

## Supporting Information

### Mortise-tenon-like ionic/electronic conductive interface facilitates long-cycle solid-state lithium metal batteries

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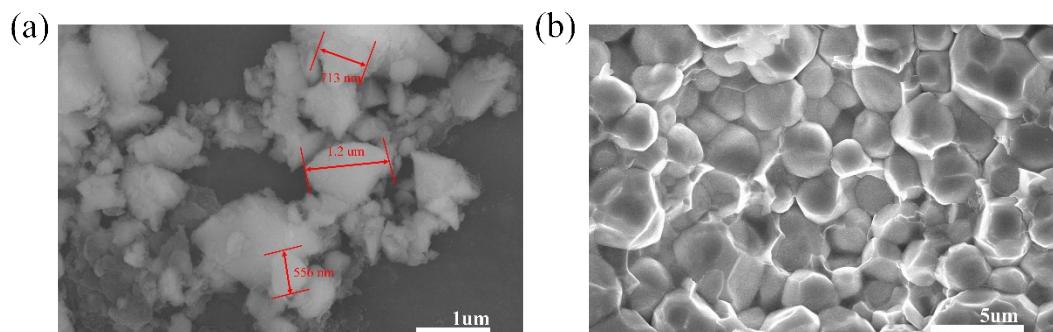


Fig. S1 (a) SEM image of LLZTO powder. (b) Cross-sectional SEM image of LLZTO pellet.

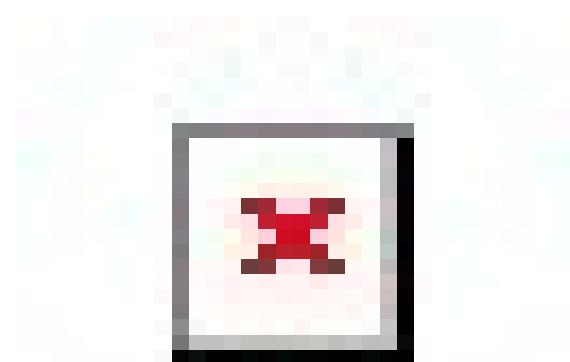


Fig. S2 XRD patterns of standard garnet cubic phase, LLZTO powder and LLZTO pellet.

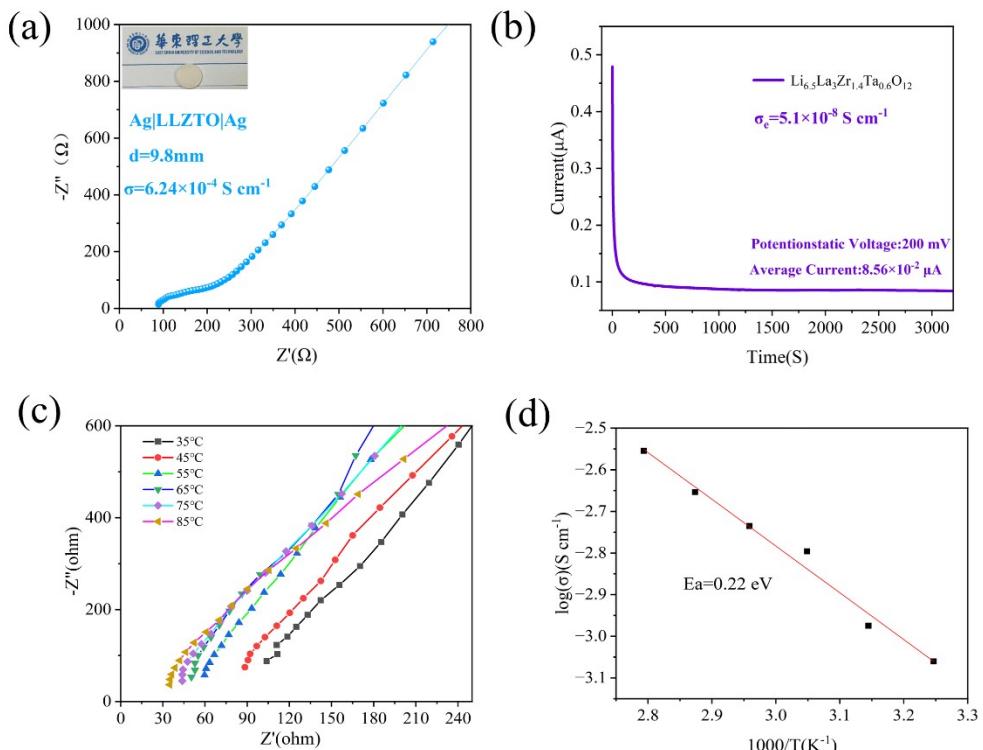


Fig. S3 (a) EIS spectrum of Ag|LLZTO|Ag at room temperature. (b) The DC Polarizatrion of the Ag|LLZTO|Ag at room temperature. (c) EIS spectra of Ag|LLZTO|Ag at different temperature. (d) Arrhenius plots of LLZTO pellet.

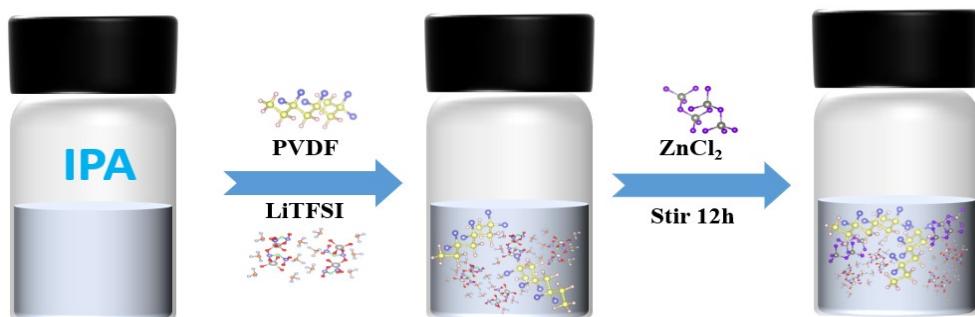


Fig. S4 Schematic diagram of the preparation process for fluorine-containing- $\text{ZnCl}_2$  solution.

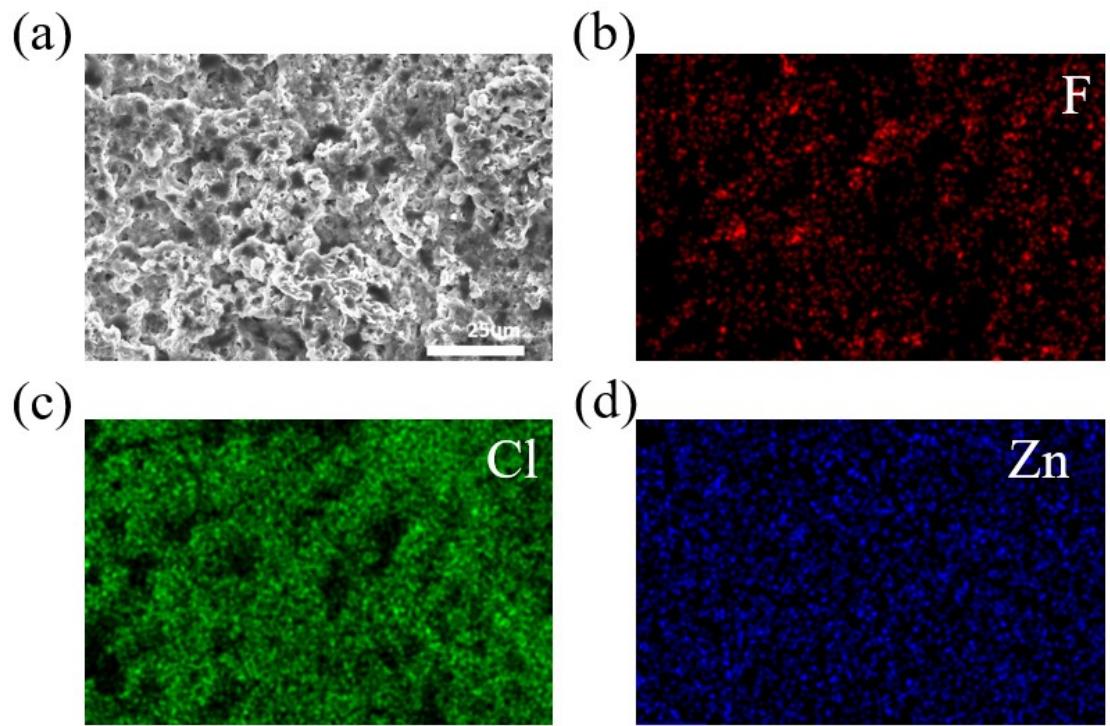


Fig. S5 (a) Surface SEM image of LLZTO after treatment with fluorine-containing-ZnCl<sub>2</sub> solution. (b-d) EDS mappings of F, Cl, Zn elements of the modified LLZTO.

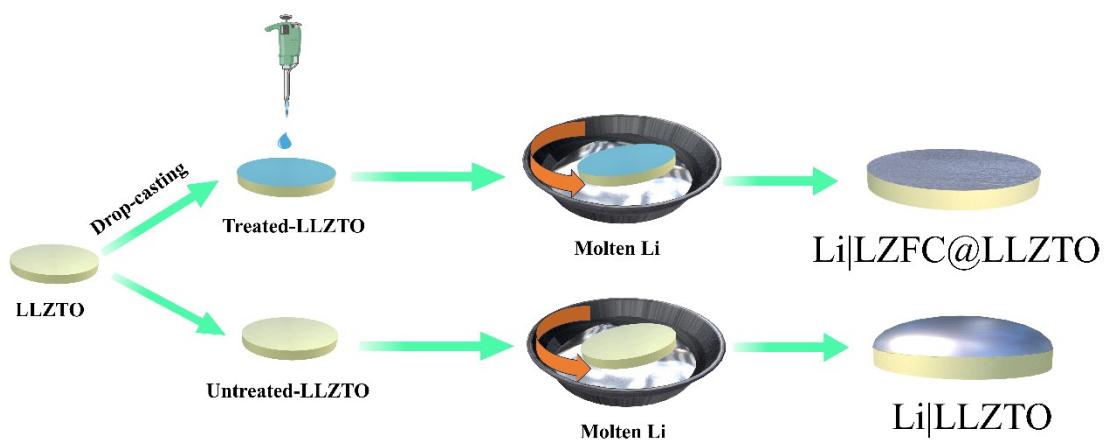


Fig. S6 Schematic diagram of preparation process of the Li/LLZTO and Li/LZFC@LLZTO interface.

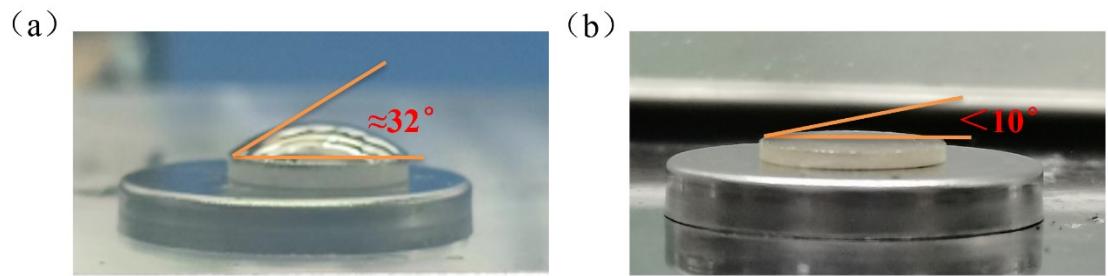


Fig. S7 Comparison of contact angle of (a) Li/LLZTO and (b) Li/LZFC@LLZTO.

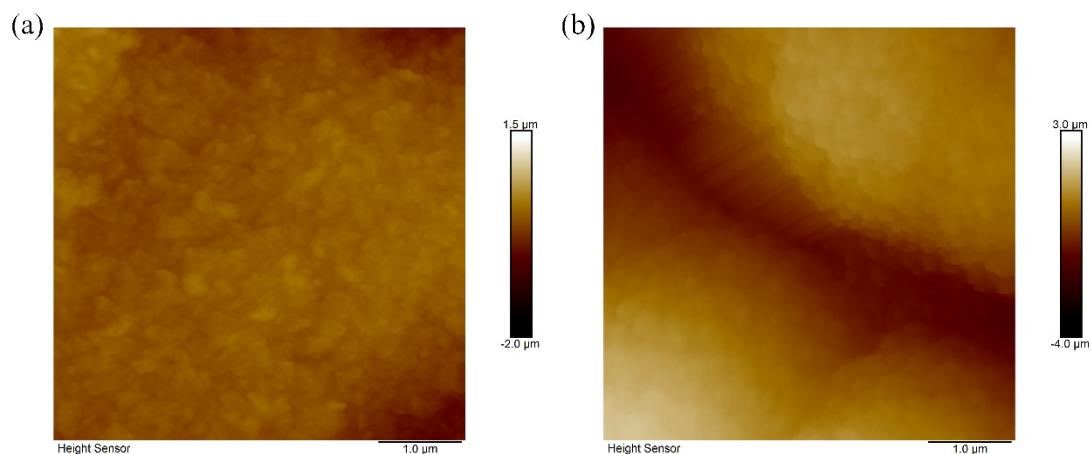


Fig. S8 (a) The 2D surface morphology of LLZTO obtained by AFM. (b) The 2D surface morphology of LZFC@LLZTO after *in situ* lithiation reaction obtained by AFM.

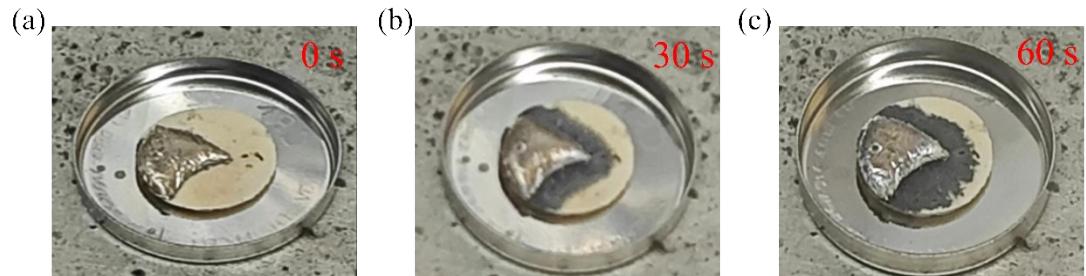


Fig. S9 (a-c) The top side photographs of Li/LZFC@LLZTO after the lithiation reaction with 0 s, 30 s and 60 s.

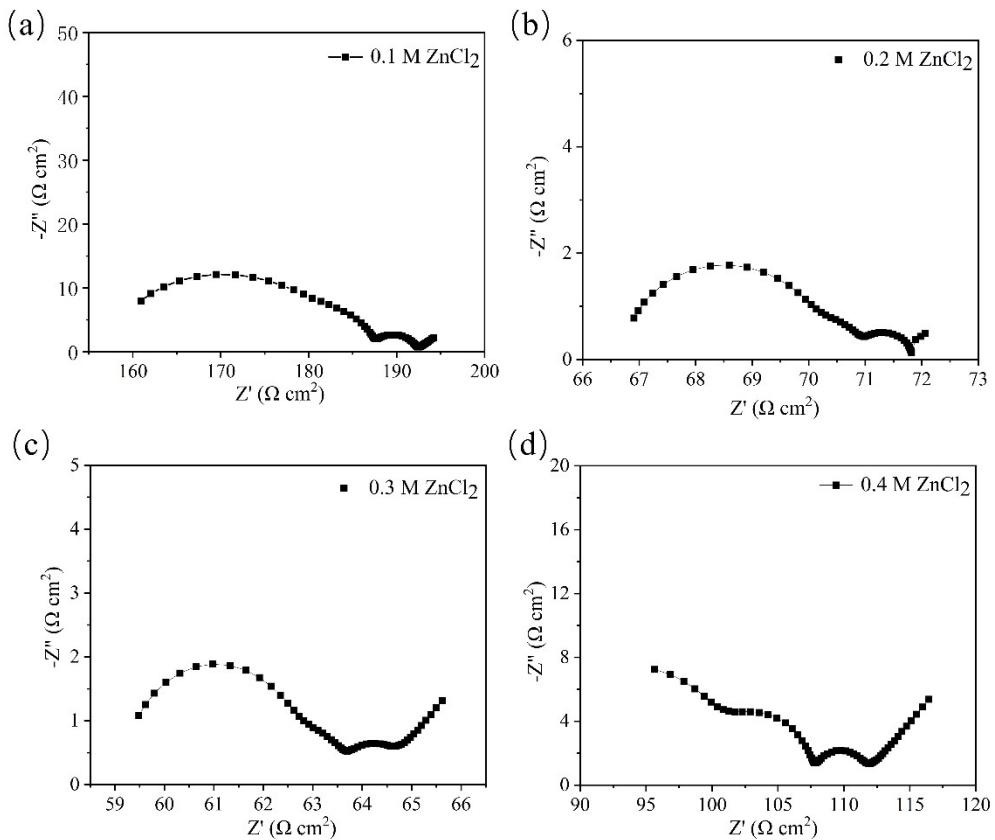


Fig. S10 (a-d) Impedance plots of  $\text{Li}|\text{LZFC@LLZTO}|\text{Li}$  with different concentrations of  $\text{ZnCl}_2$ -modified solutions.

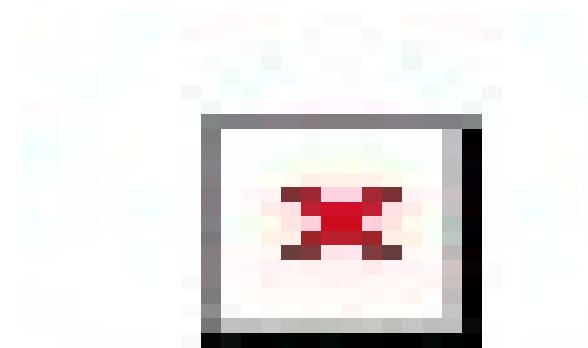


Fig. S11 The CCD of  $\text{Li}|\text{LLZTO}|\text{Li}$  cell.

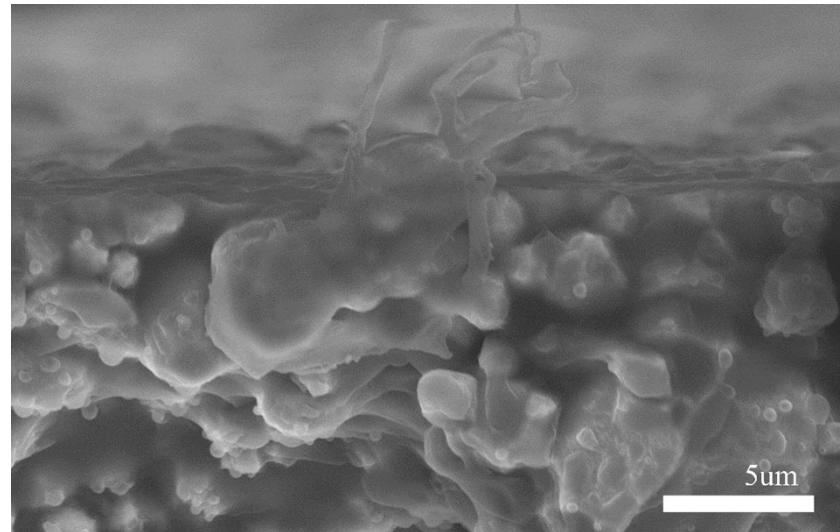


Fig. S12 Cross-section SEM of Li dendrites piercing Li|LLZTO|Li cell.

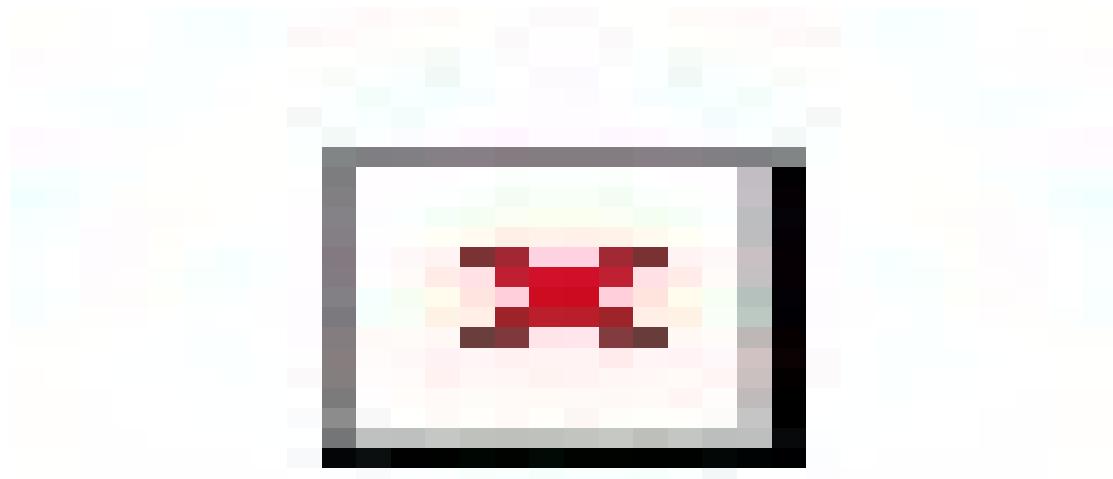


Fig. S13 The constant current cycle comparison between Li|LLZTO|Li and Li|LZFC@LLZTO|Li at  $0.2 \text{ mA cm}^{-2}$ .

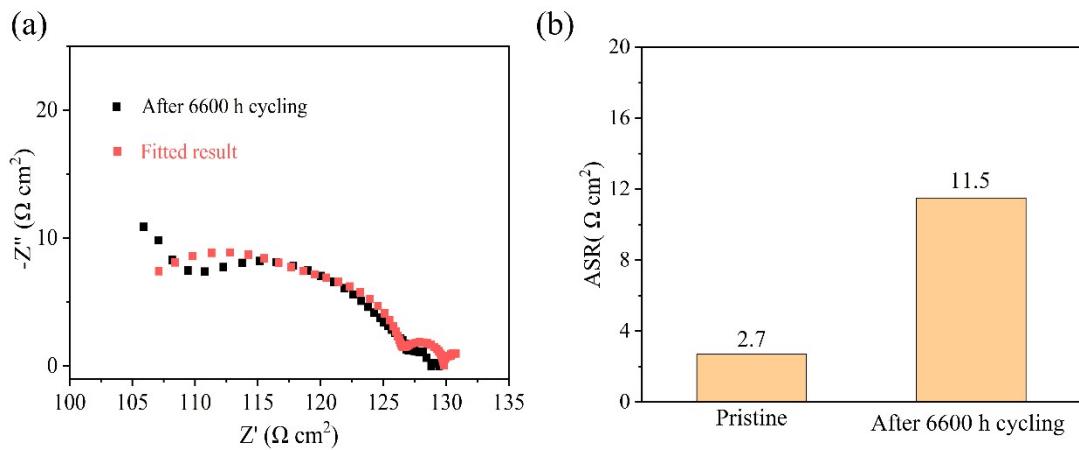


Fig. S14 (a) The EIS spectrum of Li|LZFC@LLZTO|Li after 6600 h cycling. (b) The comparison of ASR between pristine and after 6600 h cycling at  $0.2 \text{ mA cm}^{-2}$ .

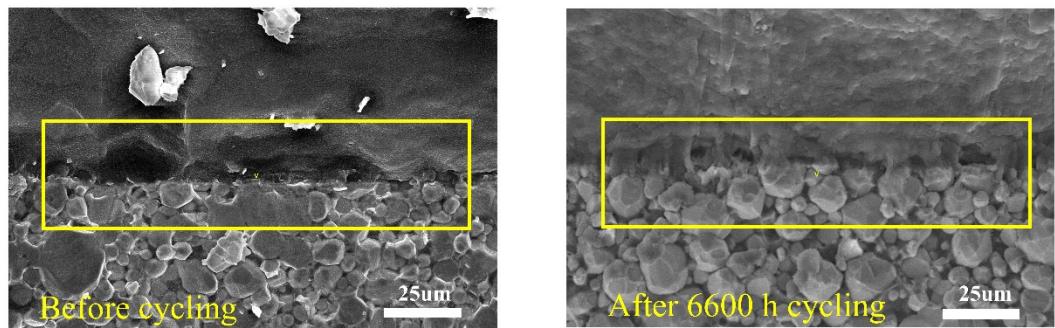


Fig. S15 (a) Cross-section SEM of Li/LZFC@LLZTO before cycling. (b) Cross-section SEM of Li/LZFC@LLZTO after 6600 h cycling at  $0.2 \text{ mA cm}^{-2}$ .

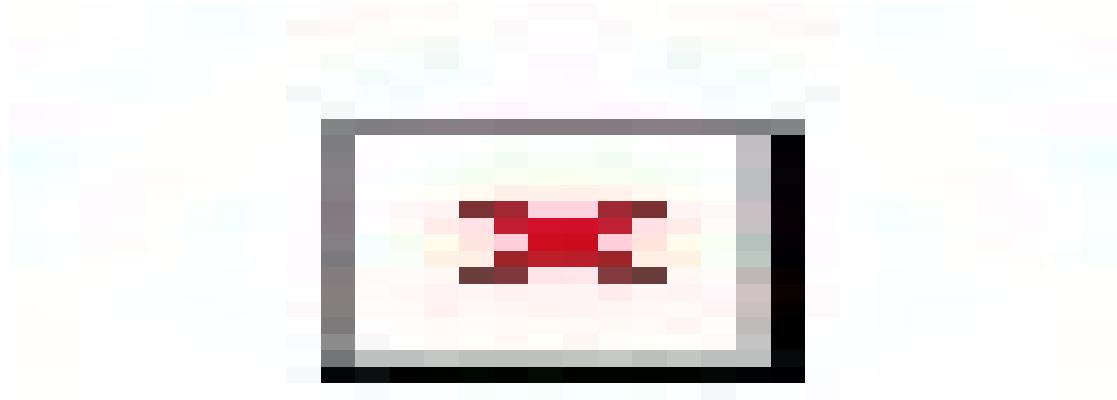


Fig. S16 Long-term galvanostatic cycles of the Li|LZFC@LLZTO|Li at  $0.5 \text{ mA cm}^{-2}$ .

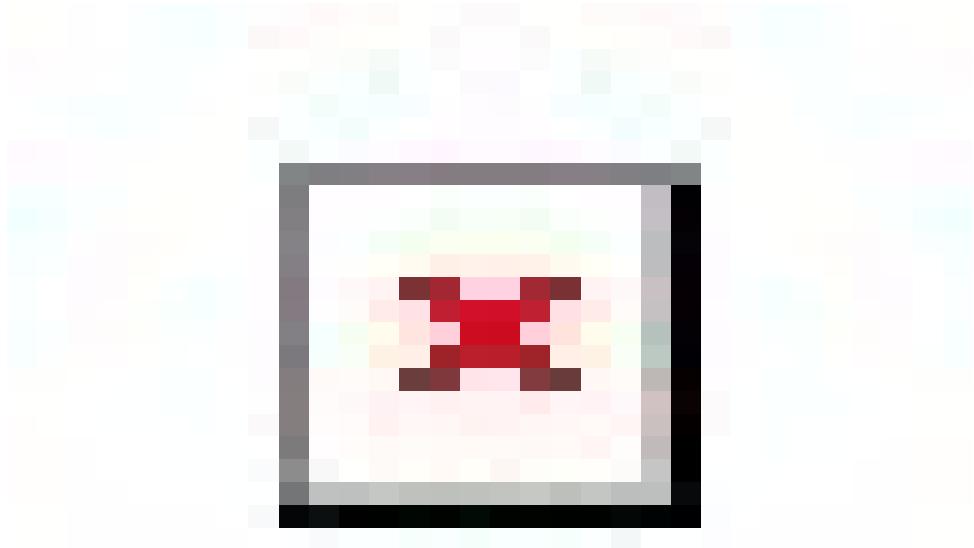


Fig. S17 XPS profile of Zn 2p orbital before and after Li melting.

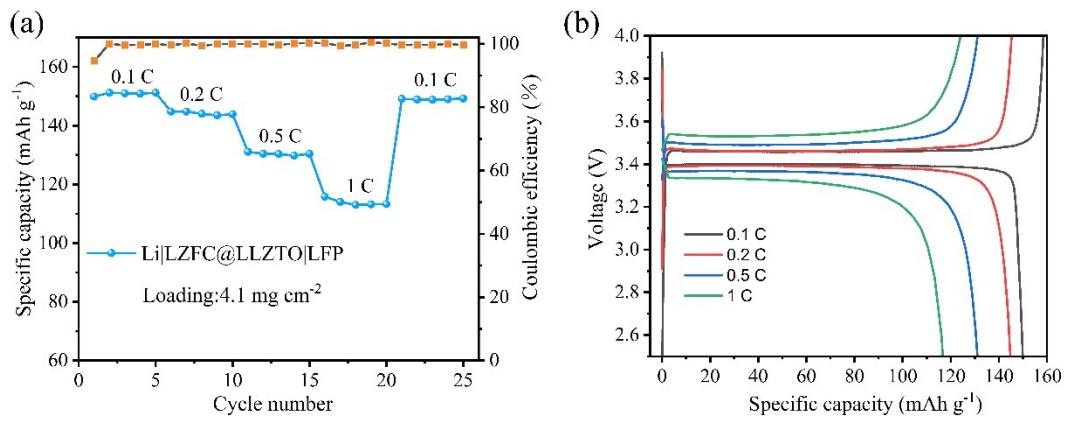


Fig. S18 The rate performance of Li|LZFC@LLZTO|LFP with high load. (b) The voltage capacity curve of Li|LZFC@LLZTO|LFP with high loading.

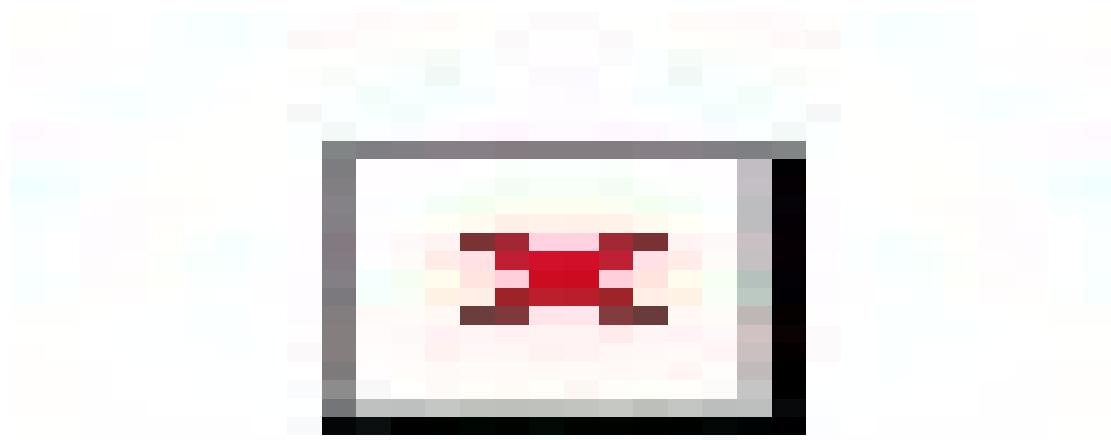


Fig. S19 Cyclic stability of the Li|LLZTO|LFP battery at 0.2 C.

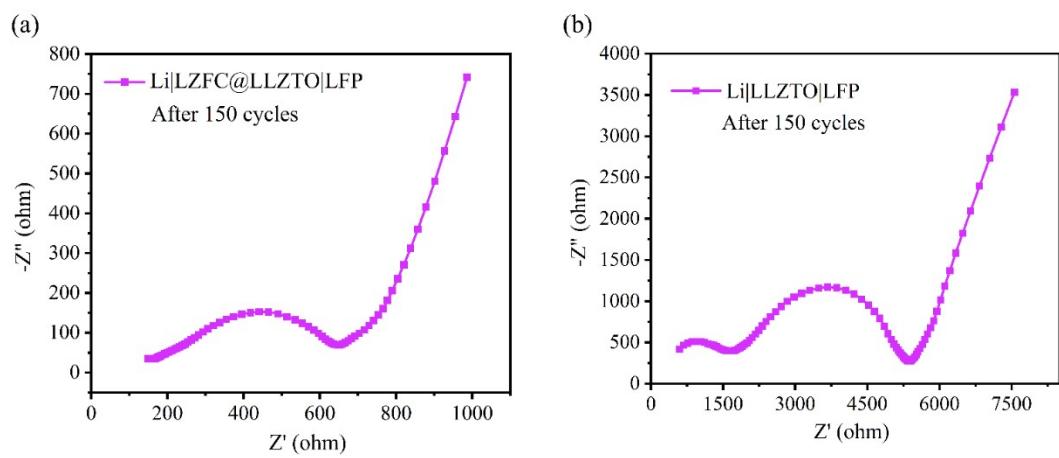


Fig. S20 (a) EIS spectra of  $\text{Li}|\text{LZFC}@\text{LLZTO}|\text{LFP}$  after 150 cycles. (b) EIS spectra of  $\text{Li}|\text{LLZTO}|\text{LFP}$  after 150 cycles.

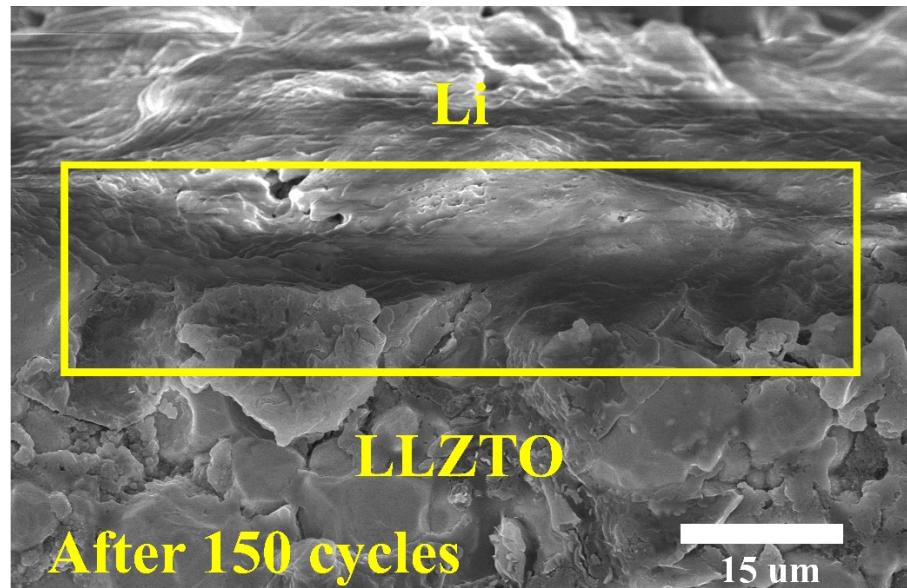


Fig. S21 The cross section SEM of  $\text{Li}|\text{LZFC}@\text{LLZTO}$  after 150 cycles of  $\text{Li}|\text{LZFC}@\text{LLZTO}|\text{LFP}$ .

Table S1 Area specific resistance (ASR), critical current density (CCD) and cycling stability of Li|LZFC@LLZTO|Li compared with recent advanced works.

Interfacial layer	ASR ( $\Omega \text{ cm}^2$ )	CCD ( $\text{mA cm}^{-2}$ )	Stability ( $\text{mA cm}^{-2}$ )/h	References
Li-Mg/LiF	25	0.65	0.3/1000	1
Li-BiF <sub>3</sub>	7.4	1.1	0.3/850	2
Li-CaCl <sub>2</sub>	7	1	0.1/1200	3
Li-Li <sub>2</sub> Se	5.1	1.4	0.5/350	4
Li-Sr-N	4.5	1.3	0.3/1200	5
Li-SnF <sub>2</sub>	11	1.3	0.5/600	6
Li-Sb	4.1	0.64	--	7
Li-LATP	6	1.5	0.3/2500	8
Li-BiOCl	27	1.1	0.5/1000	9
Li-PTLI	33.1	0.71	0.05/1000	10
Li-Al/Si	15	1.2	1/800	11
Li-CoO QDs	12.3	1.1	0.3/1700	12
Li-SnS <sub>2</sub>	47	0.5	0.2/1000	13
Li-LTO	25	0.8	0.2/450	14
Li-SnS	5.1	0.9	0.2/1500	15
Li-LiPO <sub>2</sub> F <sub>2</sub>	5.1	1.2	0.6/1500	16
Li-S	12.4	0.7	0.2/1000	17
Li-MXene	5	1.5	0.3/600	18
Li-InCl <sub>3</sub>	10	0.7	0.2/4000	19
Li-PPA	4	1.8	0.2/2500	20
<b>This Work</b>	<b>2.7</b>	<b>2.1</b>	<b>0.2/6600</b> <b>0.4/900</b> <b>0.5/520</b>	

Table S2 Liquid electrolyte content, cathode load, areal capacity and liquid/capacity ratios compared with the published literature.

Liquid electrolyte content	Cathode load ( $\text{mg cm}^{-2}$ )	Areal capacity ( $\text{mAh cm}^{-2}$ ) (0.1C)	Liquid/capacity ratios (g/Ah)	References
3uL (3.6 mg)	2.5	0.47	15	21
10uL (12 mg)	3	0.47	29	22
10uL (12 mg)	2.9	0.46	23	11
5uL (6mg)	1.5-2	0.25-0.33	16-21	23
10uL (12 mg)	2	0.29	25	2
8 uL (9.6mg)	2-3	0.3-0.44	19-28	12
3uL (3.6 mg)	1.5	0.23	31	17
3uL (3.6 mg)	1.65	0.26	28	18
10uL (12 mg)	2	0.33	32	25
5uL (6mg)	1-2	0.15-0.3	40-80	14
<b>5uL (6mg)</b>	<b>4.1</b>	<b>0.62</b>	<b>10.9</b>	<b>This work</b>

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