

Supporting Information for

**Dual Non-Noble-Metals-Immobilized Covalent Organic Frameworks for
Visible-Light-Driven Photocatalytic Hydrogen Evolution**

*Mengjiao Shao, Aodi Wang, Jiani Peng, Xueling Song, and Lei Wang**

School of Materials and Chemistry, University of Shanghai for Science and Technology, Shanghai
200093, China

Table of Contents

Figure S1. PXRD patterns of TpBpy-Cu/Co synthesized in a) CH ₃ CN and b) MeOH before and after recrystallization.	4
Table S1: ICP-MS data of TpBpy-Cu and TpBpy-Cu/Co.	5
Figure S2. Solid-state ¹³ C NMR spectrum of TpBpy-Cu.	5
Figure S3. a) SEM image and b) TEM image of TpBpy.	6
Figure S4. a) SEM image and b) TEM image of TpBpy-Cu.	6
Figure S5. EDS mapping images for TpBpy.	7
Figure S6. EDS mapping images for TpBpy-Cu.	7
Figure S7. XPS survey of TpBpy, TpBpy-Cu and TpBpy-Cu/Co.	8
Figure S8. O 1s XPS spectrum of a) TpBpy and b) TpBpy-Cu/Co.	8
Figure S9. Photocatalytic H ₂ evolution of TpBpy-Cu/Co with different content of Cu and Co under visible irradiation in 3 mL DMF/AcOH = 19/1. A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm ⁻²) was used as light source.	9
Figure S10. Photocatalytic H ₂ evolution of TpBpy-Cu/Co under visible irradiation in 3 mL DMF/AcOH = 19/1 or 3 mL DMF/H ₂ O = 19/1. A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm ⁻²) was used as light source.	10
Table S2. Photocatalytic HER experiments in 6 hours.	11
Figure S11. a) Time-dependent H ₂ production under visible light irradiation and b) photocatalytic H ₂ evolution initial rates (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm ⁻²) was used as light source.	12
Table S3. Photocatalytic HER experiments at low concentrations in 24 hours.	13
Figure S12. Proposed Perrin Jablonski diagram.	16
Figure S13. a) CVs of TpBpy-Cu/Co (0.05 mg) coated on electrode surface and Co(bpy)Cl ₂ under photocatalytic HER conditions (10 mL 0.1 M TBAPF ₆ /CH ₃ CN solution with 150 μL AcOH). b) CVs of TpBpy-Cu coated on electrode surface and [Cu(mestphen)(bpy)] ⁺ in 0.1M TBAPF ₆ /CH ₃ CN. c) Potential energy diagrams of TpBpy-Cu/Co catalyzed photocatalytic HER.	16
Figure S14. Control experiments of photocatalytic H ₂ evolution under visible irradiation. (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm ⁻²) was used as light source.	17
Figure S15. Light On-Off experiments for visible light-driven HER catalyzed by TpBpy-Cu/Co. (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm ⁻²) was used as light source.	18
Figure S16. EIS Nyquist plots of TpBpy-Cu/Co before and after 24 h photocatalysis.	19

Figure S17. a) XPS survey, b) N 1s, c) Cu 2p, and d) Co 2p spectrum of TpBpy–Cu/Co after 24 h photocatalysis.....	20
Figure S18. ^1H NMR spectrum of $[\text{Cu}(\text{mestphen})(\text{bpy})]^+$ in $\text{DMSO}-d_6$	21
Figure S19. ^{13}C NMR spectrum of $[\text{Cu}(\text{mestphen})(\text{bpy})]^+$ in $\text{DMSO}-d_6$	21
Figure S20. HRMS (ESI, positive mode) spectra of CobpyCl ₂ in CH ₃ CN.	22
References.....	23

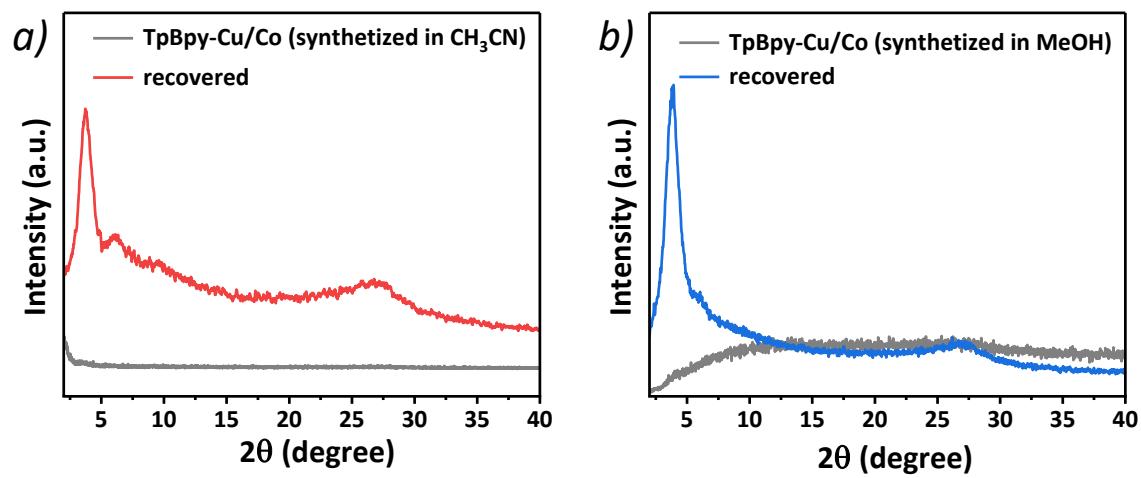


Figure S1. PXRD patterns of TpBpy-Cu/Co synthesized in a) CH_3CN and b) MeOH before and after recrystallization.

Table S1: ICP-MS data of TpBpy-Cu and TpBpy-Cu/Co.

Sample	Cu(wt%)	Co(wt%)	Cu/Co molar ratio
TpBpy-Cu	3.51	/	/
TpBpy-Cu/Co	3.45	1.77	1.8:1
TpBpy-Cu/Co-2	3.24	2.84	1.1:1
TpBpy-Cu/Co-3	5.92	1.93	2.8:1

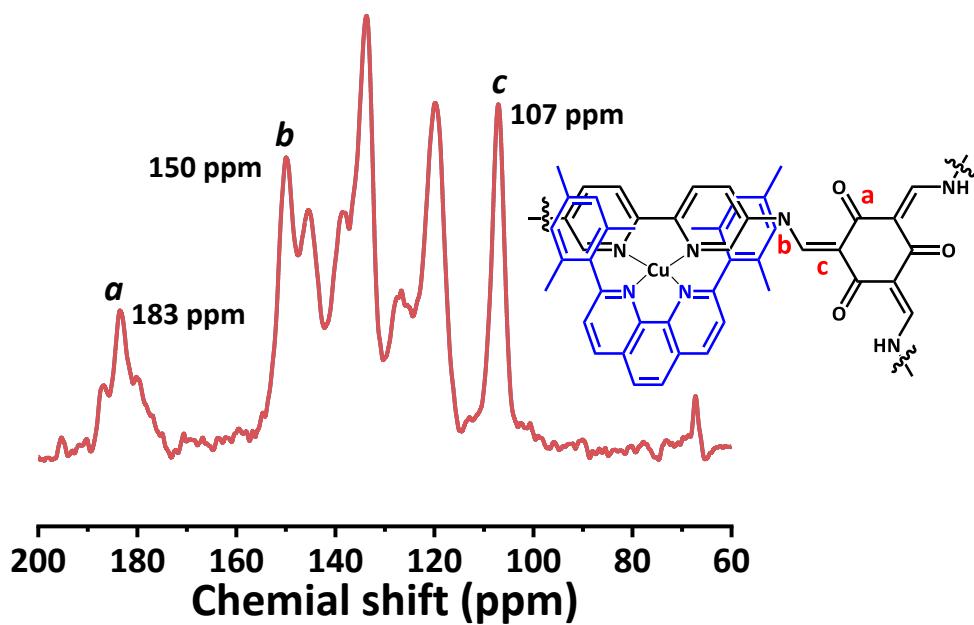


Figure S2. Solid-state ^{13}C NMR spectrum of TpBpy-Cu.

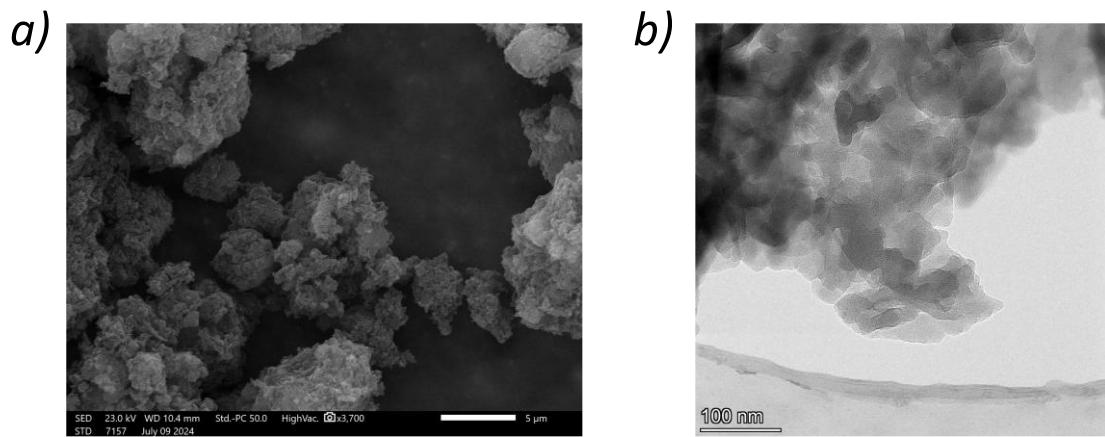


Figure S3. a) SEM image and b) TEM image of TpBpy.

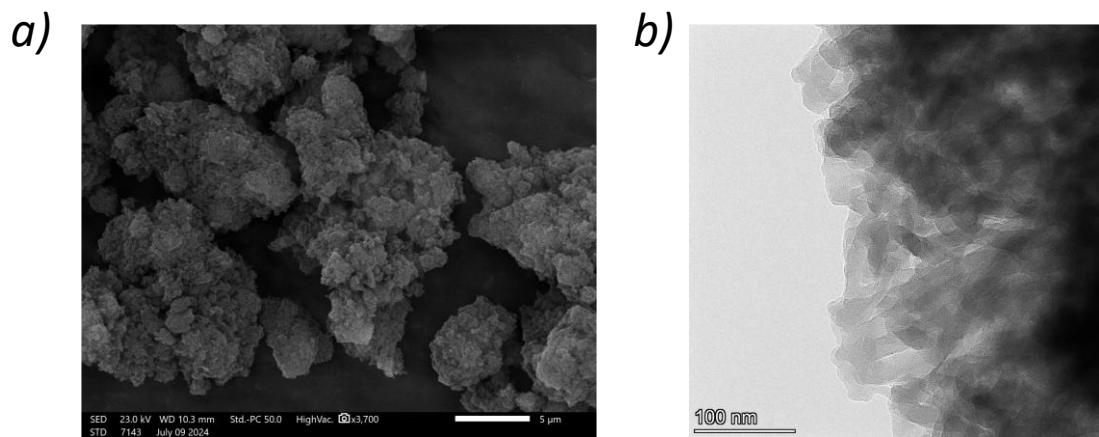


Figure S4. a) SEM image and b) TEM image of TpBpy-Cu.

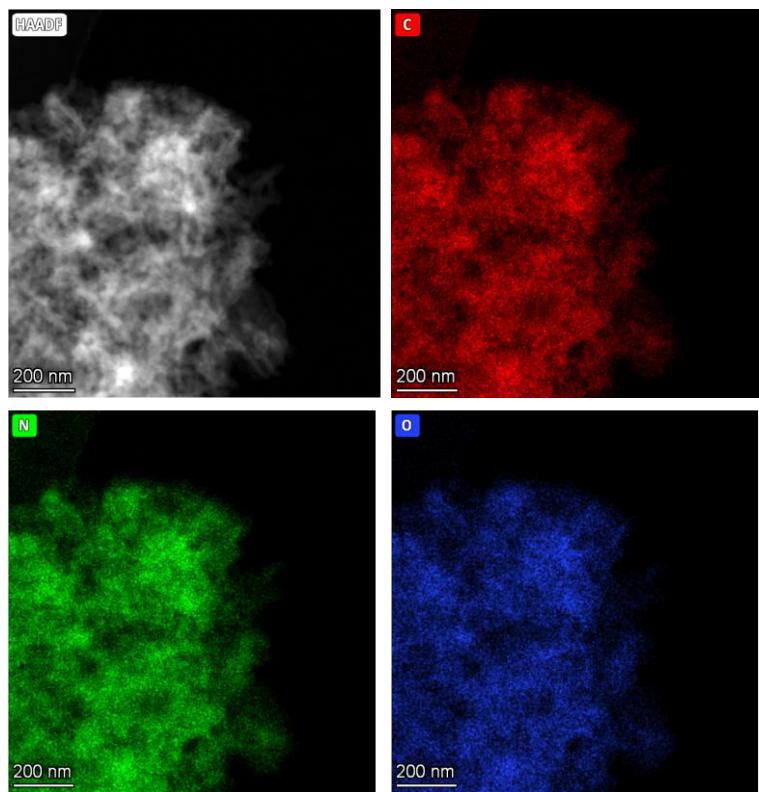


Figure S5. EDS mapping images for TpBpy.

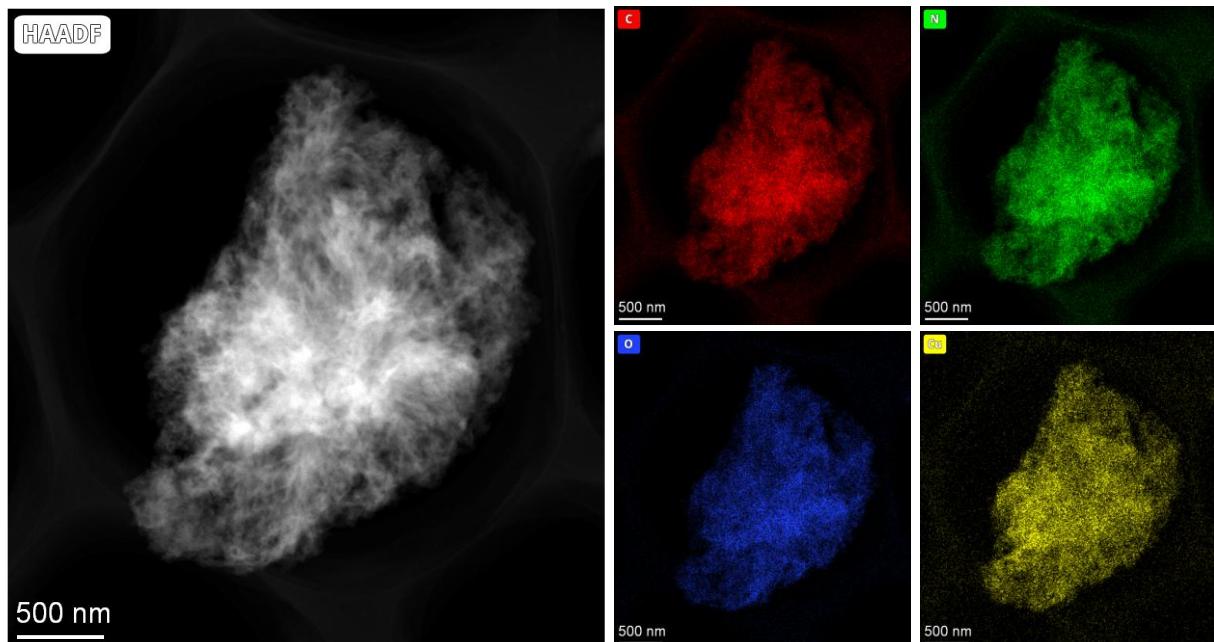


Figure S6. EDS mapping images for TpBpy-Cu.

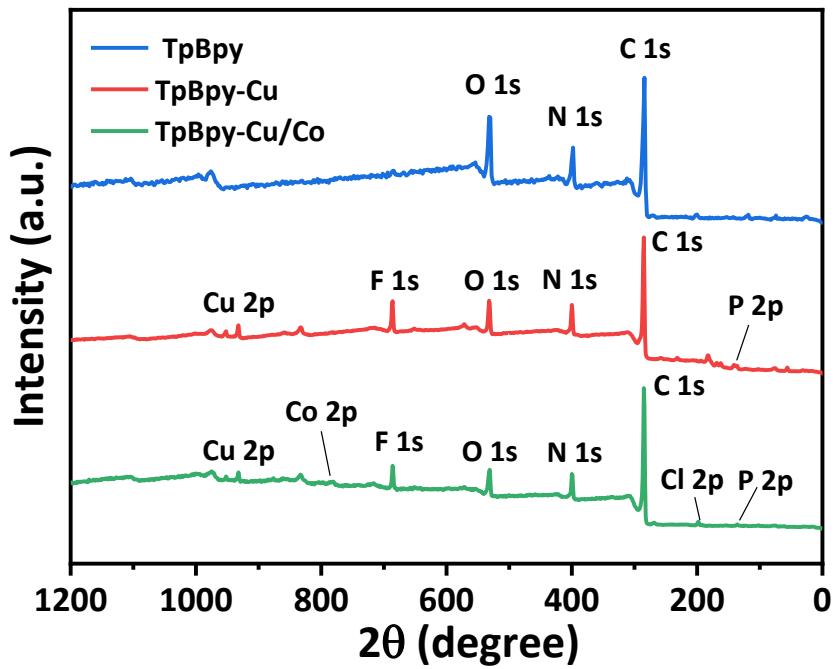


Figure S7. XPS survey of TpBpy, TpBpy–Cu and TpBpy–Cu/Co.

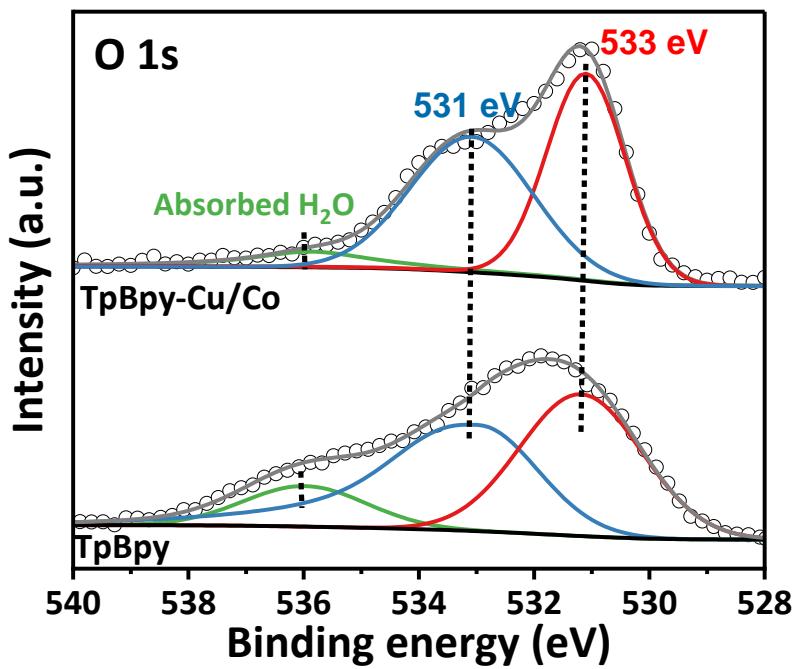


Figure S8. O 1s XPS spectrum of a) TpBpy and b) TpBpy–Cu/Co.

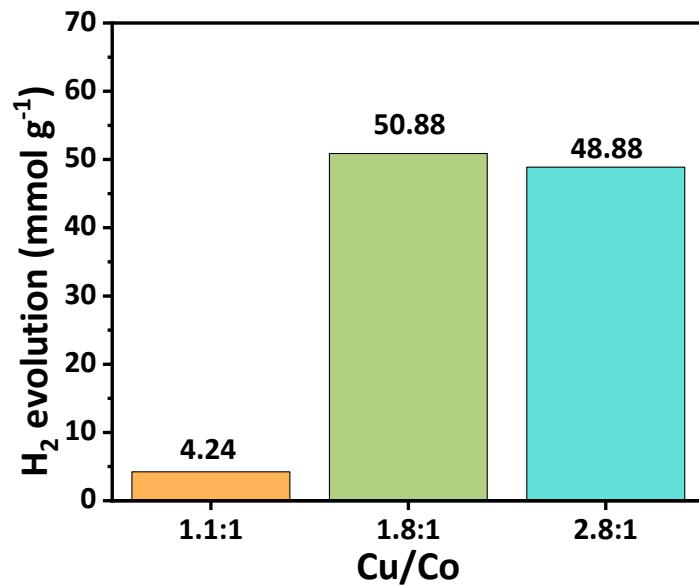


Figure S9. Photocatalytic H_2 evolution of TpBpy-Cu/Co with different content of Cu and Co under visible irradiation in 3 mL DMF/AcOH = 19/1. A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm^{-2}) was used as light source.

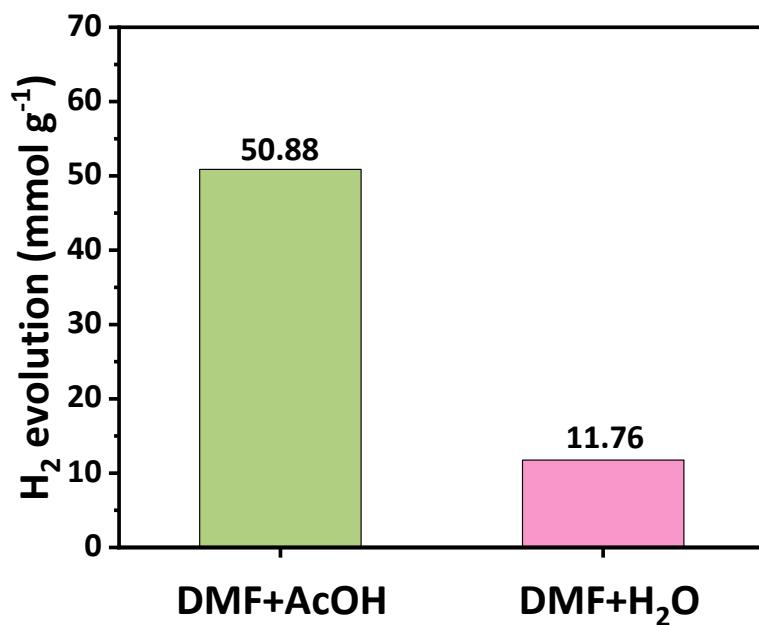


Figure S10. Photocatalytic H_2 evolution of TpBpy-Cu/Co under visible irradiation in 3 mL DMF/AcOH = 19/1 or 3 mL DMF/ H_2O = 19/1. A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm^{-2}) was used as light source.

Table S2. Photocatalytic HER experiments in 6 hours.

Entry	Catalyst	H ₂ (mmol g ⁻¹)	TON	TOF (h ⁻¹)
1	TpBpy-Cu/Co	50.88	339.2	56.5
2	TpBpy-Cu	13.44	/	/
3	TpBpy	1.60	/	/
4	TpBpy-Cu + Co(bpy)Cl ₂	32.08	213.9	35.6
5	TpBpy + [Cu(mestphen)(bpy)] ⁺ +Co(bpy)Cl ₂	11.64	77.6	12.9

Unless noted, HER reactions were conducted with COF or homogeneous catalysts containing 0.136 μmol Cu site and 0.075 μmol Co site, and 90 mg BIH, in 45 μL AcOH + 2.955 mL DMF, a 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm⁻²) was used as the light source. *TON* = $\frac{\text{no.of moles of } H_2 \text{ produced} \times 2}{\text{no.of moles of Co sites}}$; *TOF* = $\frac{TON}{\text{hours of irradiation}}$.

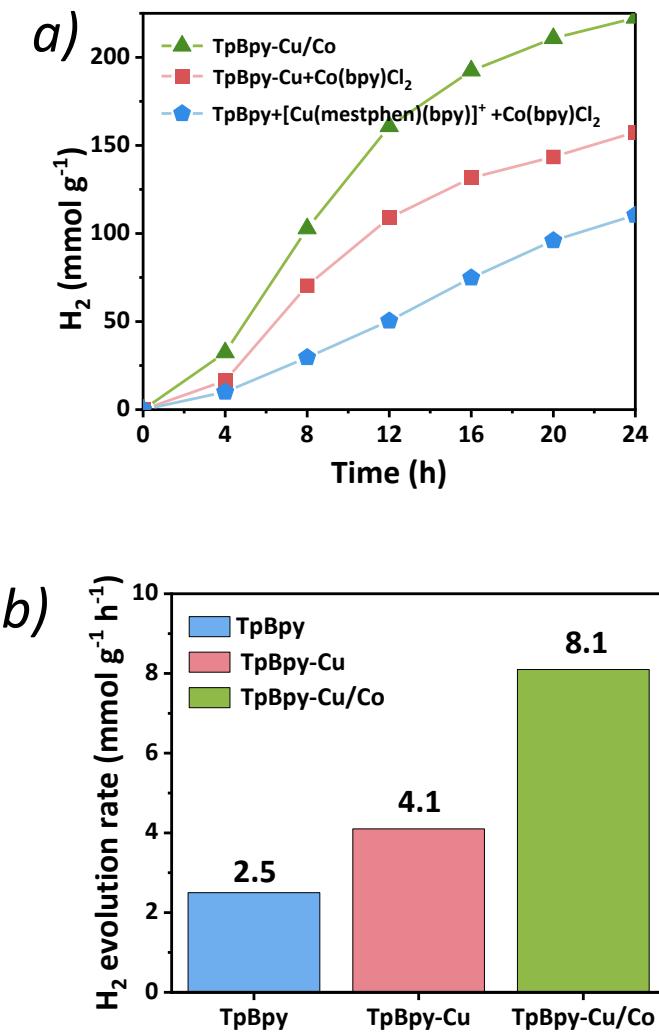


Figure S11. a) Time-dependent H_2 production under visible light irradiation and b) photocatalytic H_2 evolution initial rates (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm^{-2}) was used as light source.

Table S3. Photocatalytic HER experiments at low concentrations in 24 hours.

Entry	Catalyst	H ₂ (mmol g ⁻¹)	TON	TOF (h ⁻¹)
1	TpBpy-Cu/Co	222.40	1483	61.8
2	TpBpy-Cu + Co(bpy)Cl₂	157.60	1051	43.8
3	TpBpy + [Cu(mestphen)(bpy)]⁺+Co(bpy)Cl₂	111.60	744	31
4	[Cu(mestphen)(bpy)] ⁺ + Co(bpy)Cl ₂	78.74	267	11.1

Unless noted, HER reactions were conducted with COF or homogeneous catalysts containing 0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH + 2.955 mL DMF, a 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm⁻²) was used as the light source. *TON* = $\frac{\text{no.of moles of } H_2 \text{ produced} \times 2}{\text{no.of moles of Co sites}}$; *TOF* = $\frac{TON}{\text{hours of irradiation}}$.

Table S4. Photocatalytic HER performance of some reported metalation COFs in water.

Entry	Catalyst	Metal	Light source (300 W Xe-lamp with filter)	Experimental conditions	Hydrogen production rate ($\mu\text{mol h}^{-1} \text{g}^{-1}$)	Ref
1	TpBpy-Cu/Co	Cu(I)	>420 nm	BIH, DMF/AcOH (1.97:0.03)	12160	This work
		Co(II)				
2	Pt-Tpy-COF	Pt(II)	>400 nm	ascorbic acid, H ₂ O	7800	¹
3	Ni-Py-COF	Ni(II)	>420 nm	ascorbic acid, H ₂ O	626	²
4	TpBpy-Ir	Ir(III)	>400 nm	ascorbic acid, H ₂ O	760	³
5	CuCo ₂ O ₄ /TpPa-COF	Cu(II)	>420 nm	ascorbic acid, H ₂ O	8346	⁴
		Co(II)				
6	Co/Zn-Salen-COF	Co(II)	>420 nm	ascorbic acid, H ₂ O	1378	⁵
		Zn(II)				
7	MoS ₂ -3%/TpPa-1-COF	Mo(IV)	>420 nm	ascorbic acid, H ₂ O	5585	⁶
8	20%CdS-CTF-1	Cd(II)	>420 nm	lactic acid, H ₂ O	11430	⁷
9	Ni(OH) ₂ -2.5%/TpPa-2	Ni(II)	>420 nm	Sodium ascorbate, PBS buffer solution	1896	⁸
10	ZnPor-DETH-COF	Zn(II)	>400 nm	TEOA, phosphate buffer solution	413	⁹
		Pt(Co-catalyst)				
11	Pt ₁ @TpPa-1	Pt(II)	>420 nm	Sodium ascorbate, PBS buffer solution	719	¹⁰
12	Pd ⁰ /TpPa-1-EosinY	Pd(0)	>420 nm	TEOA, H ₂ O	10400	¹¹
13	Au100%-SAs-PAF-164	Au(δ) ($0 < \delta < 1$)	AM 1.5G filter	TEOA, MeCN/H ₂ O (3:1)	4820	¹²
14	Co0.4-SAC/Tp-Tta COF	Co(II)	320–780 nm	TEOA, H ₂ O	1798.5	¹³

15	Rh SAs/TpPa-1	Rh(0-III)	>420 nm	Sodium ascorbate, PBS buffer solution	1836.81	¹⁴
16	TpBpy-Cu/Co	Cu(I) Co(II)	>420 nm	BIH, DMF/H ₂ O (1.97:0.03)	2810	This work
17	UiO67-Ir-Cou 6/Co	Ir(III) Co(II)	>420 nm	BIH, MeCN/H ₂ O (43:7)	2440	¹⁵
18	N2-COF	Co(Co- catalyst)	AM 1.5G filter	TEOA, ACN/H ₂ O (4:1)	782	¹⁶
19	[Co-1b]-COF	Co(II)	AM 1.5G filter	TEOA, MeCN/H ₂ O (4:1)	163	¹⁷

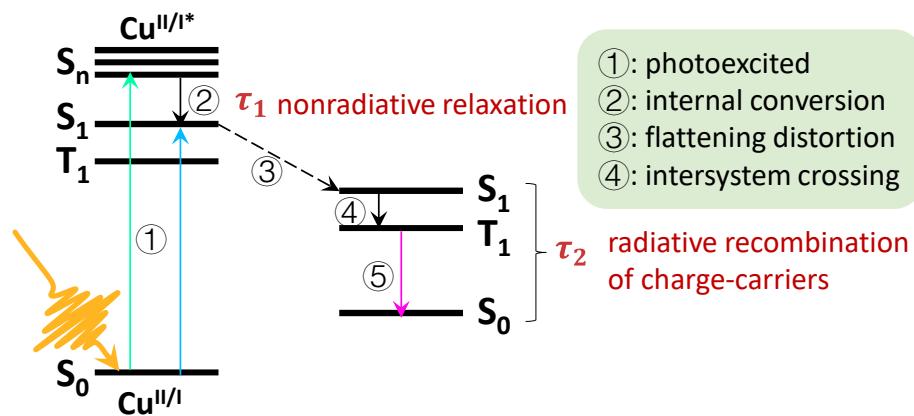


Figure S12. Proposed Perrin Jablonski diagram.

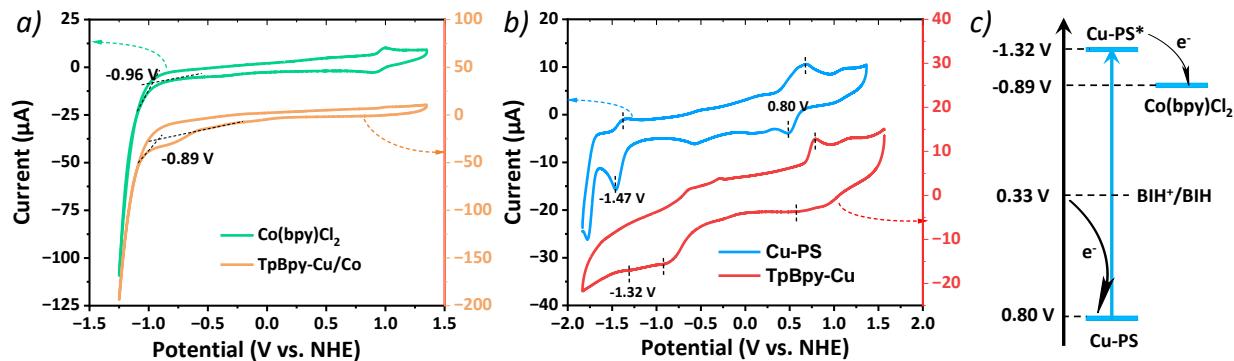


Figure S13. a) CVs of TpBpy-Cu/Co (0.05 mg) coated on electrode surface and Co(bpy)Cl₂ under photocatalytic HER conditions (10 mL 0.1 M TBAPF₆/CH₃CN solution with 150 μL AcOH).
 b) CVs of TpBpy-Cu coated on electrode surface and [Cu(mestphen)(bpy)]⁺ in 0.1M TBAPF₆/CH₃CN. c) Potential energy diagrams of TpBpy-Cu/Co catalyzed photocatalytic HER.

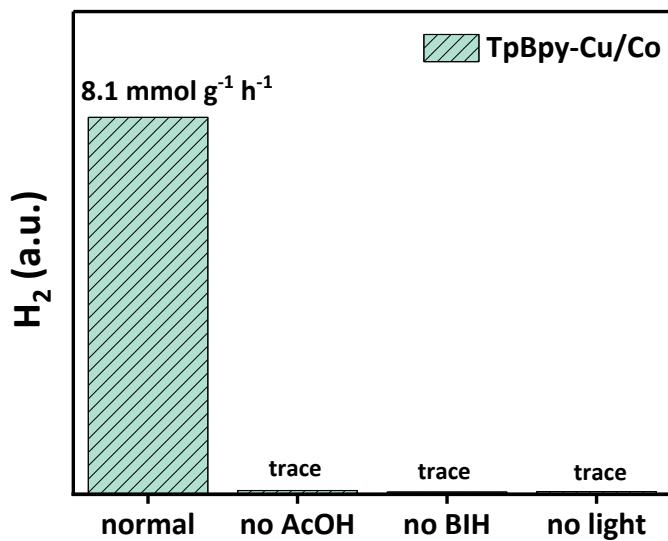


Figure S14. Control experiments of photocatalytic H₂ evolution under visible irradiation. (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm⁻²) was used as light source.

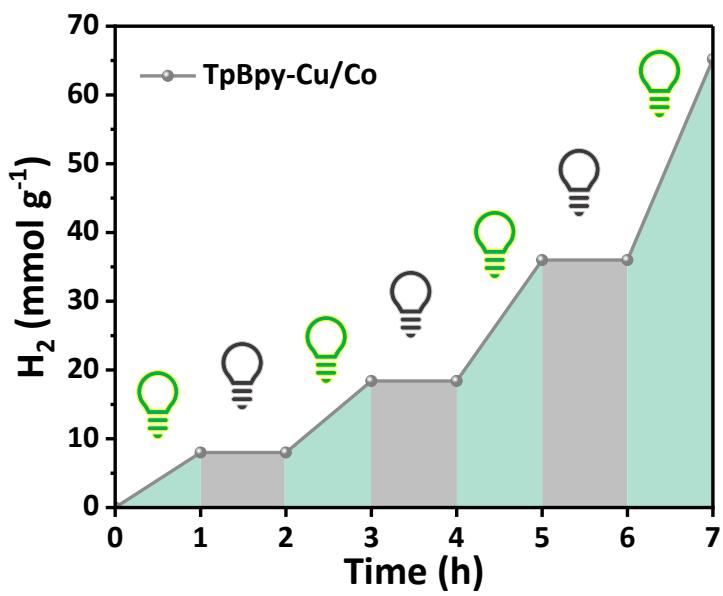


Figure S15. Light On-Off experiments for visible light-driven HER catalyzed by TpBpy-Cu/Co. (0.0136 μmol Cu site and 0.0075 μmol Co site, and 90 mg BIH, in 45 μL AcOH and 2.955 mL DMF). A 300W Xenon lamp with a 420 nm cut off filter (light intensity: 520 mW cm^{-2}) was used as light source.

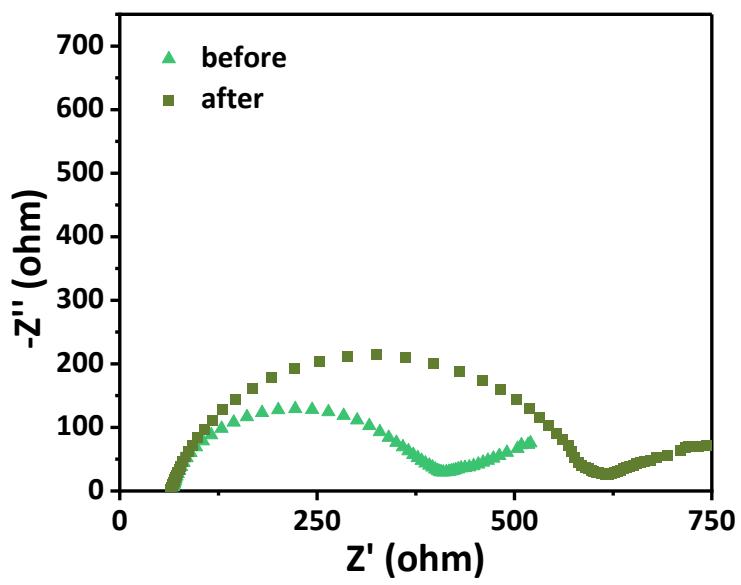


Figure S16. EIS Nyquist plots of TpBpy-Cu/Co before and after 24 h photocatalysis.

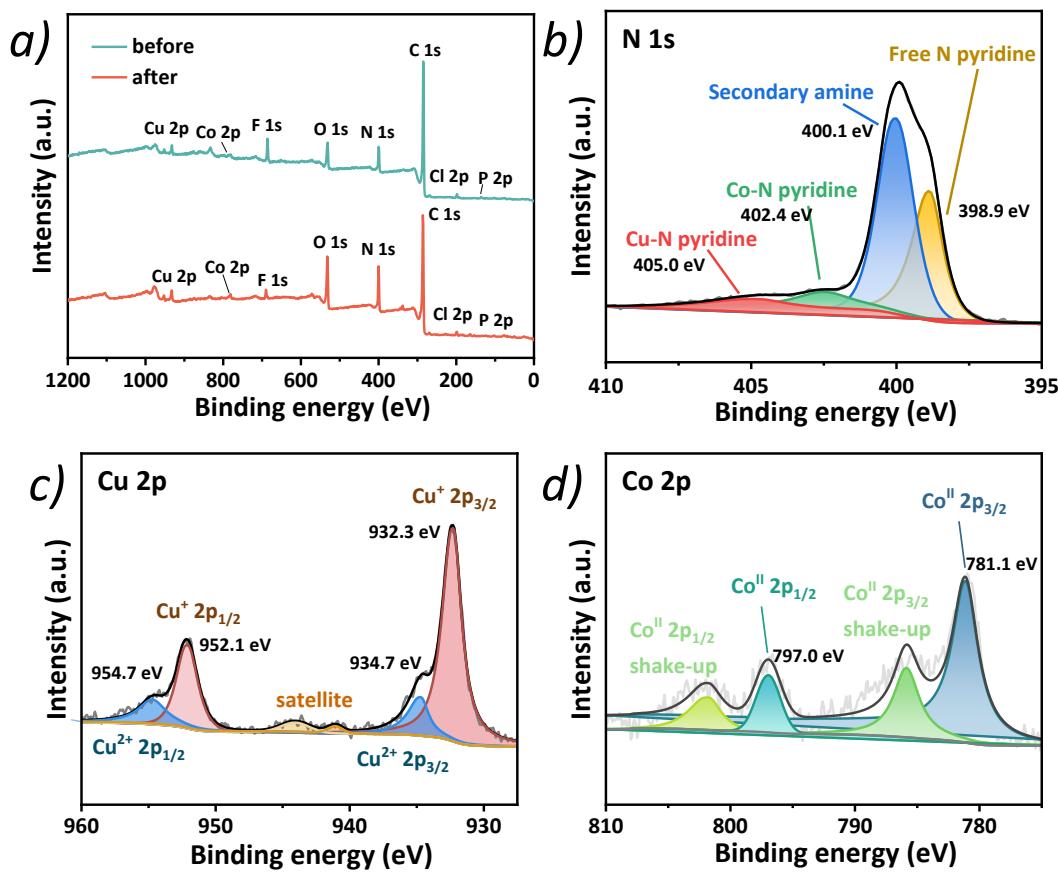
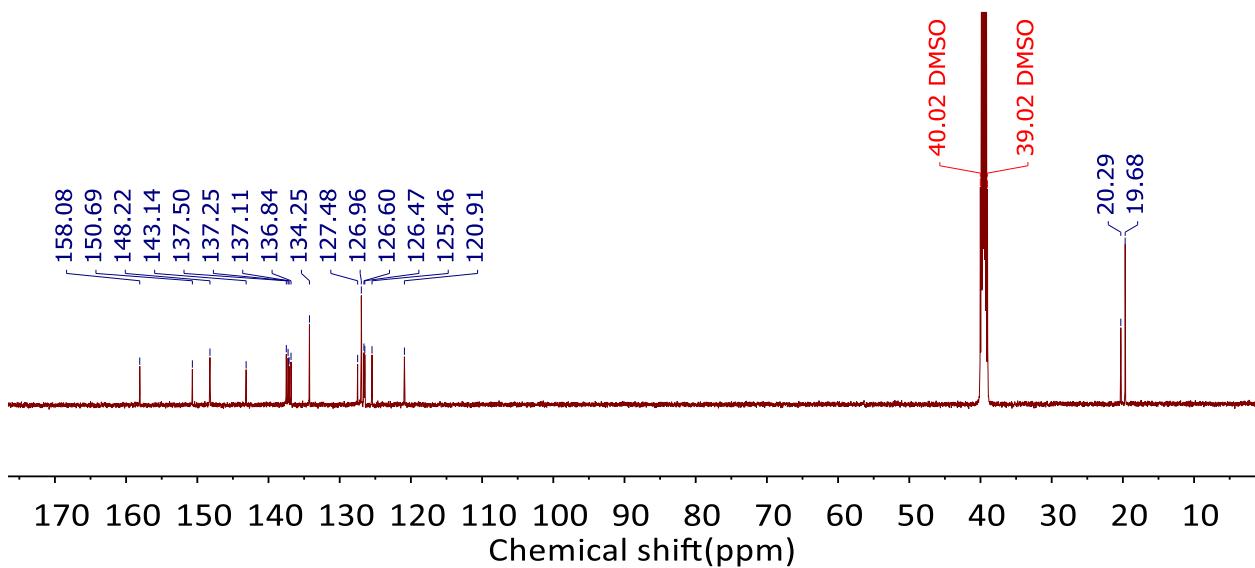
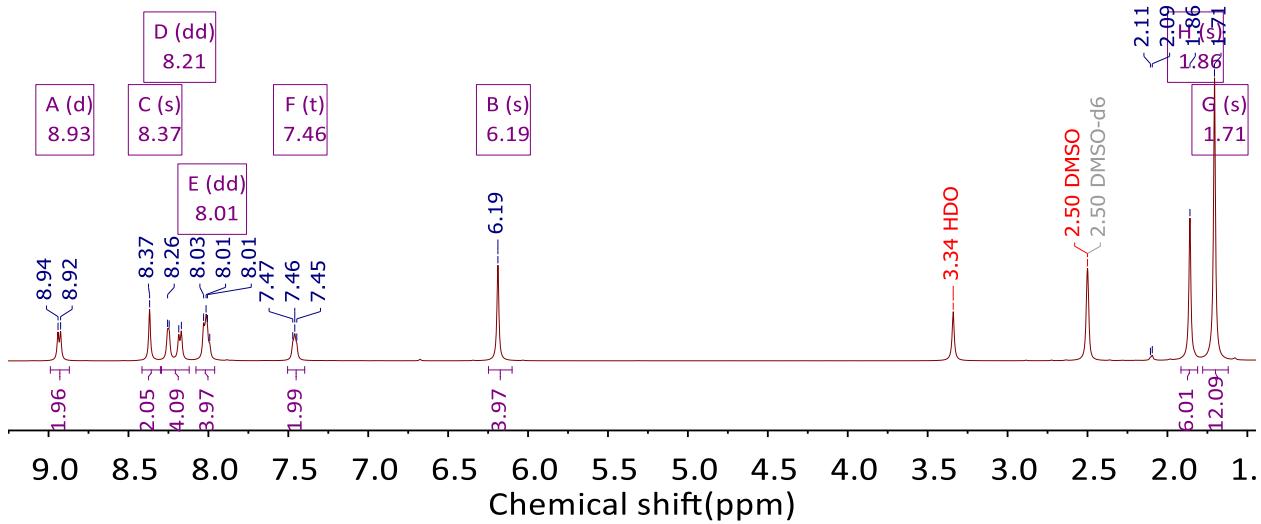


Figure S17. a) XPS survey, b) N 1s, c) Cu 2p, and d) Co 2p spectrum of TpBpy–Cu/Co after 24 h photocatalysis.



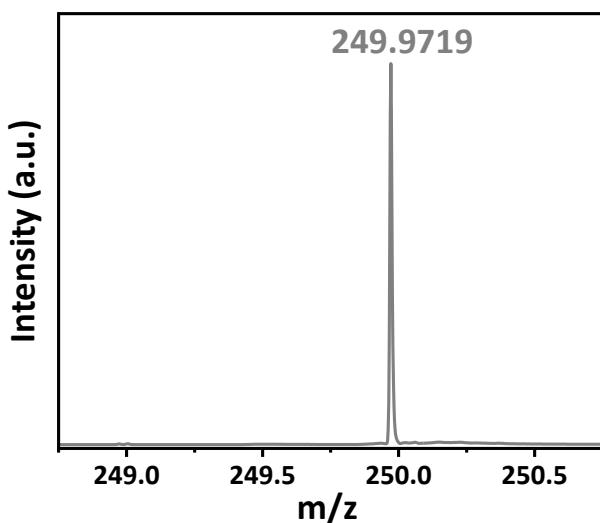


Figure S20. HRMS (ESI, positive mode) spectra of CobpyCl₂ in CH₃CN.

References

- (1) Liu, C.; Ma, D.-L.; Tian, P.-J.; Jia, C.; Qi, Q.-Y.; Jiang, G.-F.; Zhao, X. Lateral functionalization of a one-dimensional covalent organic framework for efficient photocatalytic hydrogen evolution from water. *Journal of Materials Chemistry A* **2024**.
- (2) Sun, L.; Lu, M.; Yang, Z.; Yu, Z.; Su, X.; Lan, Y. Q.; Chen, L. Nickel glyoximate based metal–covalent organic frameworks for efficient photocatalytic hydrogen evolution. *Angewandte Chemie International Edition* **2022**, *61* (30), e202204326.
- (3) Song, D.; Xu, W.; He, W.; Li, C.; Yang, J.; Li, J.; Wang, N. Selective integrating molecular catalytic units into bipyridine-based covalent organic frameworks for specific photocatalytic fuel production. *Inorganic Chemistry* **2024**, *63* (7), 3444-3451.
- (4) He, W.; Kong, K.; Wang, M.; Dong, B.; Yuan, D.; Bryliakov, K. P.; Wang, R. Photoelectron migration monitored by 3d orbital electron configuration of spinel cocatalysts for covalent organic framework-based photocatalytic hydrogen evolution. *Applied Catalysis B: Environment and Energy* **2024**, *350*, 123916.
- (5) Zhou, W.; Deng, Q. W.; He, H. J.; Yang, L.; Liu, T. Y.; Wang, X.; Zheng, D. Y.; Dai, Z. B.; Sun, L.; Liu, C. Heterogenization of salen metal molecular catalysts in covalent organic frameworks for photocatalytic hydrogen evolution. *Angewandte Chemie International Edition* **2023**, *62* (3), e202214143.
- (6) Gao, M.-Y.; Li, C.-C.; Tang, H.-L.; Sun, X.-J.; Dong, H.; Zhang, F.-M. Boosting visible-light-driven hydrogen evolution of covalent organic frameworks through compositing with MoS₂: a promising candidate for noble-metal-free photocatalysts. *Journal of Materials Chemistry A* **2019**, *7* (35), 20193-20200, 10.1039/C9TA07319A. DOI: 10.1039/C9TA07319A.
- (7) Wang, D.; Zeng, H.; Xiong, X.; Wu, M.-F.; Xia, M.; Xie, M.; Zou, J.-P.; Luo, S.-L. Highly efficient charge transfer in CdS-covalent organic framework nanocomposites for stable photocatalytic hydrogen evolution under visible light. *Science Bulletin* **2020**, *65* (2), 113-122. DOI: <https://doi.org/10.1016/j.scib.2019.10.015>.
- (8) Dong, H.; Meng, X.-B.; Zhang, X.; Tang, H.-L.; Liu, J.-W.; Wang, J.-H.; Wei, J.-Z.; Zhang, F.-M.; Bai, L.-L.; Sun, X.-J. Boosting visible-light hydrogen evolution of covalent-organic frameworks by introducing Ni-based noble metal-free co-catalyst. *Chemical Engineering Journal* **2020**, *379*, 122342.
- (9) Chen, R.; Wang, Y.; Ma, Y.; Mal, A.; Gao, X.-Y.; Gao, L.; Qiao, L.; Li, X.-B.; Wu, L.-Z.; Wang, C. Rational design of isostructural 2D porphyrin-based covalent organic frameworks for tunable photocatalytic hydrogen evolution. *Nature Communications* **2021**, *12* (1), 1354.
- (10) Dong, P.; Wang, Y.; Zhang, A.; Cheng, T.; Xi, X.; Zhang, J. Platinum single atoms anchored on a covalent organic framework: boosting active sites for photocatalytic hydrogen evolution. *ACS Catalysis* **2021**, *11* (21), 13266-13279.
- (11) Ding, S.-Y.; Wang, P.-L.; Yin, G.-L.; Zhang, X.; Lu, G. Energy transfer in covalent organic frameworks for visible-light-induced hydrogen evolution. *International Journal of Hydrogen Energy* **2019**, *44* (23), 11872-11876.
- (12) Yang, Y.; Xiao, Y.; Jiang, L.; Li, J.; Li, J.; Jia, J.; Yavuz, C. T.; Cui, F.; Jing, X.; Zhu, G. Ultrahigh Single Au Atoms Loaded Porous Aromatic Frameworks for Enhanced Photocatalytic Hydrogen Evolution. *Advanced Materials* **2024**, 2404791.

- (13) Wang, Y.; Pan, A.; Ma, Y.; Du, H. Single– atom Co anchored on covalent organic framework for remarkable photocatalytic hydrogen evolution efficiency. *International Journal of Hydrogen Energy* **2024**.
- (14) Zhang, D.; Zhang, C.; Lai, X.; Wei, X.; Zhuang, T.; Lv, Z. Engineering single-atom rhodium-C₃N sites on covalent organic frameworks for boosting photocatalytic hydrogen evolution. *Journal of Colloid and Interface Science* **2024**, 676, 691-700.
- (15) Guo, S.; Kong, L. H.; Wang, P.; Yao, S.; Lu, T. B.; Zhang, Z. M. Switching excited state distribution of metal–organic framework for dramatically boosting photocatalysis. *Angewandte Chemie International Edition* **2022**, 134 (30), e202206193.
- (16) Banerjee, T.; Haase, F.; Savascı, G. k.; Gottschling, K.; Ochsenfeld, C.; Lotsch, B. V. Single-site photocatalytic H₂ evolution from covalent organic frameworks with molecular cobaloxime cocatalysts. *Journal of the American Chemical Society* **2017**, 139 (45), 16228-16234.
- (17) Gottschling, K.; Savascı, G. k.; Vignolo-González, H.; Schmidt, S.; Mauker, P.; Banerjee, T.; Rovó, P.; Ochsenfeld, C.; Lotsch, B. V. Rational design of covalent cobaloxime–covalent organic framework hybrids for enhanced photocatalytic hydrogen evolution. *Journal of the American Chemical Society* **2020**, 142 (28), 12146-12156.