Electronic Supplementary Information

The Role of Anionic Design in Modulating Material Properties: from SHG Switching to Ferroelastic Switching

Jia-Zi She, Yun-Hui Yu, Yan-Ran Weng, Dan Lu, Yong Yu, Ying Zhou, Hui-Peng Lv, and Yong Ai *

Corresponding author: yong.ai@ncu.edu.cn

Materials

Tetramethylammonium Hydroxide Pentahydrate, Methanesulfonic acid, and trifluoromethanesulfonic acid were purchased from Aladdin and without further purifying. Methyl alcohol and acetonitrile were purchased from Merck and used as the solvent.

Crystal Growth

Crystal of $[NMe_4][CH_3SO_3H]$ (CH-NMe₄) is grown by dissolving 18.1 g (100 mmol) of tetramethylammonium hydroxide pentahydrate in 100 ml of ethanol solution, and then adding 9.8 g (100 mmol) of methanesulfonic acid. Heat and stir until the generated solid is completely dissolved. Filter the solution and slowly evaporate the solvent at a constant temperature of 30 °C. After several days, transparent crystals will precipitate. To get the crystal of $[NMe_4][CF_3SO_3H]$ (CF-NMe₄), 100 millimoles of tetramethylammonium hydroxide pentahydrate is dissolved in 100 mL of ethanol. Then 100 millimoles of trifluoromethanesulfonic acid was added. Heating and stirring make the solid dissolve, then filter the solution. After the evaporation filter at a constant temperature of 30 °C in the air for several days, plate colorless crystals are obtained. The picture of the crystal is shown in Fig. S1.

Methods

Single-crystal X-ray diffraction (SCXRD): Variable-temperature single-crystal diffraction data were collected on a Rigaku XtaLAB Synergy-S diffractometer (Mo-K α radiation, $\lambda = 0.71073$ Å) equipped with a low-temperature device. A freshly precipitated high-quality single crystal was mounted on a nylon loop under inert oil for analysis. The structure was solved by direct methods using SHELXT-2018 and refined anisotropically for all non-hydrogen atoms via full-matrix least-squares iterations in SHELXL-2018. Hydrogen atoms were geometrically positioned using AFIX instructions and refined isotropically with riding models. CCDC 2449758-2449759 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, by emailing data_request@ccdc.cam.ac.uk.

Powder X-ray diffraction: Temperature-dependent PXRD patterns were recorded on a Rigaku SmartLab SE diffractometer (Mo-K α radiation, $\lambda = 0.71073$ Å). Scans were performed in the 2θ range of 5°–40° (step size 0.02°, scan rate 2° min⁻¹) under a dynamic nitrogen atmosphere. Rietveld refinement was executed using the GSAS-II software suite.

Thermal analysis: Differential scanning calorimetry (DSC) was conducted on a PerkinElmer Diamond DSC instrument. Powdered samples $(5.0 \pm 0.2 \text{ mg})$ underwent three heating-cooling cycles (20 K/min, N₂ flow 50 mL/min) to ensure reproducibility. Thermogravimetric analysis (TGA) utilized the TA Q50 system system (20 K/min, 300–950 K, N₂ atmosphere), with α -Al₂O₃ as a reference standard.

Second harmonic generation (SHG): SHG measurements were performed using an Edinburgh Instruments FLS 920 fluorescence spectrometer system coupled with a collimated pulsed laser source. The laser source adopted was a Vibrant 355 II laser generator (OPOTEK) excited by pulsed Nd:YAG at a wavelength of 1064 nm, 5 ns pulse duration, 1.6 MW peak power, and a 10 Hz repetition rate.

Dielectric constant measurement: For the dielectric measurements, the polycrystalline samples were ground into powder and pressed into a sheet of about 0.6 mm in thickness, and then the silver glue was evenly spaced on both sides, covering an area of about 15 mm². Then copper wire and IC six-hole socket were used to make a capacitor. The Tong hui TH2828A instrument was used to measure the compound dielectric constant ε ($\varepsilon = \varepsilon'$ -*i* ε'').

Hirshfeld surface analysis: The Hirshfeld surfaces and their associated 2D-fingerprint plots were conducted using the Crystal Explorer software, utilizing CIF format structure files as input. The morphology of these surfaces is contingent on the interactions both between molecules within the crystal and between atoms within the molecule. All Hirshfeld surfaces were generated at a standard high surface resolution. The intensity of molecular interactions is visually represented on the Hirshfeld surface through a color scheme comprising red, blue, and white regions. Specifically, white regions correspond precisely to van der Waals contact distances, blue regions indicate longer contacts and red regions signify closer contacts.

The normalized contact distance, denoted as d_{norm} , is derived from the parameters de, di, and the van der Waals (vdW) radii of the two atoms, one external (r_e^{vdW}) and one internal (r_i^{vdW}) to the surface. Mathematically, d_{norm} is defined as:

$$\mathbf{d}_{\text{norm}} = \frac{\frac{d_i - r^{vdW}}{r^{vdW}}}{r^{vdW}} + \frac{\frac{d_e - r}{r^{vdW}}}{r^{vdW}}$$

The d_{norm} value is a valuable metric for identifying close intermolecular interactions. A smaller d_{norm} value indicates stronger intermolecular interactions within the system.

Ferroelastic measurement and spontaneous strain tensor calculation: The evolution of domain structures was observed by using an Olympus BX51TRF polarizing microscope at different temperatures controlled by INSTEC HCC602 cooling/heating device. The calculation particular matrix is as follows:

 $\varepsilon_{ij} =$



Fig. S1. TGA curves of (a) CH-NMe₄, (b) CF-NMe₄.



Fig. S2. Piezoelectric coefficient of CH-NMe₄ measured via the quasi-static method..



Fig. S3. Simulated and experimental PXRD patterns of (a) CH-NMe₄ at 293 K and (b) CF-NMe₄ at 260 K.



Fig. S4. Two-dimensional tightly connected packing structures of (a) CH-NMe₄ at 293 K, (b) CF-NMe₄ at 260 K from c axis.



Fig. S5. Visualization map of the distribution of the interactions and the two-dimensional fingerprint plots of CH-NMe₄.



Fig. S6. Visualization map of the distribution of the interactions and the two-dimensional fingerprint plots of CF-NMe.



Fig. S7. Temperature-dependent dielectric constants were measured at 1, 5, 10, and 100 kHz for the (a) CH-NMe₄ and (b) CF-NMe₄.



Fig. S8. Variable-temperature PXRD patterns of (a) CH-NMe₄, (b) CF-NMe₄.



Fig. S9. The final Rietveld refinement plot of CH-NMe₄ structure at 330 K in the high-temperature phase: experimental pattern (red line), calculated pattern (blue line), difference profile (yellow line), background (pink line), and background profile (green dot). Through the Rietveld refinements of the PXRD data, we obtained the tetragonal crystal system, among which the most possible point group is 4/m.



Fig. S10. The final Rietveld refinement plot of CF-NMe₄ structure at 503 K in the high-temperature phase: experimental pattern (red line), calculated pattern (blue line), difference profile (yellow line), background (pink line), and background profile (green dot). Through the Rietveld refinements of the PXRD data, we obtained the tetragonal crystal system, among which the most possible point groups are 4/mmm and 4/m.

Compound	CH-NMe ₄	CF-NMe ₄
Temperature/K	293	260
Crystal system	Orthorhombic	Monoclinic
Empirical formula	$C_5H_{15}NO_3S$	$C_5H_{12}F_3NO_3S$
Formula weight	169.24	223.22
Space group	$Pca2_1$	$P2_{1}/m$
a/Å	11.994(2)	10.3722(3)
$b/\text{\AA}$	6.1359(13)	8.6225(2)
$c/\text{\AA}$	11.923(3)	11.6086(3)
$\alpha/^{\circ}$	90	90
$eta/^{\circ}$	90	103.824(3)
$\gamma^{\prime \circ}$	90	90
Volume/Å ³	877.5(3)	1008.13(5)
Ζ	4	4
Radiation	<i>MoK</i> α(λ=0.71073)	<i>MoKa</i> (λ=0.71073)
$ ho_{ m calc} { m g/cm^3}$	1.281	1.471
GOF	1.040	1.076
Final R indexes [I>=2 σ (I)]	$R_1 = 0.0807 \ wR_2 = 0.1467$	$R_1 = 0.0795, wR_2 = 0.2161$

Table S1. Variable-temperature crystallographic data of the single-crystal of CH-NMe $_4$ and CF-NMe $_4$.

Table S2. Non-classical Hydrogen Bonds for CH-NMe₄ at 293 K

D	Н	А	d(<i>D</i> -H)/Å	d(H-A)/Å	d(D-A)/Å	<i>D</i> –H–A/°
C1	H1C	02	0.96	2.52	3.457	167
C3	H3B	01	0.96	2.58	3.511	164
C4	H4B	02	0.96	2.40	3.349	170
C4	H4C	01	0.96	2.60	3.539	167

D	Н	А	d(<i>D</i> −H)/Å	d(H−A)/Å	d(D-A)/Å	<i>D</i> –H–A/°
C2	H2C	03	0.96	2.56	3.420	149
C3	H3C	03	0.96	2.50	3.373	151
C5	H5A	01	0.96	2.56	3.414	148
C5	H5B	01	0.96	2.53	3.414	153
C6	H6A	02	0.96	2.57	3.442	151
C6	H6B	01	0.96	2.56	3.420	150

Table S3. Non-classical Hydrogen Bonds for CF-NMe₄ at 260 K