

## Electronic Supplementary Information (ESI)

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Title: Molecular dynamics study on thermal conductivity of KH<sub>2</sub>PO<sub>4</sub> with vacancy defects

Table S1. Full Force-Field Parameters for the KDP (KH<sub>2</sub>PO<sub>4</sub>) Potential

The potential function used in this study is based on the model developed and validated by Yang et al. [Comput. Mater. Sci. 187 (2021) 110122]. The total potential energy is expressed as:  $E_{\text{pot}} = E_{\text{bond}} + E_{\text{angle}} + E_{\text{vdW}} + E_{\text{Coulomb}} + E_{\text{H-bond}}$

### (a) Atomic Partial Charges

Atom Type	Charge $q_i$ (e)
K	+0.835
P	+0.461
O <sub>1</sub>	-0.454
O <sub>2</sub>	-0.506
H	+0.312

**Note:** O<sub>1</sub> and O<sub>2</sub> represent oxygen atoms within the hydroxyl group (P–OH) and the phosphoryl group (P=O) of the [PO<sub>4</sub>]<sup>3-</sup> tetrahedra, respectively, as defined in the potential model by Yang et al. [1].

### (b) Covalent Bond and Angle Parameters (Harmonic)

$$\text{Form: } E_{\text{bond}} = K_r(r - r_0)^2; E_{\text{angle}} = K_\theta(\theta - \theta_0)^2$$

Interaction	$r_0$ (Å) / $\theta_0$ (°)	$K$ [kcal/(mol·Å <sup>2</sup> ) or kcal/(mol·rad <sup>2</sup> )]
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Interaction	$r_0$ (Å) / $\theta_0$ (°)	$K$ [kcal/(mol·Å <sup>2</sup> ) or kcal/(mol·rad <sup>2</sup> )]
Bond : P–O	1.600	430
Bond : O–H	1.037	500
Angle : O–P–O	109.5	25.0
Angle : P–O–H	110.5	148.8

(c) Hydrogen Bond Parameters (Specific H-bond part)[2]

$$\text{Form: } E_{\text{hyd}} = \epsilon [5(\sigma/r)^{12} - 6(\sigma/r)^{10}] \cos^4 \theta$$

Parameter	Value
Equilibrium well depth $\epsilon$	17.5 kcal/mol
Zero-crossing distance $\sigma$	2.5 Å

(d) 9-6 Lennard-Jones Parameters (Van der Waals)

$$\text{Form: } E_{\text{vdW}} = \epsilon_{ij} [2(\sigma_{ij}/r_{ij})^9 - 3(\sigma_{ij}/r_{ij})^6]$$

Interaction ( $i - j$ )	$r_{0,ij}$ (Å)	$\epsilon_{0,ij}$ (kcal/mol)
O <sub>1</sub> – O <sub>1</sub>	3.500	0.0385
O <sub>1</sub> – O <sub>2</sub>	3.665	0.0451
O <sub>1</sub> – K	3.544	0.0882

Interaction ( - )	$i$ $j$	$r_{0,ij}$ (Å)	$\epsilon_{0,ij}$ (kcal/mol)
O <sub>2</sub> – O <sub>2</sub>		3.800	0.0560
O <sub>2</sub> – K		3.700	0.1050
H – H		1.098	0.0091
H – P		4.009	0.0012
H – K		3.194	0.0025
P – P		4.500	0.1750
P – K		4.164	0.1518
K – K		3.585	0.2030

## References

- [1] S. Yang, L. Zhang, H. Xie and W. Liu, Interaction potential function for the deformation analysis of potassium dihydrogen phosphate using molecular dynamics simulation, *Comput. Mater. Sci.*, 2021, 187, 110122.
- [2] S. L. Mayo, B. D. Olafson and W. A. Goddard III, DREIDING: a generic force field for molecular simulations, *J. Phys. Chem.*, 1990, 94, 8897–8909.