

Supporting Information

Upcycling sodium lignosulfonate into carbon anode with an inorganic-rich interphase by potential regulation for lithium-ion batteries

Kuo Li^{1,2}, Shuhao Ruan¹, Tu Ran¹, Yawei Li¹, Bo Li^{1,2}, Xiaofei Yang³, Haisong Wang², Yi Cheng^{1*}

¹Institute of Advanced Studies, Zhejiang Normal University, 1108 Gengwen Road, Hangzhou 311231, P. R. China.

²School of Light Industry and Chemical Engineering, Dalian Polytechnic University, Dalian, 116034, China

³Division of Energy Storage, Dalian Institute of Chemical Physics, Chinese Academy of Sciences, Zhongshan Road 457, Dalian 116023, PR China

**Corresponding author: Yi Cheng (chengyi@zjnu.edu.cn)*

Correlated Calculations:

$$i(V) = av^b \quad \text{Eq.S1}$$

$$\text{Log}(i) = \text{log}(a) + b\text{log}(v) \quad \text{Eq.S2}$$

where i represents the peak current, v denotes the scan rate, a is a constant, and b signifies the slope of the line as a function of $\text{log}(i)$ and $\text{log}(v)$.

$$i = k_1v + k_2v^{0.5} \quad \text{Eq.S3}$$

where k_1v stands for the current response contributed by the pseudo capacitive process, and $k_2v^{0.5}$ represents for the diffusion-controlled portion of the current response.

$$D = \frac{4}{\pi\tau} \cdot \left(\frac{n_m \cdot V_m}{S}\right)^2 \cdot \left(\frac{\Delta E_s}{\Delta E_\tau}\right)^2 = \frac{4}{\pi\tau} \cdot \left(L \frac{\Delta E_s}{\Delta E_\tau}\right)^2 \quad \text{Eq.S4}$$

Where τ is the relaxation time; n_m (mol) is the number of moles of the active material; V_m (cm³ mol⁻¹) is the molar volume; S (cm²) is the contact area between the active material with the electrolyte; ΔE_s (V) is the change in voltage induced by the pulses; and ΔE_τ (V) is voltage variation under constant current puls; L is the average migration path length for lithium ions within the active material particles

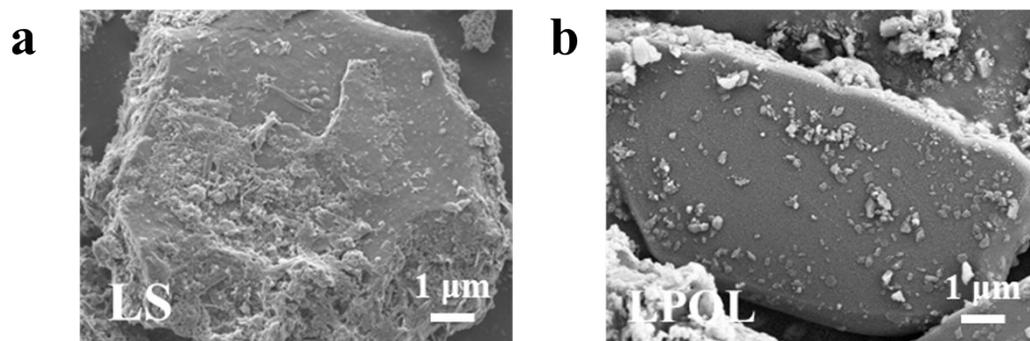


Fig. S1. (a) SEM image of the pristine LS powders; (b) SEM image of the fabricated LPOL powders.

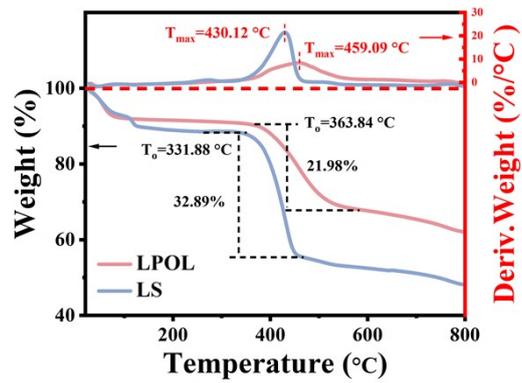


Fig. S2 TG-DTG curve of the two samples.

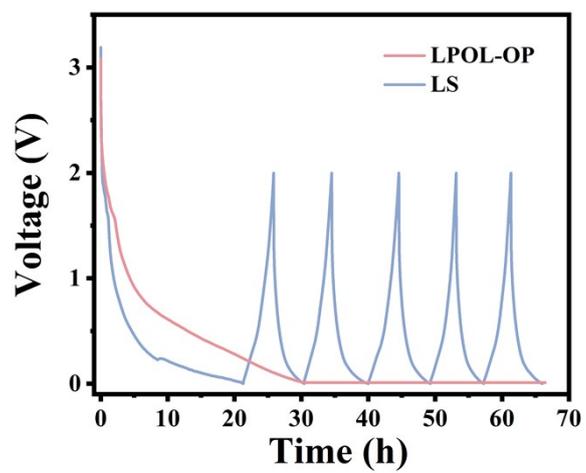


Fig. S3. Potential change during SEI formation process using continuous galvanostatic charge-discharge (CGCD) method.

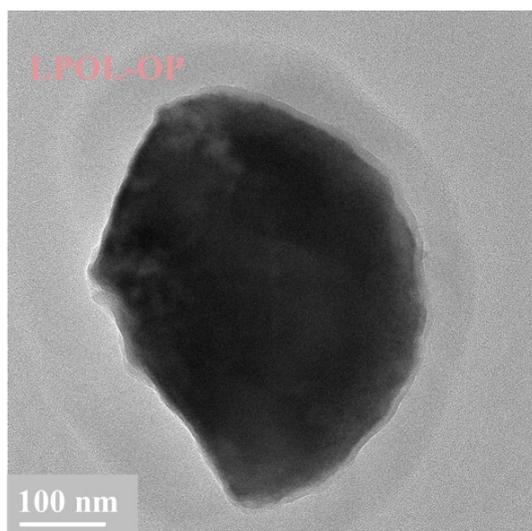


Fig. S4. HRTEM images of LPOL-OP anode;

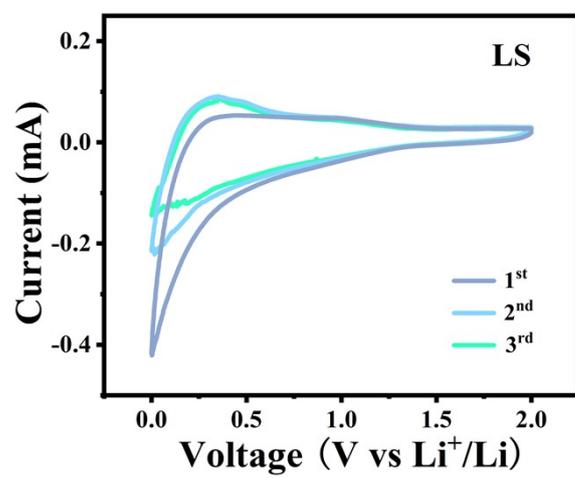


Fig. S5. The CV curves of LS electrodes in initial three cycles at 0.2 mV s^{-1}

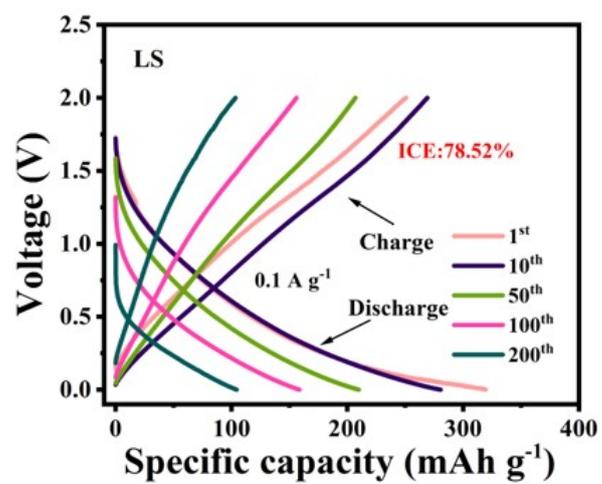


Fig. S6. Charge-discharge curves of LS after different cycles at 100 mA g⁻¹.

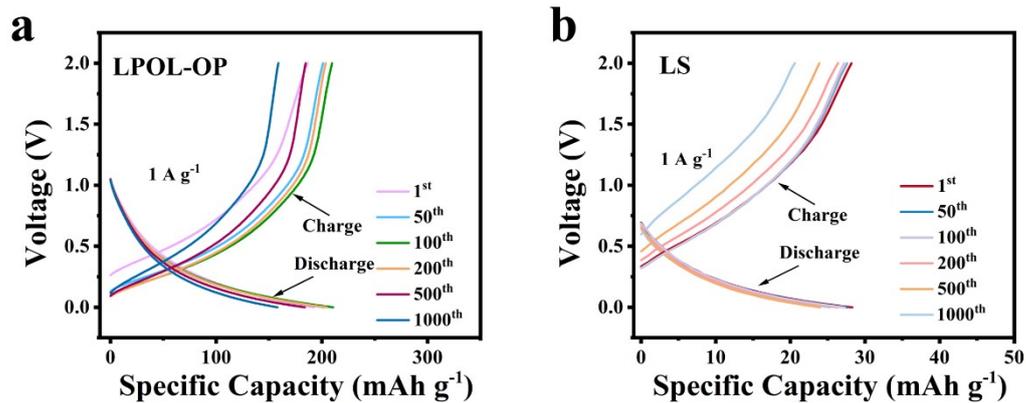


Fig. S7. (a) Charge-discharge curves of LPOL-OP after different cycles at 1000 mA g⁻¹; (b) Charge-discharge curves of LS after different cycles at 1000 mA g⁻¹.

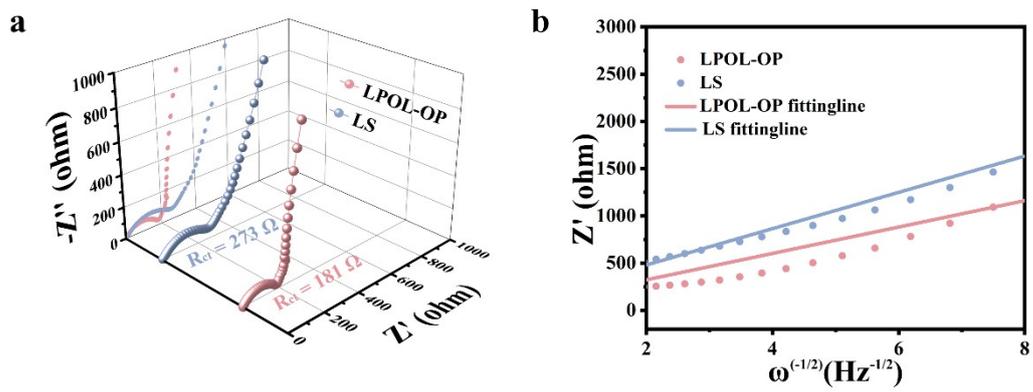


Fig. S8. (a) EIS of the LPOL-OP and LS; (b) Linear fitting of inclined line at low frequency.