

Boosting organic pollutants degradation through synergistic photocatalysis and peroxymonosulfate activation by bismuth oxybromide

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Table S1. Comparison of recent studies on BiOBr based photocatalysts for PMS activation.

Catalyst	Pollutant	Reaction condition	Removal efficiency	Ref.
Zn _x Ni _{1-x} Fe ₂ O ₄ @BiOBr	Paracetamol, 10 mg L ⁻¹	[Catalyst] = 0.5 g L ⁻¹ , [PMS] = 2 mM, UV-A	About 100%, 100 min	42
BiOBr	Sulfamethox azole, 0.04 mM	[Catalyst] = 2.0 g L ⁻¹ , [PMS] = 0.4 mM, ambient conditions	About 100%, 60 min	43
BiOBr- cyclodextrin polymer	Acid Orange 7, 0.2 mM	[Catalyst] = 1 g L ⁻¹ , [PMS] = 1 mM, 500 W xenon lamp, wavelength > 420 nm	99.2%, 60 min	44
AgBr/BiOBr/F e ₃ O ₄	Carbamazepi ne, 10 mg L ⁻¹	[Catalyst] = 0.3 g L ⁻¹ , [PMS] = 0.6 mM, 500 W xenon lamp, wavelength > 400 nm	96.84%, 30 min	45
BiOBr/FeOOH	TH, 20 ppm	[Catalyst] = 0.2 g L ⁻¹ , [PMS] = 0.24 mM, xenon lamp, wavelength > 420 nm	96%, 40 min	46
BiOBr/BiOF	Levofloxacin , 20 mg L ⁻¹	[Catalyst] = 0.4 g L ⁻¹ , [PMS] = 1 mM, 500 W xenon lamp	89.8%, 180 min	47
Fe-BiOBr	Atrazine, 0.5 mg L ⁻¹	[Catalyst] = 38 mg L ⁻¹ , [PMS] = 0.1 mM, simulated sunlight	97.04%, 30 min	48

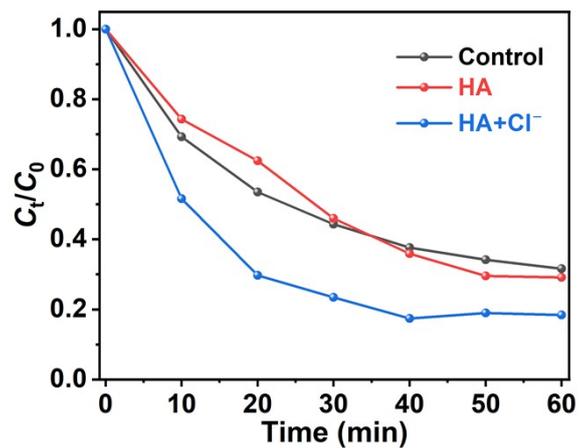


Figure S1. Effect of HA and the co-present of HA and Cl^- on TH removal in BiOBr/PMS/Vis (conditions: $[BiOBr] = 1.0 \text{ g L}^{-1}$; $[TH] = 10 \text{ mg L}^{-1}$; $[PMS] = 1.5 \text{ mM}$, $[HA] = 5 \text{ mg L}^{-1}$, $[NaCl] = 5.0 \text{ g L}^{-1}$).