## Supplementary Information: Machine Learning-Assisted Profiling of a Kinked Ladder Polymer Structure using Scattering

Lijie Ding,<sup>a</sup> Chi-Huan Tung,<sup>a</sup> Zhiqiang Cao,<sup>a</sup> Zekun Ye,<sup>b</sup> Xiaodan Gu,<sup>c</sup> Yan Xia,<sup>b</sup> Wei-Ren Chen,<sup>a</sup> and Changwoo Do<sup>\*a</sup>

**Effect of temperature variation** We confirm the sample is fully dissolved and without aggregation by comparing the scattering function I(Q) measured under different temperatures: 25 °C, 75 °C and 125 °C. As shown in Fig. S1, the I(Q) spectrum are consistent cross all three temperatures, indicating that the polymer sample we used is fully dissolved.



Fig. S1 Scattering intensity of the synthesized CANAL polymer under three different temperature: 25 °C, 75 °C and 125 °C.



Fig. S2 Nuclear magnetic resonance spectra for the CANAL ladder polymer. (a) Low molecular weight. (b) high molecular weight

**Characterization of synthesized ladder polymer** All chemicals, including 1,4-diethylbenzene, were purchased from commercial sources and used as received unless otherwise noted. 1,4-dibromo-2,5-diethylbenzene was prepared following known literature procedures (Sugamata et al., 2021). All reactions were performed under nitrogen in oven- or flame-dried glassware unless otherwise noted. Flash column chromatography was carried out with Silica 60 (230–400 mesh; Fisher). Analytical thin-layer chromatography (TLC) was carried out using 0.2 mm silica gel plates (silica gel 60, F254, EMD Chemical).

Proton nuclear magnetic resonance (<sup>1</sup>H NMR) and carbon nuclear magnetic resonance (<sup>13</sup>C NMR) spectra were recorded in CDCl<sub>3</sub> using 500 MHz Varian NMR spectrometers. Chemical shifts are reported in parts per million (ppm)

<sup>&</sup>lt;sup>a</sup> Neutron Scattering Division, Oak Ridge National Laboratory, Oak Ridge, TN 37831, USA

<sup>&</sup>lt;sup>b</sup> Department of Chemistry, Stanford University, Stanford, CA 94305, USA

<sup>&</sup>lt;sup>c</sup> School of Polymer Science and Engineering, Center for Optoelectronic Materials and Devices, The University of Southern Mississippi, Hattiesburg, MS 39406, USA

<sup>\*</sup> E-mail: doc1@ornl.gov

relative to residual protonated solvent for <sup>1</sup>H (CHCl<sub>3</sub> =  $\delta$ 7.26, DCM =  $\delta$ 5.30, methanol-d<sub>4</sub> =  $\delta$ 3.31) and relative to carbon resonances of the solvent for <sup>13</sup>C (CDCl<sub>3</sub> =  $\delta$ 77.0). Data are reported as follows: chemical shift, multiplicity (s = singlet, d = doublet, dd = doublet of doublets, t = triplet, q = quartet, m = multiplet, br = broad signal, and associated combinations), coupling constant(s) (Hz), and integration.