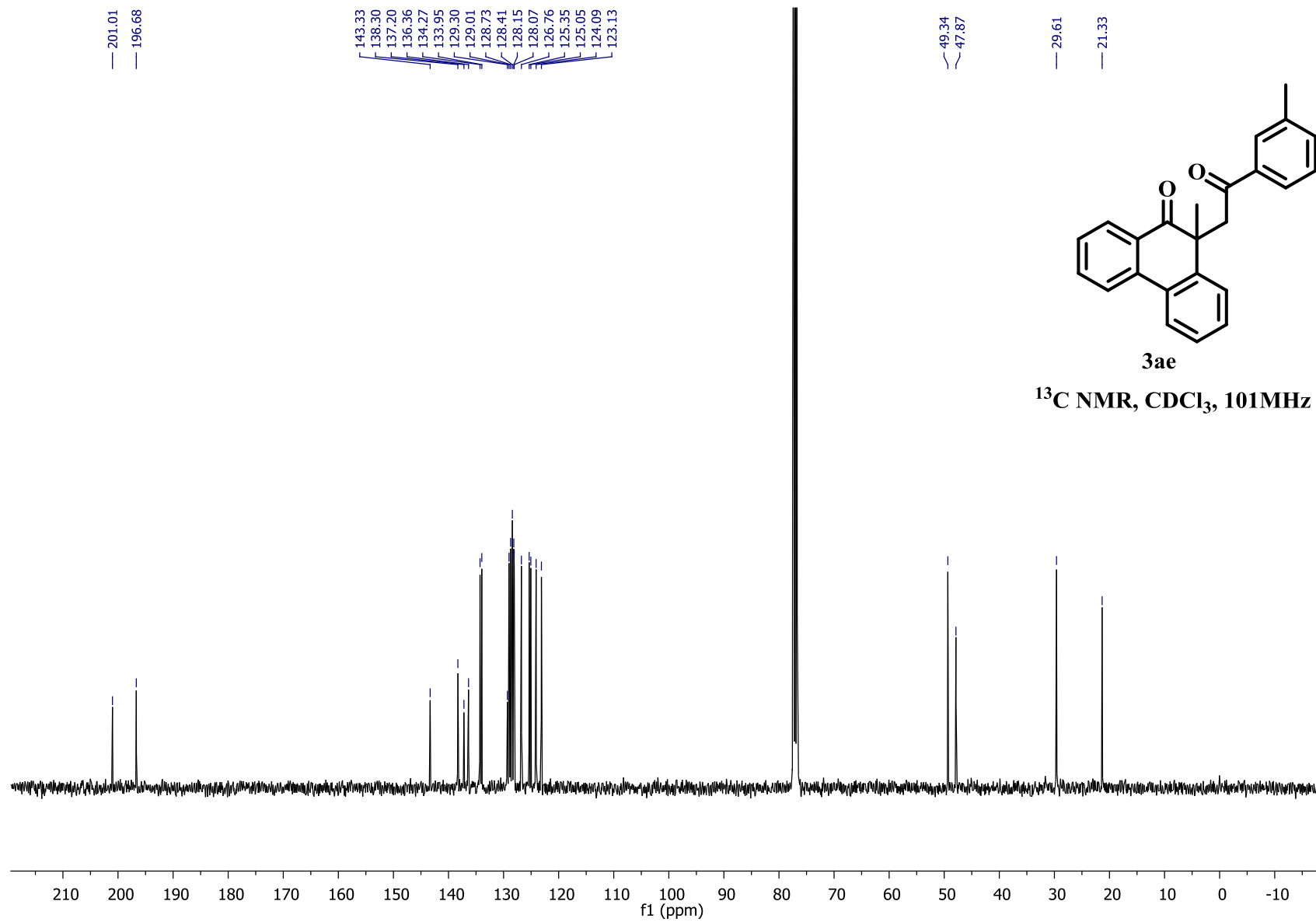


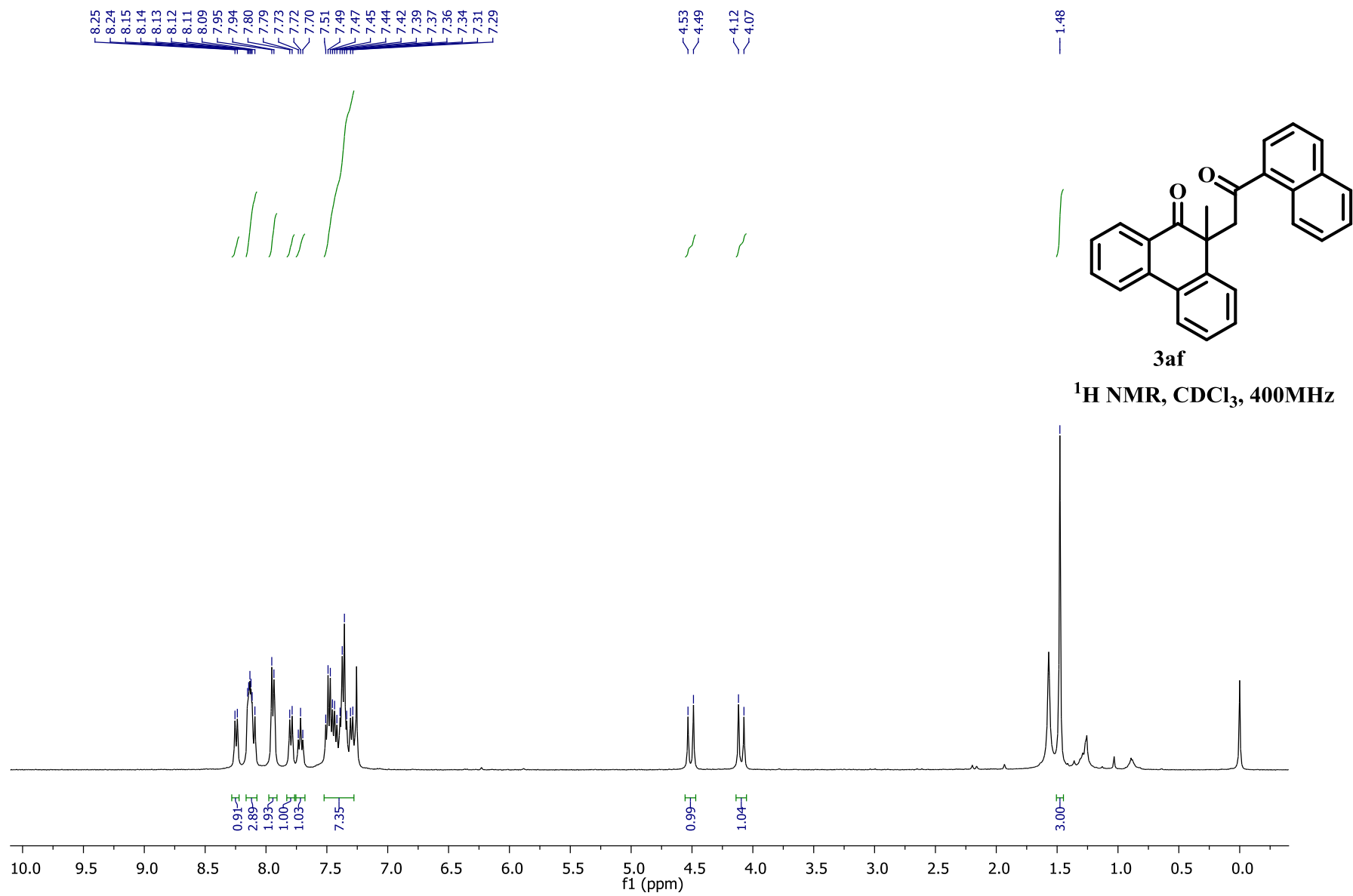


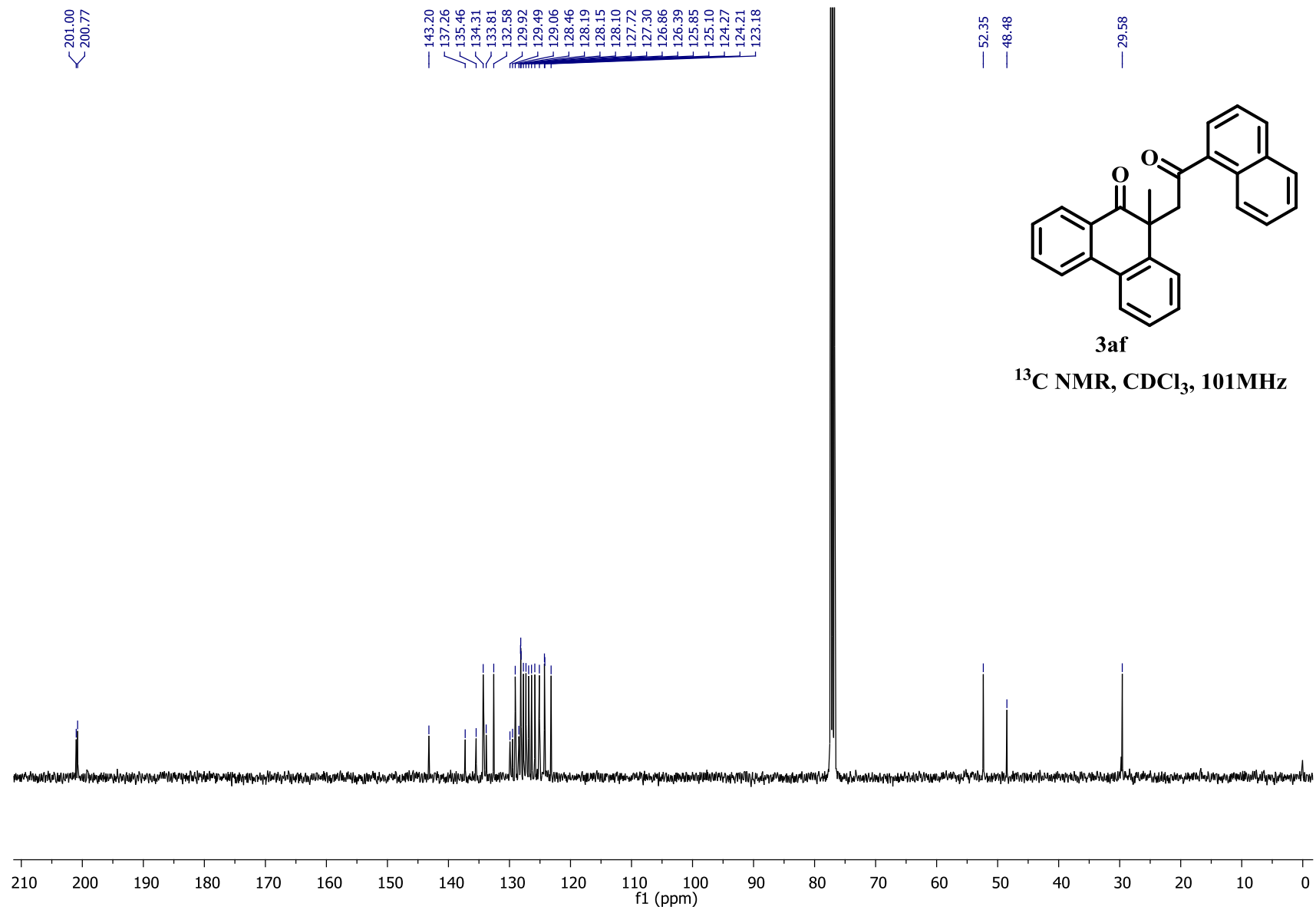


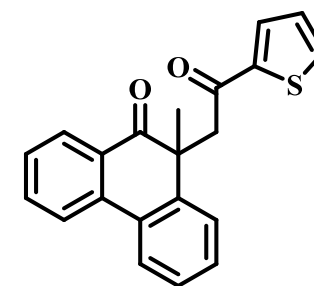
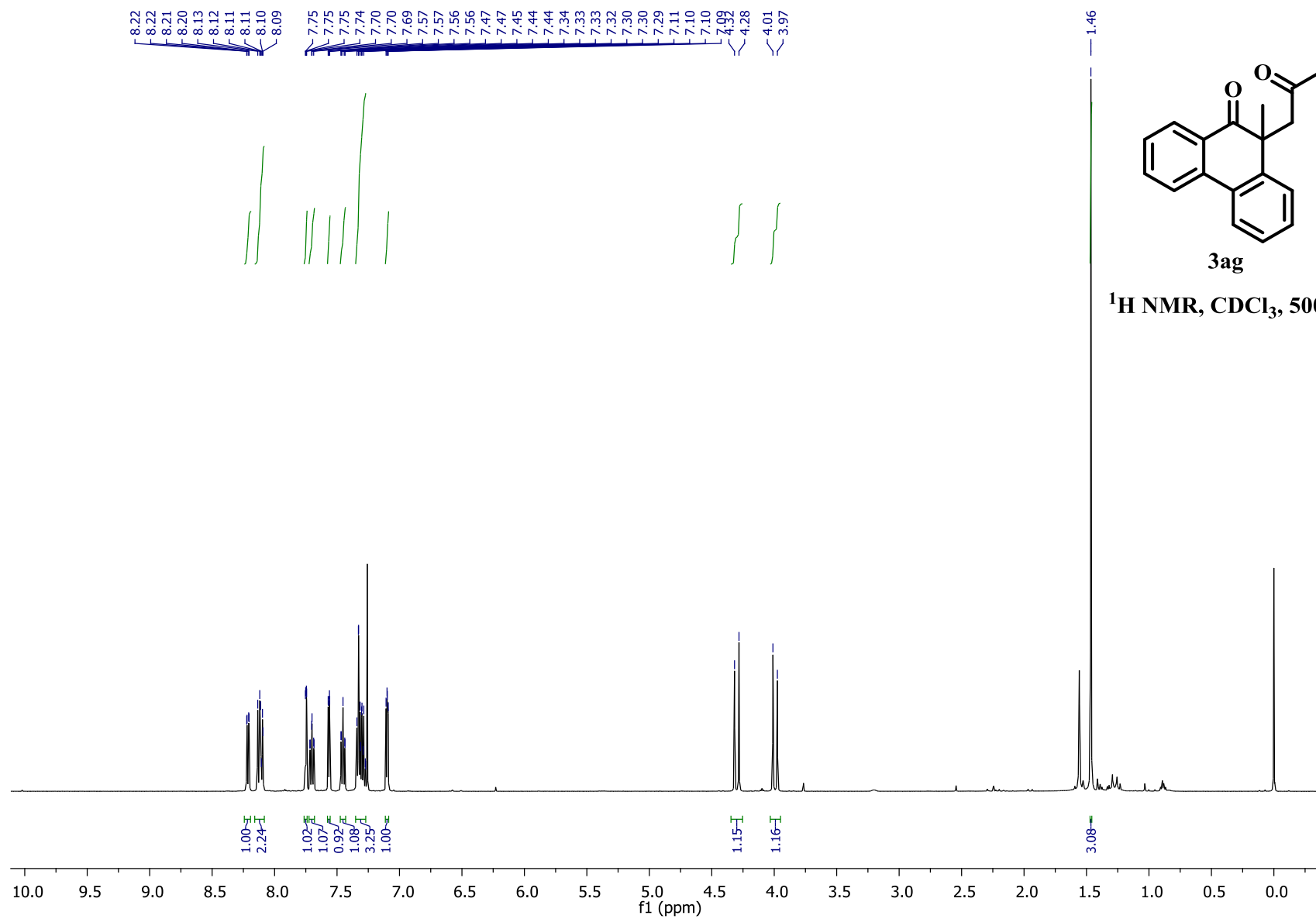






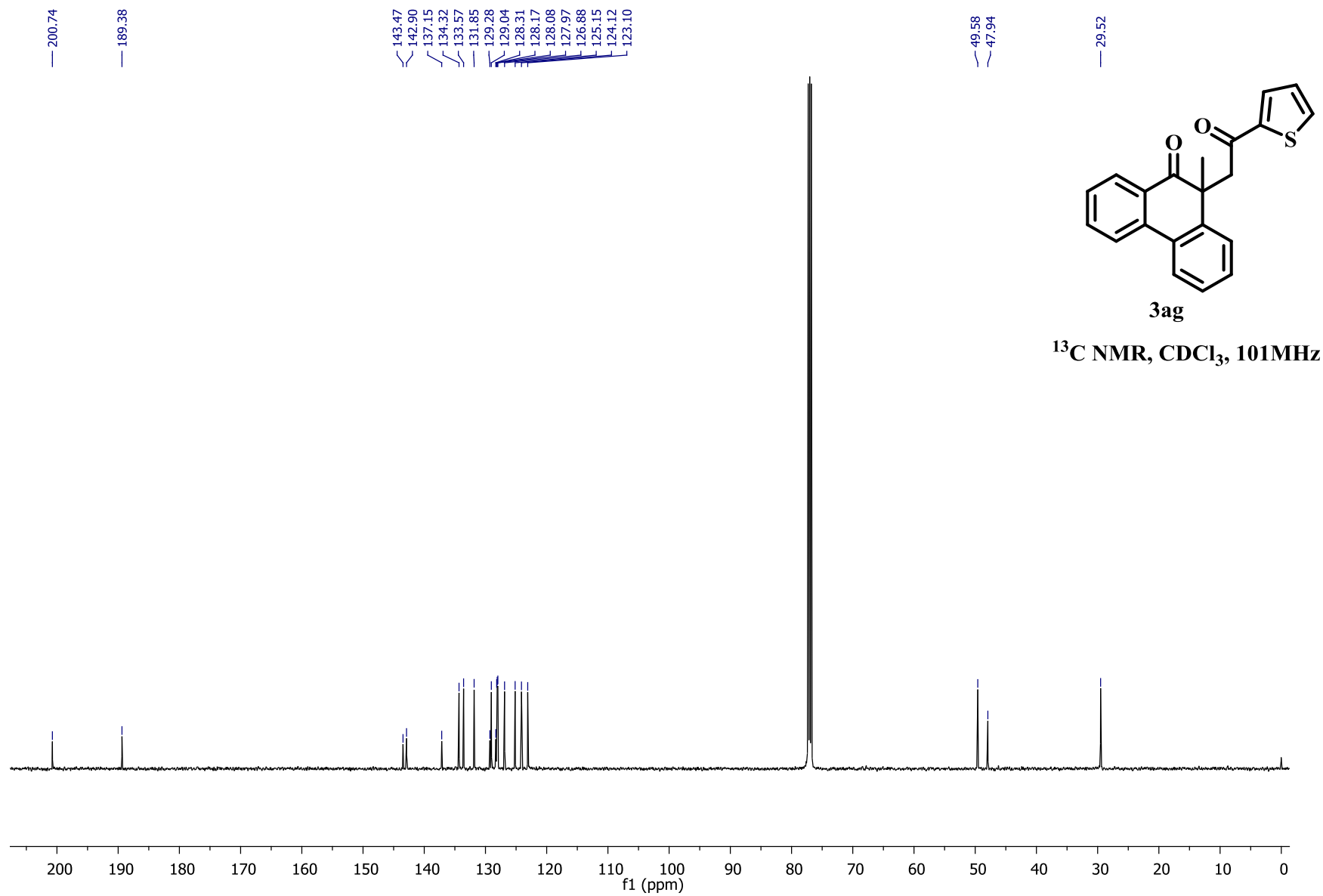


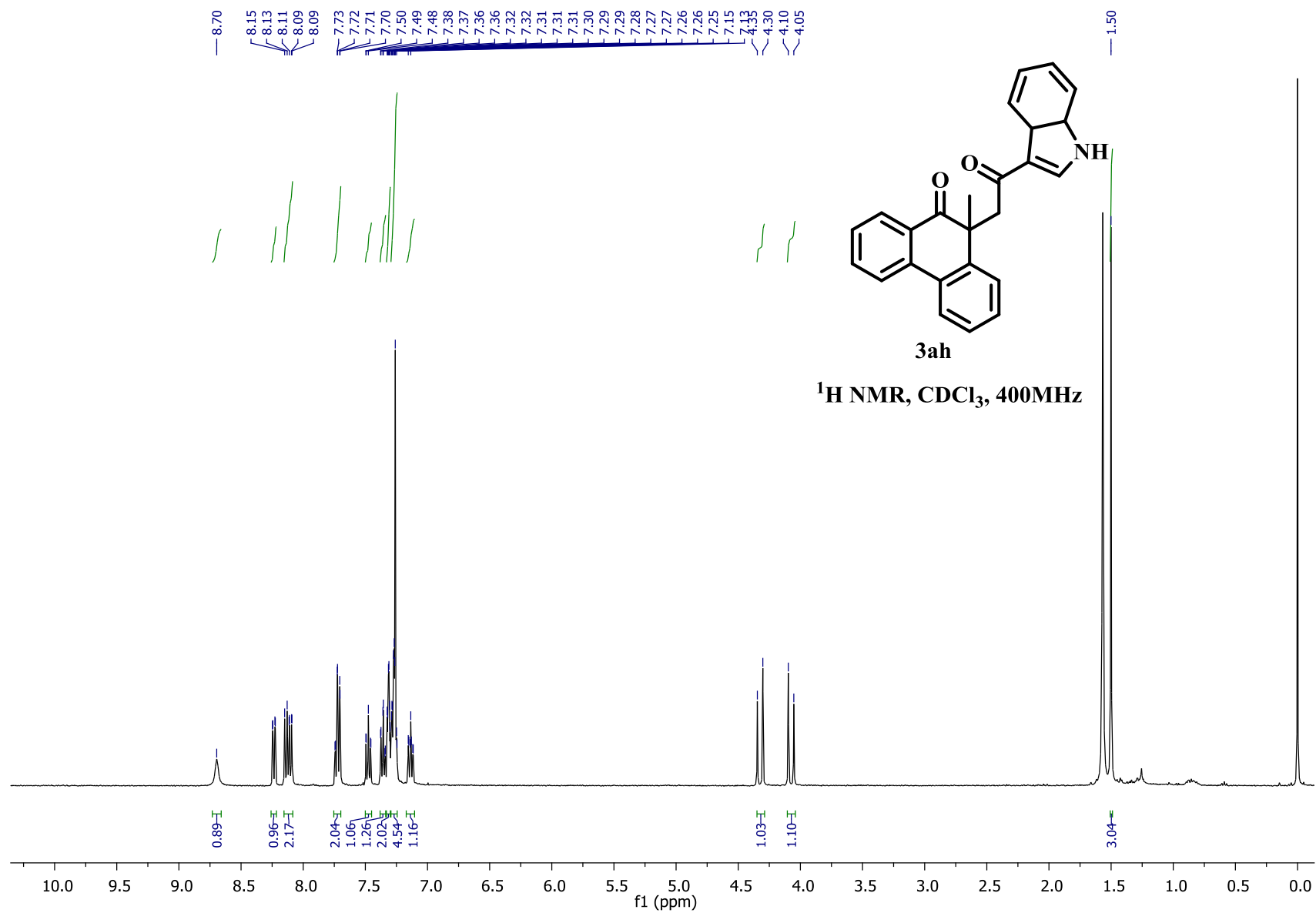


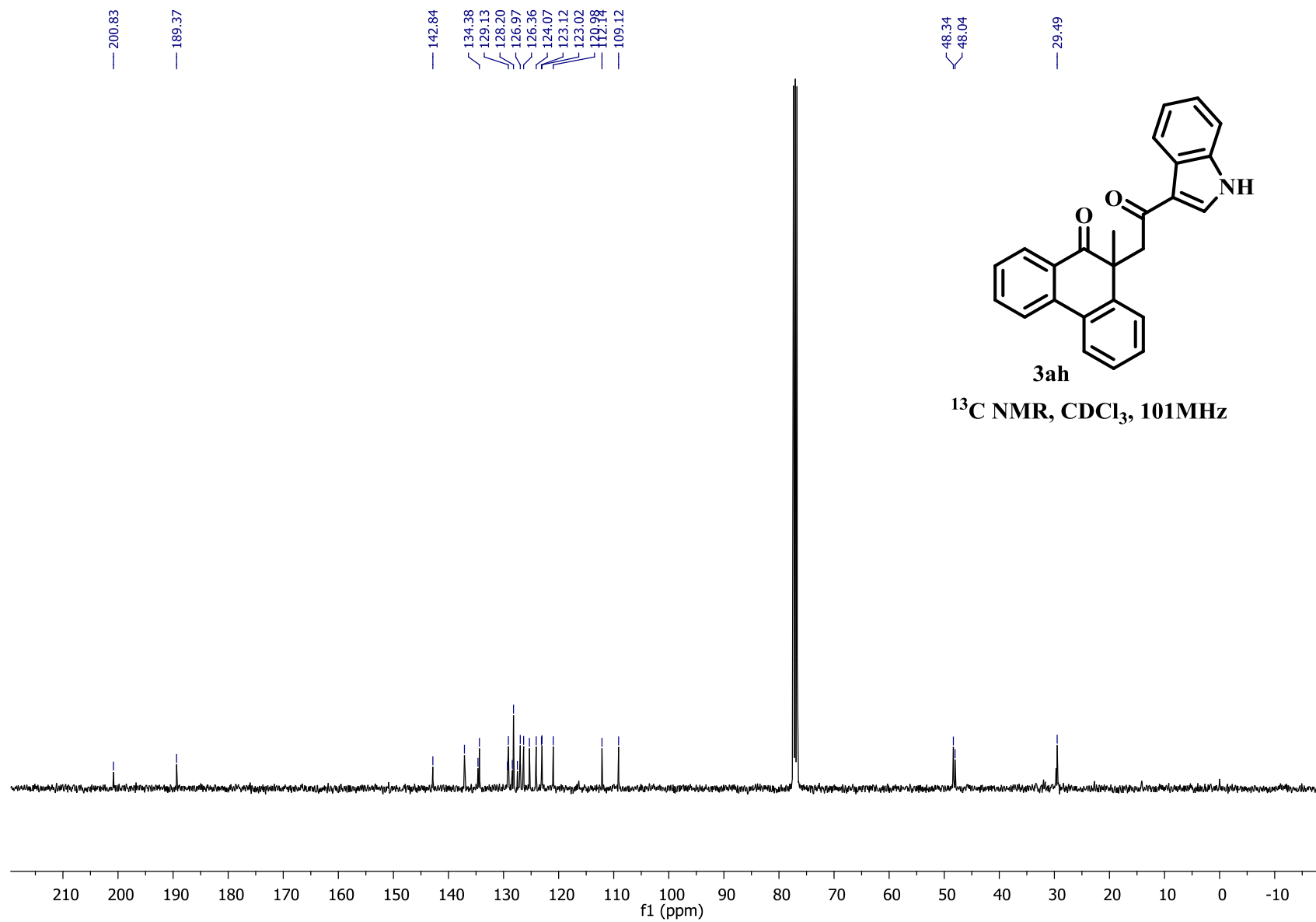


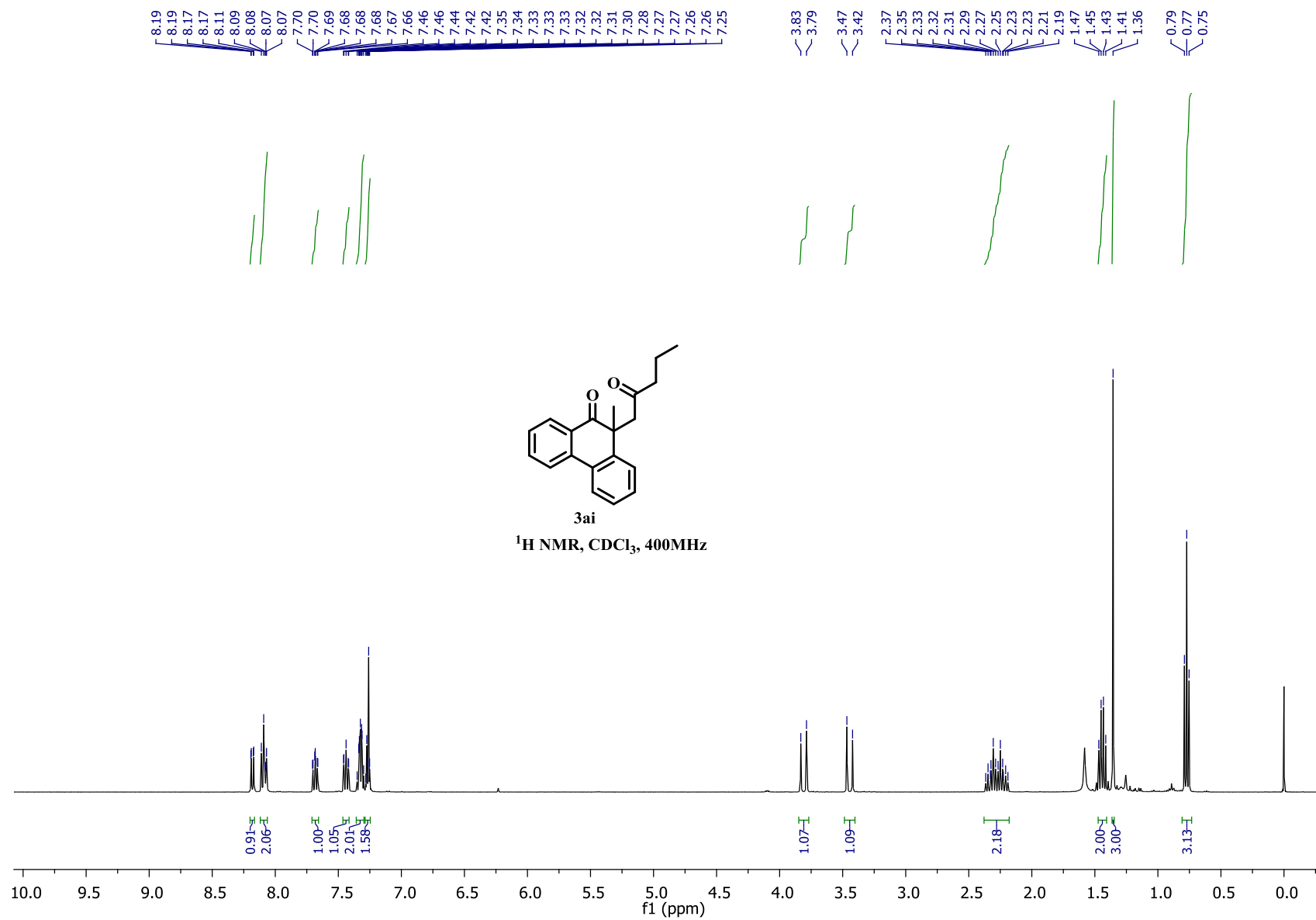
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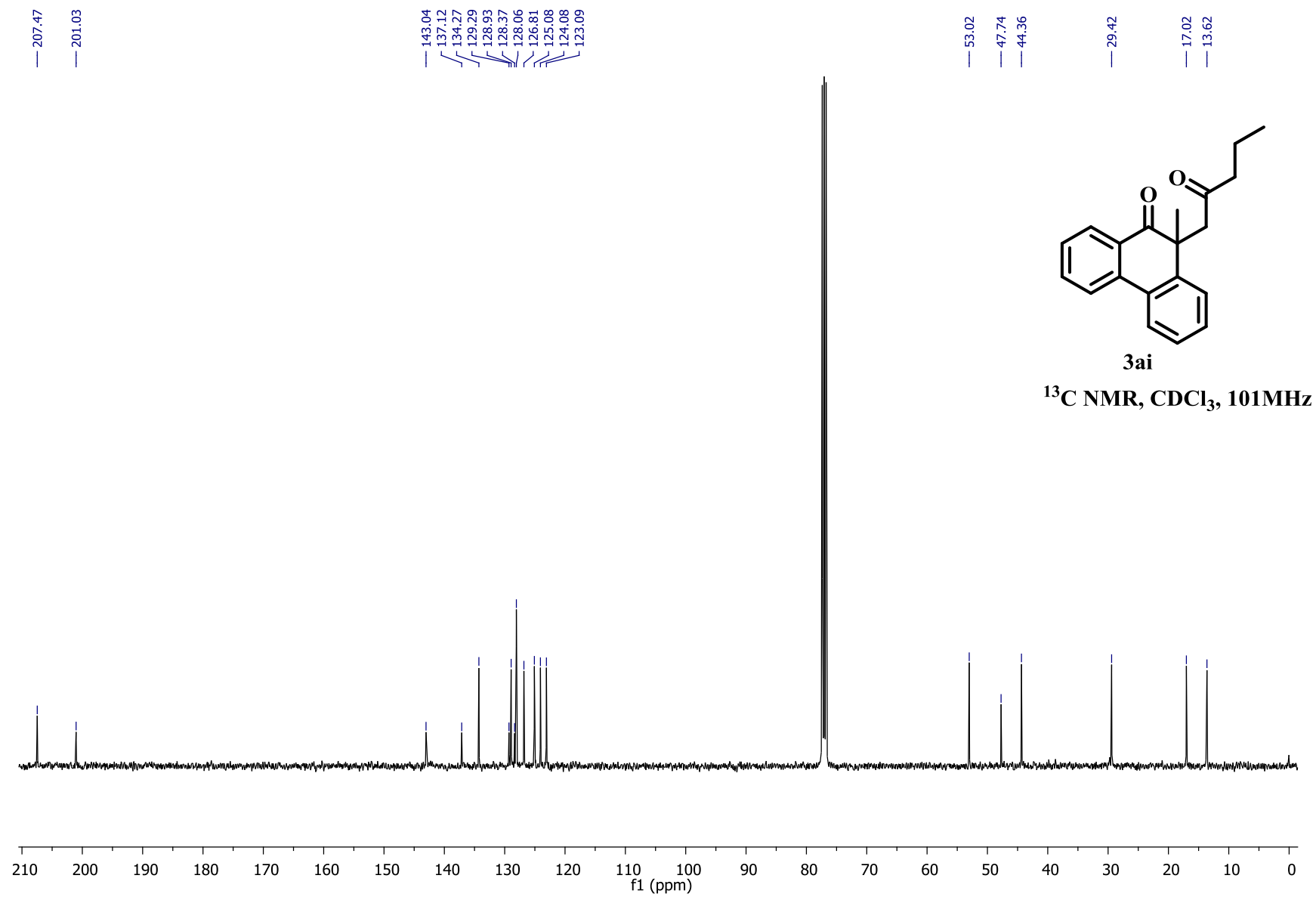
¹H NMR, CDCl₃, 500MHz

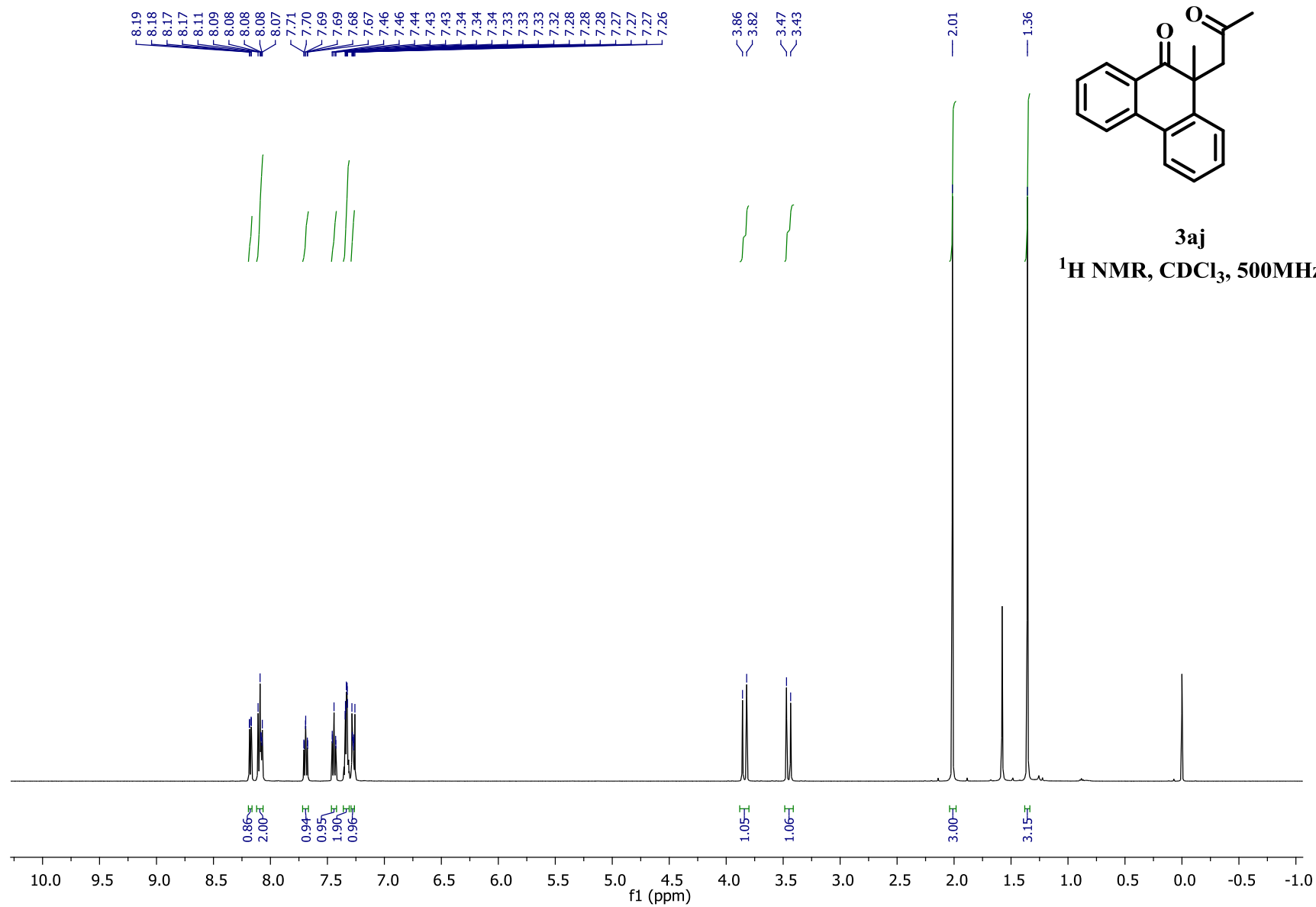


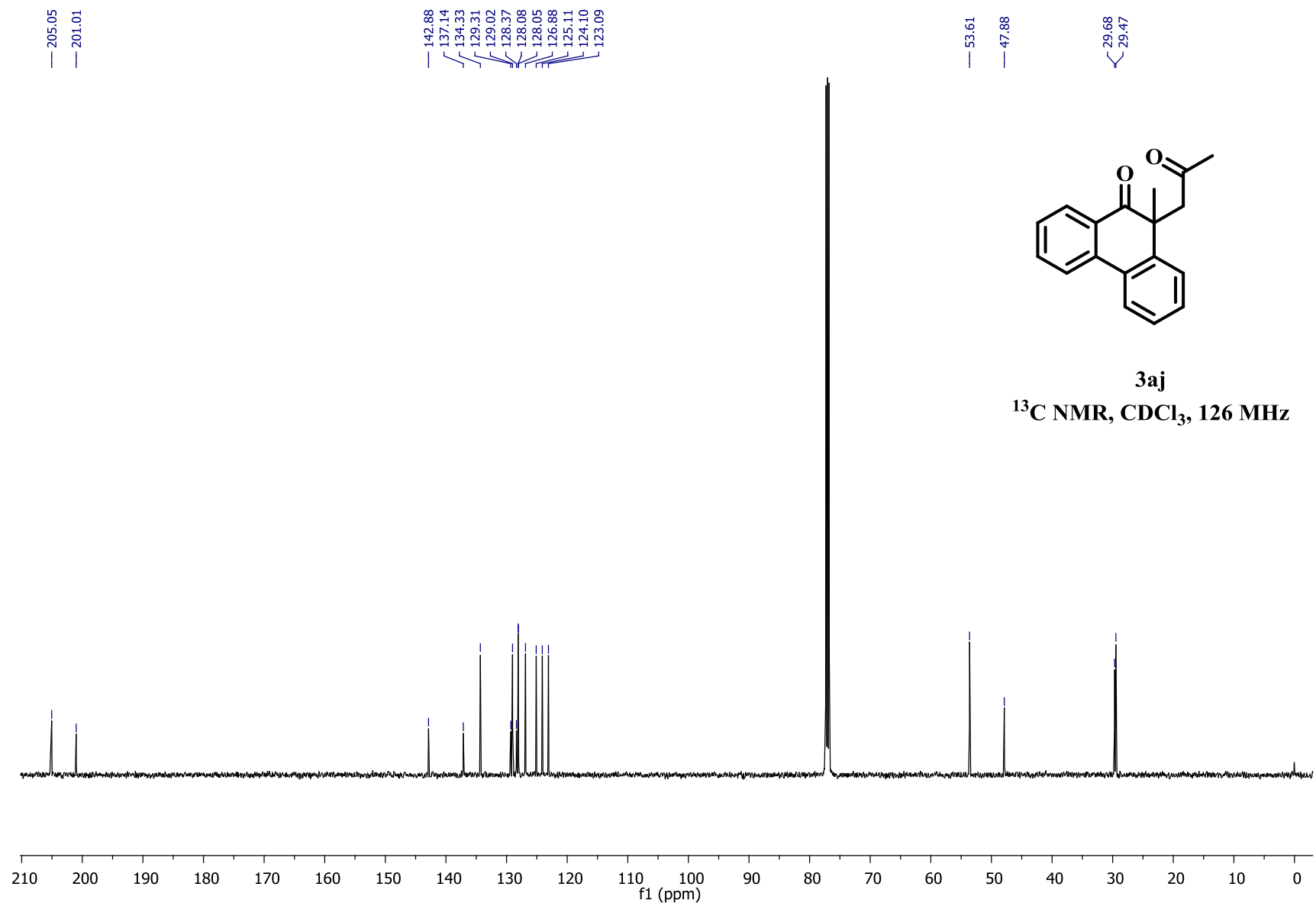


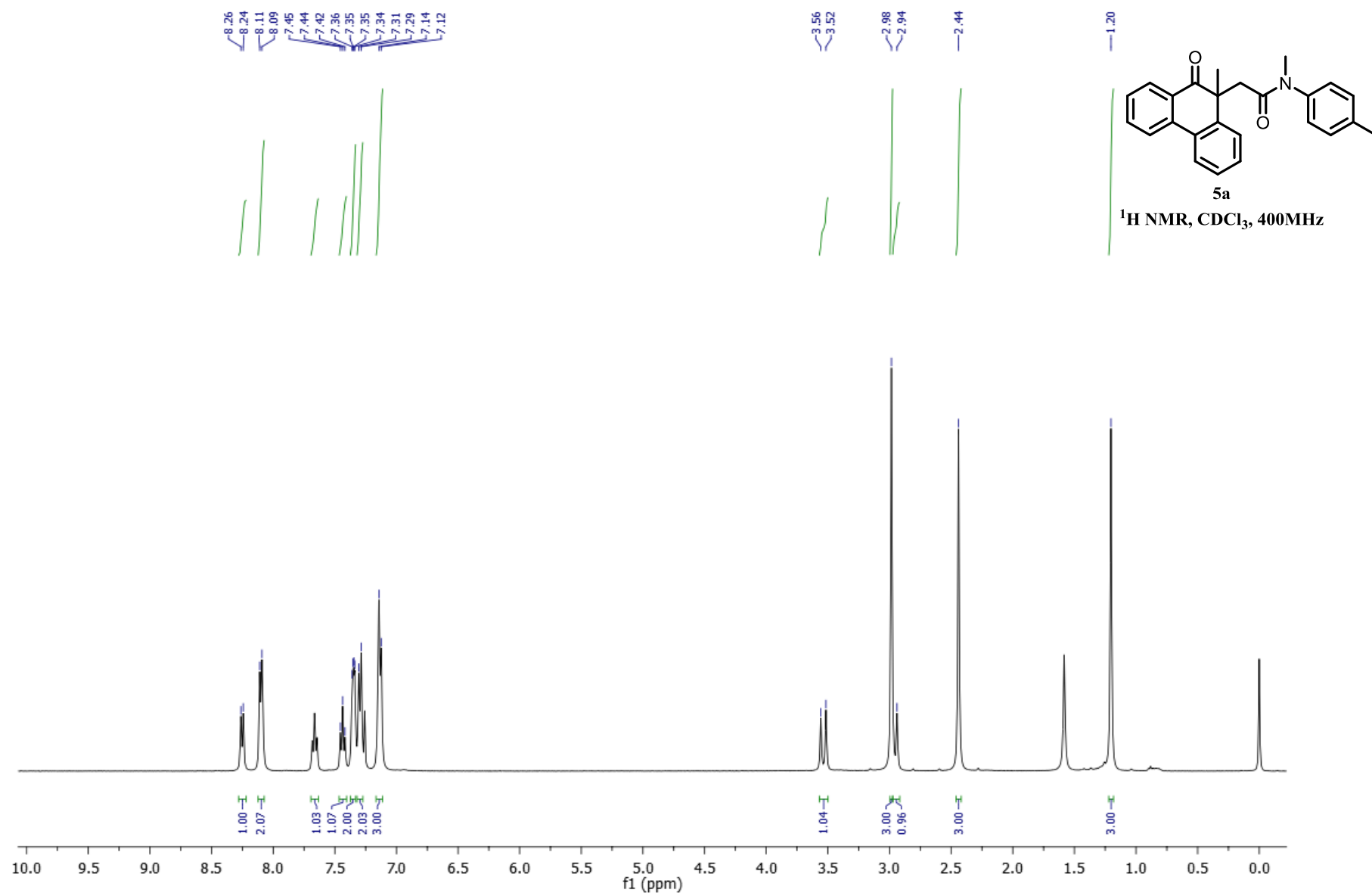


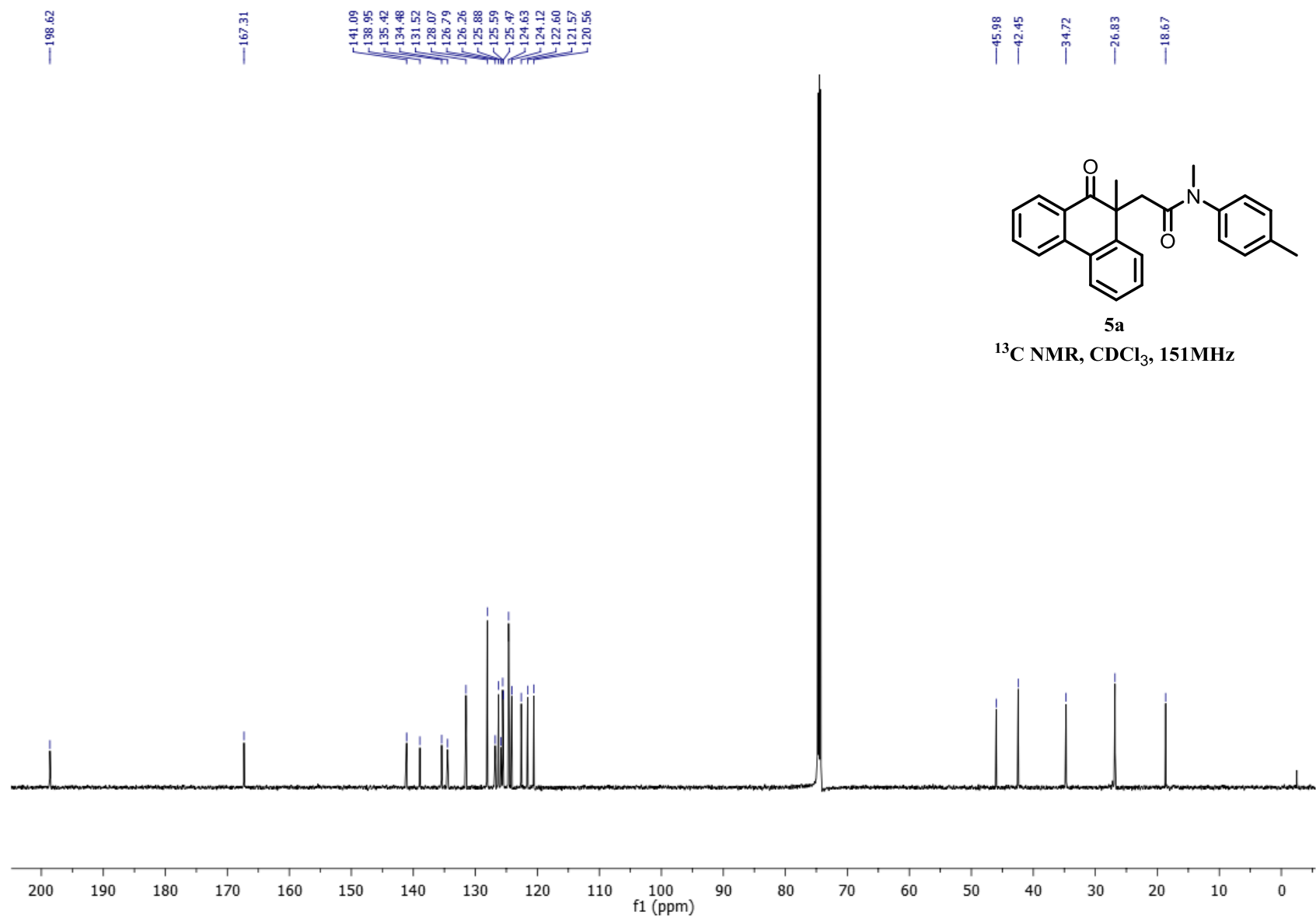


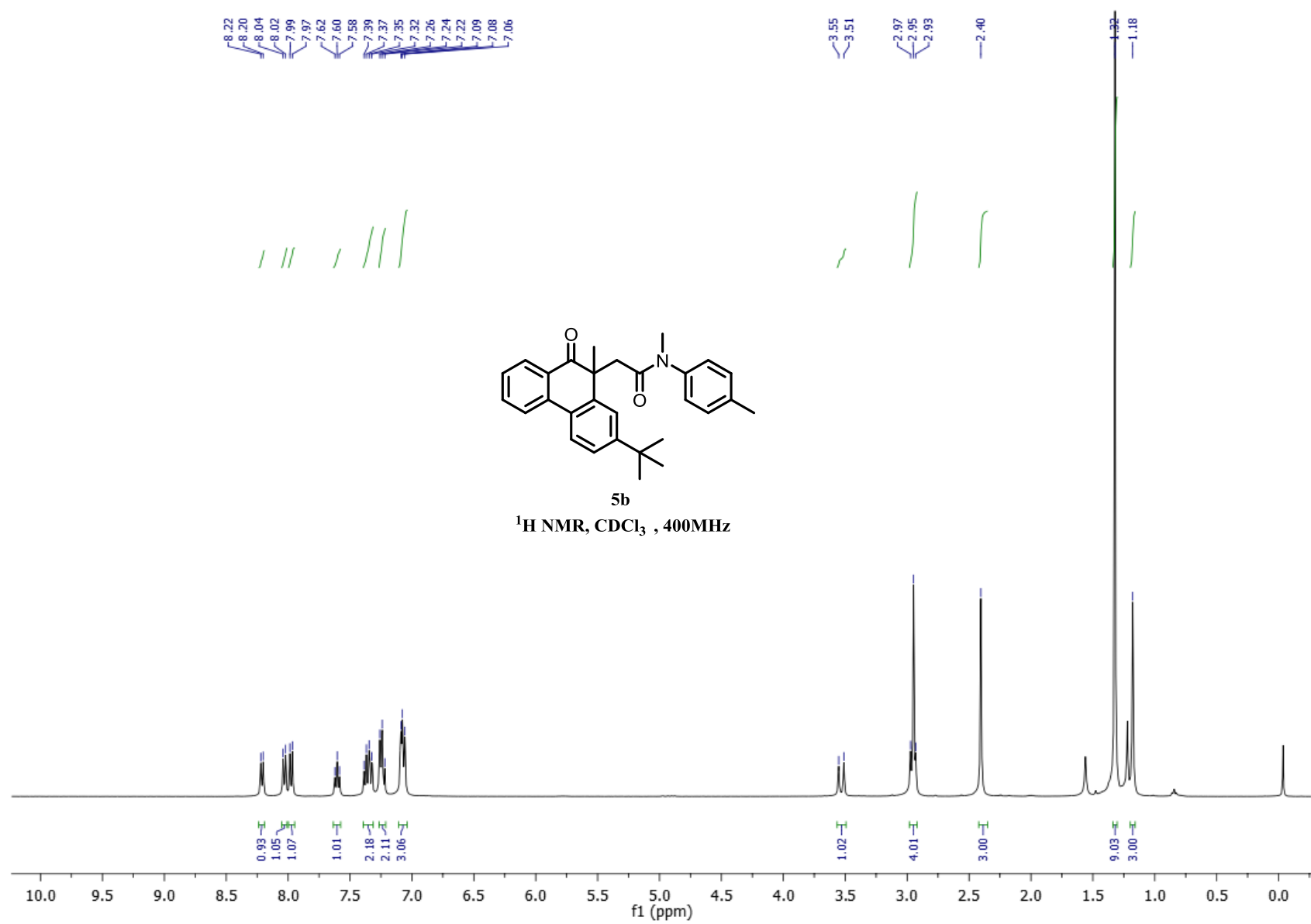


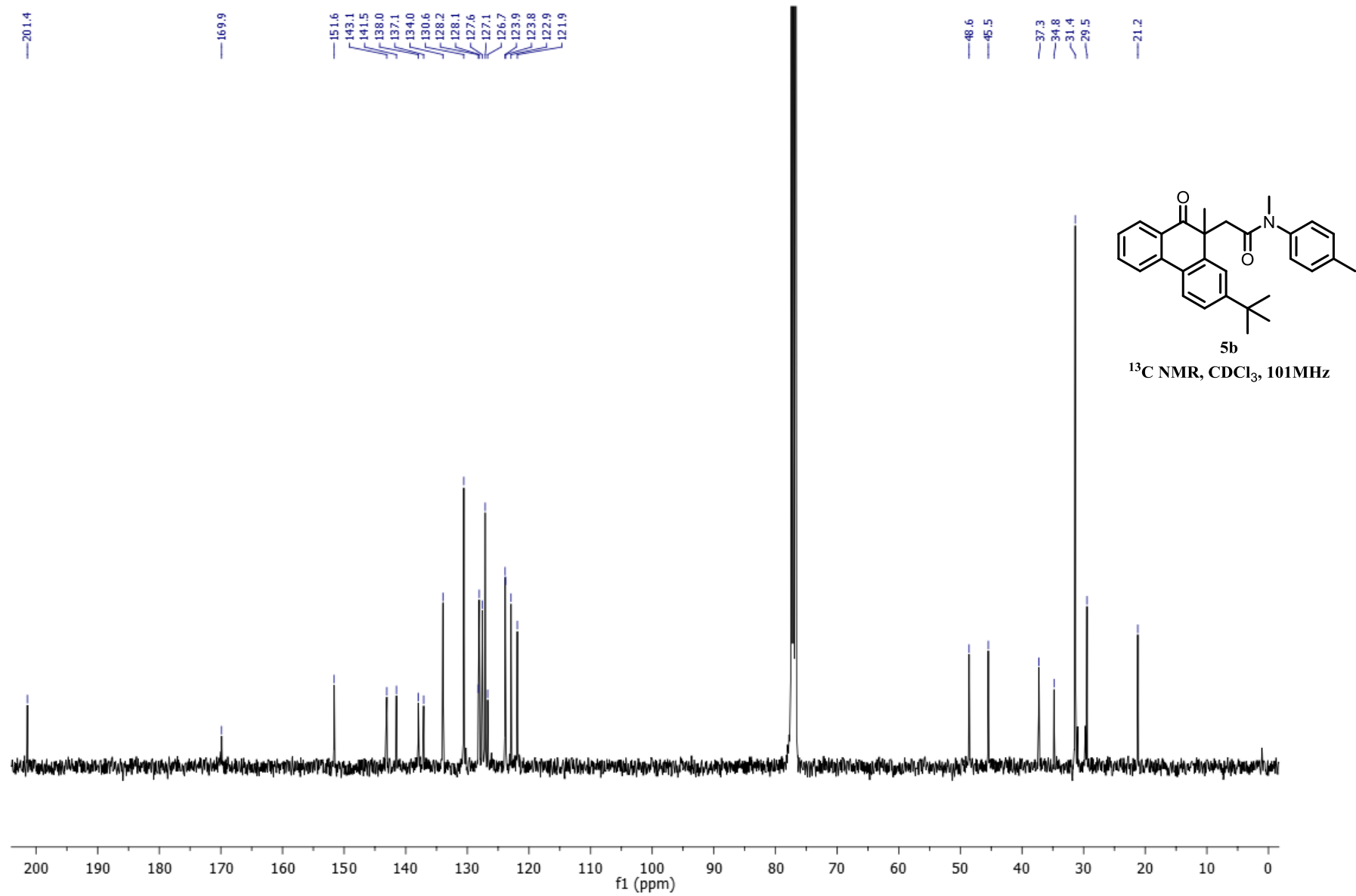


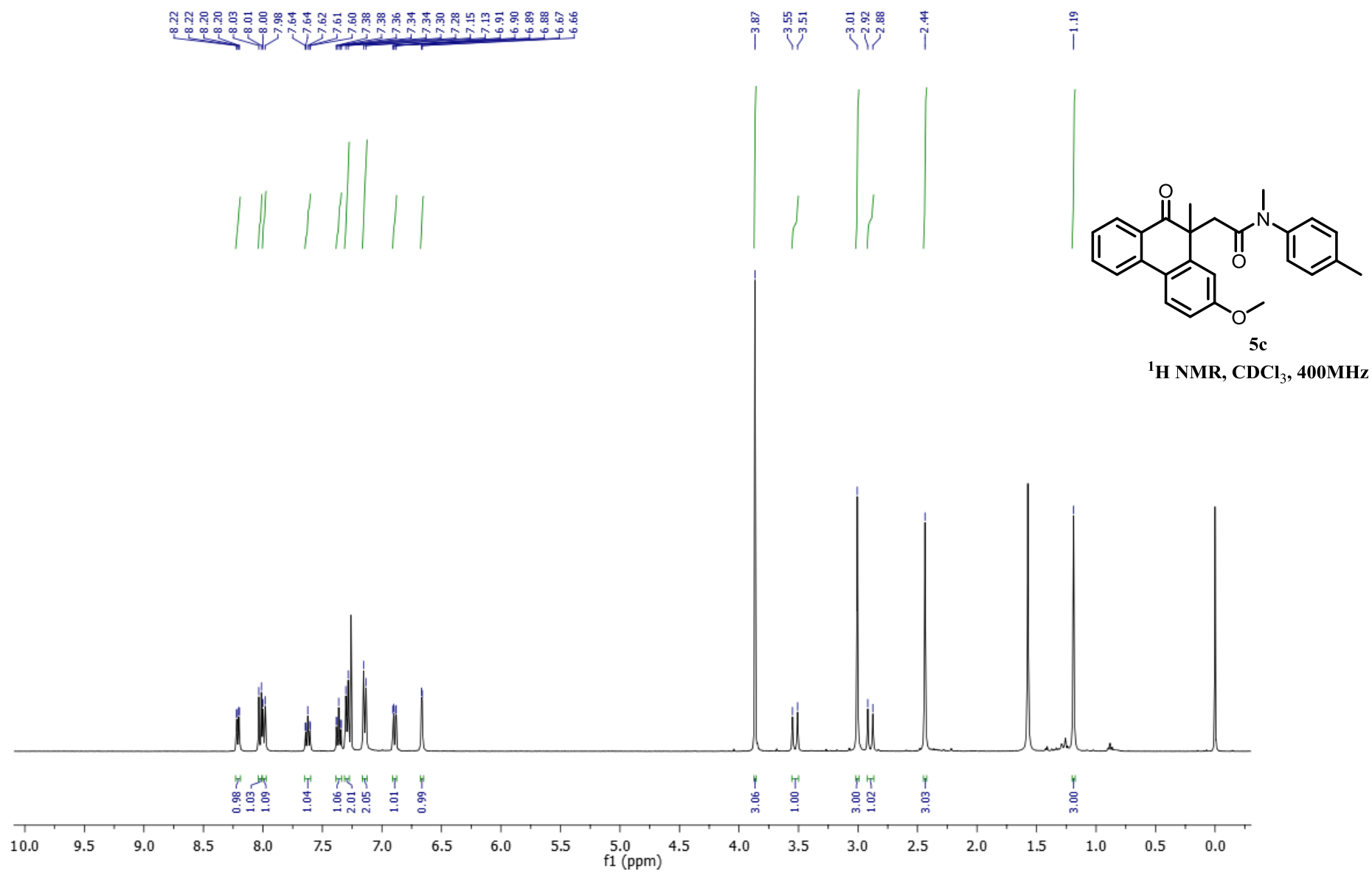


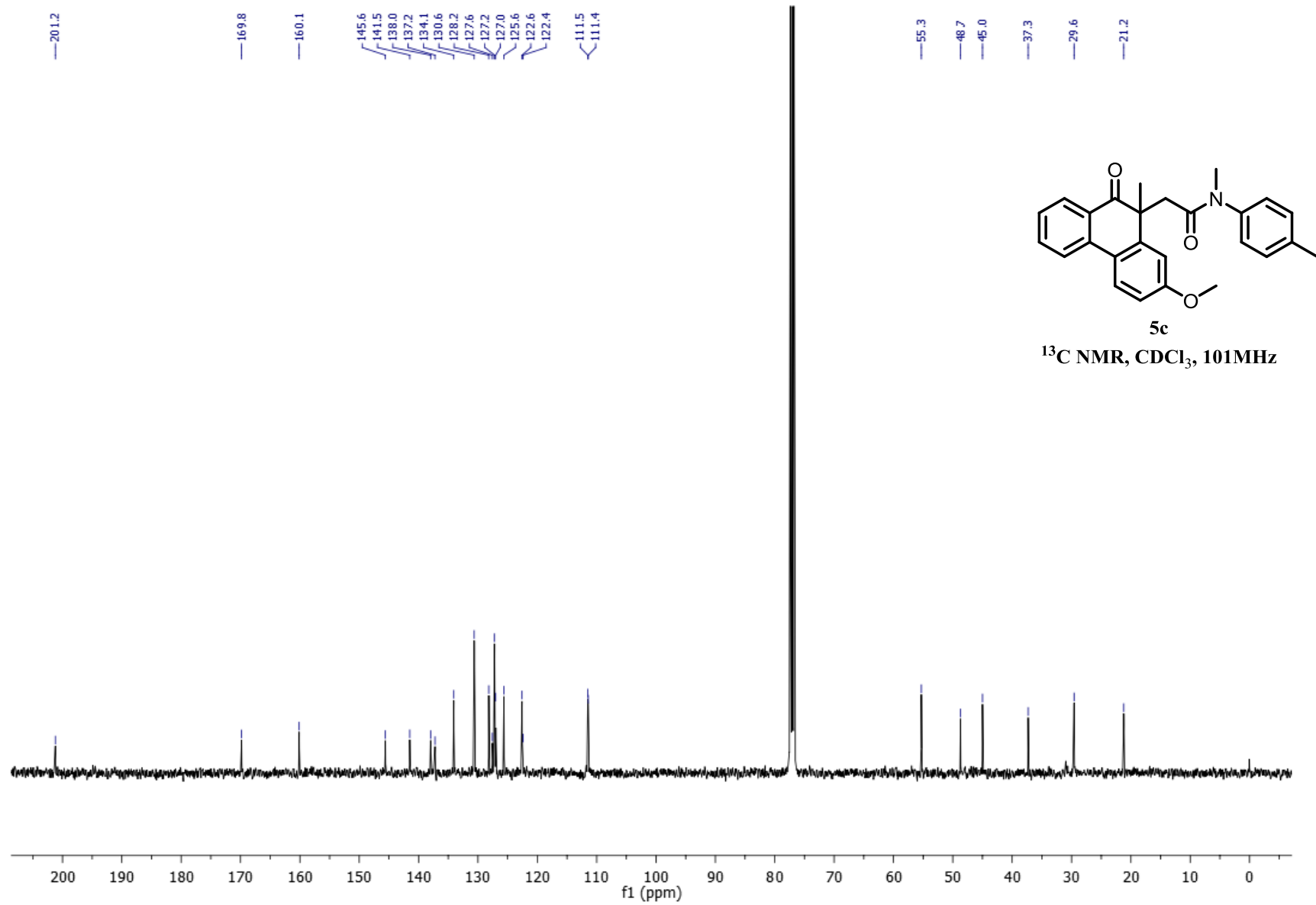


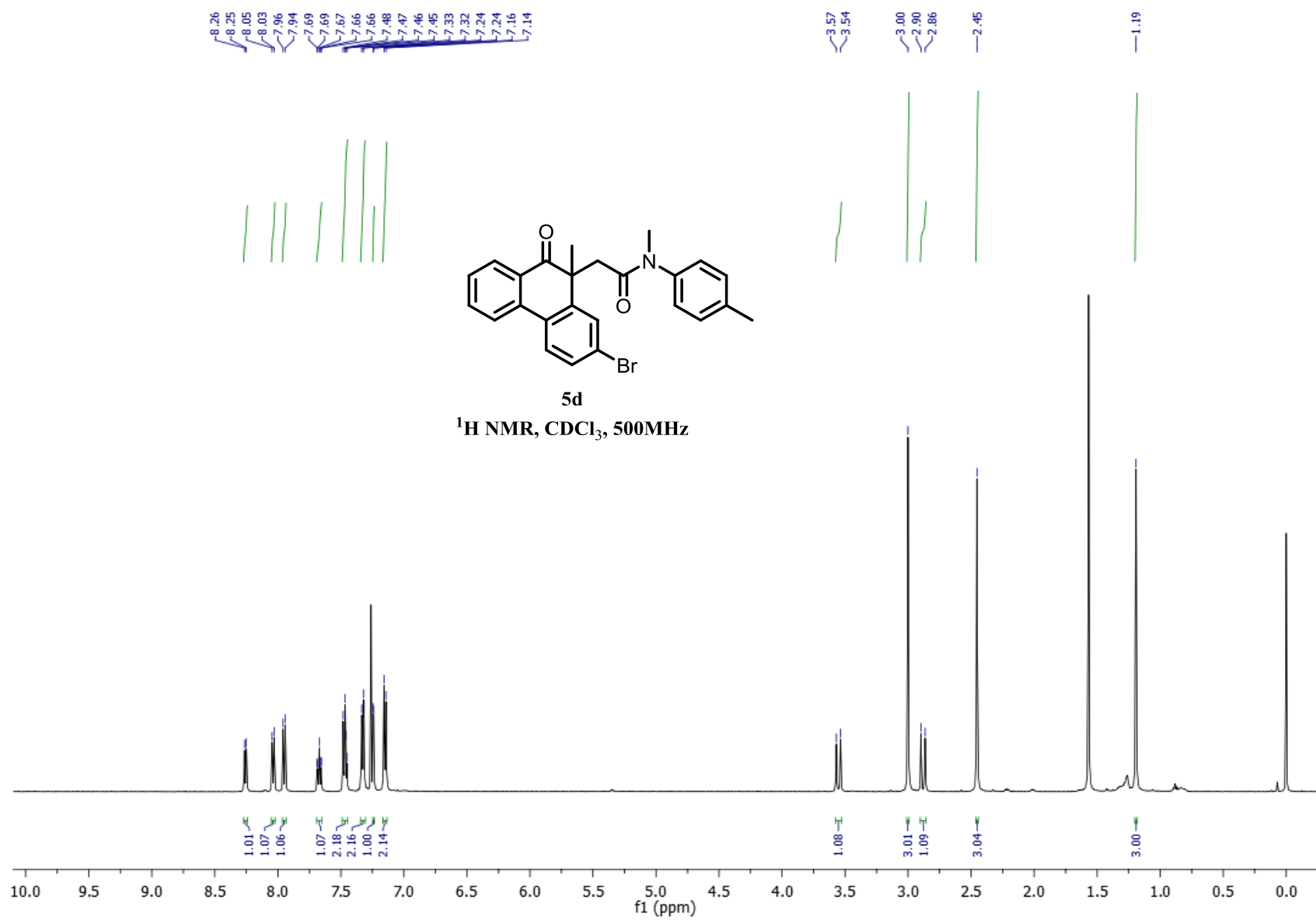


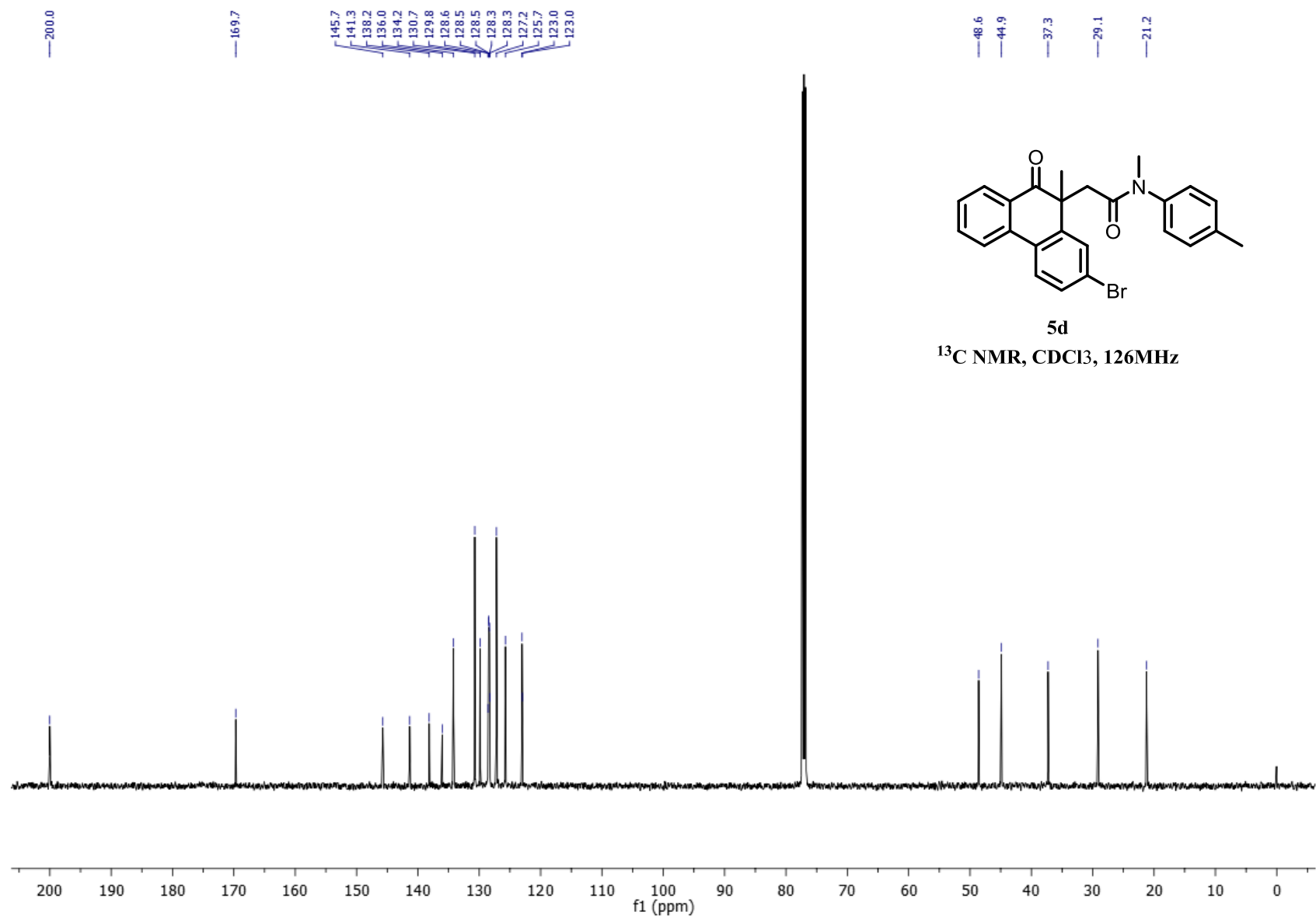


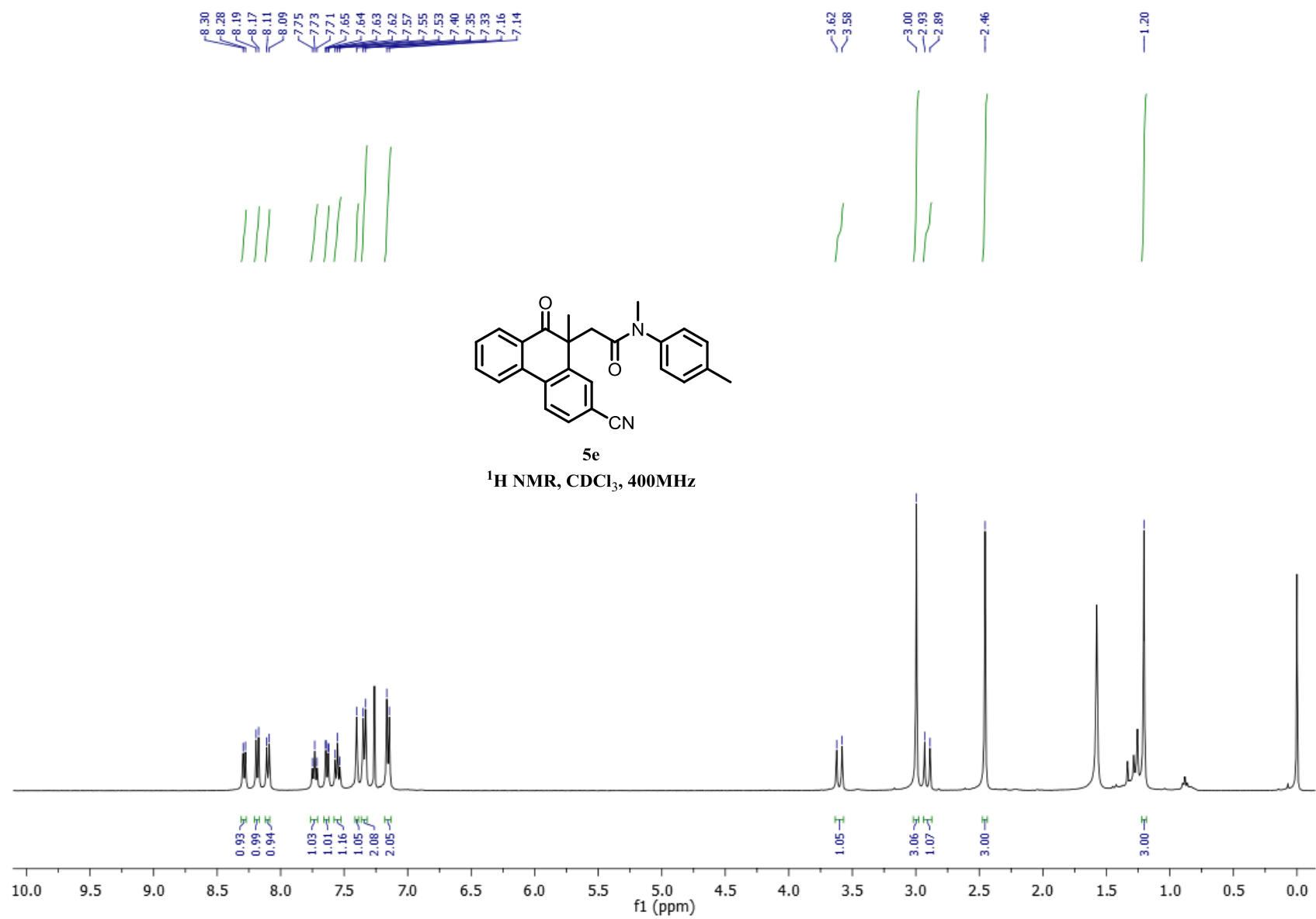


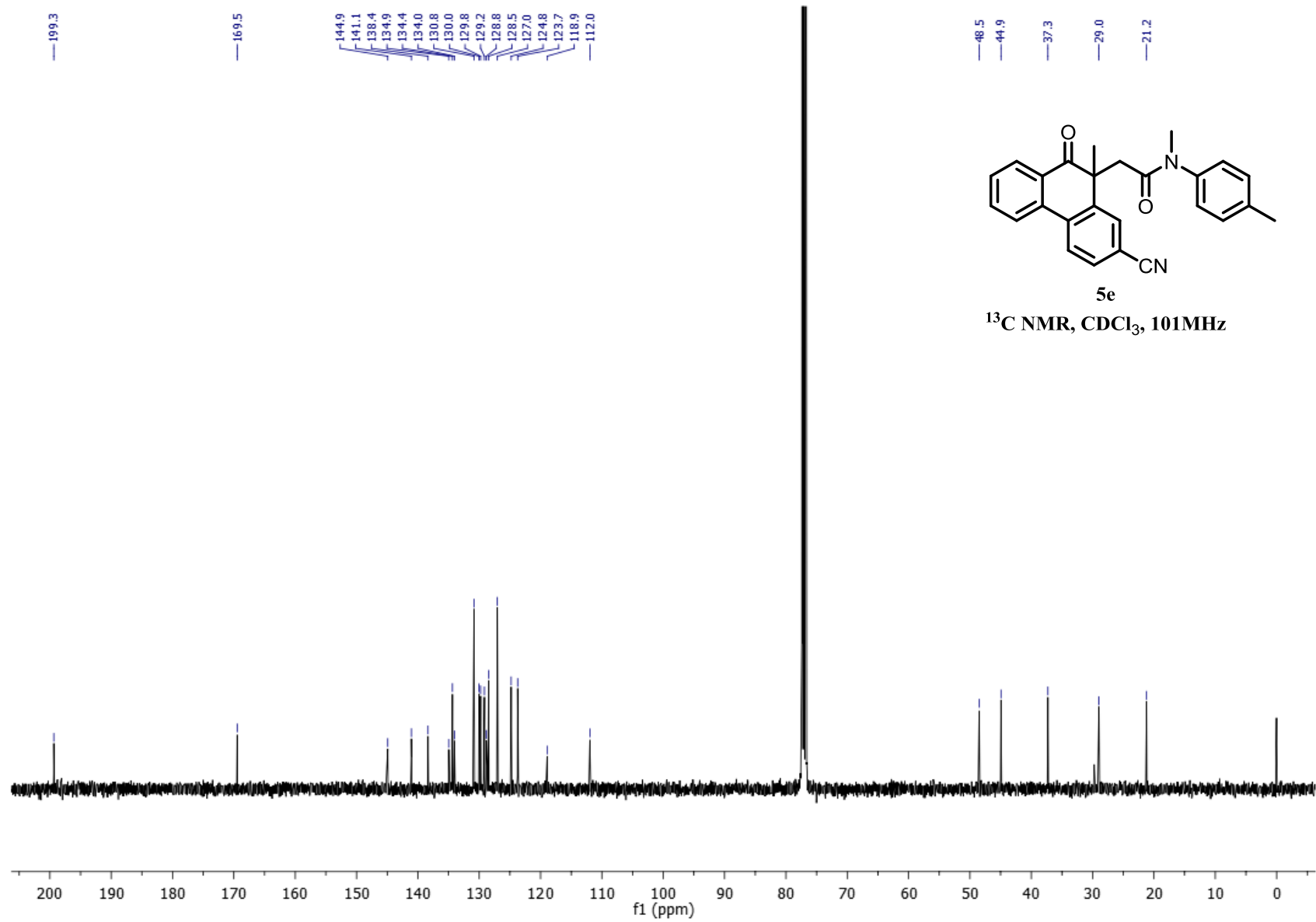


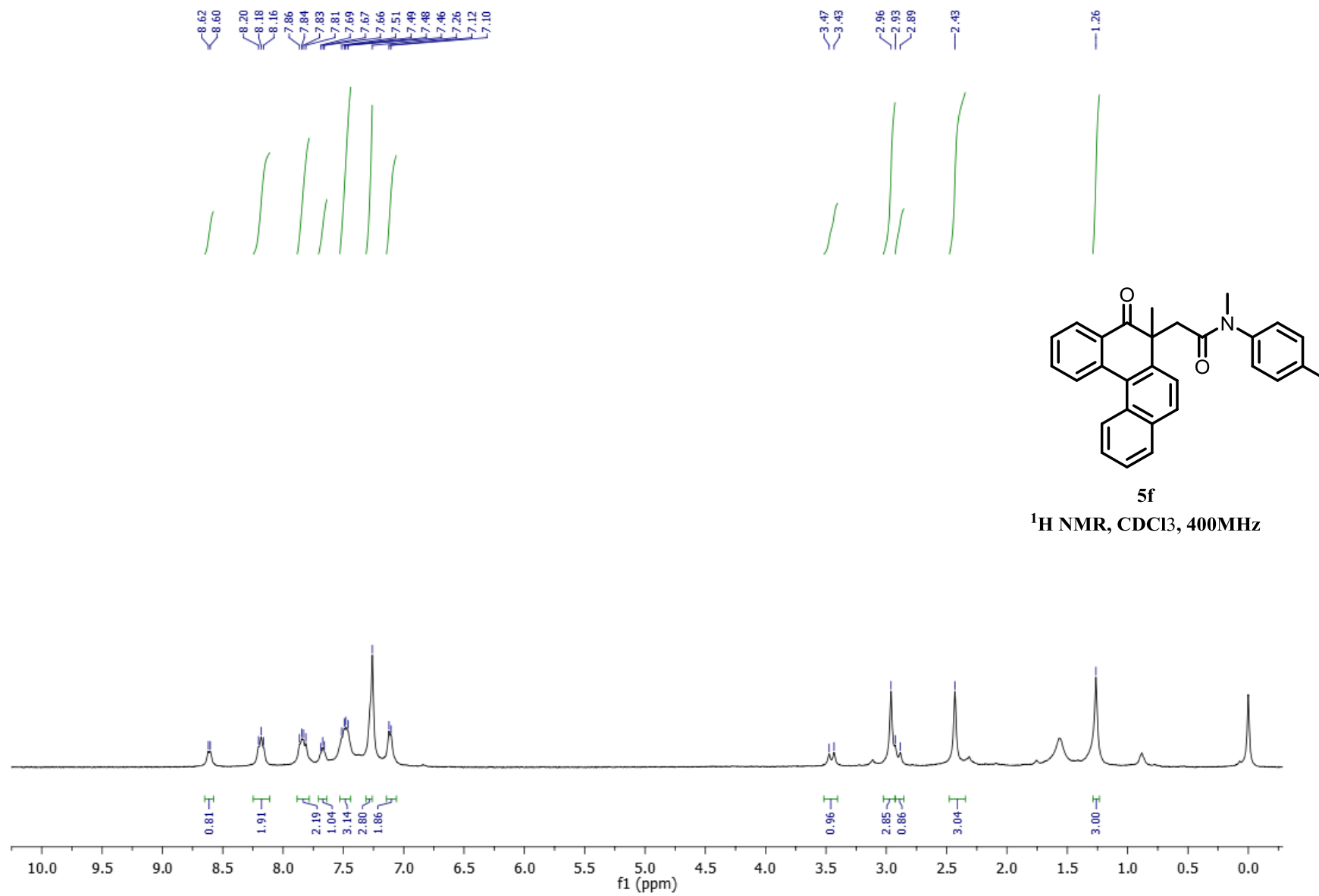


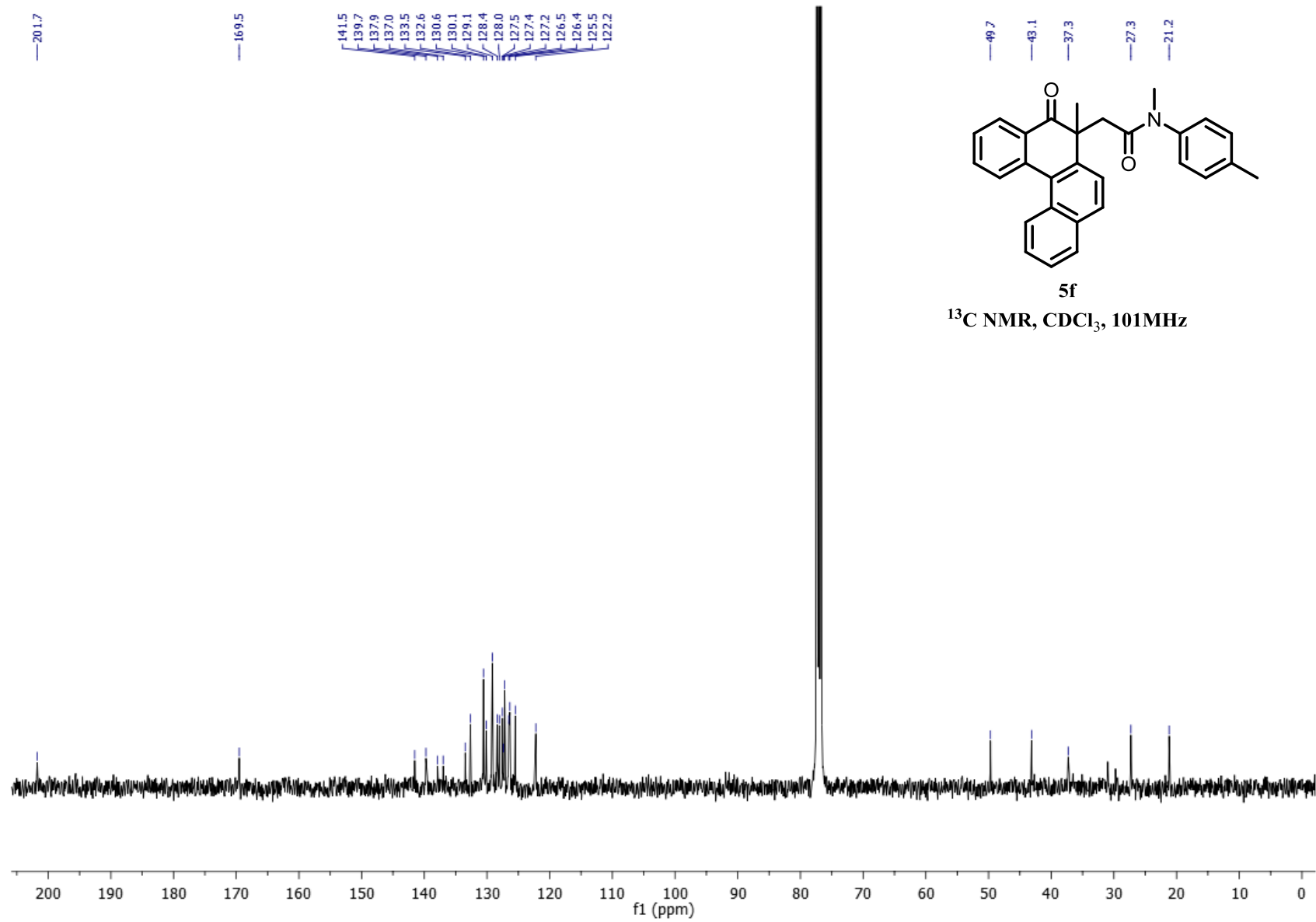


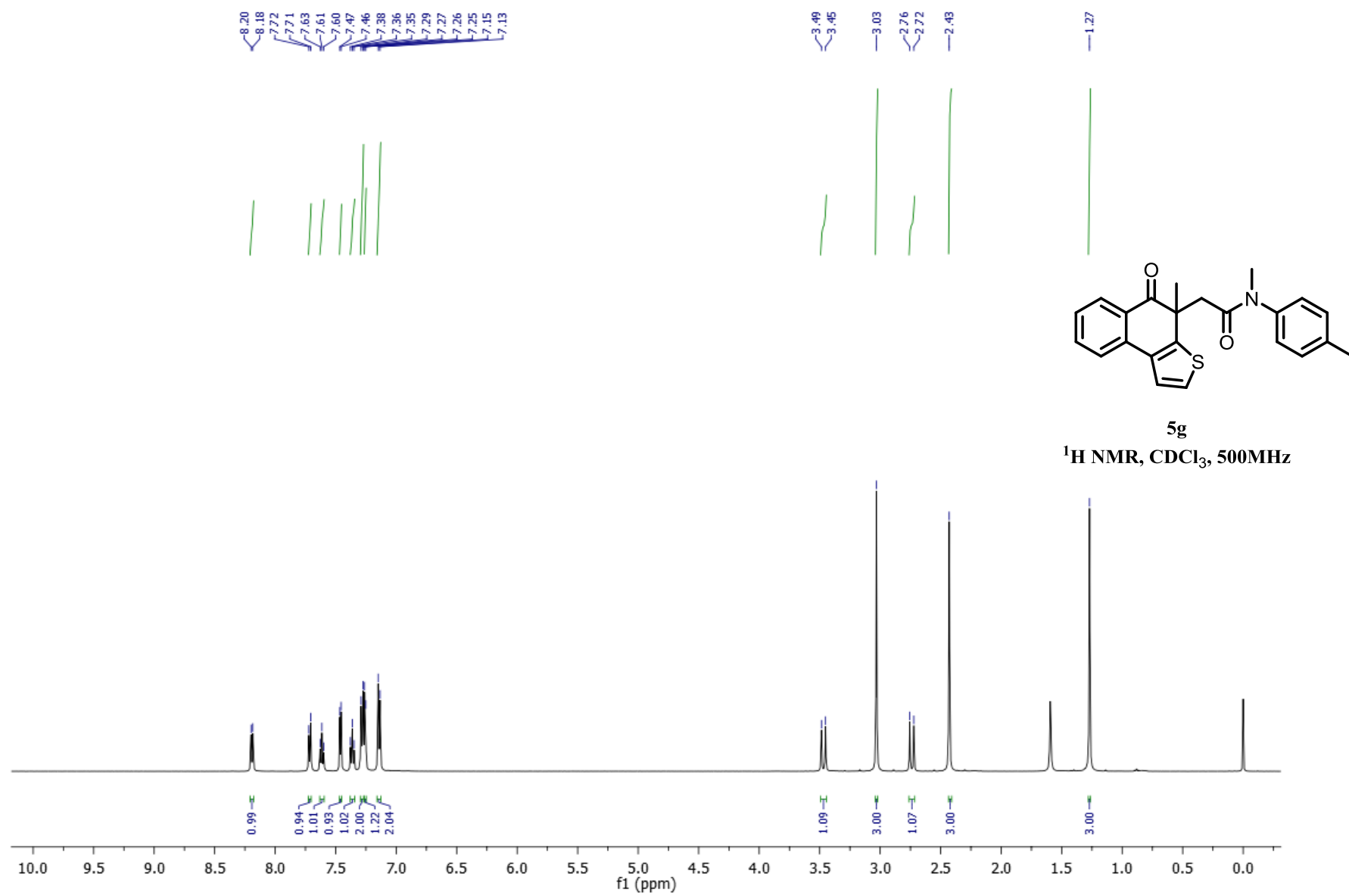


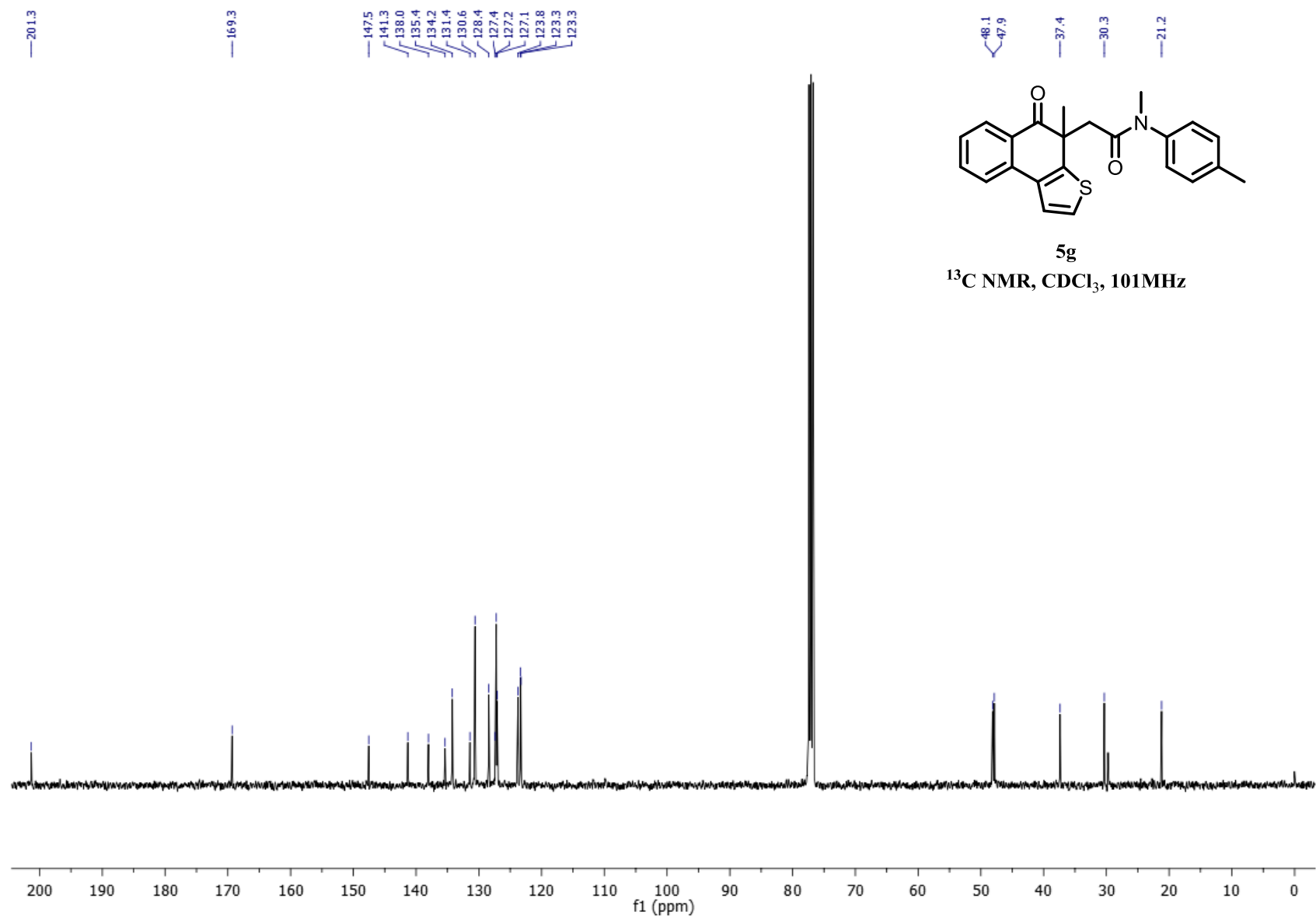


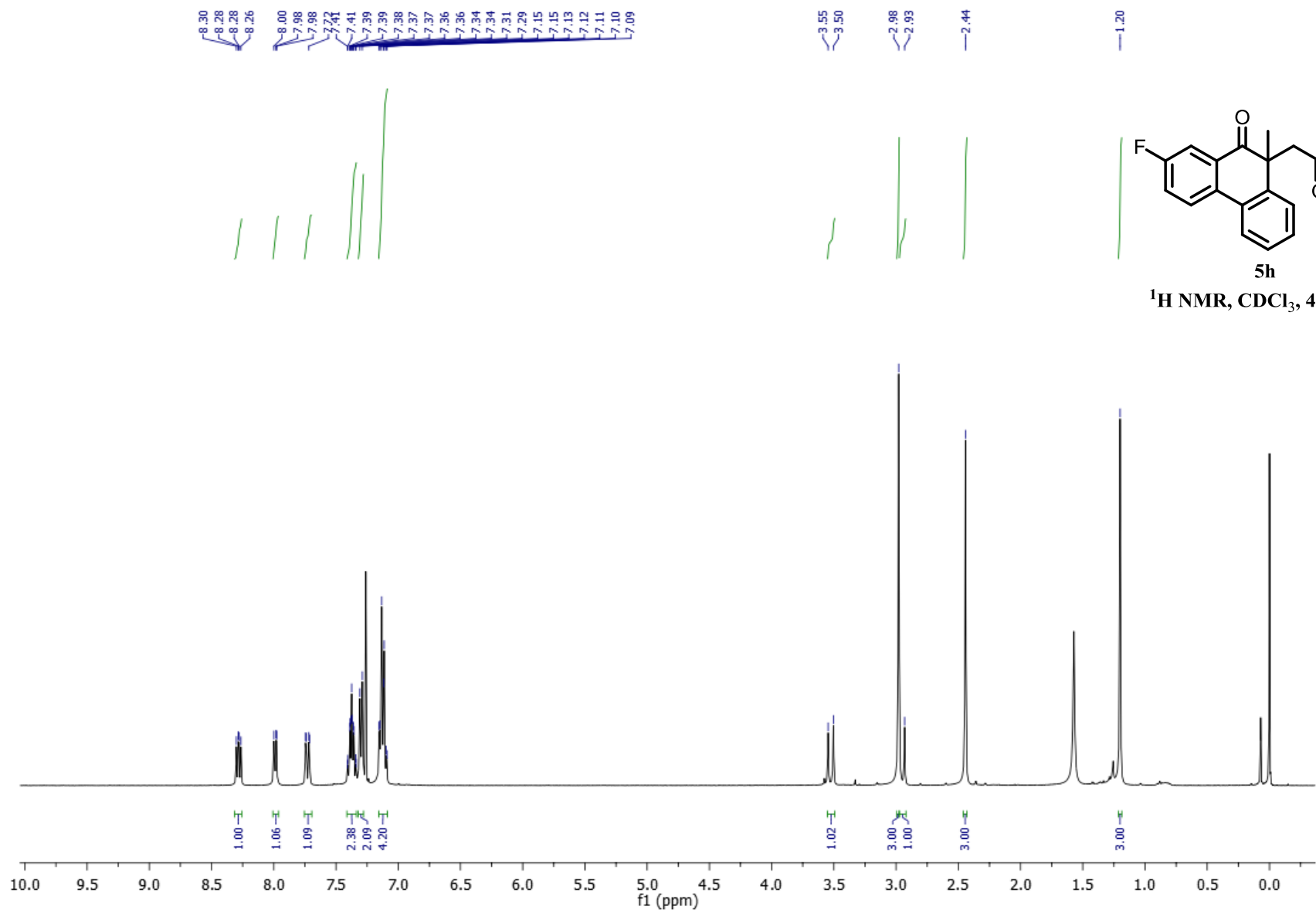


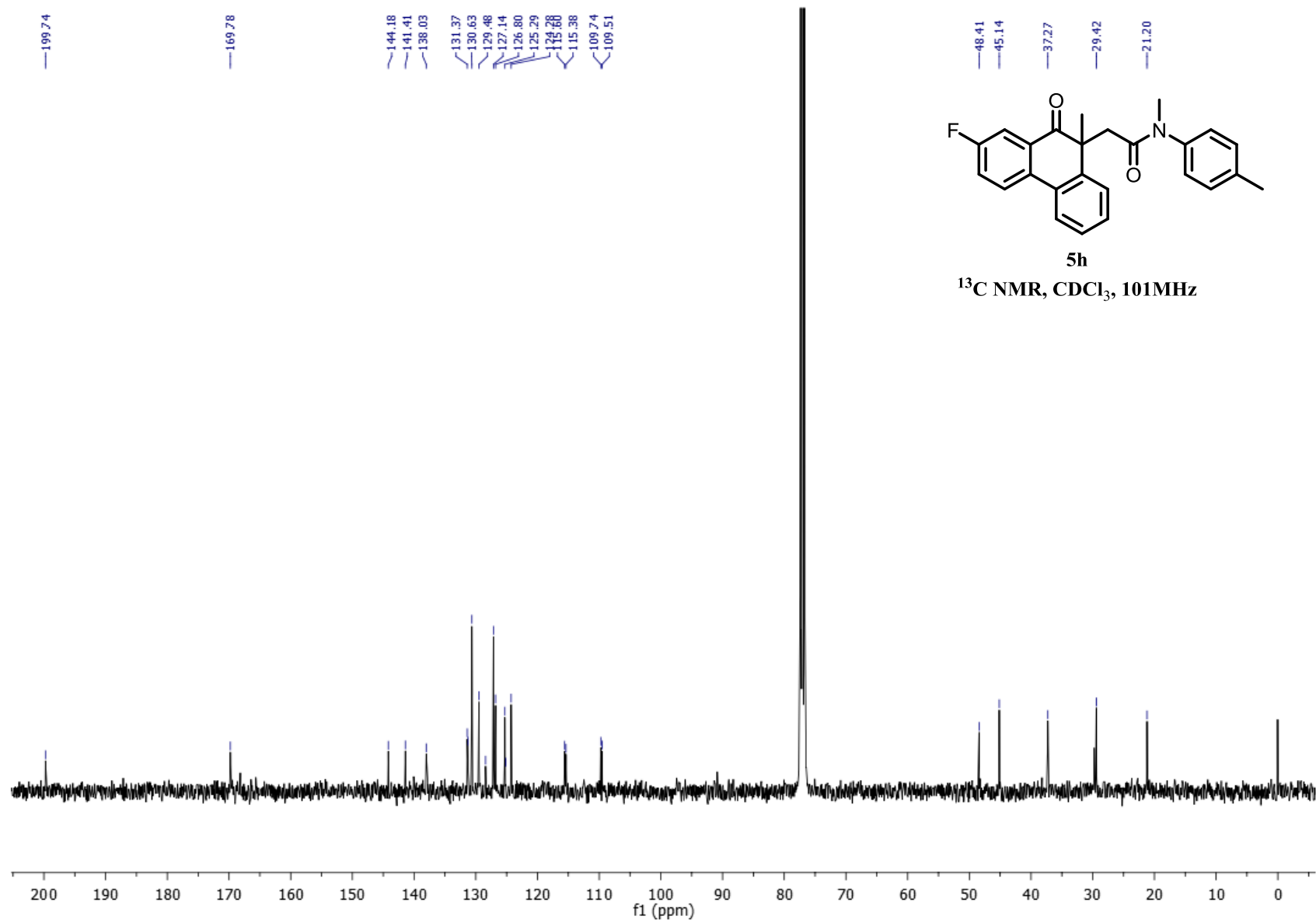


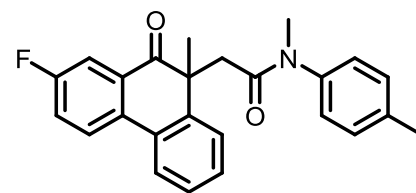






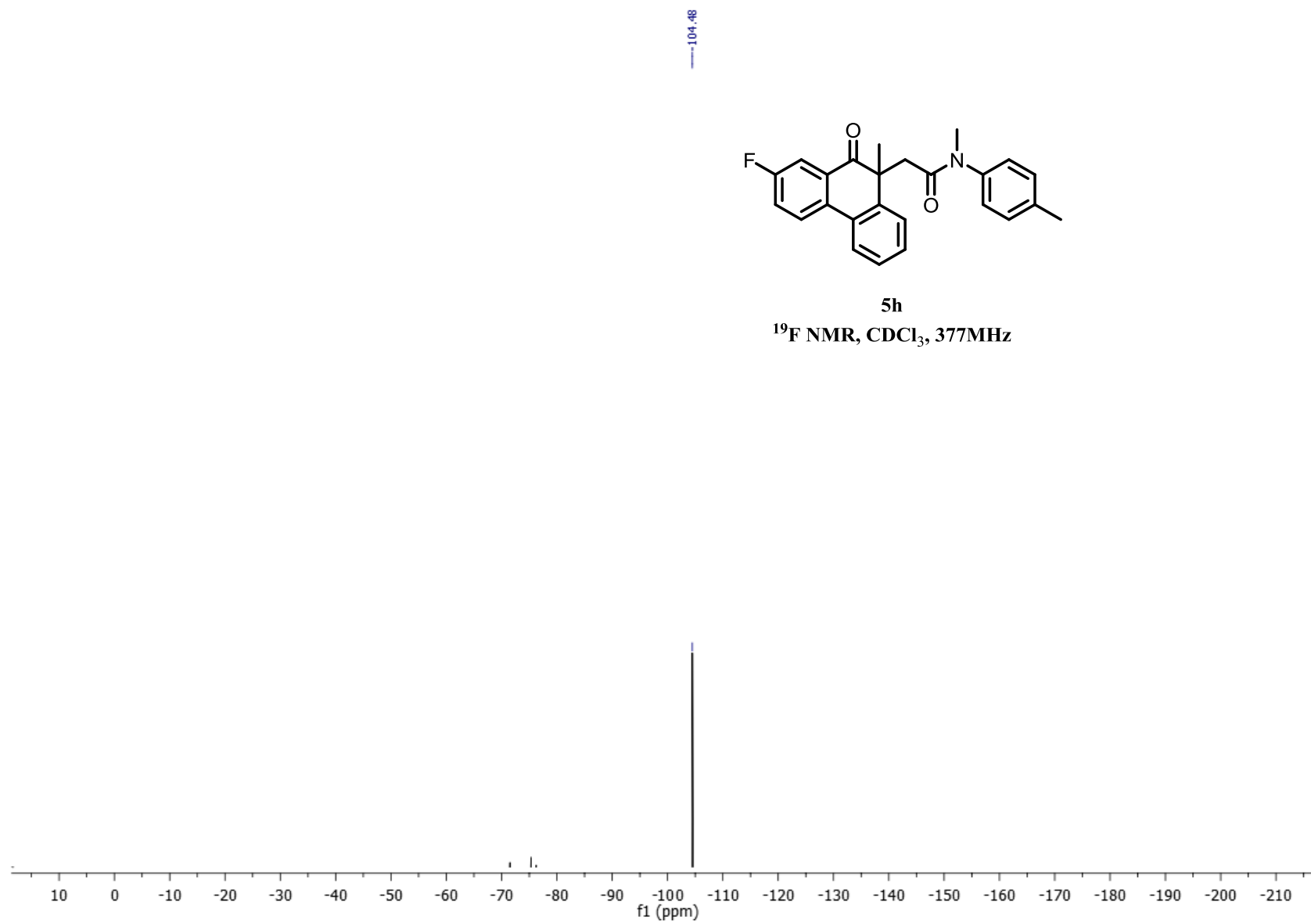






5h

^{19}F NMR, CDCl_3 , 377MHz

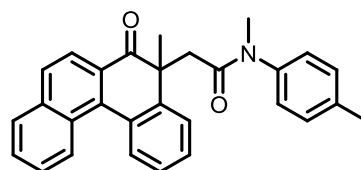


8.70
8.68
8.19
8.18
8.16
8.15
8.14
7.94
7.92
7.87
7.85
7.63
7.62
7.60
7.59
7.57
7.56
7.40
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7.39
7.31
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7.16
7.15

3.43
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3.07
2.94
2.91

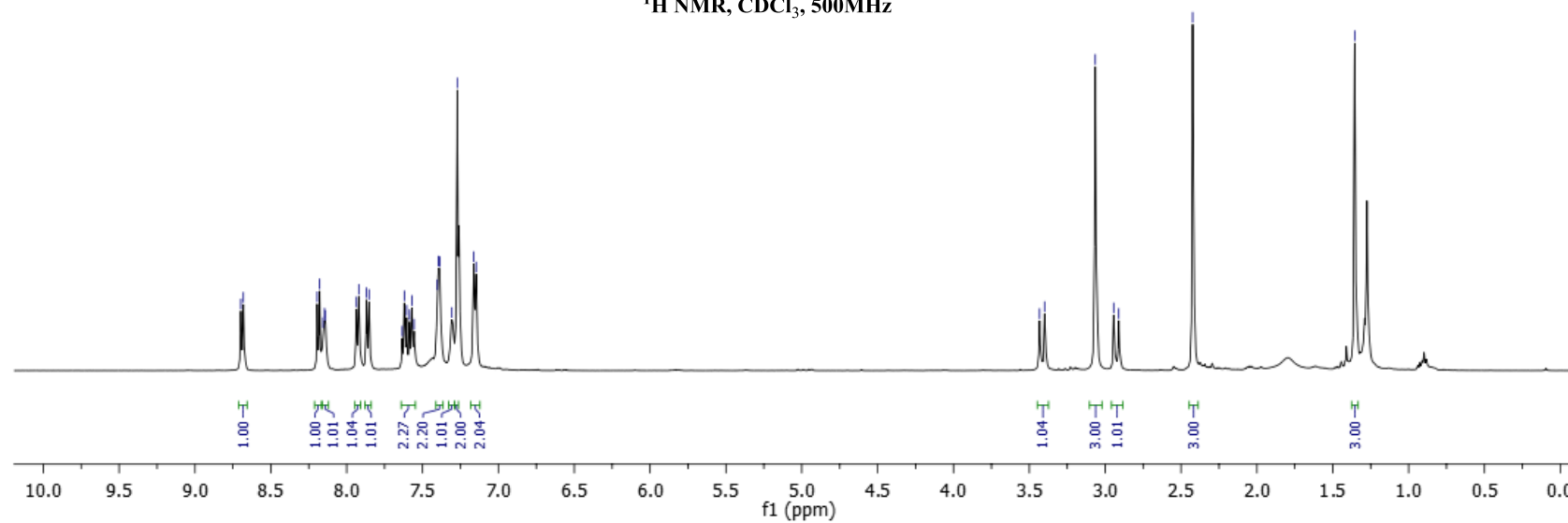
2.42

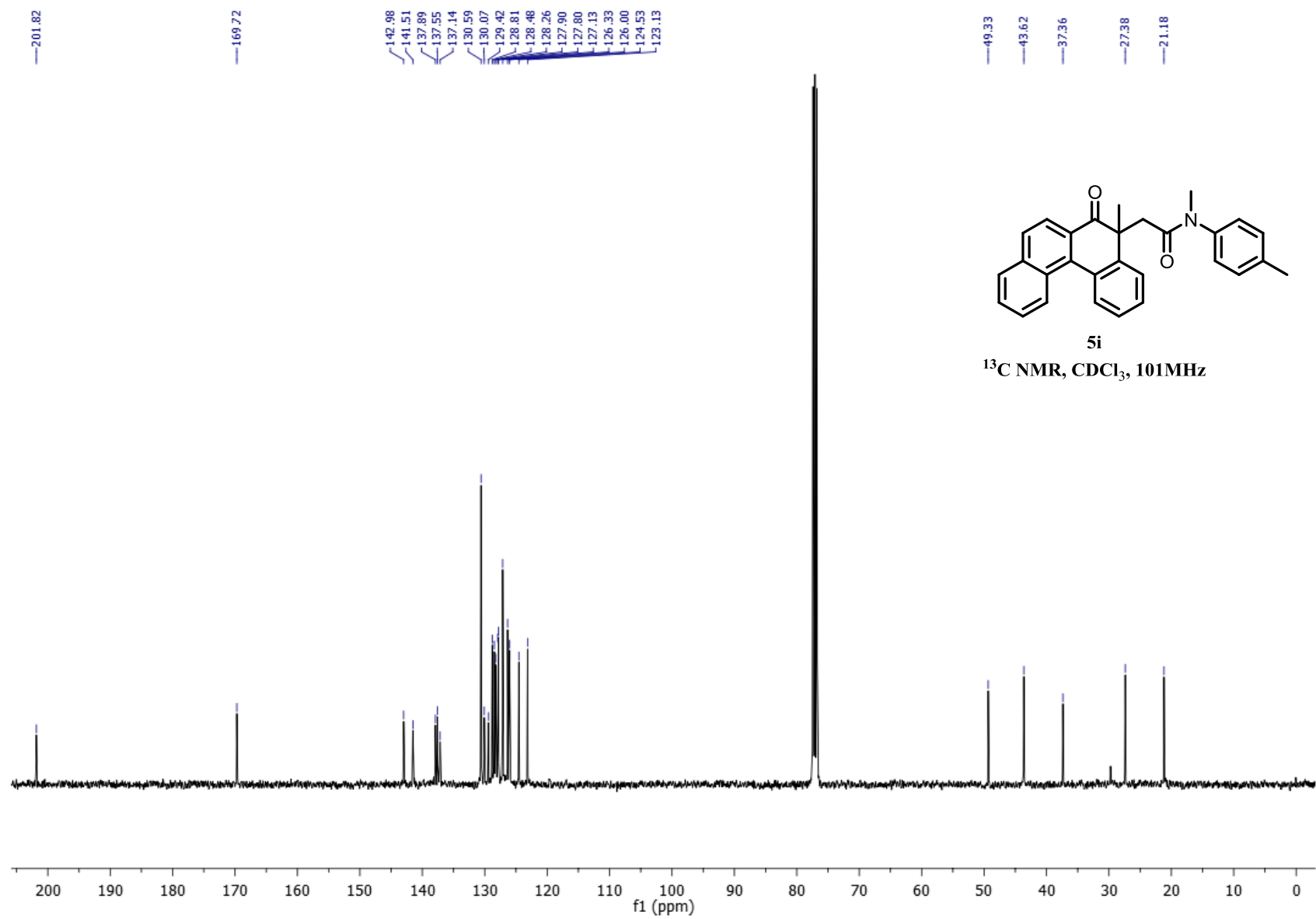
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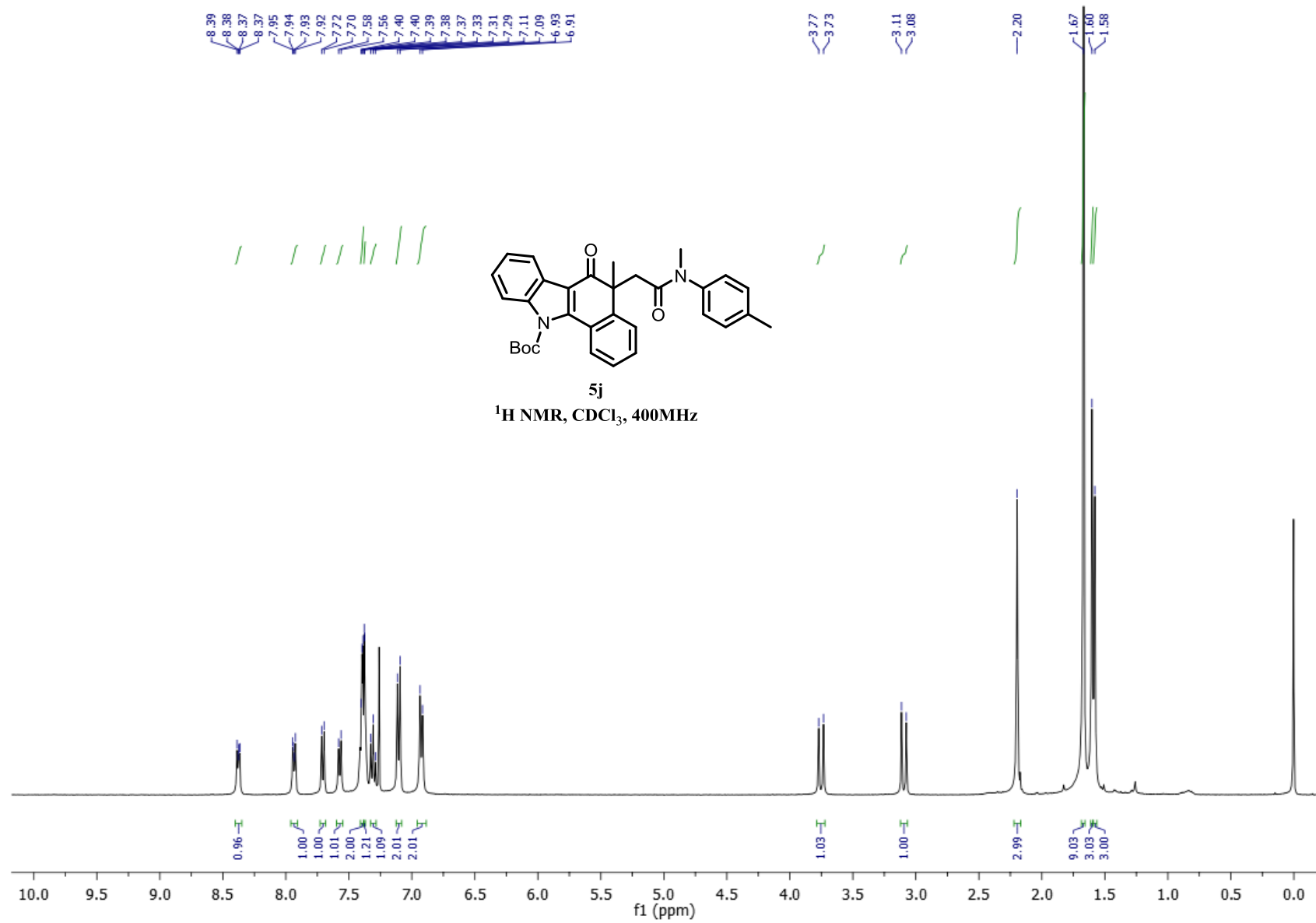


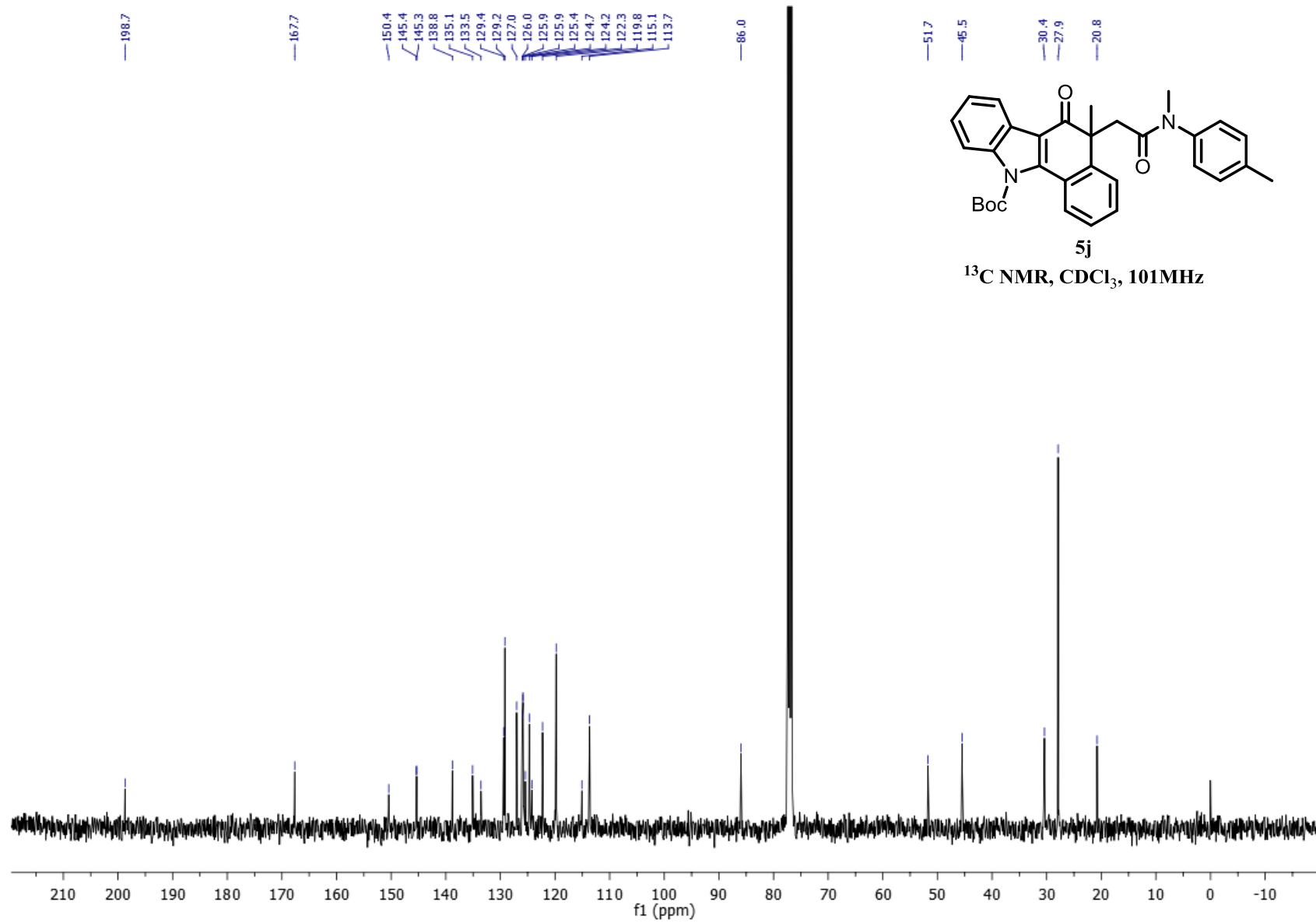
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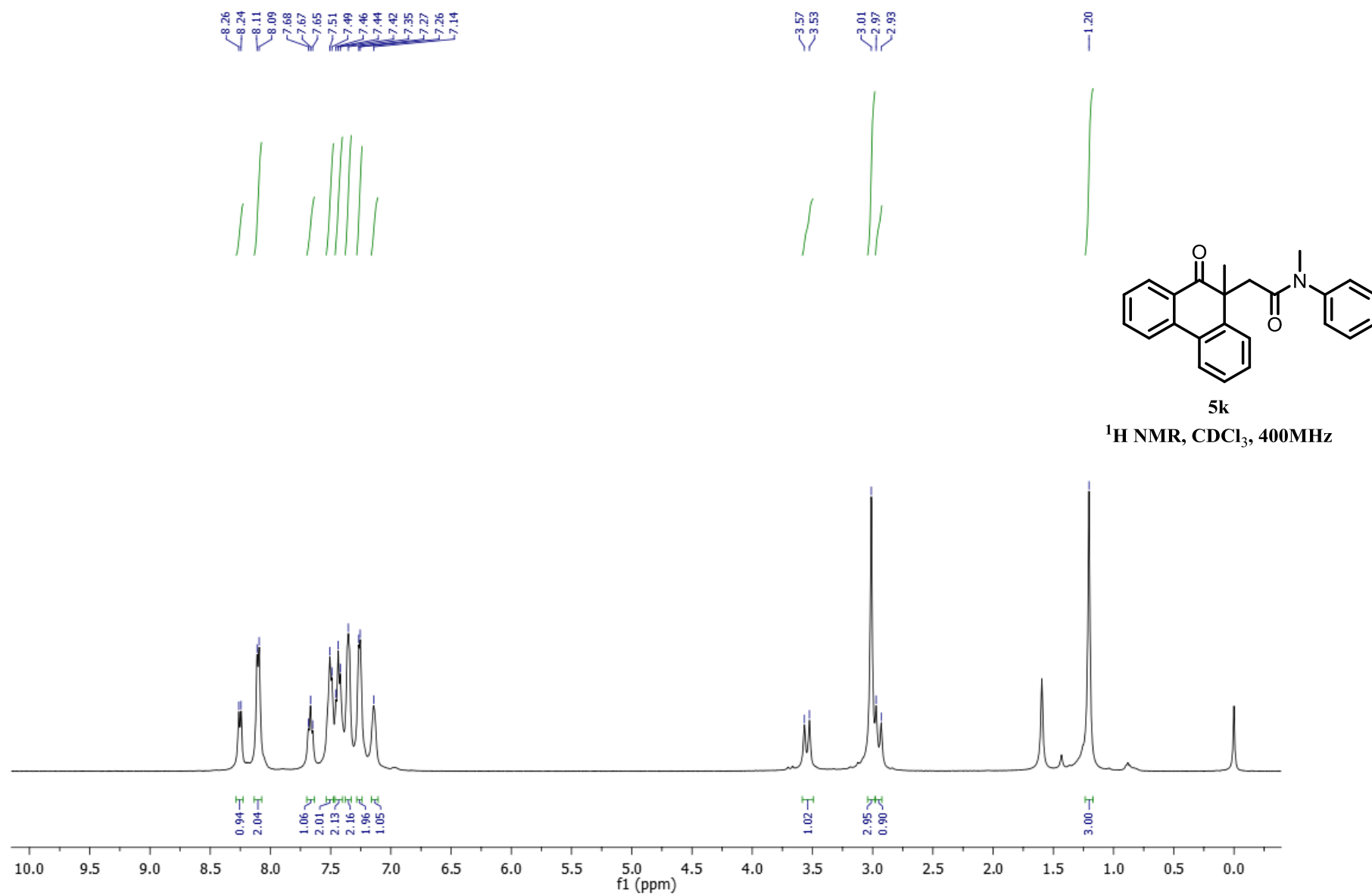
¹H NMR, CDCl₃, 500MHz

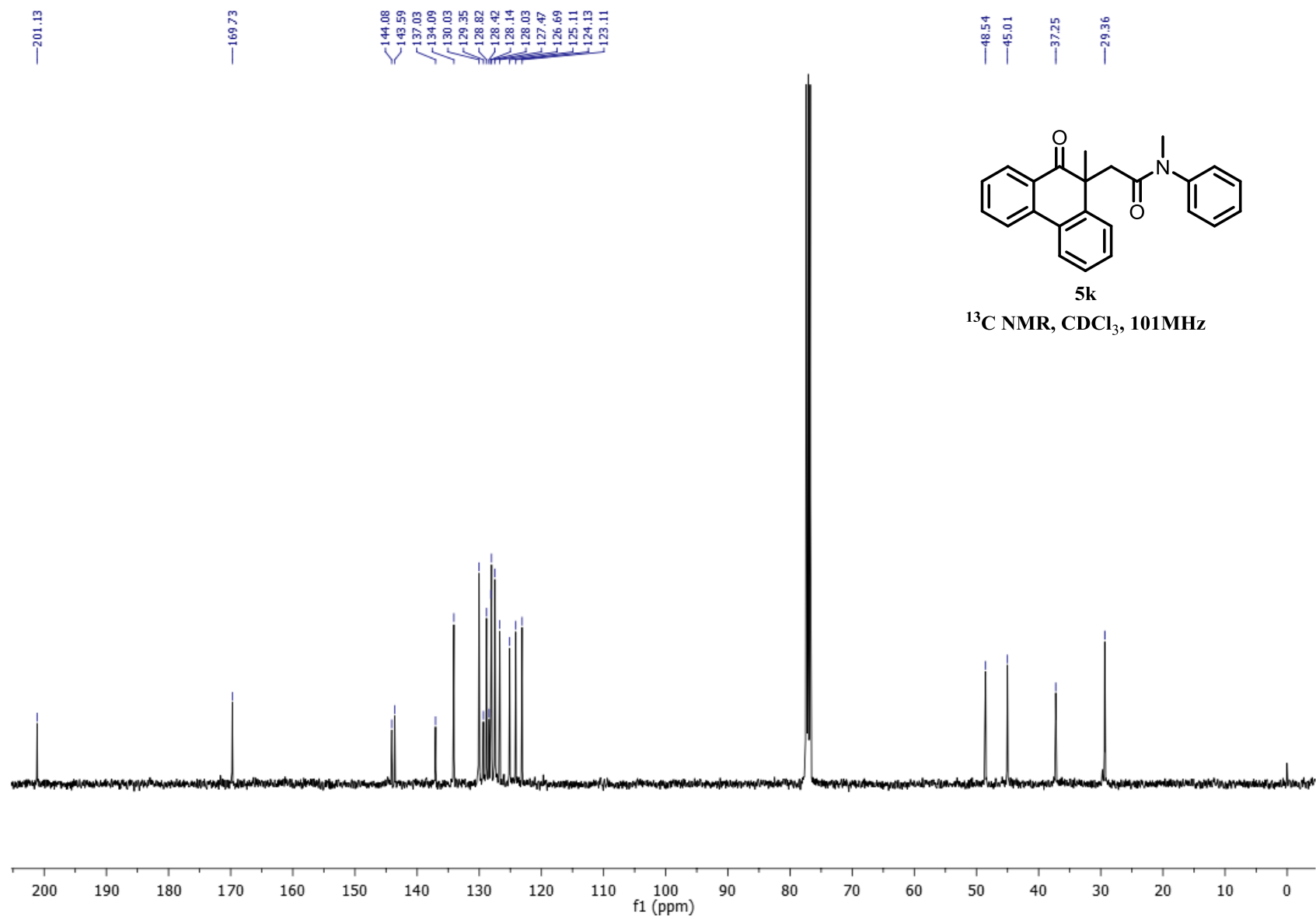


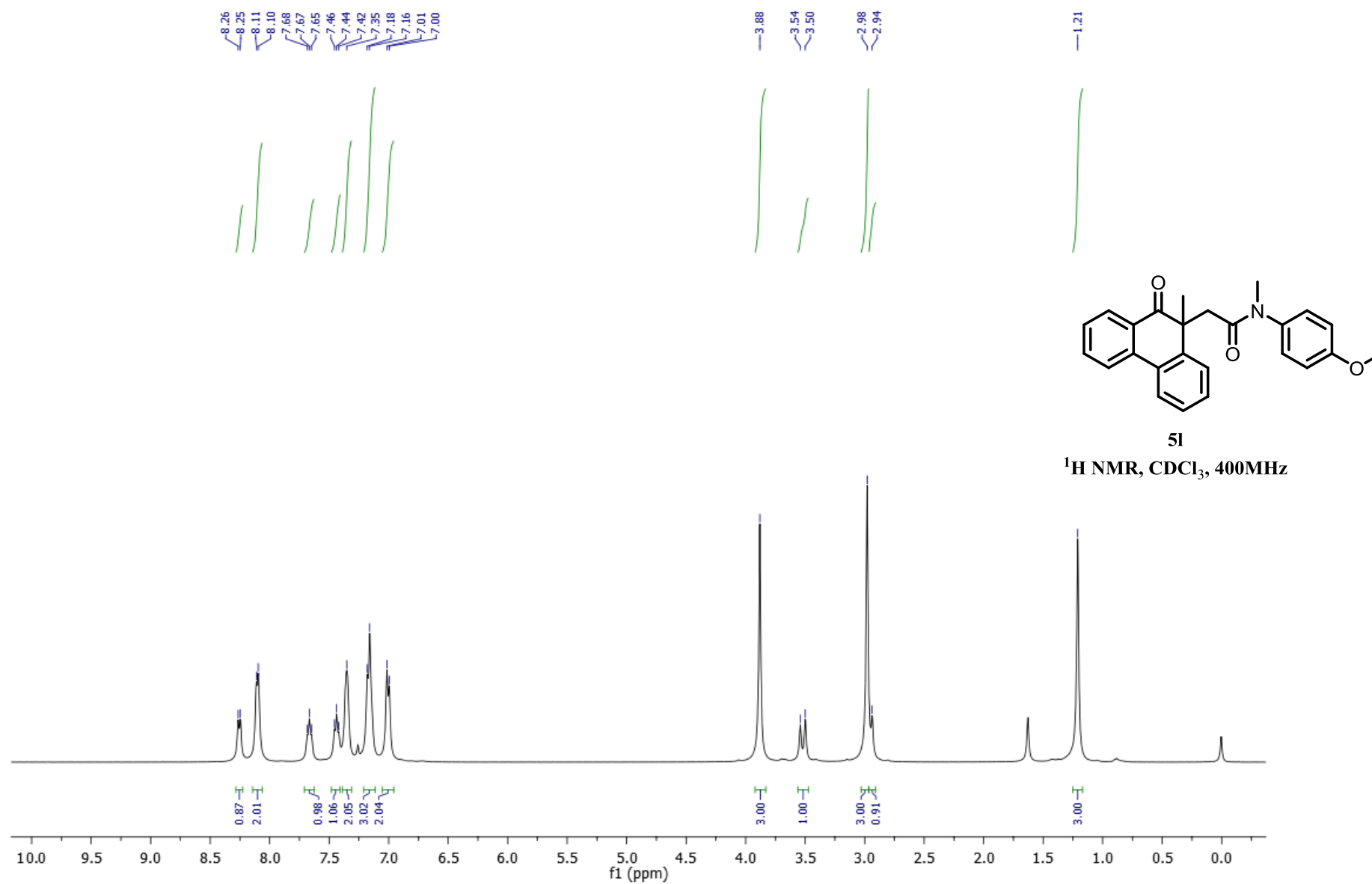


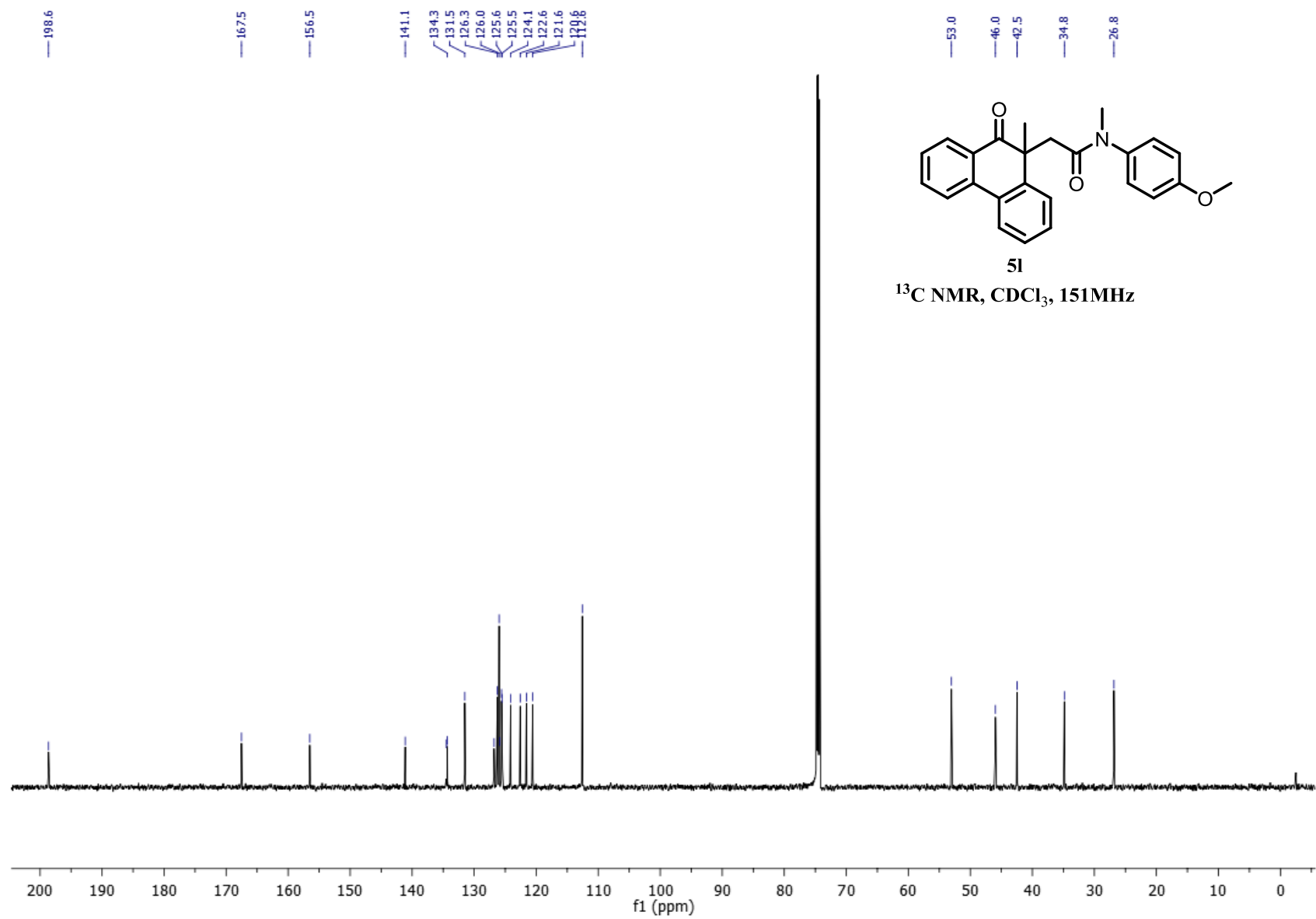


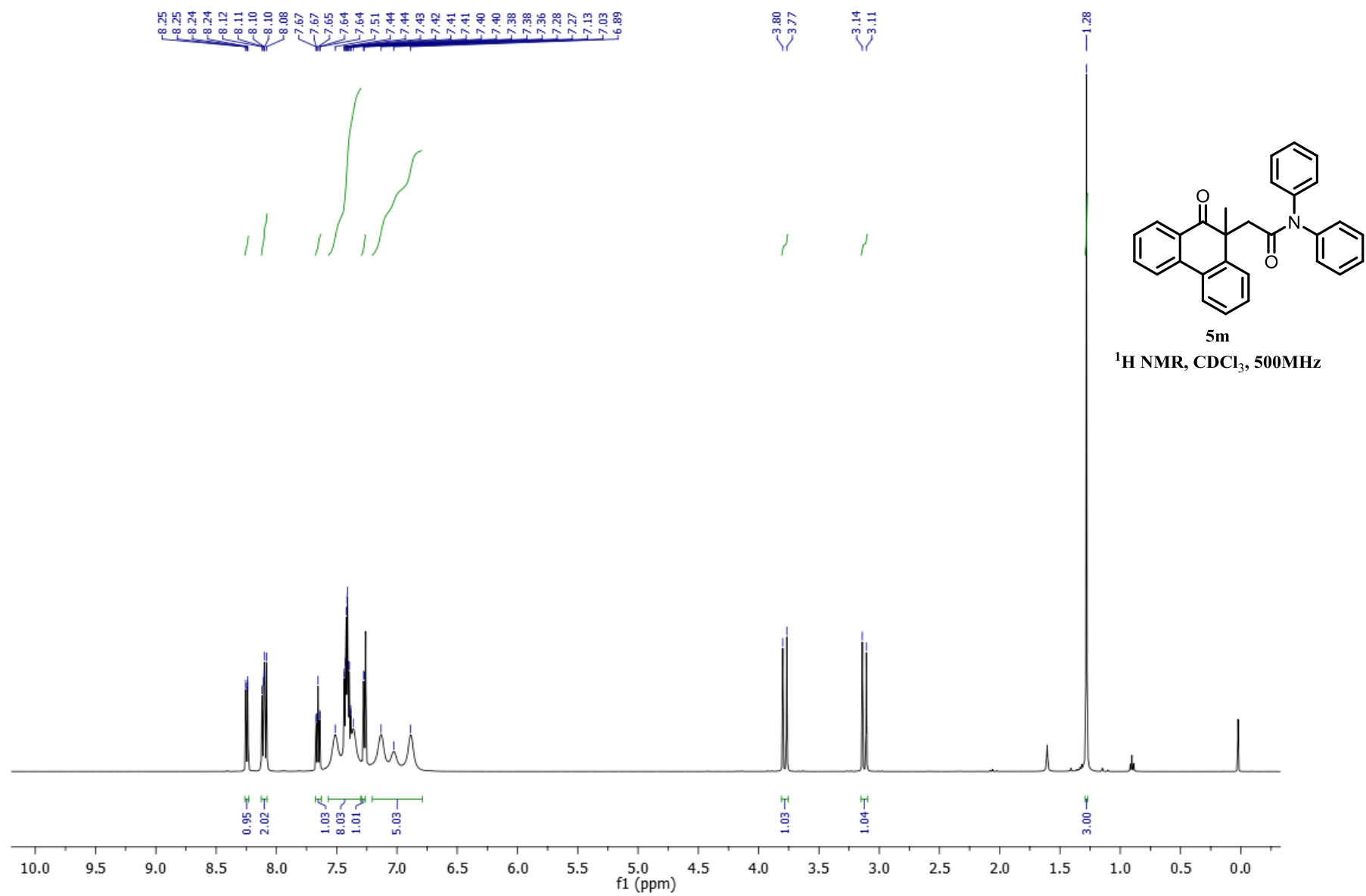


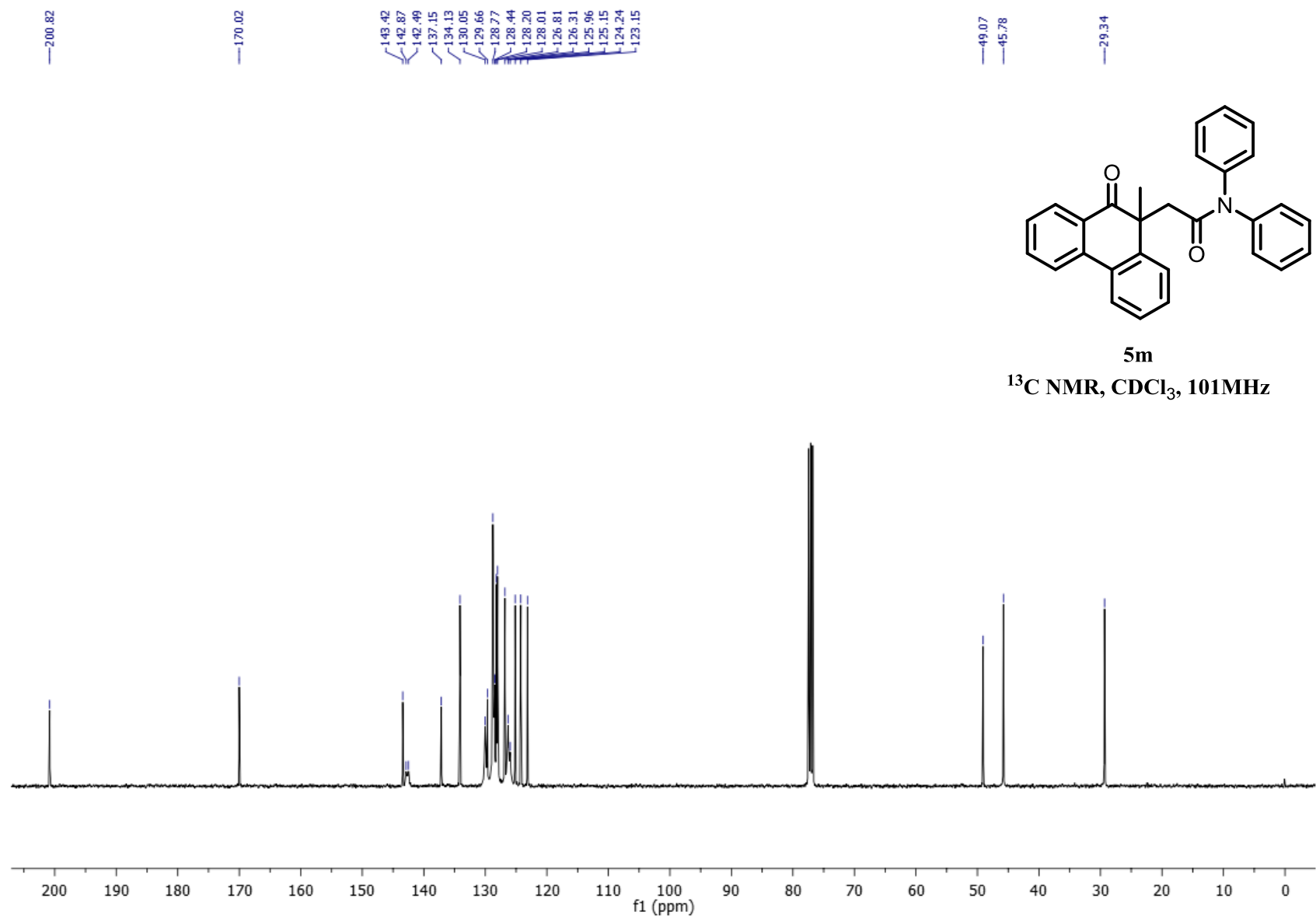


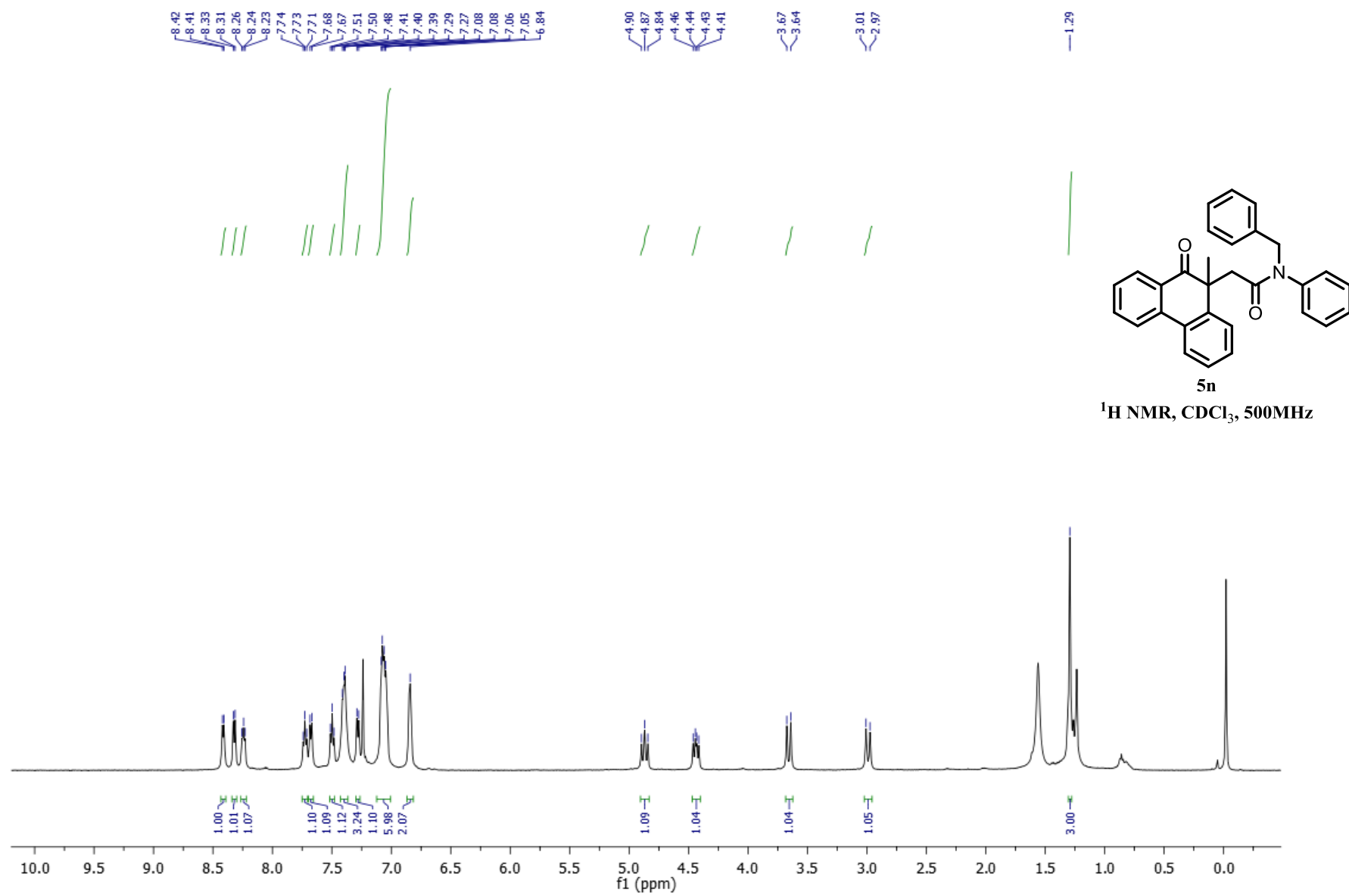


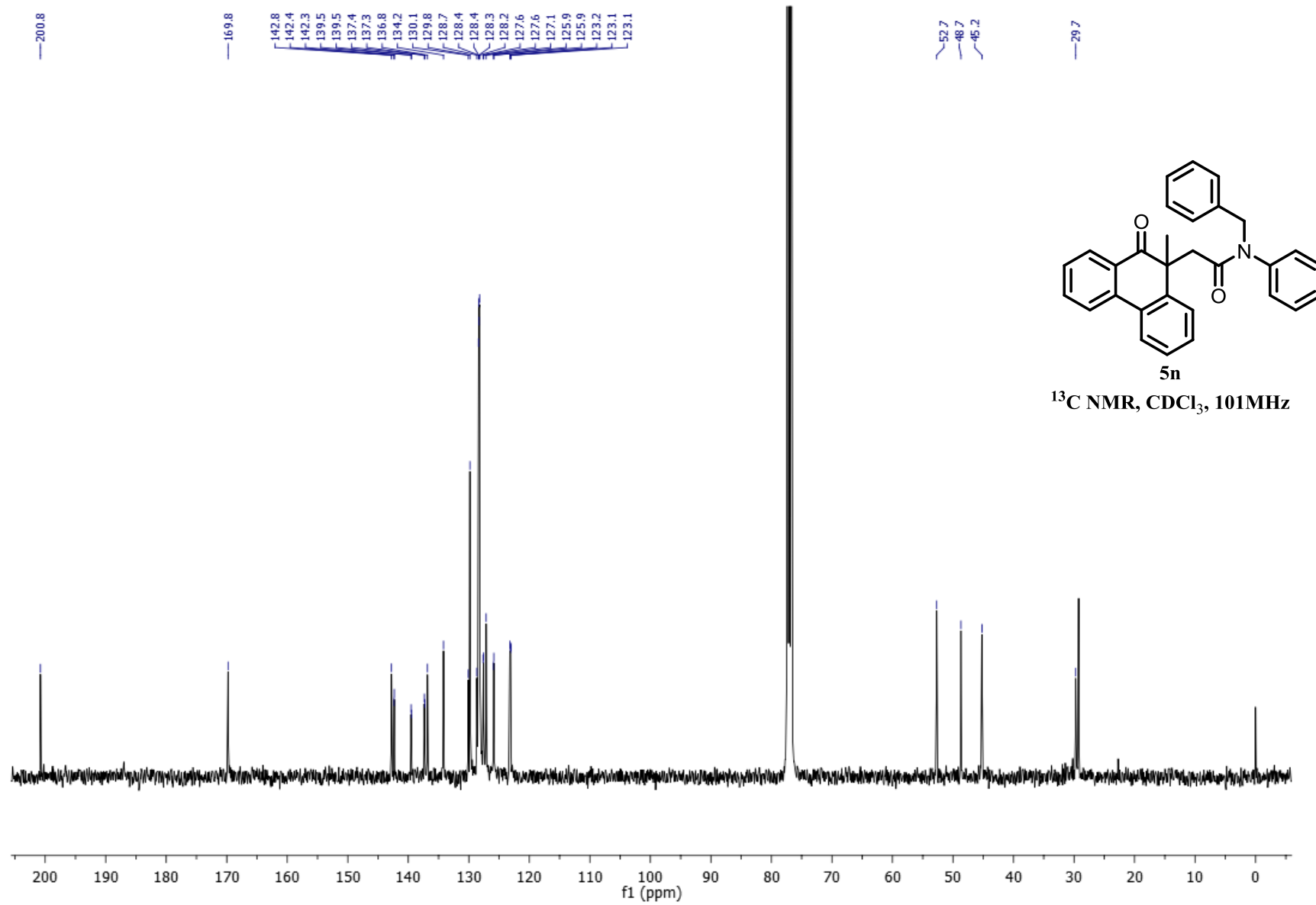


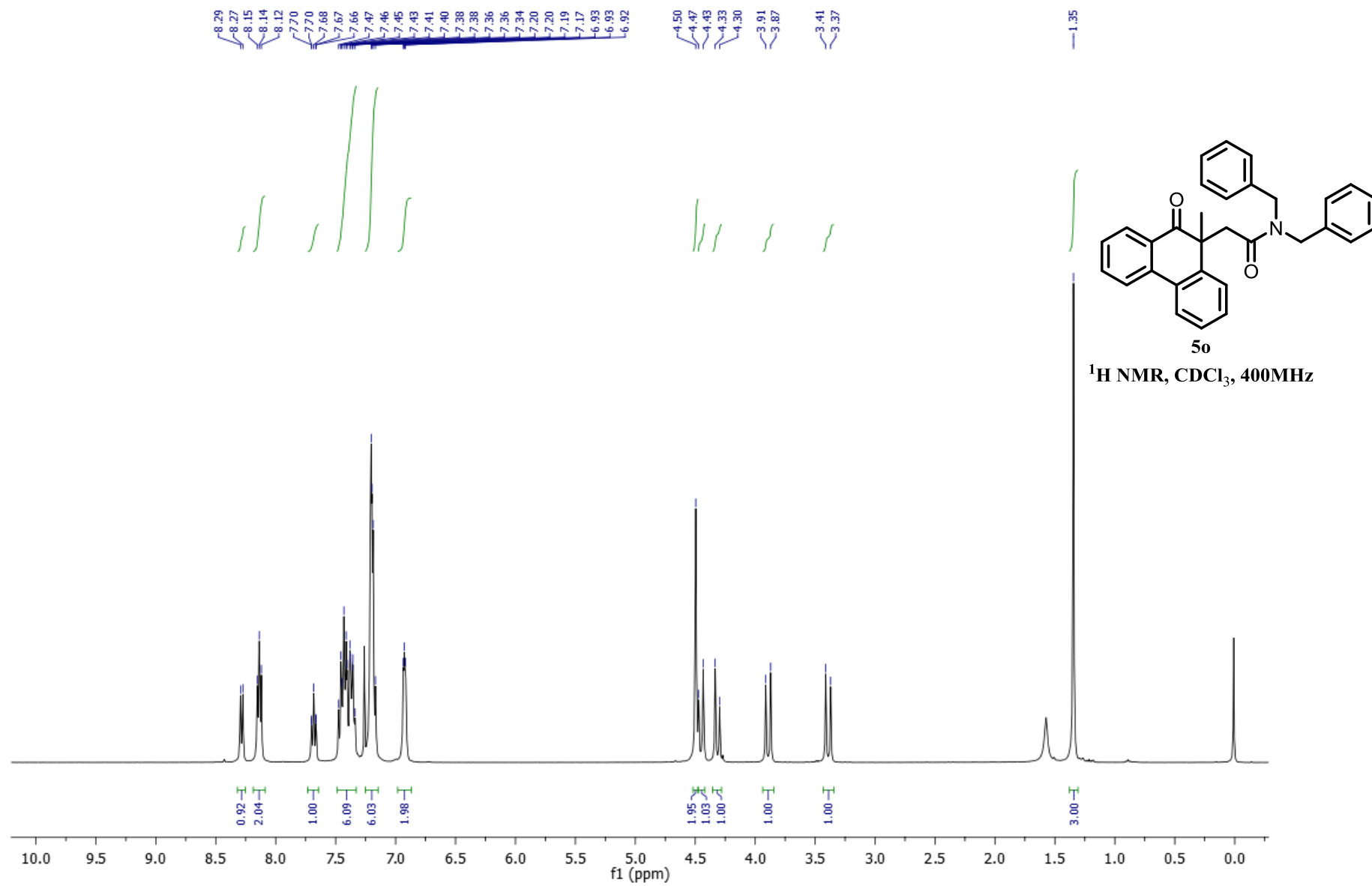


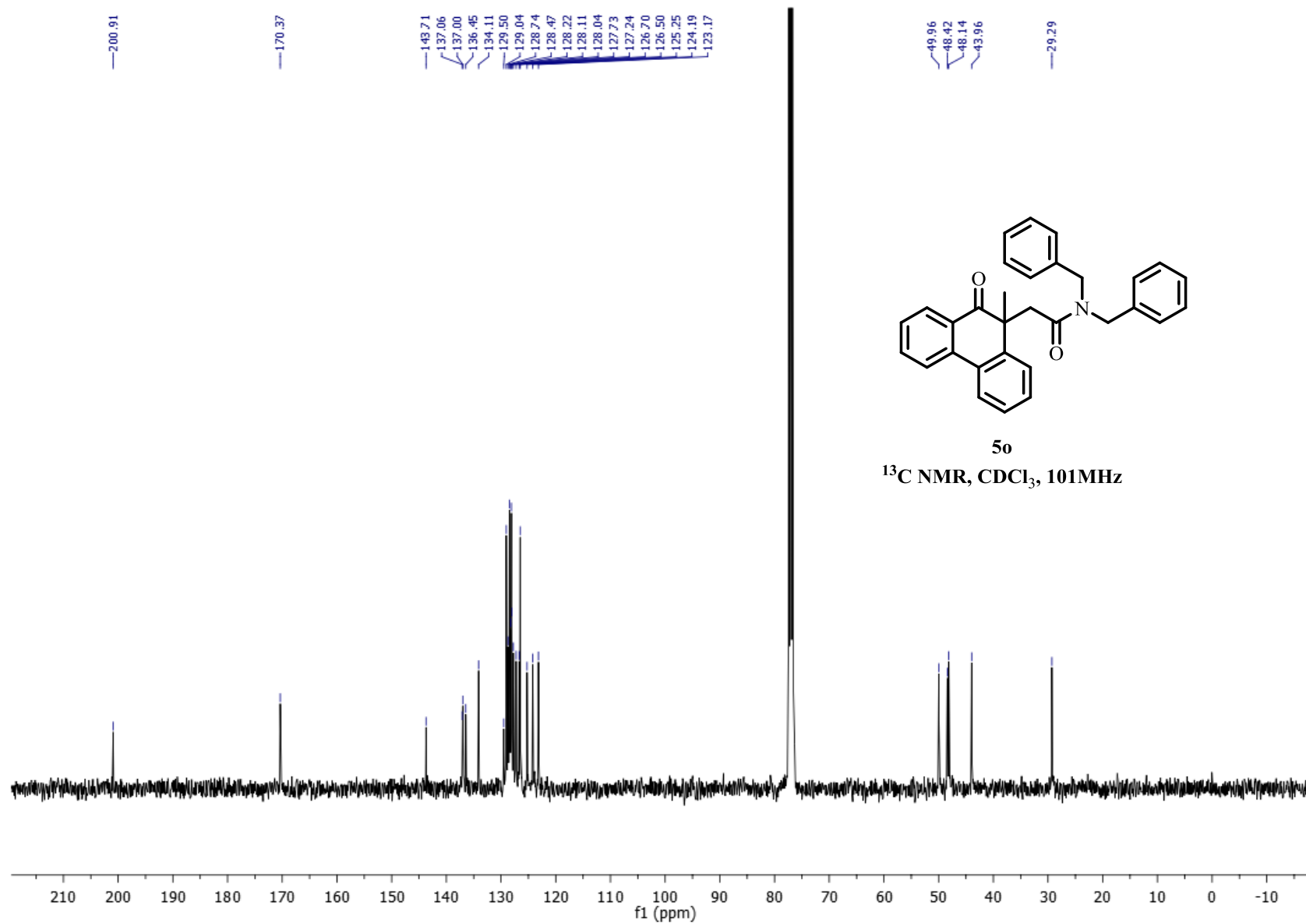


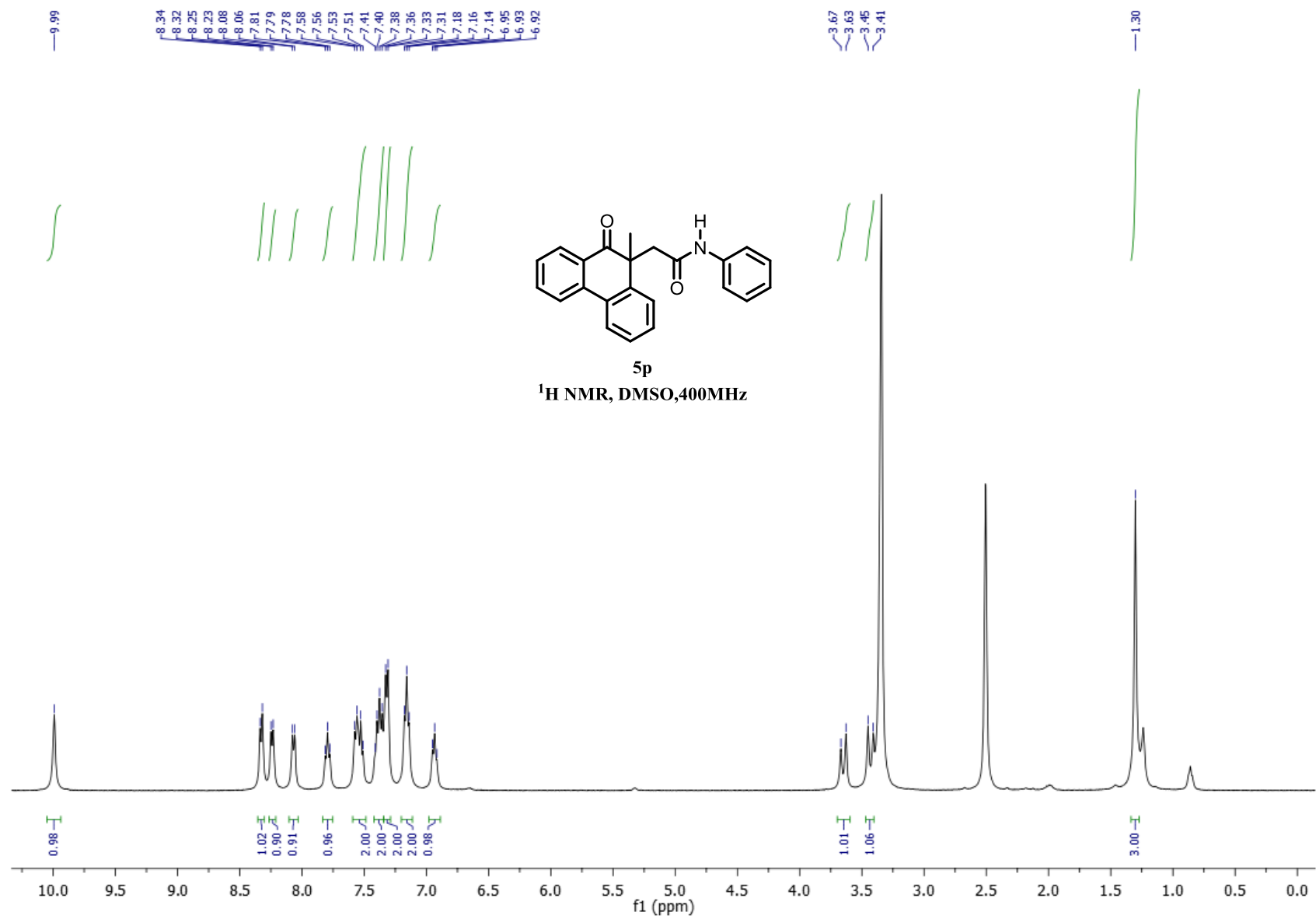


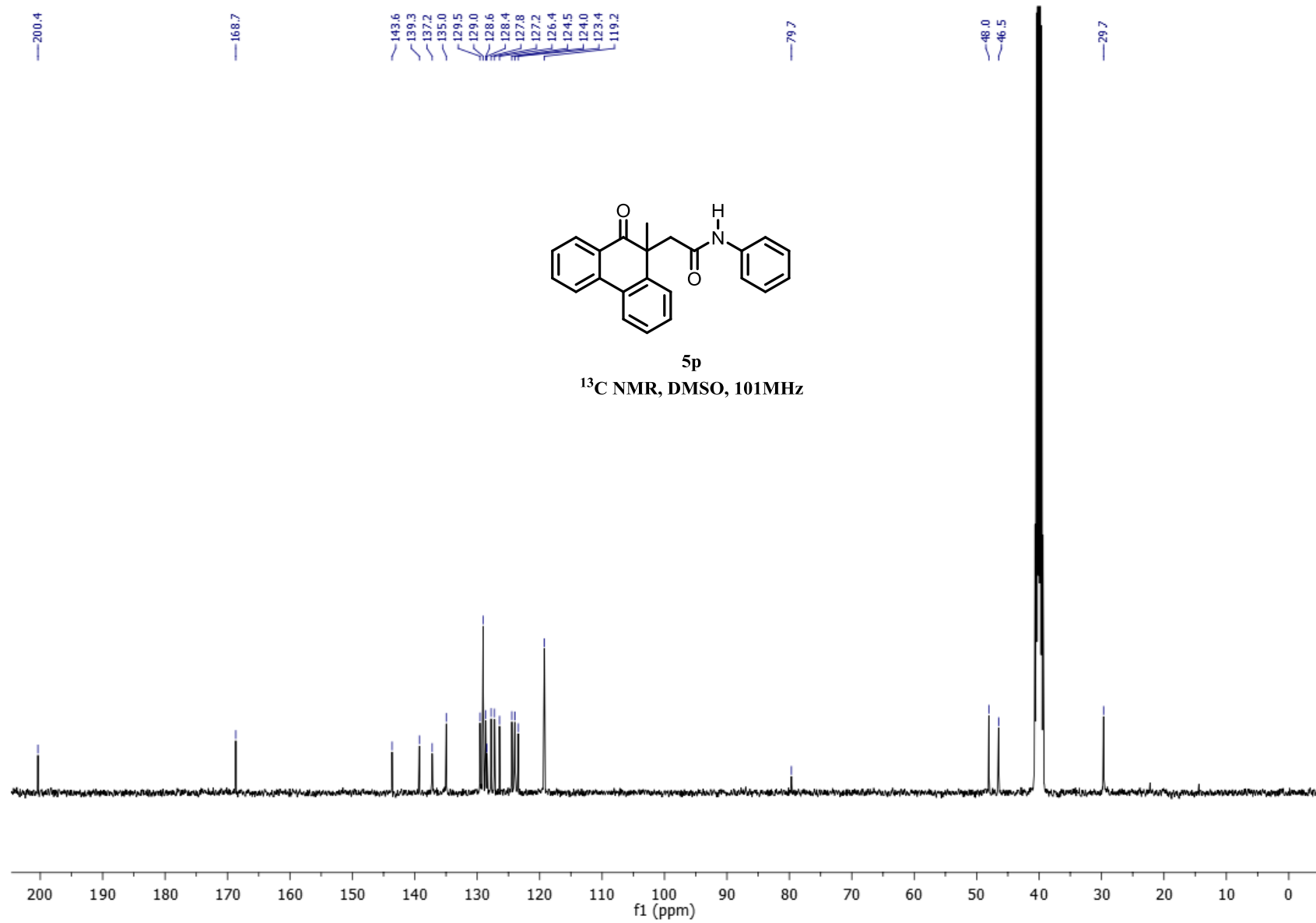


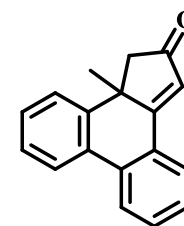
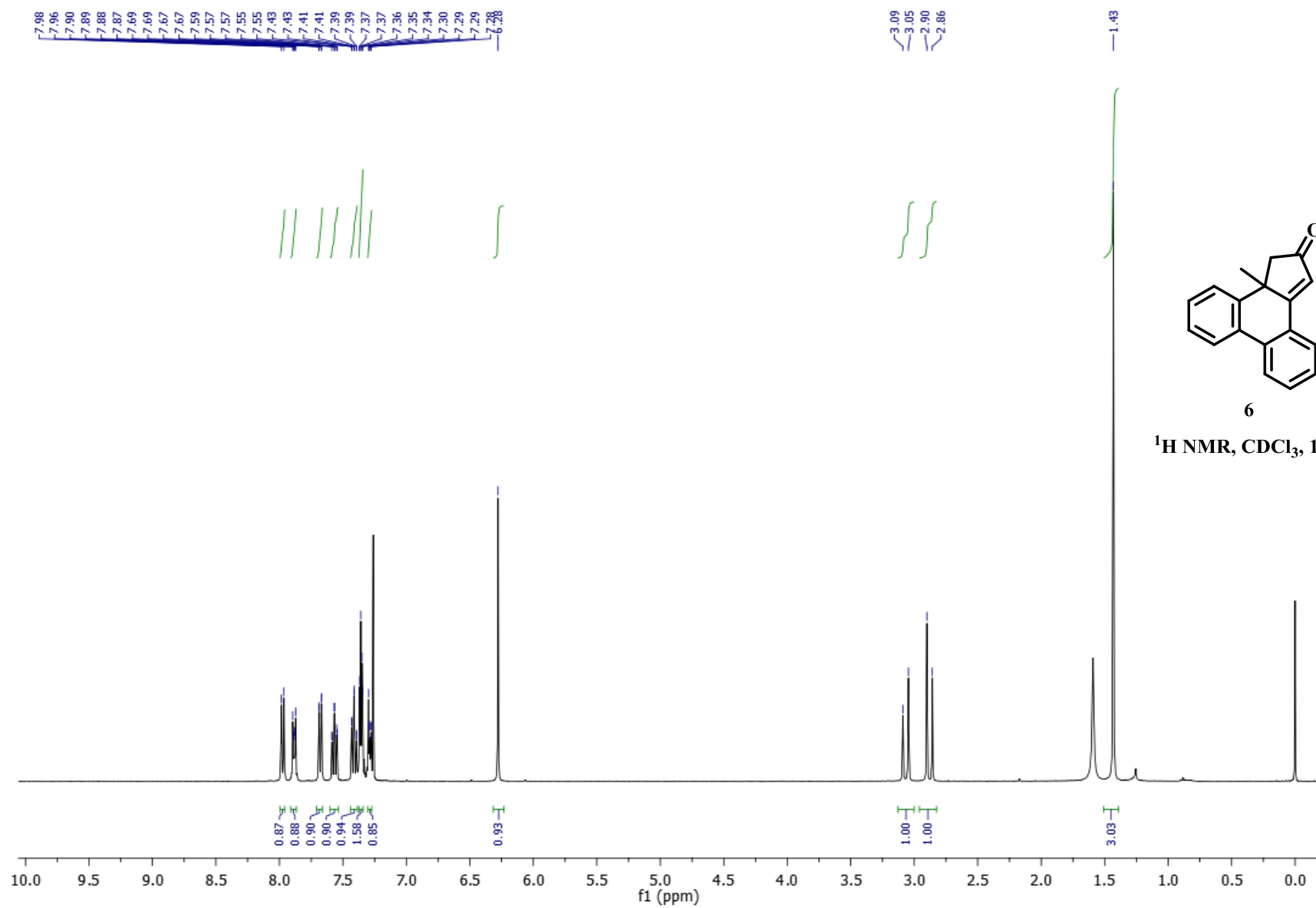






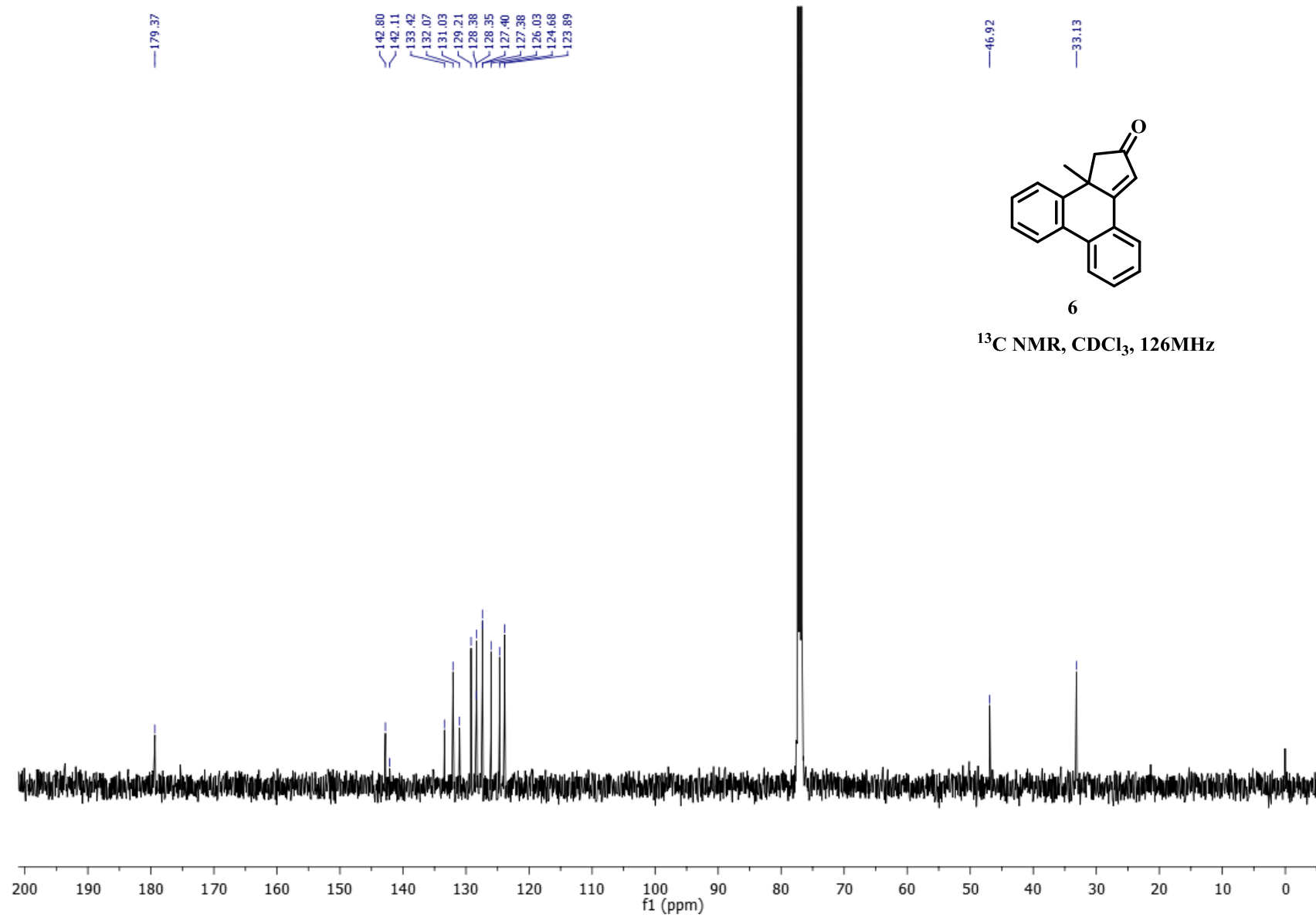


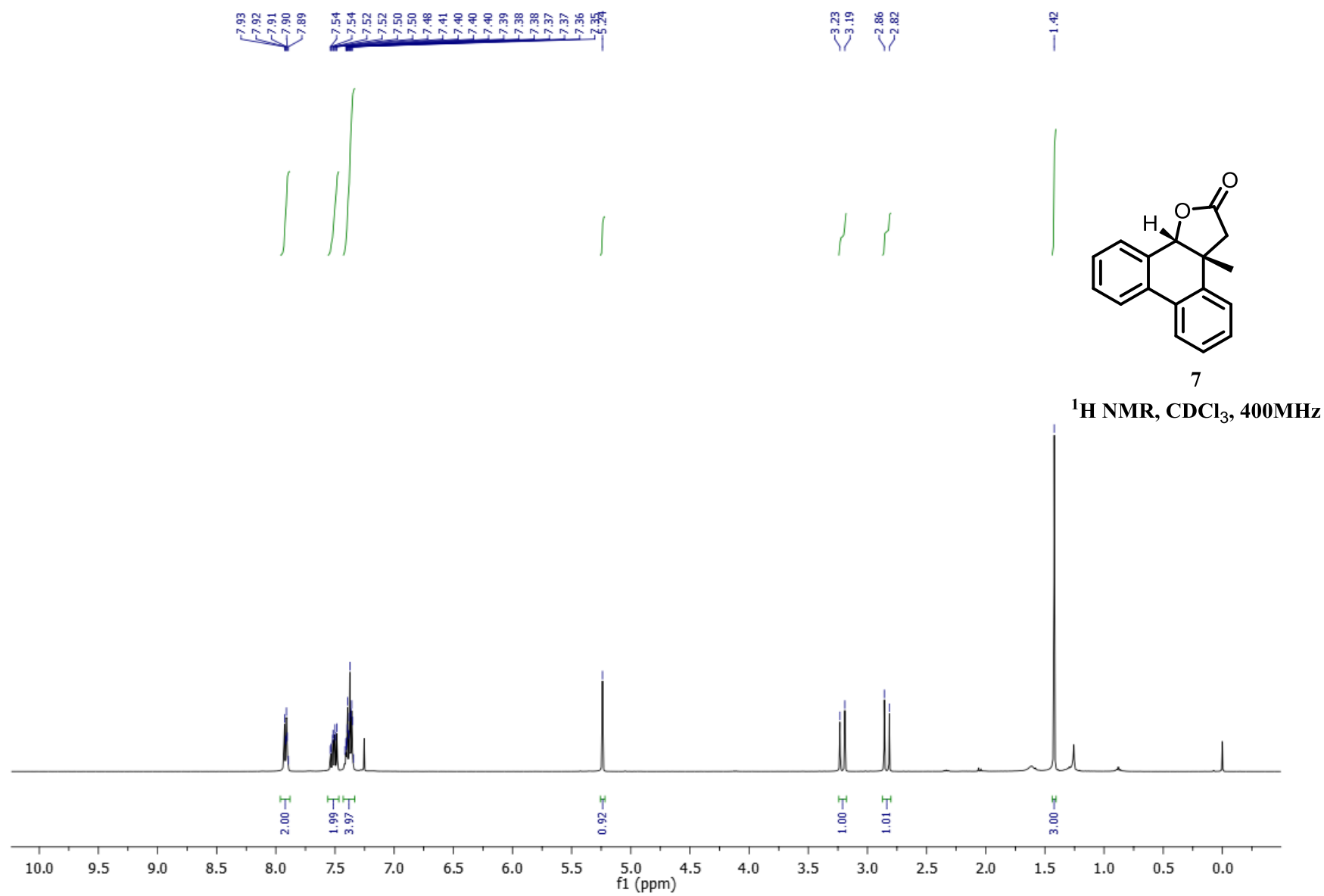


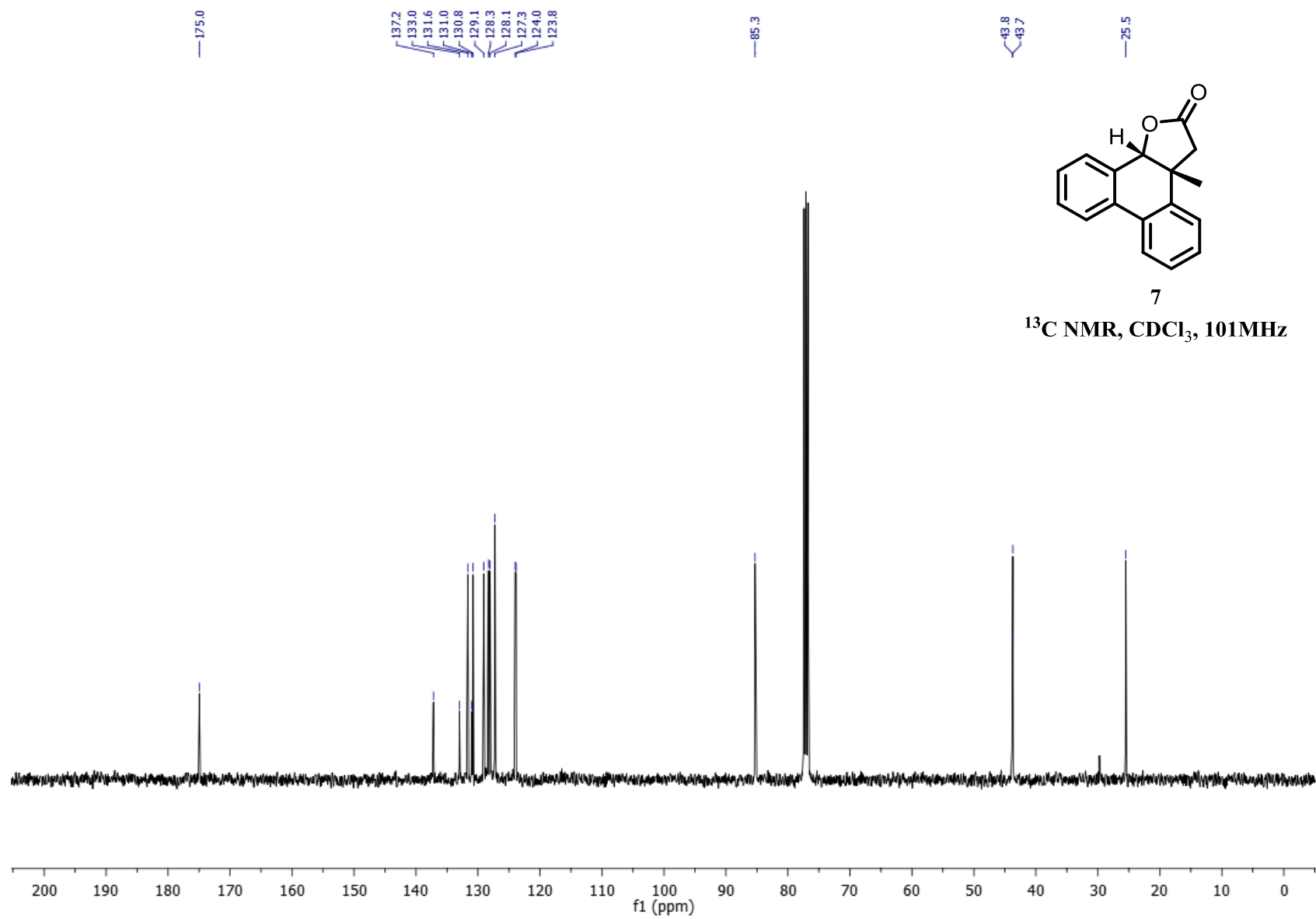


6

¹H NMR, CDCl₃, 126MHz







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DEPARTMENT OF ANALYTICAL & STRUCTURAL CHEMISTRY

05/20/22 15:16:59

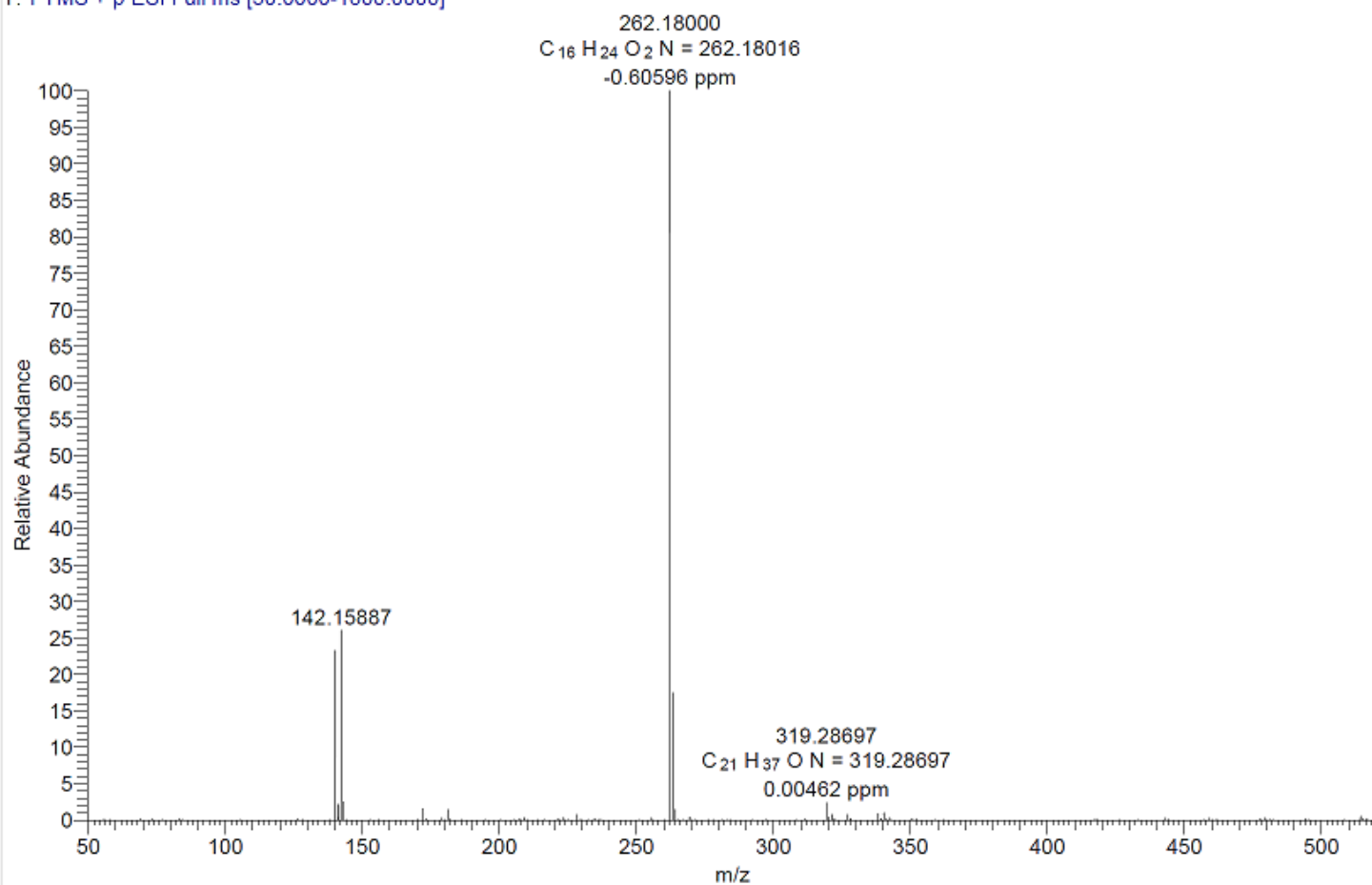
1180133589

Thermo Scientific Orbitrap Exploris 120

Analysed by G SAIKRISHNA

CRR-DK-2 #1-49 RT: 0.00-0.11 AV: 49 SB: 382 0.32-1.20 NL: 1.93E9

T: FTMS + p ESI Full ms [50.0000-1800.0000]



HRMS spectrum of TEMPO-benzoyl adduct **X-2a**