

Supplementary Information

Selective aromatic halogenation by a manganese compound I model: A mimic of chloroperoxidase

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Electrophilic (C₁-HAT) and nucleophilic (C₁- and C₂-hydroxylation) reaction of [Mn^{IV}OH(Por^{•*})Cl]⁺ or [Mn^{IV}OH(Por^{•*})OTf]⁺, naphthalene and chloride

Table S1 Relative energies (in kcal mol⁻¹) of the electrophilic (C₁-HAT) and nucleophilic (C₁ and C₂ hydroxylation) reaction of [Mn^{IV}(OH)(Por^{•*})Cl]⁺, naphthalene and chlorine.

[Mn ^{IV} (OH)(Por ^{•*})Cl] ⁺	ΔDef2-SVP	ΔΔDef2-TZVPP	ΔE ^a	ΔZ ₀	ΔE _{thermal} ^b	-TΔS ^b	ΔDisp	ΔCplx ^c	ΔG ^d
Separated (3 components)	30.83	-21.02	9.81	-2.30	-1.01	-14.99	+10.61	+3.78	5.89
Electrophilic (C₁-HAT)									
⁵ [Reactant Complex]	0.04	-0.08	-0.05	-0.03	+0.00	-0.08	-0.11	+0.00	-0.27
⁵ [Transition State]	9.09	+1.38	10.46	-2.66	-0.53	+2.04	+0.65	+0.00	9.97
⁵ [Product Complex]	-48.84	+0.61	-48.24	+0.04	+0.42	-1.13	-3.89	+0.00	-52.80
Nucleophilic (C₁-hydroxylation)									
⁵ [Reactant Complex]	1.21	-1.05	0.16	-0.14	+0.06	-0.61	+1.50	+0.00	1.58
⁵ [Transition State]	10.40	-0.49	9.91	-0.97	-0.37	+0.72	-0.13	+0.00	8.43
⁵ [Product Complex]	-34.45	+4.01	-30.44	+2.19	-0.64	+1.89	-6.51	+0.00	-35.39
Nucleophilic (C₂-hydroxylation)									
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	7.67	+0.01	7.68	-0.87	-0.49	+0.00	-1.17	+0.00	5.15
⁵ [Product Complex]	-34.11	+3.94	-30.18	+2.14	-0.64	+0.00	-6.20	+0.00	-34.88

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b T = 298.15 K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S2 Mulliken spin density distribution of the electrophilic (C₁-HAT) and nucleophilic (C₁ and C₂ hydroxylation) reaction of [Mn^{IV}(OH)(Por^{•*})Cl]⁺, naphthalene and chlorine.

[Mn ^{IV} (OH)(Por ^{•*})Cl] ⁺	Mn	OH	Cl _{axial}	Porphine	Naphthalene	Cl
Separated (3 components)	3.06	0.10	0.04	0.80	0.00	0.00
Electrophilic (C₁-HAT)						
⁵ [Reactant Complex]	3.15	0.04	0.01	-0.20	0.56	0.44
⁵ [Transition State]	3.59	-0.08	-0.04	-0.20	0.62	0.04
⁵ [Product Complex]	3.97	0.02	0.15	-0.14	0.00	0.00
Nucleophilic (C₁-hydroxylation)						
⁵ [Reactant Complex]	3.17	0.06	-0.01	-0.22	1.01	-0.01
⁵ [Transition State]	3.50	-0.10	0.04	-0.22	0.77	0.01
⁵ [Product Complex]	3.96	0.01	0.16	-0.13	0.00	-0.00
Nucleophilic (C₂-hydroxylation)						
⁵ [Reactant Complex]	3.17	0.05	-0.01	-0.21	0.48	0.52
⁵ [Transition State]	3.54	-0.12	0.05	-0.22	0.64	0.12
⁵ [Product Complex]	3.96	0.01	0.16	-0.13	0.00	0.00

Table S3 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{Cl}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{Cl}]^+$	Mn-O	Mn-Cl	HO-H _{Naphthalene}	Mn-N _{eq} (average)	C-Cl
Separated (3 components)	1.80	2.29	∞	2.04	∞
Electrophilic (C_1-HAT)					
⁵ [Reactant Complex]	1.83	2.30	2.32	2.03	2.58
⁵ [Transition State]	1.95	2.38	1.57	2.04	1.86
⁵ [Product Complex]	2.45	2.42	0.97	2.05	1.76
Nucleophilic (C_1-hydroxylation)					
⁵ [Reactant Complex]	1.83	2.31	3.34	2.04	2.15
⁵ [Transition State]	1.91	2.36	2.80	2.04	2.15
⁵ [Product Complex]	2.53	2.41	2.42	2.04	2.15
Nucleophilic (C_2-hydroxylation)					
⁵ [Reactant Complex]	1.83	2.30	4.50	2.03	2.59
⁵ [Transition State]	1.92	2.37	2.35	2.04	1.95
⁵ [Product Complex]	2.53	2.41	1.43	2.04	1.85

Table S4 Relative energies (in kcal mol⁻¹) of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^a	ΔZ_0	$\Delta E_{\text{thermal}}^b$	$-T\Delta S^b$	ΔDisp	ΔC_{plex}^c	ΔG^d
Separated (3 components)	33.65	-20.64	13.01	-1.14	-0.76	-16.24	+10.38	+3.79	9.04
⁵ [3 + Naphthalene] + Cl ⁻ (1)	31.58	-19.08	12.49	-0.91	-0.44	-9.28	+2.65	+1.89	6.40
⁵ [3 + Naphthalene] + Cl ⁻ (2)	30.87	-17.52	13.35	-0.94	+0.11	-6.89	+0.85	+1.89	8.38
Electrophilic (C_1-HAT)									
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	6.26	+0.81	7.06	-1.92	-0.36	+0.69	+1.11	+0.00	6.58
⁵ [Product Complex]	-55.36	+1.20	-54.16	+0.20	+0.33	-2.25	-1.92	+0.00	-57.80
³ [Reactant Complex]	-0.06	+0.14	0.08	-0.09	+0.04	-0.28	-0.93	+0.00	-1.18
³ [Transition State]	21.69	+2.41	24.10	-5.11	-0.08	+0.79	+1.11	+0.00	20.81
³ [Product Complex]	-42.10	+6.23	-35.87	+1.67	-0.51	+1.70	+0.14	+0.00	-32.86
Nucleophilic (C_1-hydroxylation)									
⁵ [Reactant Complex]	0.80	-0.57	0.23	-0.13	+0.06	-0.58	+0.34	+0.00	-0.09
⁵ [Transition State]	8.24	-1.17	7.07	-1.18	-0.20	-0.36	+0.51	+0.00	5.84
⁵ [Intermediate Complex]	-39.23	+4.45	-34.78	+2.18	-0.57	+1.41	-6.83	+0.00	-38.60
Nucleophilic (C_2-hydroxylation)									
⁵ [Reactant Complex]	-0.06	+0.15	0.09	-0.07	+0.04	-0.17	-0.45	+0.00	-0.56
⁵ [Transition State]	5.08	-0.70	4.38	-1.13	-0.34	+0.55	-0.60	+0.00	2.85
⁵ [Intermediate Complex]	-39.03	+4.41	-34.62	2.16	-0.61	+0.96	-5.75	+0.00	-37.85

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b $T = 298.15$ K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S5 Mulliken spin density distribution of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$	Mn	OH	$\text{OTf}_{\text{axial}}$	Porphine	Naphthalene	Cl
Separated (3 components)	3.09	0.06	0.02	0.83	0.00	0.00
⁵ [3 + Naphthalene] + Cl ⁻ -(1)	3.05	0.09	0.02	0.82	0.02	0.00
⁵ [3 + Naphthalene] + Cl ⁻ -(2)	3.09	0.06	0.02	0.58	0.25	0.00
Electrophilic (C_1-HAT)						
⁵ [Reactant Complex]	3.11	0.04	0.01	-0.16	0.56	0.44
⁵ [Transition State]	3.55	-0.09	0.02	-0.22	0.68	0.06
⁵ [Product Complex]	4.06	0.03	0.05	-0.13	0.00	0.00
³ [Reactant Complex]	3.10	0.04	0.01	-0.16	-0.56	-0.44
³ [Transition State]	3.90	-0.23	0.04	-0.29	-1.24	-0.17
³ [Product Complex]	2.05	-0.01	0.02	-0.06	0.00	0.00
Nucleophilic (C_1-hydroxylation)						
⁵ [Reactant Complex]	3.10	0.04	0.01	-0.16	1.00	0.44
⁵ [Transition State]	3.40	-0.08	0.02	-0.20	0.86	0.21
⁵ [Intermediate Complex]	4.06	0.01	0.05	-0.13	0.01	0.00
Nucleophilic (C_2-hydroxylation)						
⁵ [Reactant Complex]	3.11	0.04	0.01	-0.16	0.47	0.52
⁵ [Transition State]	3.47	-0.09	0.02	-0.22	0.69	0.13
⁵ [Intermediate Complex]	4.05	0.02	0.06	-0.14	0.00	0.00

Table S6 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+•})\text{OTf}]^+$	Mn-O	Mn-OTf	$\text{HO}-(\text{H}, {}^a \text{C})_{\text{Naphthalene}}$	Mn-N _{eq} (average)	C-Cl
Separated (3 components)	1.79	1.97	∞	2.03	∞
⁵ [3 + Naphthalene] + Cl ⁻ -(1)	1.79	2.01	3.72	2.02	∞
⁵ [3 + Naphthalene] + Cl ⁻ -(2)	1.79	1.97	3.86	2.03	∞
Electrophilic (C_1-HAT)					
⁵ [Reactant Complex]	1.79	2.00	2.39	2.02	2.58
⁵ [Transition State]	1.89	2.18	1.75	2.04	1.88
⁵ [Product Complex]	2.32	2.11	0.97	2.04	1.77
³ [Reactant Complex]	1.79	1.80	2.38	2.02	2.58
³ [Transition State]	2.07	2.07	1.24	2.04	1.87
³ [Product Complex]	2.03	2.04	0.98	2.03	1.77
Nucleophilic (C_1-hydroxylation)					
⁵ [Reactant Complex]	1.79	2.00	3.28	2.02	2.57
⁵ [Transition State]	1.85	2.07	2.00	2.03	2.79
⁵ [Intermediate Complex]	2.39	2.16	1.44	2.03	2.76
Nucleophilic (C_2-hydroxylation)					
⁵ [Reactant Complex]	1.80	2.00	3.42	2.02	3.42
⁵ [Transition State]	1.85	2.08	1.97	2.04	1.97
⁵ [Intermediate Complex]	2.40	2.16	1.40	2.03	1.40

^a OH-H_{Naphthalene} distance for the electrophilic reaction. ^b OH-C_{Naphthalene} distance for the nucleophilic reaction.

Electrophilic (C_1 -HAT) and nucleophilic (C_1 - and C_2 -hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$ and naphthalene in absence of chloride

Table S7 Relative Energies (in kcal mol⁻¹) of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$ and naphthalene.

$[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^a	ΔZ_0	$\Delta E_{\text{thermal}}^b$	$-T\Delta S^b$	ΔDisp	ΔCplx^c	ΔG^d
Separated RC (2 components)	2.60	-1.87	0.73	-0.26	-0.88	-8.89	+12.21	+1.89	4.80
Electrophilic (C_1-HAT)									
⁵ [Reactant Complex]-1	+0.53	-0.32	0.20	-0.02	+0.03	-1.93	+2.20	+1.89	2.37
⁵ [Reactant Complex]-2	-0.18	+1.24	1.07	-0.05	-0.01	+0.46	+1.22	+0.00	2.68
⁵ [Transition State]	+32.16	+4.81	36.97	-3.43	-0.39	+1.85	+0.29	+0.00	35.30
⁵ [Product Complex]	+29.20	+4.66	33.86	-0.41	+0.10	-0.24	+2.20	+0.00	35.50
Nucleophilic (C_1-hydroxylation) α-electron transfer									
⁵ [Reactant Complex]	0.51	-0.35	0.17	+0.05	+0.01	-1.65	+2.97	+0.00	1.55
⁵ [Transition State]	10.22	+0.94	11.16	-1.72	+0.08	+0.37	-1.83	+0.00	8.05
⁵ [Intermediate Complex]	-16.76	+4.03	-12.73	+1.67	-0.15	+0.46	-2.02	+0.00	-12.78
Nucleophilic (C_1-hydroxylation) β-electron transfer									
⁵ [Reactant Complex]	+0.53	-0.32	0.20	+0.01	+6.46	-1.93	+3.01	+0.00	7.75
⁵ [Transition State]	+11.18	+3.42	14.59	+0.16	-7.94	+2.32	-0.59	+0.00	8.55
⁵ [Intermediate Complex]	-16.75	+4.05	-12.69	+1.65	-1.62	-0.01	-2.10	+0.00	-14.77
Nucleophilic (C_2-hydroxylation) α-electron transfer									
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	12.40	+0.71	13.11	-1.80	+0.08	-0.02	-0.61	+0.00	10.75
⁵ [Intermediate Complex]	-11.77	+3.19	-8.58	+1.20	-0.14	+0.85	-0.75	+0.00	-7.41
Nucleophilic (C_2-hydroxylation) β-electron transfer									
⁵ [Reactant Complex]	1.14	-0.36	0.78	-0.01	+0.12	-2.52	+6.84	+0.00	5.21
⁵ [Transition State]	14.20	+3.21	17.41	+0.26	-0.54	+2.36	+1.00	+0.00	20.49
⁵ [Intermediate Complex]	-12.10	+3.47	-8.62	+1.26	-0.17	+1.12	-0.61	+0.00	-7.01

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b $T = 298.15 \text{ K}$. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S8 Mulliken spin density distribution of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$ and naphthalene.

$[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$	Mn	OH	$\text{OTf}_{\text{axial}}$	Porphine	Naphthalene
Separated RC (2 components)	3.09	0.06	0.02	0.83	0.00
Electrophilic (C_1-HAT)					
⁵ [Reactant Complex]-1	3.05	0.09	0.02	0.82	0.02
⁵ [Reactant Complex]-2	3.09	0.07	0.03	0.55	0.27
⁵ [Transition State]	3.04	0.06	0.02	0.00	0.88
⁵ [Product Complex]	3.02	0.02	0.02	-0.01	0.95
Nucleophilic (C_1) α-transfer					
⁵ [Reactant Complex]	3.05	0.10	0.03	0.79	0.02
⁵ [Transition State]	3.87	-0.23	0.06	0.82	-0.51
⁵ [Intermediate Complex]	4.05	0.01	0.07	-0.13	0.00
Nucleophilic (C_1) β-transfer					
⁵ [Reactant Complex]	3.05	0.09	0.02	0.82	0.02
⁵ [Transition State]	3.31	0.00	0.03	-0.11	0.77
⁵ [Intermediate Complex]	4.06	0.00	0.06	-0.13	0.01
Nucleophilic (C_2) α-transfer					
⁵ [Reactant Complex]	3.06	0.08	0.02	0.74	0.09
⁵ [Transition State]	3.88	-0.23	0.05	0.83	-0.54
⁵ [Intermediate Complex]	4.05	0.00	0.06	-0.13	0.01
Nucleophilic (C_2) β-transfer					
⁵ [Reactant Complex]	3.07	0.08	0.03	0.81	0.01
⁵ [Transition State]	3.33	0.01	0.03	-0.12	0.75
⁵ [Intermediate Complex]	4.05	0.01	0.07	-0.13	0.00

Table S9 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$ and naphthalene.

$[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$	Mn-O	Mn-OTf	HO-HNaphthalene	Mn-N _{eq} (average)
Separated RC (2 components)	1.79	1.97	3.72	2.03
Electrophilic (C_1-HAT)				
⁵ [Reactant Complex]-1	1.79	2.01	3.72	2.02
⁵ [Reactant Complex]-2	1.79	1.97	2.50	2.03
⁵ [Transition State]	1.95	1.89	1.12	1.99
⁵ [Product Complex]	2.04	1.86	1.00	1.99
Nucleophilic (C_1) α-transfer				
⁵ [Reactant Complex]	1.77	1.98	4.11	2.03
⁵ [Transition State]	2.03	2.15	2.11	2.04
⁵ [Intermediate Complex]	2.50	2.13	1.41	2.03
Nucleophilic (C_1) β-transfer				
⁵ [Reactant Complex]	1.77	1.98	4.14	2.03
⁵ [Transition State]	1.92	1.96	1.92	2.02
⁵ [Intermediate Complex]	2.5	2.13	1.41	2.03
Nucleophilic (C_2) α-transfer				
⁵ [Reactant Complex]	1.77	1.98	3.33	2.03
⁵ [Transition State]	2.06	2.15	2.03	2.04
⁵ [Intermediate Complex]	2.49	2.13	1.41	2.04
Nucleophilic (C_2) β-transfer				
⁵ [Reactant Complex]	1.79	1.97	2.50	2.03
⁵ [Transition State]	1.95	1.89	1.12	2.01
⁵ [Intermediate Complex]	2.04	1.86	1.00	2.03

Electrophilic (C₁- and C₂-HAT) and nucleophilic (C₁- and C₂-hydroxylation) reaction of [Mn^{IV}(OH)(TDCPP⁺⁺)OTf]⁺, naphthalene and chloride with [OTf]⁻ hydrogen bonded to Mn^{IV}(OH) (large model).

Table S10 Relative Energies (in kcal mol⁻¹) of the electrophilic (C₁- and C₂-HAT) and nucleophilic (C₁ and C₂ hydroxylation) reaction of [Mn^{IV}(OH)(TDCPP⁺⁺)OTf]⁺, naphthalene and chlorine with [OTf]⁻ hydrogen bonded to Mn^{IV}(OH).

[Mn ^{IV} (OH)(TDCPP ⁺⁺)OTf] ⁺	ΔDef2-SVP	ΔΔDef2-TZVPP	ΔE ^a	ΔZ ₀	ΔE _{thermal} ^b	-TΔS ^b	ΔDisp	ΔCplx ^c	ΔG ^d
Separated (4 components)	47.46	-30.79	16.67	-2.20	-0.95	-30.61	+22.45	+5.68	11.04
{ ⁵ 3 + [OTf] ⁻ } + Naphthalene + Cl ⁻	27.38	-20.83	6.55	-1.58	-0.70	-18.36	+12.93	+3.79	2.62
{ ⁵ 3 + [OTf] ⁻ } + Naphthalene } + Cl ⁻	23.20	-15.49	7.71	-1.30	+0.25	-9.33	+4.75	+1.89	3.97
Electrophilic (C₁-HAT)									
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	8.05	+1.16	9.21	-1.79	-0.35	+1.66	-2.71	+0.00	6.01
⁵ [Product Complex]	-57.63	+1.17	-56.46	+0.27	+0.39	-2.09	-0.68	+0.00	-58.57
Electrophilic (C₂-HAT)									
⁵ [Reactant Complex]	0.35	+0.99	1.33	-0.07	-0.04	+0.42	-1.40	+0.00	0.25
⁵ [Transition State]	27.77	+1.29	29.06	-5.56	+0.35	-1.20	-5.74	+0.00	16.92
⁵ [Product Complex]	-58.57	+1.11	-57.46	+0.31	+0.32	-1.40	-1.21	+0.00	-59.44
Nucleophilic (C₁-hydroxylation)									
⁵ [Reactant Complex]	1.79	+0.72	2.51	-0.04	-0.04	+0.45	-0.74	+0.00	2.14
⁵ [Transition State]	11.71	+1.11	12.83	-0.91	-0.22	+2.08	-7.17	+0.00	6.62
⁵ [Intermediate Complex]	-38.49	+5.84	-32.65	+2.12	-0.57	+2.59	-13.07	+0.00	-41.58
Nucleophilic (C₂-hydroxylation)									
⁵ [Reactant Complex]	3.43	-1.38	2.05	-0.20	+0.04	+0.29	-1.50	+0.00	0.67
⁵ [Transition State]	7.74	+2.07	9.81	-0.94	-0.46	+3.69	-5.66	+0.00	6.43
⁵ [Intermediate Complex]	-38.05	+6.95	-31.10	+2.20	-0.70	+3.26	-12.80	+0.00	-39.15

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b T = 298.15 K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S11 Mulliken spin density distribution of the electrophilic (C_1 - and C_2 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^*)\text{OTf}]^+$, naphthalene and chlorine with $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{IV}(\text{OH})$.

$[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^*)\text{OTf}]^+$	Mn	OH	$\text{OTf}_{\text{axial}}$	$\text{OTf}_{\text{H-bond}}$	TDCPP	Naphthalene	Cl
Separated (4 components)	2.97	0.15	0.02	0.00	0.84	0.00	0.00
$\{\text{53} + [\text{OTf}]^-\} + \text{Naphthalene} + \text{Cl}^-$	2.97	0.15	0.02	0.01	0.85	0.00	0.00
$\{\text{53} + [\text{OTf}]^-\} + \text{Naphthalene } + \text{Cl}^-$	3.05	0.07	0.03	0.01	0.85	0.00	0.00
Electrophilic (C_1-HAT)							
${}^5[\text{Reactant Complex}]$	3.04	0.11	0.01	0.00	-0.16	0.48	0.52
${}^5[\text{Transition State}]$	3.60	-0.08	0.02	0.01	-0.21	0.62	0.05
${}^5[\text{Product Complex}]$	4.04	0.04	0.05	0.00	-0.13	0.00	0.00
Electrophilic (C_2-HAT)							
${}^5[\text{Reactant Complex}]$	3.05	0.11	0.00	0.00	-0.16	0.48	0.52
${}^5[\text{Transition State}]$	3.93	-0.30	0.03	0.00	-0.16	0.04	0.46
${}^5[\text{Product Complex}]$	4.04	0.03	0.04	0.00	-0.11	0.00	0.00
Nucleophilic (C_1-hydroxylation)							
${}^5[\text{Reactant Complex}]$	3.04	0.09	0.00	0.01	-0.14	0.57	0.43
${}^5[\text{Transition State}]$	3.60	-0.17	0.02	0.00	-0.21	0.41	0.34
${}^5[\text{Intermediate Complex}]$	4.05	0.03	0.05	0.00	-0.14	0.01	0.00
Nucleophilic (C_2-hydroxylation)							
${}^5[\text{Reactant Complex}]$	3.05	0.09	0.00	0.01	-0.15	0.60	0.40
${}^5[\text{Transition State}]$	3.47	-0.09	0.02	0.00	-0.20	0.65	0.15
${}^5[\text{Intermediate Complex}]$	4.05	0.01	0.04	0.00	-0.12	0.01	0.00

Table S12 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 - and C_2 -HAT) and nucleophilic (C_1 and C_2 hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^*)\text{OTf}]^+$, naphthalene and chlorine with $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{IV}(\text{OH})$.

$[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^*)\text{OTf}]^+$	Mn-O	$\text{Mn-OTf}_{\text{axial}}$	$\text{HO-H}_{\text{Naphthalene}}$	$\text{Mn-N}_{\text{eq}} (\text{average})$	$\text{OH-OTf}_{\text{H-bond}}$	C-Cl
Separated (4 components)	1.74	2.01	∞	2.03	∞	∞
$\{\text{53} + [\text{OTf}]^-\} + \text{Naphthalene} + \text{Cl}^-$	1.74	2.01	∞	2.01	1.55	∞
$\{\text{53} + [\text{OTf}]^-\} + \text{Naphthalene } + \text{Cl}^-$	1.79	1.97	2.68	2.01	1.60	∞
Electrophilic (C_1-HAT)						
${}^5[\text{Reactant Complex}]$	1.76	2.05	2.66	2.01	1.66	2.56
${}^5[\text{Transition State}]$	1.88	2.17	1.78	2.03	1.80	1.88
${}^5[\text{Product Complex}]$	2.25	2.21	0.97	2.03	1.66	1.76
Electrophilic (C_2-HAT)						
${}^5[\text{Reactant Complex}]$	1.75	2.06	2.51	2.01	1.66	2.95
${}^5[\text{Transition State}]$	2.11	2.20	1.16	2.02	1.68	2.95
${}^5[\text{Product Complex}]$	2.25	2.22	0.97	2.03	1.69	1.76
Nucleophilic (C_1-hydroxylation)						
${}^5[\text{Reactant Complex}]$	1.75	2.06	3.46	2.02	1.66	2.57
${}^5[\text{Transition State}]$	1.88	2.16	2.39	2.03	1.79	2.84
${}^5[\text{Intermediate Complex}]$	2.35	2.18	1.43	2.04	1.67	2.77
Nucleophilic (C_2-hydroxylation)						
${}^5[\text{Reactant Complex}]$	1.75	2.06	3.42	2.02	1.67	2.56
${}^5[\text{Transition State}]$	1.84	2.16	2.60	2.03	1.80	1.99
${}^5[\text{Intermediate Complex}]$	2.37	2.21	1.43	2.03	1.60	1.86

Electrophilic (C₁-HAT) reaction of Mn^{IV}-oxo species [Mn^{IV}(O)(Por⁺)OTf], naphthalene and chloride

Table S13 Relative energies (in kcal mol⁻¹) of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(O)(Por⁺)OTf], naphthalene and chlorine.

[Mn ^{IV} (O)(Por ⁺)OTf]	ΔDef2-SVP	ΔΔDef2-TZVPP	ΔE ^a	ΔZ ₀	ΔE _{thermal} ^b	-TΔS ^b	ΔDisp	ΔCplx ^c	ΔG ^d
⁵ [3 + Naphthalene] + Cl ⁻	16.77	-18.60	-1.83	-2.14	-2.01	+4.94	+14.28	+1.89	15.14
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	12.83	+1.11	13.94	-4.27	-0.52	+1.39	-0.59	+0.00	9.96
⁵ [Product Complex]	-43.55	-1.50	-45.05	-0.76	+0.24	-1.46	+1.40	+0.00	-45.63

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b T = 298.15 K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S14 Mulliken spin density distribution of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(O)(Por⁺)OTf], naphthalene and chlorine.

[Mn ^{IV} (O)(Por ⁺)OTf]	Mn	O	OTf _{axial}	Porphine	Naphthalene	Cl
⁵ [3 + Naphthalene] + Cl ⁻	2.56	0.57	0.01	0.78	0.00	0.00
⁵ [Reactant Complex]	2.65	0.48	0.00	-0.14	0.00	0.43
⁵ [Transition State]	3.16	0.22	0.01	-0.16	0.77	0.05
⁵ [Product Complex]	3.93	0.18	0.01	-0.13	0.01	0.00

Table S15 Selected DFT optimized geometries (in Å) of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(O)(Por⁺)OTf], naphthalene and chlorine.

[Mn ^{IV} (O)(Por ⁺)OTf]	Mn-O	Mn-OTf	O-H _{Naphthalene}	Mn-N _{eq} (average)	C-Cl
⁵ [3 + Naphthalene] + Cl ⁻	1.64	2.14	∞	2.05	∞
⁵ [Reactant Complex]	1.65	2.21	2.22	2.04	2.60
⁵ [Transition State]	1.71	2.26	1.31	2.05	1.86
⁵ [Product Complex]	1.98	2.51	0.97	2.05	1.76

Electrophilic (C_1 - and C_2 -HAT) and nucleophilic (C_1 - and C_2 -hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$ or $[\text{Mn}^{\text{IV}}(\text{OH})(\text{TDCPP}^{+\bullet})\text{OTf}]^+$, naphthalene, chloride and/or $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{\text{IV}}(\text{OH})$ in other combinations than above.

Table S16 Relative energies (in kcal mol⁻¹) of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$, naphthalene and chlorine with $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{\text{IV}}(\text{OH})$.

$[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^a	ΔZ_0	$\Delta E_{\text{thermal}}^b$	$-T\Delta S^b$	ΔDisp	ΔC_{plex}^c	ΔG^d
Separated (4 components)	50.21	-31.39	18.82	-2.11	-2.93	-27.99	+18.96	+5.68	12.74
{ ⁵ 3 + Naphthalene + $[\text{OTf}]^-$ } + Cl^-	24.46	-16.48	7.98	-1.06	-0.17	-7.51	+2.18	+1.89	3.65
Electrophilic (C_1-HAT)									
⁵ [Reactant Complex]	3.07	-1.45	1.62	-0.22	+0.08	-0.21	-0.98	+0.00	0.28
⁵ [Transition State]	10.48	+0.08	10.56	-2.34	-0.40	+2.57	-1.87	+0.00	8.52
⁵ [Product Complex]	-54.83	+0.51	-54.32	+0.15	+0.32	-1.30	-1.68	+0.00	-56.83
Nucleophilic (C_2-hydroxylation)									
⁵ [Reactant Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	8.98	-0.43	8.55	-1.25	-0.39	+2.50	-3.74	+0.00	5.67
⁵ [Intermediate Complex]	-35.34	3.39	-31.94	+1.70	-0.52	+1.77	-8.99	+0.00	-37.98

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b $T = 298.15$ K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S17 Mulliken spin density distribution of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$, naphthalene and chlorine with $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{\text{IV}}(\text{OH})$.

$[\text{Mn}^{\text{IV}}(\text{OH})(\text{Por}^{+\bullet})\text{OTf}]^+$	Mn	OH	OTf _{axial}	OTf _{H-bond}	Porphine	Naphthalene	Cl
Separated (4 components)	3.09	0.06	0.02	0.00	0.83	0.00	0.00
{ ⁵ 3 + Naphthalene + $[\text{OTf}]^-$ } + Cl^-	3.02	0.13	0.02	0.01	0.81	0.02	0.00
Electrophilic (C_1-HAT)							
⁵ [Reactant Complex]	3.06	0.09	0.01	0.00	-0.16	0.57	0.43
⁵ [Transition State]	3.58	-0.09	0.02	0.00	-0.22	0.66	0.05
⁵ [Product Complex]	4.06	0.03	0.04	0.00	-0.13	0.00	0.00
Nucleophilic (C_2-hydroxylation)							
⁵ [Reactant Complex]	3.06	0.09	0.01	0.00	-0.17	0.60	0.40
⁵ [Transition State]	3.49	-0.10	0.02	0.00	-0.21	0.67	0.13
⁵ [Intermediate Complex]	4.06	0.02	0.04	0.00	-0.13	0.01	0.00

Table S18 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^+)^*\text{OTf}]^+$, naphthalene and chlorine with $[\text{OTf}]^-$ hydrogen bonded to $\text{Mn}^{IV}(\text{OH})$.

$[\text{Mn}^{IV}(\text{OH})(\text{Por}^+)^*\text{OTf}]^+$	Mn-O	$\text{Mn}_1\text{-OTf}_{\text{axial}}$	$\text{HO-H}_{\text{Naphthalene}}$	$\text{Mn-N}_{\text{eq}} \text{(average)}$	$\text{OH-OTf}_{\text{H-bond}}$	C-Cl
Separated (4 components)	1.79	1.90	∞	2.03	∞	∞
$\{^5\text{3} + \text{Naphthalene} + [\text{OTf}]^- + \text{Cl}^-\}$	1.75	2.01	2.66	2.03	1.60	∞
Electrophilic (C_1-HAT)						
${}^5[\text{Reactant Complex}]$	1.76	2.05	2.46	2.02	1.68	2.57
${}^5[\text{Transition State}]$	1.87	2.16	1.69	2.03	1.77	1.87
${}^5[\text{Product Complex}]$	2.25	2.22	0.97	2.04	1.69	1.76
Nucleophilic (C_2-hydroxylation)						
${}^5[\text{Reactant Complex}]$	1.76	2.05	4.92	2.02	1.66	2.60
${}^5[\text{Transition State}]$	1.84	2.14	2.50	2.03	1.78	1.96
${}^5[\text{Intermediate Complex}]$	2.34	2.20	1.43	2.03	1.65	1.86

Table S19 Relative energies (in kcal mol⁻¹) of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{+*})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{+*})\text{OTf}]^+$	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^a	ΔZ_0	$\Delta E_{\text{thermal}}^b$	$-\text{T}\Delta S^b$	ΔDisp	ΔCplx^c	ΔG^d
Separated (3 components)	34.55	-20.42	14.13	-1.44	-0.61	-17.56	+11.57	+3.79	9.89
$\{^5\text{3} + \text{Naphthalene} + \text{Cl}^-\}$	31.68	-17.00	14.68	-0.87	+0.19	-8.36	+5.34	+1.89	12.86
Electrophilic (C_1-HAT)									
${}^5[\text{Reactant Complex}]$	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	0.00	0.00
${}^5[\text{Transition State}]$	5.74	+1.48	7.23	-1.94	-0.32	+1.32	-0.21	+0.00	6.08
${}^5[\text{Product Complex}]$	-54.76	+0.67	-54.09	-0.01	+0.54	-1.89	-3.83	+0.00	-59.28
Nucleophilic (C_2-hydroxylation)									
${}^5[\text{Reactant Complex}]$	0.84	-0.69	0.15	+0.02	+0.00	-0.57	+0.42	+0.00	0.03
${}^5[\text{Transition State}]$	3.92	+0.66	4.58	-1.07	-0.35	+1.74	-2.11	+0.00	2.79
${}^5[\text{Intermediate Complex}]$	-40.38	+4.55	-35.83	+2.15	-0.53	+0.68	-7.01	+0.00	-40.53

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b $T = 298.15$ K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S20 Mulliken spin density distribution of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{+*})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{+*})\text{OTf}]^+$	Mn	OH	$\text{OTf}_{\text{axial}}$	TDCPP	Naphthalene	Cl
Separated (3 components)	2.97	0.15	0.02	0.84	0.00	0.00
$\{^5\text{3} + \text{Naphthalene} + \text{Cl}^-\}$	3.05	0.07	0.03	0.85	0.00	0.00
Electrophilic (C_1-HAT)						
${}^5[\text{Reactant Complex}]$	3.07	0.04	0.01	-0.12	0.57	0.43
${}^5[\text{Transition State}]$	3.55	-0.09	0.02	-0.20	0.65	0.06
${}^5[\text{Product Complex}]$	4.04	0.02	0.05	-0.11	0.00	0.00
Nucleophilic (C_2-hydroxylation)						
${}^5[\text{Reactant Complex}]$	3.07	0.04	0.04	-0.13	0.56	0.44
${}^5[\text{Transition State}]$	3.44	-0.09	0.02	-0.20	0.69	0.14
${}^5[\text{Intermediate Complex}]$	4.04	0.01	0.05	-0.12	0.01	0.00

Table S21 Selected DFT optimized geometries (in Å) of the electrophilic (C_1 -HAT) and nucleophilic (C_2 -hydroxylation) reaction of $[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{\bullet\bullet})\text{OTf}]^+$, naphthalene and chlorine.

$[\text{Mn}^{IV}(\text{OH})(\text{TDCPP}^{\bullet\bullet})\text{OTf}]^+$	Mn-O	Mn-OTf	$\text{HO-H}_{\text{Naphthalene}}$	Mn-N _{eq} (average)	C-Cl
Separated (3 components)	1.74	2.01	∞	2.03	∞
$^5[3 + \text{Naphthalene}] + \text{Cl}^-$	1.79	1.97	2.68	2.01	∞
Electrophilic (C_1-HAT)					
$^5[\text{Reactant Complex}]$	1.80	2.04	1.91	2.02	2.57
$^5[\text{Transition State}]$	1.89	2.11	1.76	2.02	1.88
$^5[\text{Product Complex}]$	2.08	2.14	1.00	2.03	1.75
Nucleophilic (C_2-hydroxylation)					
$^5[\text{Reactant Complex}]$	1.80	2.00	4.57	2.01	2.56
$^5[\text{Transition State}]$	1.85	2.07	2.59	2.03	1.96
$^5[\text{Intermediate Complex}]$	2.39	2.15	1.44	2.03	1.85

Possible reaction step after the initial nucleophilic (C₂-hydroxylation) reaction

Table S22 Relative energies (in kcal mol⁻¹) of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(OH)(TDCPP^{•+})OTf]⁺ and 1-chloro-1,2-dihydroronaphthalen-2-ol (naphthalene-OH) with [OTf]⁻ hydrogen bonded to Mn^{IV}(OH).

[Mn ^{IV} (OH)(TDCPP ^{•+})OTf] ⁺	ΔDef2-SVP	ΔΔDef2-TZVPP	ΔE ^a	ΔZ ₀	ΔE _{thermal} ^b	-TΔS ^b	ΔDisp	ΔCplx ^c	ΔG ^d
⁵ [Intermediate-1 Complex]	0.00	+0.00	0.00	+0.00	+0.00	+0.00	+0.00	+0.00	0.00
⁵ [Transition State]	12.04	+1.59	13.62	-3.59	+0.03	+0.71	-6.94	+0.00	3.83
⁵ [Intermediate-2 Complex]	-35.58	-2.13	-37.71	-0.81	+1.02	-4.00	+3.22	+0.00	-38.28

^a Def2-TZVPP electronic energy = sum of the two previous columns. ^b T = 298.15 K. ^c Complexation energy due to change of standard states of the solvent. ^d Gibb's free energy = sum of the previous six columns.

Table S23 Mulliken spin density distribution of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(OH)(TDCPP^{•+})OTf]⁺ and 1-chloro-1,2-dihydroronaphthalen-2-ol (naphthalene-OH) with [OTf]⁻ hydrogen bonded to Mn^{IV}(OH).

[Mn ^{IV} (OH)(TDCPP ^{•+})OTf] ⁺	Mn	OH	OTf _{axial}	OTf _{H-bond}	TDCPP	NaphthaleneOH	Cl
⁵ [Intermediate-1 Complex]	3.00	0.12	0.02	0.00	0.85	0.00	0.00
⁵ [Transition State]	3.78	-0.21	0.04	0.00	0.85	-0.42	-0.03
⁵ [Intermediate-2 Complex]	4.04	0.02	0.05	0.00	-0.11	0.00	0.00

Table S24 Selected DFT optimized geometries (in Å) of the electrophilic (C₁-HAT) reaction of [Mn^{IV}(OH)(TDCPP^{•+})OTf]⁺ and 1-chloro-1,2-dihydroronaphthalen-2-ol (naphthalene-OH) with [OTf]⁻ hydrogen bonded to Mn^{IV}(OH).

[Mn ^{IV} (OH)(TDCPP ^{•+})OTf] ⁺	Mn-O	Mn-OTf _{axial}	HO-HNaphthalene	Mn-N _{eq} (average)	OH-OTf _{H-bond}
⁵ [Intermediate-1 Complex]	1.76	2.00	2.78	2.02	1.68
⁵ [Transition State]	2.00	2.16	1.47	2.03	1.92
⁵ [Intermediate-2 Complex]	2.30	2.19	0.97	2.03	1.77

Ionization and Electron Affinities

Table S25 Ionisation energies of naphthalene and Cl^- , both separate and complexed (in kcal mol $^{-1}$).

Compound (unionized) ^a	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^b	ΔZ_0	$\Delta E_{\text{thermal}}^c$	$-T\Delta S^c$	ΔDisp	ΔG^d
Naphthalene	142.41	+0.55	142.96	-0.25	+0.15	-0.68	+0.00	142.18
Cl^- (anion)	121.15	+20.29	141.44	+0.00	+0.00	-0.41	+0.00	141.03
[Cl^- + Naphthalene] (in complex)	107.12	+14.50	121.62	+0.01	-0.06	+0.23	-1.27	120.54

^a Ionization energies are expresses as the energy of the structurally optimized ionized compound minus the unionized compound. ^b Def2-TZVPP electronic energy = sum of the two previous columns. ^c $T = 298.15$ K. ^d Gibb's free energy = sum of the previous five columns.

Table S26 Mulliken spin density distribution of the ionised compounds.

Compound	Naphthalene	Cl
Naphthalene ⁺ *	1.00	—
Cl^*	—	1.00
[Cl + Naphthalene] [•]	0.44	0.56

Table S27 Electron Affinity of $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+*})\text{OTf}]^+$ (in kcal mol $^{-1}$).

Compound (unreduced) ^a	$\Delta\text{Def2-SVP}$	$\Delta\Delta\text{Def2-TZVPP}$	ΔE^b	ΔZ_0	$\Delta E_{\text{thermal}}^c$	$-T\Delta S^c$	ΔDisp	ΔG^d
⁵ $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+*})\text{OTf}]^+$	-136.36	+2.08	-134.28	+0.35	+0.01	+7.59	-0.53	-126.85

^a Electron affinity is expresses as the energy of the structurally optimized reduced compound minus the unreduced compound. ^b Def2-TZVPP electronic energy = sum of the two previous columns. ^c $T = 298.15$ K. ^d Gibb's free energy = sum of the previous five columns.

Table S28 Mulliken spin density distribution of the unreduced $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+*})\text{OTf}]^+$ and reduced catalyst (**2OH-OTf**).

Compound	Mn	OH	OTf _{axial}	Porphine
⁵ $[\text{Mn}^{IV}(\text{OH})(\text{Por}^{+*})\text{OTf}]^+$	3.06	0.08	0.03	0.83
⁴ $[\text{Mn}^{IV}(\text{OH})(\text{Por})\text{OTf}]$	3.10	0.05	0.01	-0.17

Table S29 One-electron oxidation potentials of substrates, second-order rate constants (k_2), and $\log(k_2)$ values for the aromatic ring activation of various substrates by **3**/Cl⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C.

substrate	E_{ox} vs. SCE, V	k_2 , M ⁻¹ s ⁻¹	$\log(k_2)$
1-Me-naphthalene	1.57	1.4(1) × 10 ²	2.1
naphthalene	1.76	2.0(2)	0.30
1-Br-naphthalene	1.81	8.3(6) × 10 ⁻¹	-0.081
1-NO ₂ -naphthalene	2.10	6.5(5) × 10 ⁻³	-2.2
anisole	1.76	2.3(2)	0.36
toluene	2.10	2.5(2) × 10 ⁻²	-1.6
benzene	2.23	4.7(4) × 10 ⁻³	-2.3

Table S30 One-electron oxidation potentials of substrates, second-order rate constants (k_2), and $\log(k_2)$ values for the OAT reaction of various substrates by **3**/Cl⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -40 °C.

substrate	E_{ox} vs. SCE, V	k_2 , M ⁻¹ s ⁻¹	$\log(k_2)$
4-CN-thioanisole	1.67	6.8(5) × 10 ⁻²	2.83
4-CHO-thioanisole	1.67	1.9(1) × 10 ⁻²	2.27
4-NO ₂ -thioanisole	1.73	82(6)	1.91

Table S31 One-electron oxidation potentials of substrates, second-order rate constants (k_2), and $\log(k_2)$ values for the oxidation reaction of styrene derivatives by **3**/Cl⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C.

substrate	E_{ox} vs. SCE, V	k_2 , M ⁻¹ s ⁻¹	$\log(k_2)$
4-Me-styrene	1.80	1.9(2)	0.28
styrene	1.86	4.7(4) × 10 ⁻¹	-0.33
4-Cl-styrene	1.93	2.4(2) × 10 ⁻¹	-0.62
3-Cl-styrene	2.04	6.8(5) × 10 ⁻²	-1.17
4-NO ₂ -styrene	2.38	2.5(2) × 10 ⁻³	-2.60

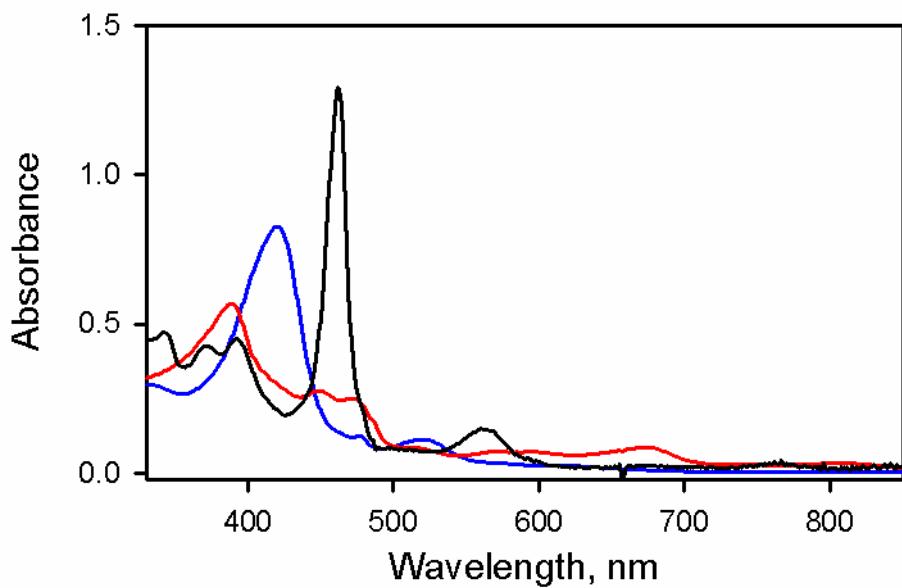
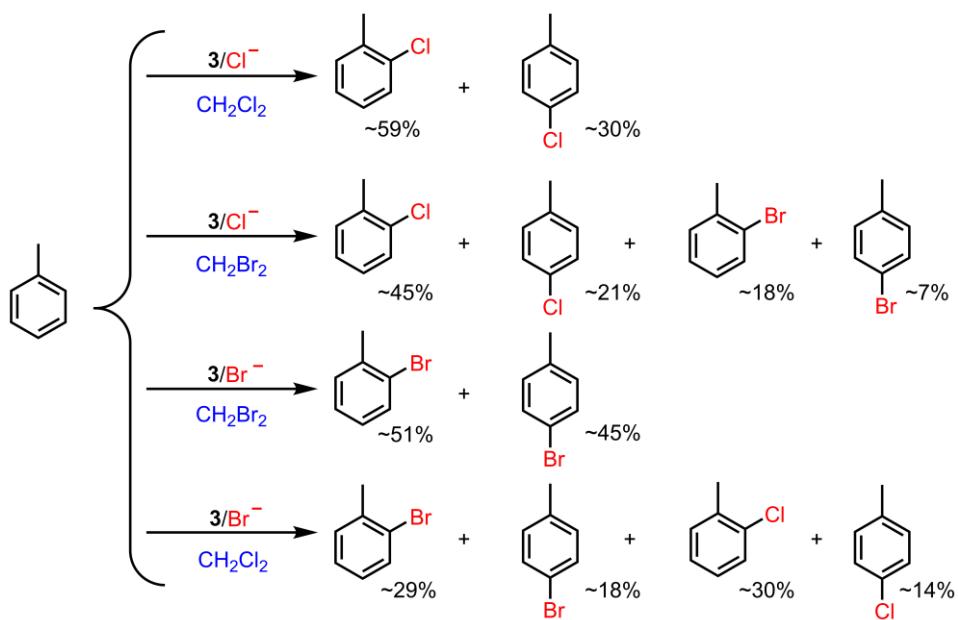
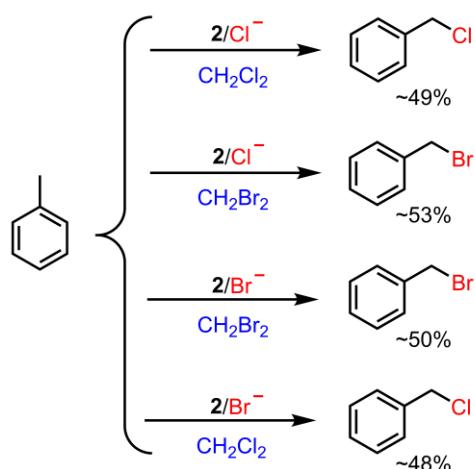


Fig. S1 UV-vis spectra of **1**/Cl⁻ (0.10 mM; black line), **2**/Cl⁻ (0.10 mM; blue line) and **3**/Cl⁻ (0.10 mM; red line) in CH₂Cl₂ at -10 °C.

(a) Reaction of $[\text{Mn}^{\text{IV}}(\text{OH})(\text{TDCPP}^{+•})]^{2+}$ ($\mathbf{3}/\text{Cl}^-$)



(b) Reaction of $[\text{Mn}^{\text{IV}}(\text{O})(\text{TDCPP})]^{(2/\text{X}^-)}$



(c) Reaction of $[\text{Mn}^{\text{III}}(^{\text{s}}\text{PhIO})(\text{TDCPP})]^+ (1/\text{X}^-)$

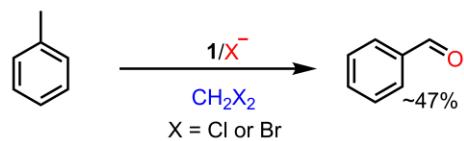


Fig. S2 Products produced in the oxidation of toluene by $[\text{Mn}^{\text{IV}}(\text{OH})(\text{TDCPP}^{+•})]^{2+}$, $[\text{Mn}^{\text{IV}}(\text{O})(\text{TDCPP})]$ and $[\text{Mn}^{\text{III}}(^{\text{s}}\text{PhIO})(\text{TDCPP})]^+$ with different counter anion in different solvent system at -10°C .

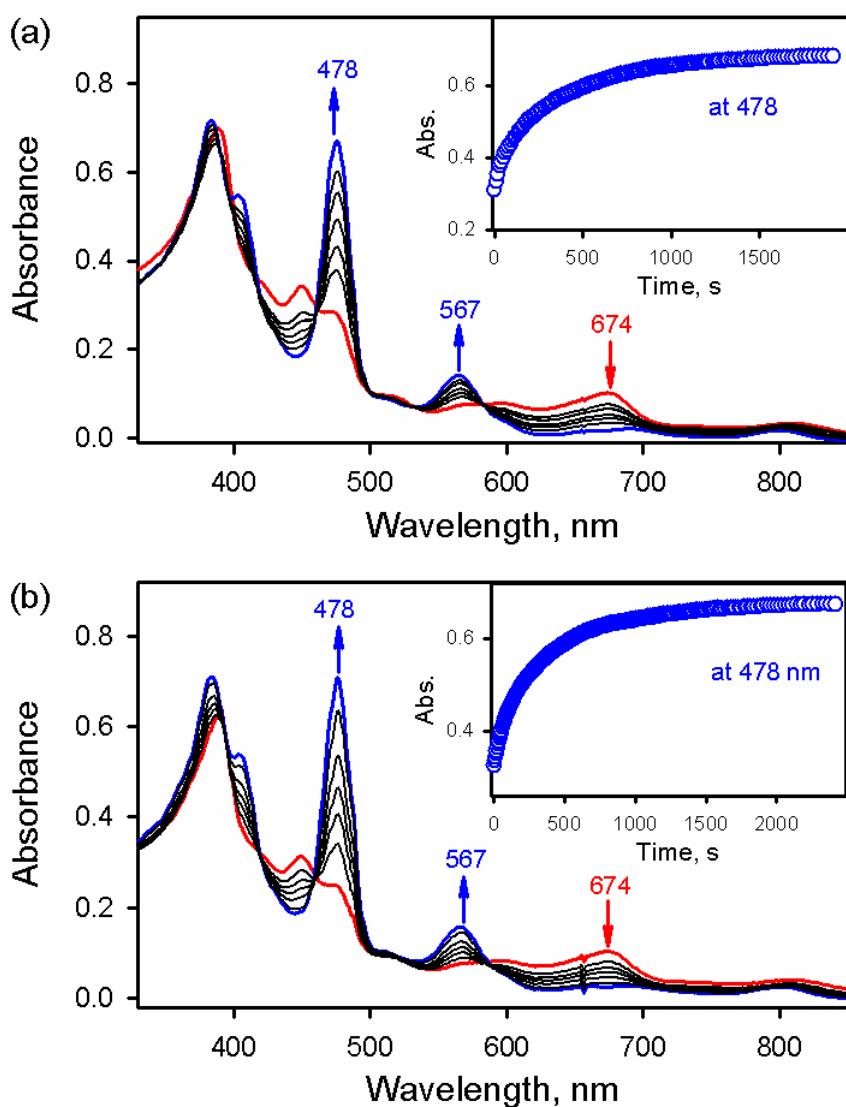


Fig. S3 UV-vis spectral changes observed in the reaction of **3**/Cl⁻ (0.10 mM; red line) with (a) naphthalene (2.0 mM) and (b) benzene (0.60 M) in CH₂Cl₂/CH₃CN (*v/v* 20:1) at -10 °C.

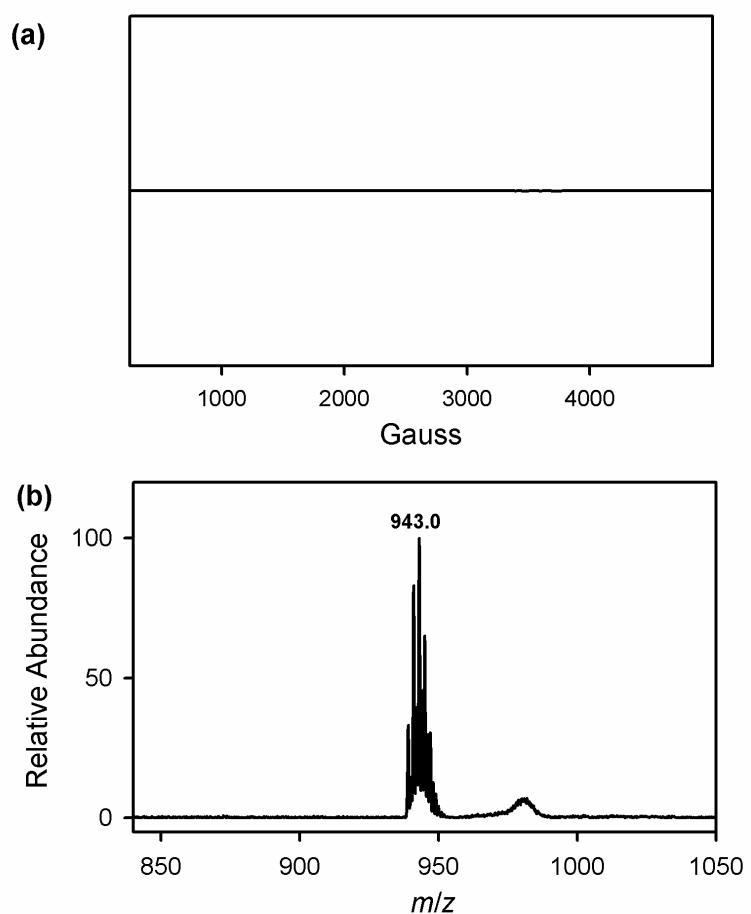


Fig. S4 (a) X-band CW-EPR spectrum of the complete reaction solution obtained in the oxidation of naphthalene (5.0 mM) by **3**/Cl⁻ (0.50 mM) in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C. Spectrum was recorded at 5 K. A silent EPR indicates that Mn(III) complex is the final product. (b) ESI-MS spectrum of the complete reaction solution obtained in the oxidation of naphthalene (0.30 mM) by **3**/Cl⁻ (0.030 mM) in CH₂Cl₂/CH₃CN (v/v 20:1) at -40 °C. The peak at *m/z* = 943.0 corresponds to [Mn^{III}(TDCPP)]⁺ (*calcd.* *m/z* = 943.2).

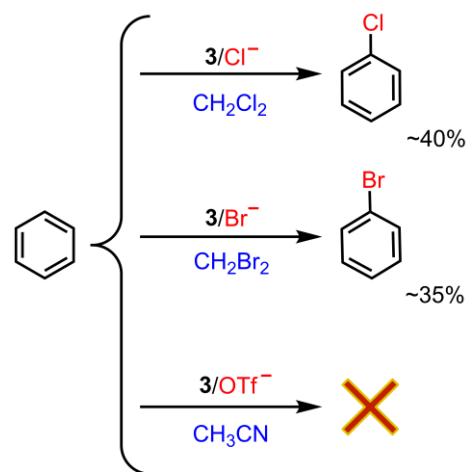


Fig. S5 Products obtained in the oxidation of benzene by **3**/X⁻ in different solvent system at -10 °C.

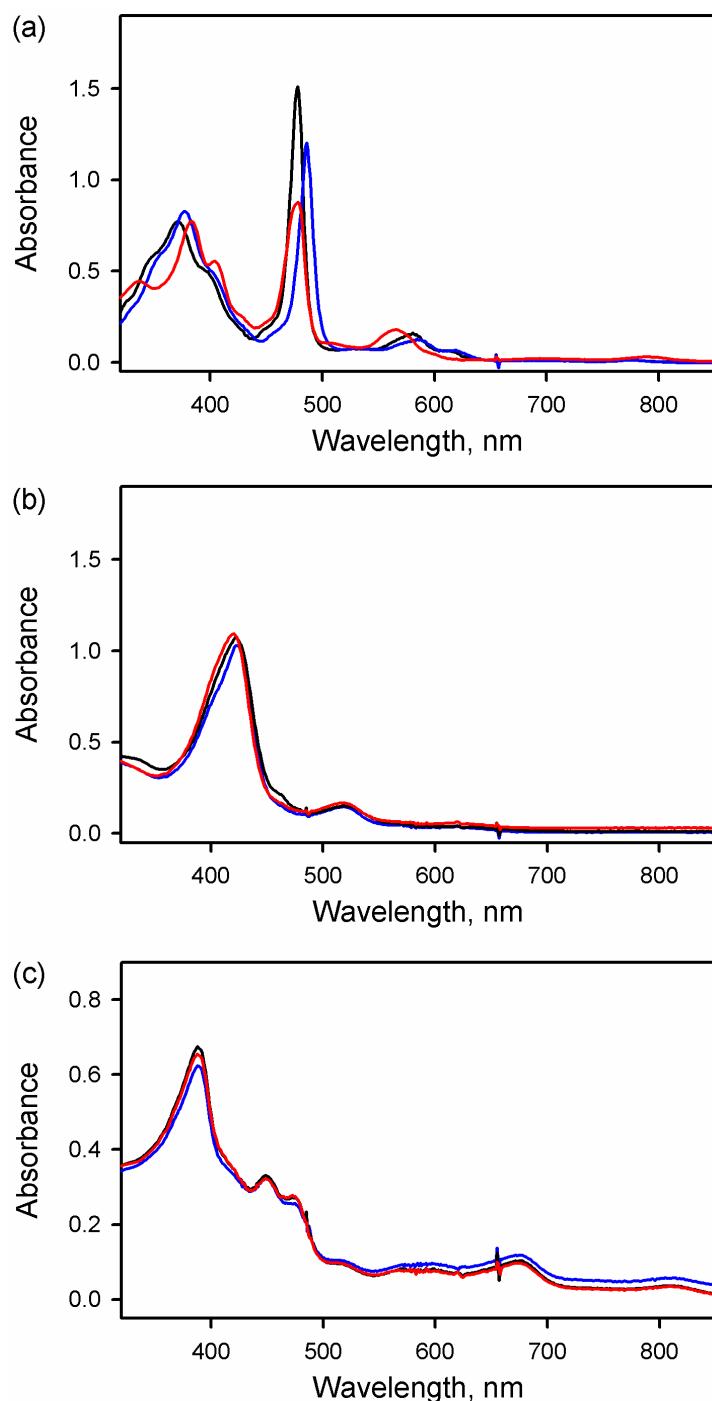


Fig. S6 (a) UV-vis spectra of $\text{Mn}^{\text{III}}(\text{TDCPP})(\text{Cl})$ (0.10 mM; black line), $\text{Mn}^{\text{III}}(\text{TDCPP})(\text{Br})$ (0.10 mM; blue line) and $\text{Mn}^{\text{III}}(\text{TDCPP})(\text{OTf})$ (0.10 mM; red line) in CH_2Cl_2 at $-10\text{ }^{\circ}\text{C}$. (b) UV-vis spectra of **2**/ Cl^- , **2**/ Br^- and **2**/ OTf^- generated by using corresponding starting compound in CH_2Cl_2 at $-10\text{ }^{\circ}\text{C}$. (c) UV-vis spectra of **3**/ Cl^- , **3**/ Br^- and **3**/ OTf^- generated by using corresponding starting compound in CH_2Cl_2 at $-10\text{ }^{\circ}\text{C}$.

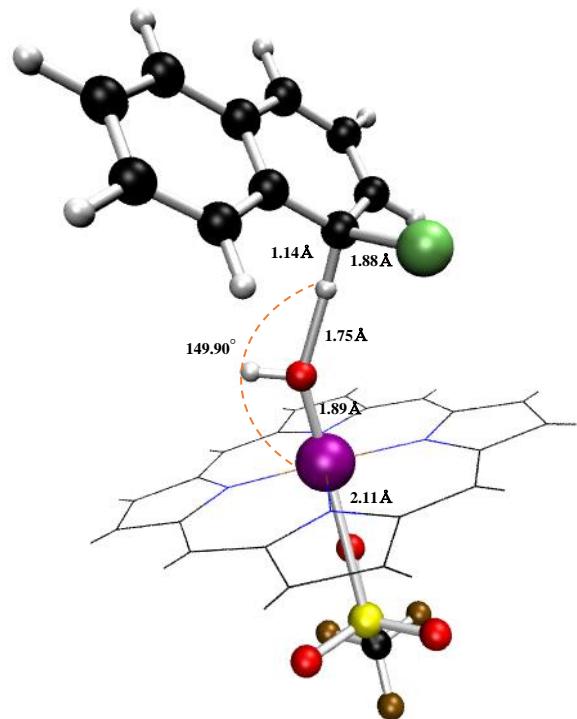


Fig. S7 Schematic representation of the transition state structure of C-H activation alongside with chlorine addition with important bond length and bond angles.

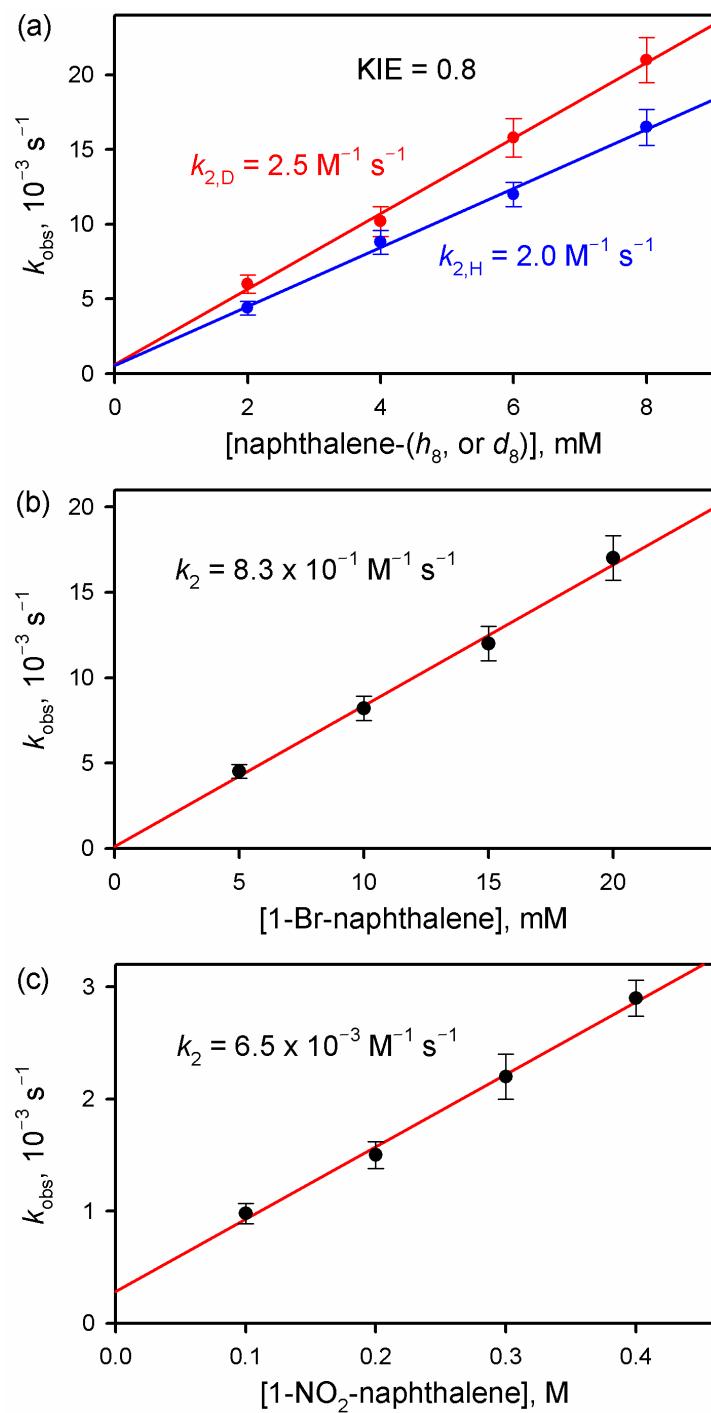


Fig. S8 Plots of the first-order rate constants (k_{obs}) against concentrations of (a) naphthalene (blue circles) and naphthalene- d_8 (red circles), (b) 1-Br-naphthalene, and (c) 1-NO₂-naphthalene for the aromatic halogenation by **3**/Cl⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C to determine the second-order rate constants (k_2).

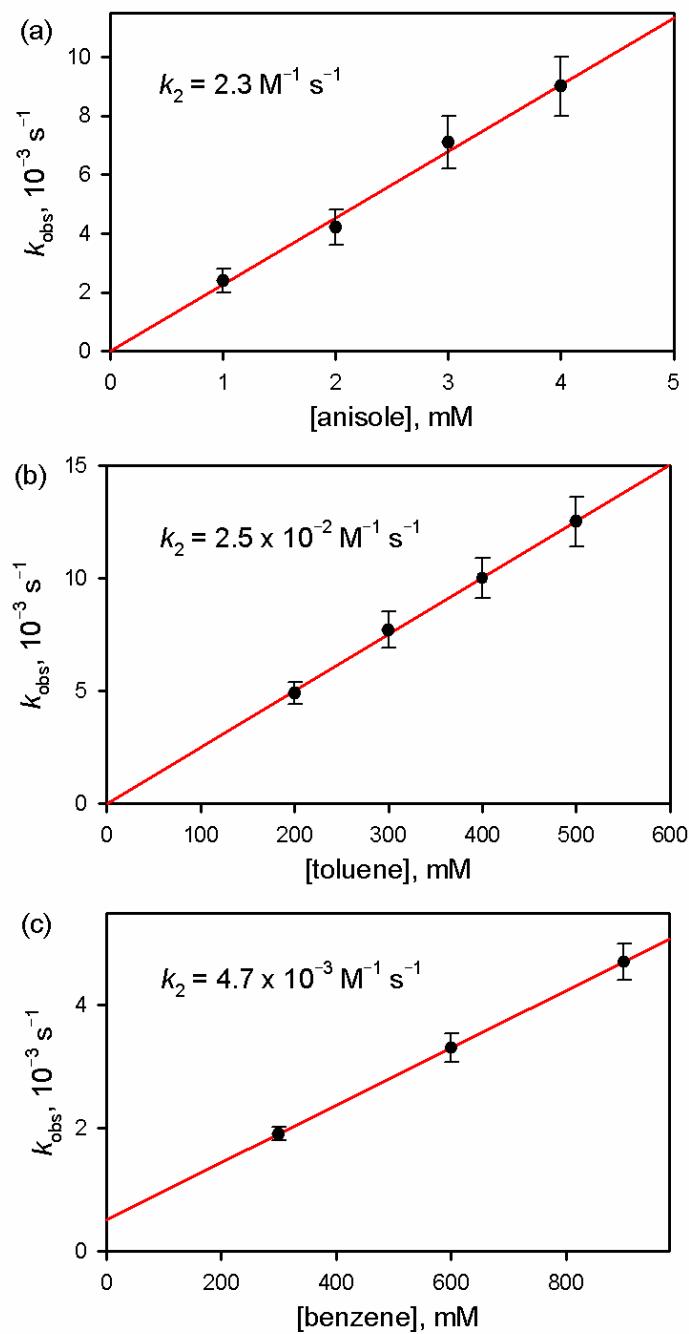


Fig. S9 Plots of the first-order rate constants (k_{obs}) against concentrations of (a) anisole, (b) toluene, and (c) benzene for the oxidation of aromatic substrates by **3**/Cl⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C to determine the second-order rate constants (k_2).

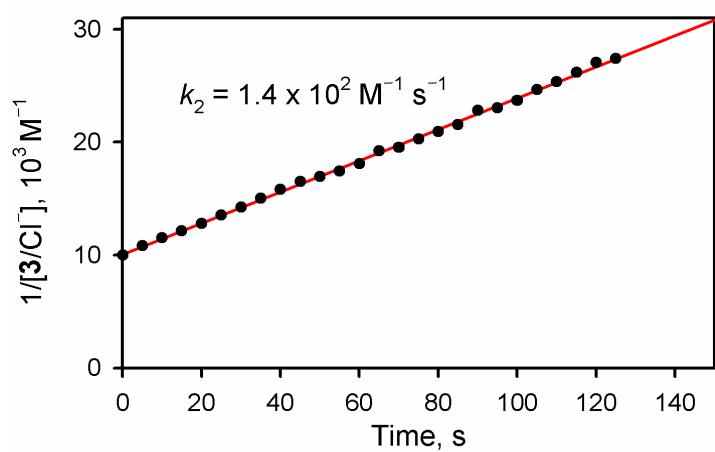


Fig. S10 Second-order plot of $1/[3/\text{Cl}^-]$ against time for the oxidation of 1-Me-naphthalene (0.10 mM) by **3**/Cl⁻ (0.10 mM) in CH₂Cl₂/CH₃CN (*v/v* 20:1) at -10 °C.

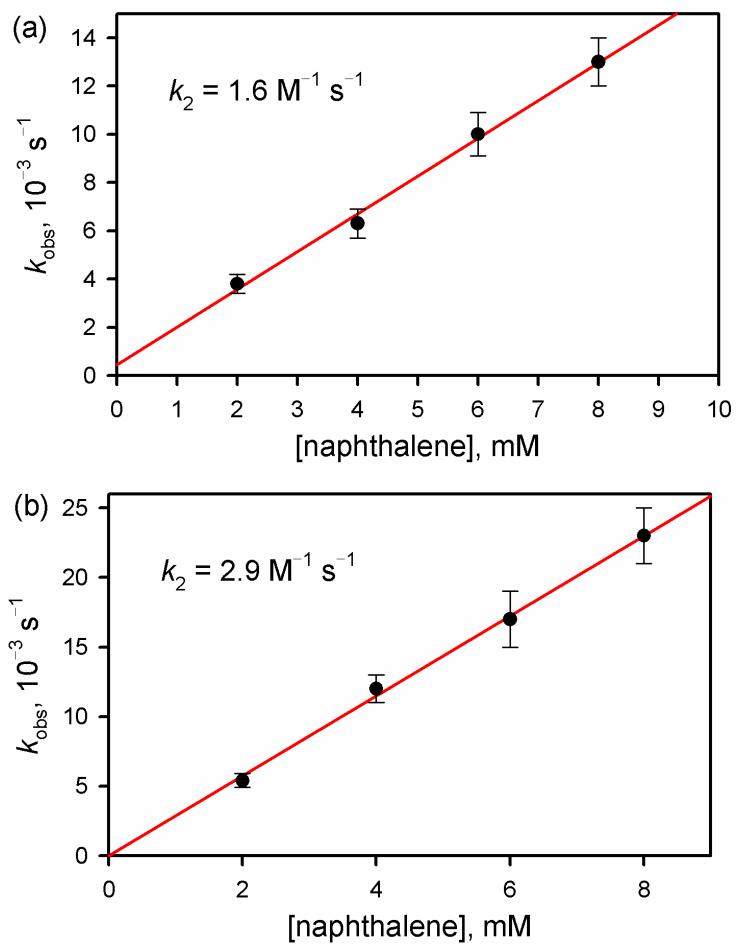


Fig. S11 Plots of the first-order rate constants (k_{obs}) against concentrations of naphthalene for the oxidation of naphthalene by (a) **3**/Br⁻ and (b) **3**/OTf⁻ in CH₂Cl₂/CH₃CN (v/v 20:1) at -10 °C to determine the second-order rate constants (k_2).

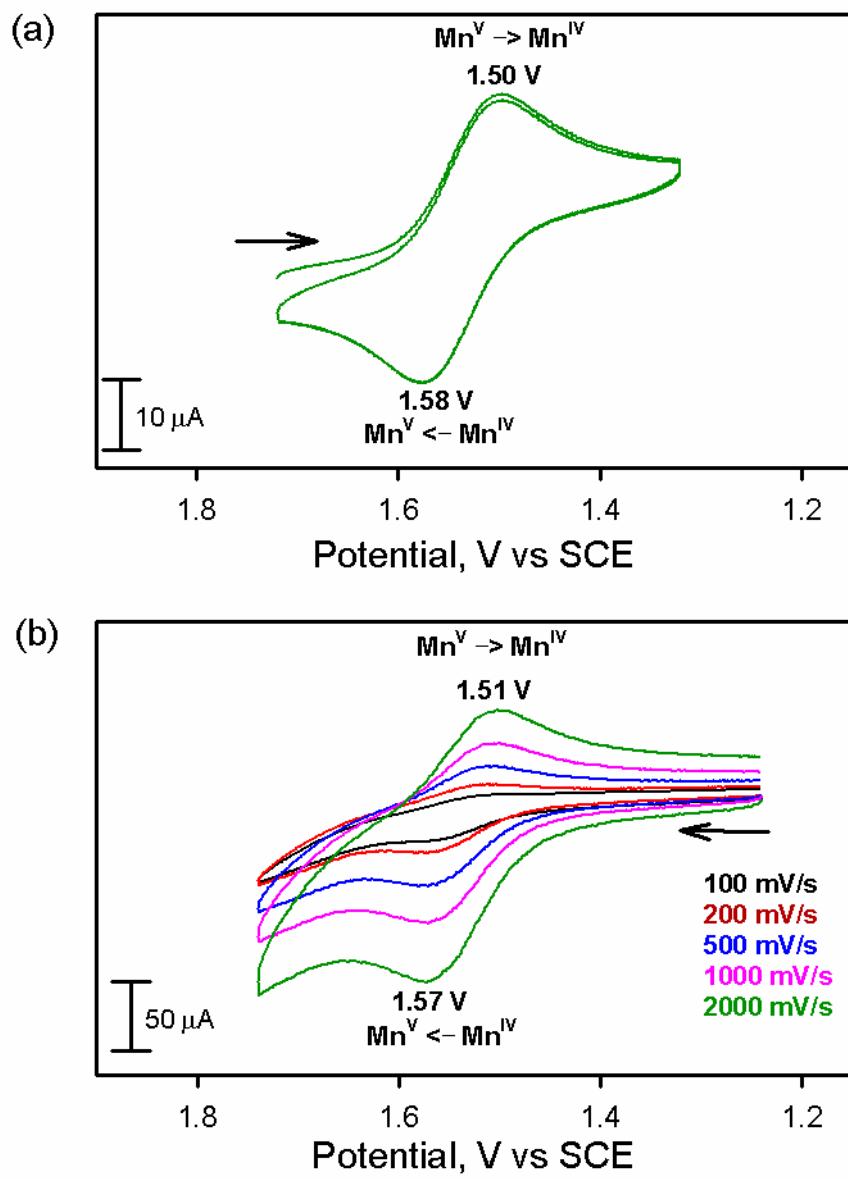


Fig. S12 Cyclic voltammograms of (a) $[\text{Mn}^{\text{IV}}(\text{OH})(\text{TDCPP}^{+\bullet})]^{2+}$ (**3**/ Cl^- ; 0.50 mM) and (b) $[\text{Mn}^{\text{IV}}(\text{O})(\text{TDCPP})]$ (**2**/ Cl^- ; 0.50 mM) in the presence of $n\text{-Bu}_4\text{NPF}_6$ (0.10 M) as a supporting electrolyte in $\text{CH}_3\text{CN}/\text{CH}_2\text{Cl}_2/\text{TFE}$ (v/v/v 1:1:0.01) at -40°C . A Pt electrode was used as a working electrode. The scan rate used in Fig. S12a was 100 mV/s. The one-electron reduction potential (E_{red}) of **3**/ Cl^- was determined to be 1.54 V vs SCE.

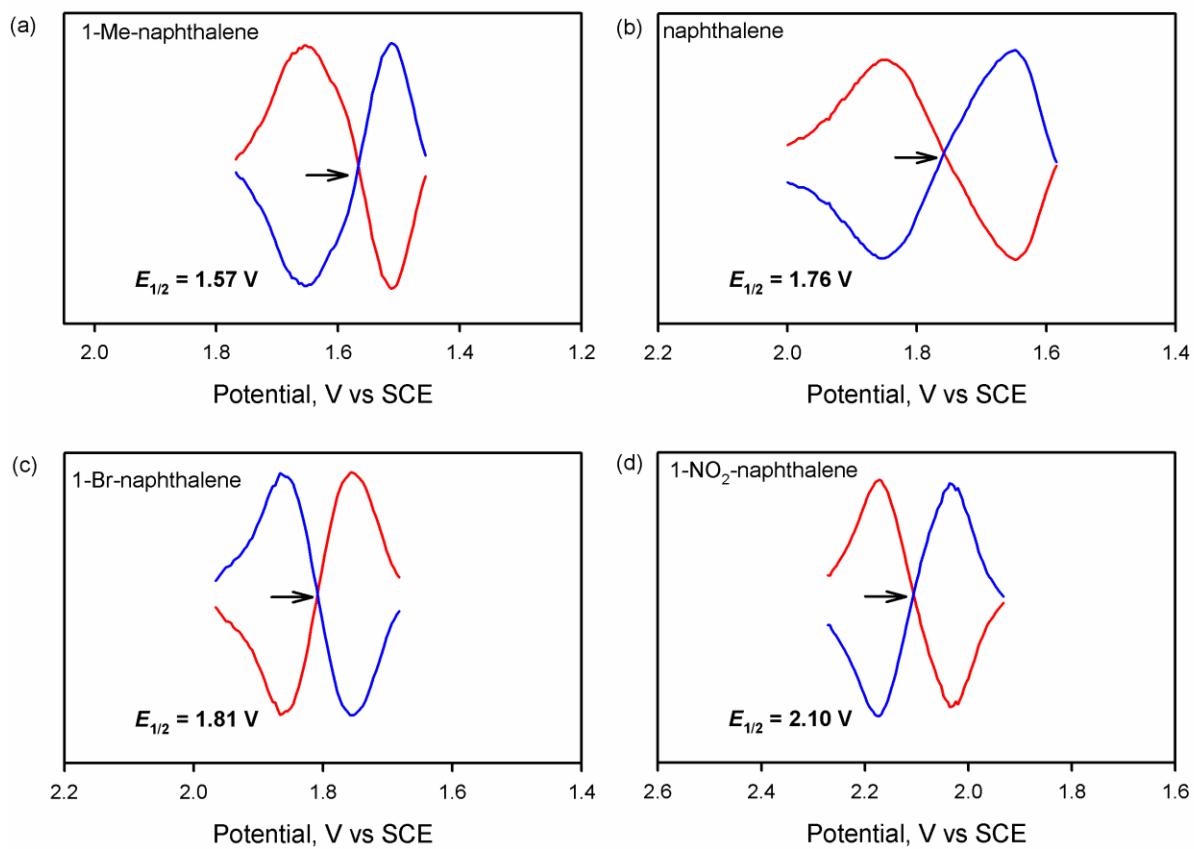


Fig. S13 SHACV voltammograms of (a) 1-Me-naphthalene (2.0 mM), (b) naphthalene (2.0 mM), (c) 1-Br-naphthalene (2.0 mM), and (d) 1-NO₂-naphthalene (2.0 mM) in the presence of *n*-Bu₄NPF₆ (0.10 M) as a supporting electrolyte in CH₃CN/CH₂Cl₂ (*v/v* 1:1) at 25 °C.

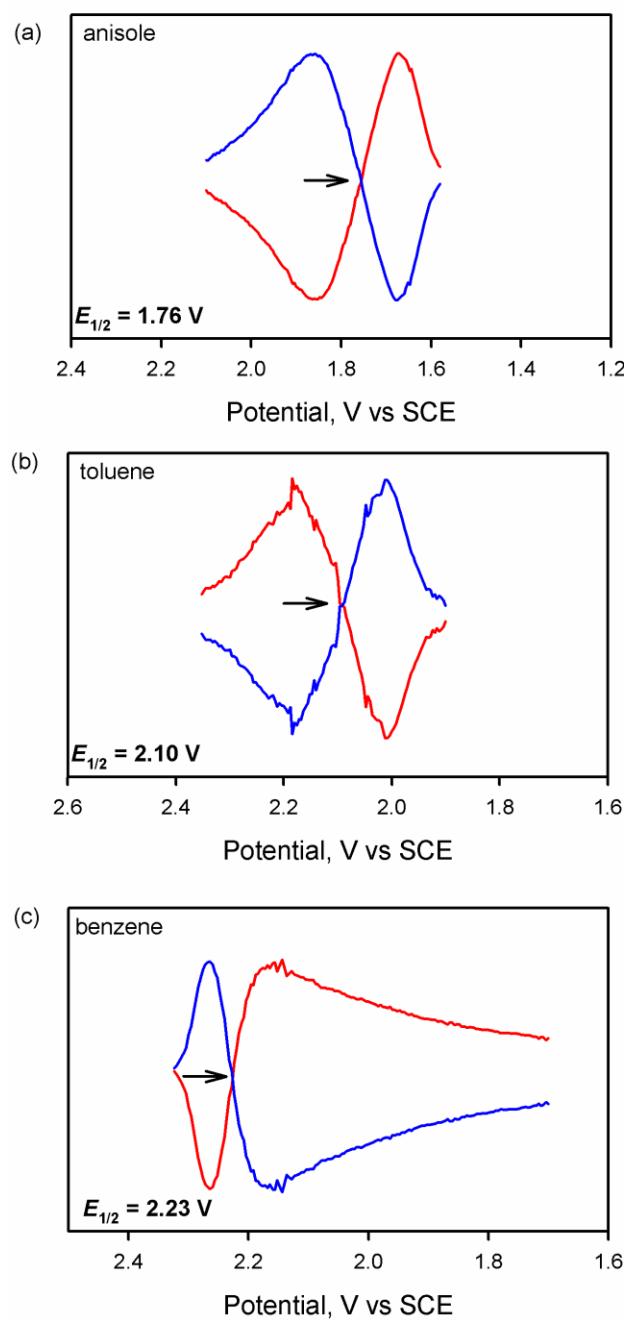


Fig. S14 SHACV voltammograms of (a) anisole (2.0 mM), (b) toluene (2.0 mM) and (c) benzene (2.0 mM) in the presence of $n\text{-Bu}_4\text{NPF}_6$ (0.10 M) as a supporting electrolyte in $\text{CH}_3\text{CN}/\text{CH}_2\text{Cl}_2$ (v/v 1:1) at 25 °C.

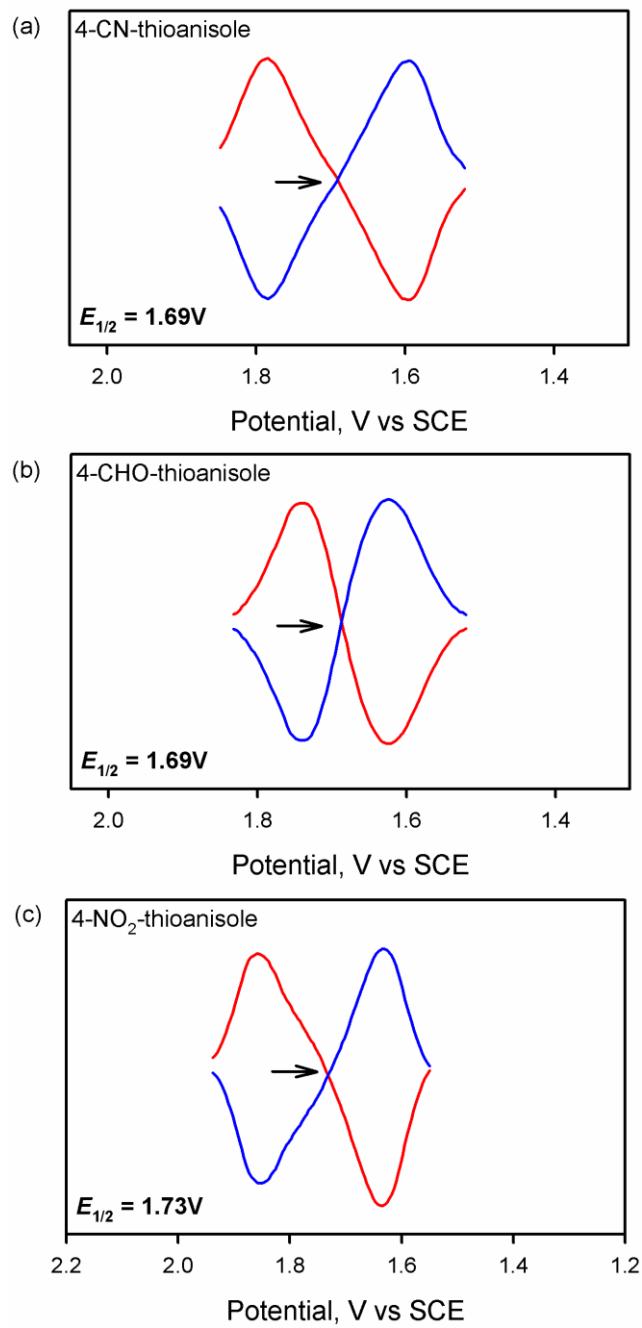


Fig. S15 SHACV voltammograms of (a) 4-CN-thioanisole (2.0 mM), (b) 4-CHO-thioanisole (2.0 mM) and (c) 4-NO₂-thioanisole (2.0 mM) in the presence of *n*-Bu₄NPF₆ (0.10 M) as a supporting electrolyte in CH₃CN/CH₂Cl₂ (*v/v* 1:1) at 25 °C.

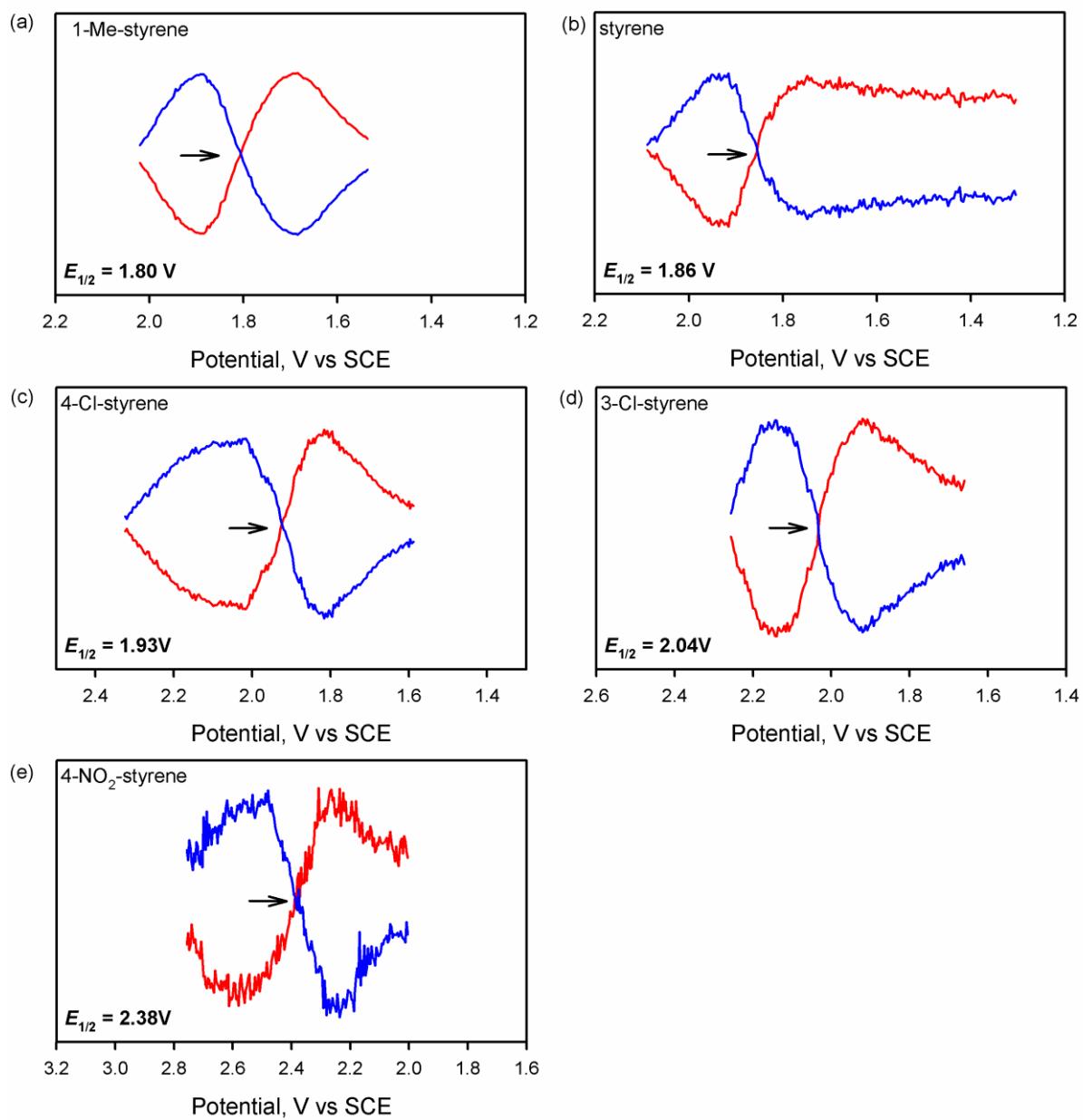


Fig. S16 SHACV voltammograms of (a) 1-Me-styrene (2.0 mM), (b) styrene (2.0 mM), (c) 4-Cl-styrene (2.0 mM), (d) 3-Cl-styrene (2.0 mM), and (e) 4-NO₂-styrene (2.0 mM) in the presence of *n*-Bu₄NPF₆ (0.10 M) as a supporting electrolyte in CH₃CN/CH₂Cl₂ (v/v 1:1) at 25 °C.

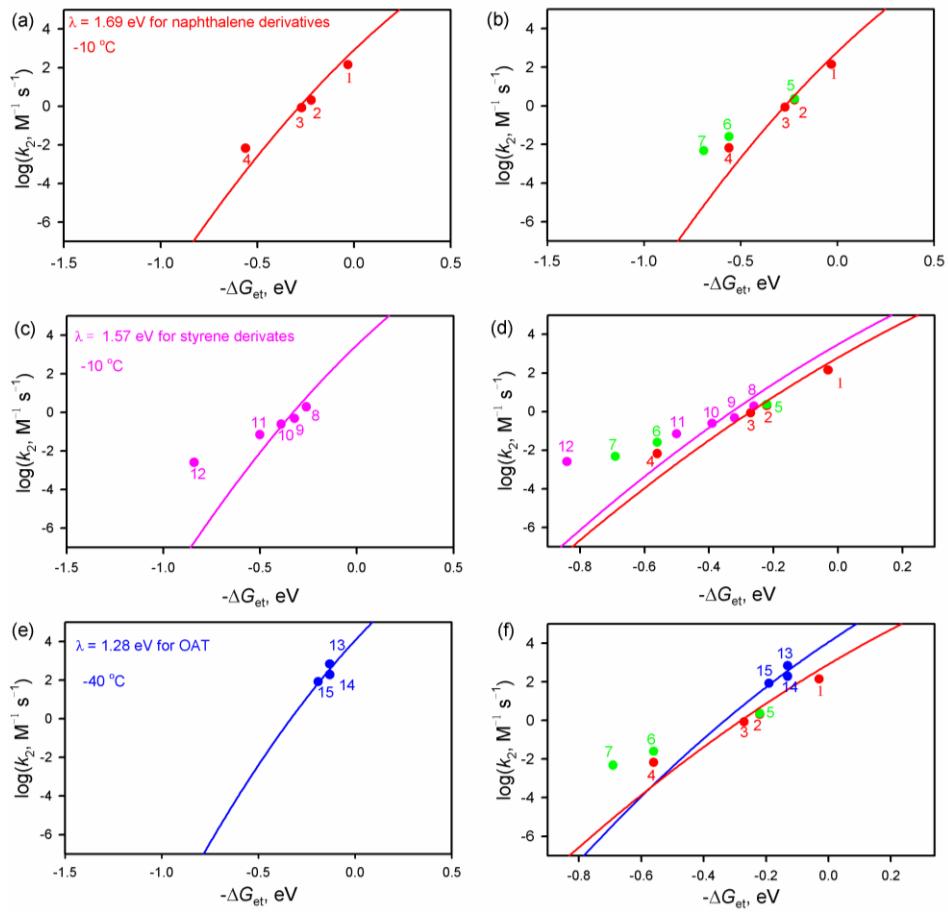


Fig. S17 Plots of $\log k_2$ for oxidation of different substrates [(1) 1-Me-naphthalene, (2) naphthalene, (3) 1-Br-naphthalene, (4) 1-NO₂-naphthalene, (5) anisole, (6) toluene, (7) benzene, (8) 4-Me-styrene, (9) styrene, (10) 4-Cl-styrene, (11) 3-Cl-styrene, and (12) 4-NO₂-naphthalene] at -10°C and [(13) 4-CN-thioanisole, (14) 4-CHO-thioanisole, and (15) 4-NO₂-thioanisole] at -40°C vs. the driving force of electron transfer [$-\Delta G_{\text{et}} = -e(E_{\text{ox}} - E_{\text{red}})$].

