1,2-Diaminocyclohexane-derived chiral tetradentate ligands for Mn(I)-catalyzed

asymmetric hydrogenation of ketones

Yi Su ^{a,b‡}, Dongzhi Zhu ^{a,c‡,*}, Zhifeng Ma ^{c,d,‡,*}, Yizhou Wang^c, Zechen Wang,^b Zheng Wang,^{b,c*}, Yanping Ma^c, Wen-Hua Sun^{c,*}

^aGuangxi Key Laboratory of Advanced Structural Materials and Carbon Neutralization, School of Materials and Environment, Guangxi Minzu University, Nanning 530105, China

^bCollege of Science, Hebei Agricultural University, Baoding 071001, China.

^cKey Laboratory of Engineering Plastics and Beijing National Laboratory for Molecular Science, Institute of Chemistry, Chinese Academy of Sciences, Beijing 100190, China.

^dSchool of Chemistry & Environment, Yunnan Key Laboratory of Chiral Functional Substance Research and Application, Yunnan Minzu University, Kunming 650504 Yunnan, China.

⁺ These authors contributed equally to this work.

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1 · General information

All manipulations and their complexes were carried out under a nitrogen atmosphere using standard Schlenk techniques. All solvents were dried and distilled under nitrogen prior to use. All the liquid substrates and solid substrates (Table S1) were used directly without further purification. MnBr(CO)₅ (CAS: 14516-54-2) can be purchased from Shijiazhuang chiral chemical CO., LTD., 2-(diphenylphosphino)benzaldehyde (CAS: 50777-76-9), (1R,2R)-(-)-1,2-diaminocyclohexane (CAS: 20439-47-8).¹H, ¹³C and ³¹P NMR spectra were recorded on Bruker AV²400 NMR and Bruker AV²500 NMR spectrometers. Chemical shift values in ¹H and ¹³C NMR spectra were referenced internally to the residual solvent resonances, whereas ³¹P NMR spectra

were referenced externally to H₃PO₄. Elemental analysis was carried out with a Vario EL III CHN microanalyzer. Infrared spectroscopy was performed in the solid state on a Bruker ALPHA. A nitrogen-hydrogen-air integrated machine (GX-300A, ZhongXingHuiLi) provides 99% H2, 99% N2 and air as the carrier make up gas used for GC analysis. The GC analysis was carried out on a FuLi GC-9790Plus instrument (Zhe Jiang FuLi Analytical Instrument) using a PB-FFAP column ($30m*0.32mm*0.25\mum$, Wuhan Puli Technology Co., Ltd.) or RB-WAX ($30m*0.25mm*0.25\mum$, Wuhan Puli Technology Co., Ltd.) or RB-WAX ($30m*0.25mm*0.25\mum$, Wuhan Puli Technology Co., Ltd.): injector temp. = 300 °C, flame ionization detector (FID) temp. = 300 °C, column temp. = 80 °C, withdraw time 2 min, then 20 °C /min to 240 °C keeping for 5 min, then 20 °C /min to 280 °C, withdraw time for 5 min; Enantioselectivities (ee, %) were determined by an FuLi GC-9790Plus instrument using a chiral column [CP-Chirasil-Dex CB column ($25 m \times 0.25 mm \times 0.25 \mum$)], FID temp. = 280 °C, Injector temp. = 250 °C, column temp. = 120 °C isothermal for 2 minutes, 5 °C /min, 180 °C for 30 minutes.

2. Syntheses and characterization of ligands and complexes

2.1 Synthesis of L1



Under nitrogen, a mixture of 2-(diphenylphosphino)benzaldehyde (9.86 g, 34 mmol), (R, R)-1,2diaminocyclohexane (1.98 g, 17 mmol), anhydrous Na_2SO_4 (20 g, 140 mmol) and 60 mL CH_2Cl_2 in 250 mL Schlenk flask was stirred for 48 h at 25 °C. A pale-orange solution was obtained. The solution was filtered, and then concentrated under reduced pressure to 5 mL. To the solution was added 40 mL of ethanol (Dilute and dissolve the sample) and the resulting solution was cooled to -20 °C (Low temperature reduces the solubility of samples in solution) to give a pale-yellow solid, filtered and dried in vacuo (11.19 g, 90% yield), M.p. 60-62 °C.

¹H NMR (400 MHz, CDCl₃) δ 8.70 (d, *J* = 4.1 Hz, 2H), 7.75 (m, *J* = 7.0, 3.5 Hz, 2H), 7.38 – 7.25 (m, 14H), 7.21 (d, *J* = 4.0 Hz, 8H), 6.82 (m, *J* = 6.6, 4.7 Hz, 2H), 3.19 – 3.07 (m, 2H), 1.68 (d, *J* = 8.6 Hz, 2H), 1.48 (d, *J* = 9.4 Hz, 2H), 1.37 (d, *J* = 13.5 Hz, 2H), 1.28 (t, *J* = 9.8 Hz, 2H);¹³C{H} NMR (100 MHz, CDCl₃) δ 159.5, 159.3, 140.2, 140.0, 137.2, 137.1, 137.0, 136.9, 134.1, 134.0, 133.9, 133.8, 133.3, 129.7, 128.8, 128.6, 128.5, 128.4, 128.1, 140.0, 137.2, 128.4, 128.1, 128.4, 128.1, 128.4, 128.1, 128.4, 128.1, 128.4, 128.1, 128.4, 128.1, 128.4, 128.4, 128.1, 128.4,

73.6, 32.5, 24.3; ³¹P NMR (243 MHz, CDCl₃) δ -13.68; IR (ATR, cm⁻¹, KBr): 1636 (m, $\nu_{C=N}$), 2857 (s), 2929 (s), 3050 (s), t. The spectroscopic data correspond to those reported in the literature.¹

2.2 Synthesis of L2



Under nitrogen, a solution of compound (R, R)-L1 (5.0 g, 7.6 mmol) and NaBH₄ (2.8 g, 76 mmol) in absolute ethanol (30 mL) was refluxed with stirring for 48 h. The solution was cooled to room temperature and H₂O (10 mL) was added to destroy excess NaBH₄. The mixture solution was extracted with CH₂Cl₂ (30 mL × 3). The combined extract was washed with 10% aqueous NH₄Cl (10 mL × 2), H₂O (10 mL × 2) and the organic layer was dried over anhydrous Na₂SO₄, filtered (Remove the sediments) and concentrated to 5 mL. Then 15 ml of ethanol (Dilute and dissolve the sample) was added and cooled to -30°C (Low temperature reduces the solubility of samples in solution) to give cream-white crystals (4.72 g, 93% yield). M.p. 51-52°C.

¹H NMR (400 MHz, CDCl₃) δ 7.52 (m, *J* = 7.8, 4.6 Hz, 2H), 7.29 (tt, *J* = 8.4, 5.6 Hz, 14H), 7.23 – 7.17 (m, 8H), 7.14 (t, *J* = 7.5 Hz, 2H), 6.83 (m, *J* = 7.7, 4.5, 1.3 Hz, 2H), 4.04 (d, *J* = 13.5 Hz, 2H), 3.85 (d, *J* = 13.5 Hz, 2H), 2.19 (br, 2H), 1.99 (d, *J* = 12.9 Hz, 2H), 1.61 (d, *J* = 9.2 Hz, 3H), 1.26 (d, *J* = 6.1 Hz, 2H), 1.09 (t, *J* = 5.7 Hz, 2H); ¹³C{H} NMR (100 MHz, CDCl₃) δ 133.9, 133.9, 133.8, 133.8, 133.4, 129.0, 128.6, 128.5, 128.5, 128.5, 127.1, 60.8, 48.9, 31.5, 24.8, 22.6, 14.1; ³¹P NMR (243 MHz, CDCl₃) δ -15.92; IR (ATR, cm⁻¹, KBr): 1684 (w, $v_{C=N}$), 2852 (s), 2924 (s), 3051 (s, v_{NH}).

The spectroscopic data correspond to those reported in the literature.¹

2.3 Synthesis of L3



A mixture of 2-thenaldehyde (3.82 g, 34 mmol), (R,R)-1,2-diaminocyclohexane (1.98g, 17 mmol) and anhydrous Na₂SO₄ (20 g, 140 mmol) in CH₂Cl₂ (60 mL) was stirred for 48 h at 25 °C. A pale-orange solution was obtained. The solution was filtered, and then concentrated under reduced pressure to 5 ml. To the solution was added 40 ml of ethanol (Dilute and dissolve the sample) and the resulting solution was cooled to -35°C (Low temperature reduces the solubility of samples in solution) to give a pale-brown solid, filtered and dried in vacuo (5.14 g, 75 % yield). M.p. 75-76 °C.

¹H NMR (400 MHz, CDCl₃) δ 8.30 (s, 1H), 7.33 (d, *J* = 5.0 Hz, 1H), 7.18 (m, *J* = 3.5, 0.8 Hz, 1H), 7.00 (m, *J* = 5.0, 3.7 Hz, 1H), 3.41 – 3.31 (m, 2H), 1.94 – 1.74 (m, 3H), 1.48 (t, *J* = 9.4 Hz, 1H); ¹³C{H} NMR (100 MHz, CDCl₃) δ 154.3, 142.4, 130.1, 128.7, 128.2, 127.3, 127.2, 73.4, 33.6, 33.0, 32.8, 25.0, 24.7, 24.4. The spectroscopic data correspond to those reported in the literature²

2.4 Synthesis of L4



Under nitrogen, 2-diphenylphosphinylbenzoic acid (10 g, 32.6mmol) and CDI (5.57 g, 33.3mmol) were charged into a clean and dry 250 mL Schlenk flask. After CH₃CN (20 mL) had been added to give a slurry, the resulting mixture was agitated for 2 h at 25 °C to afford a clear and homogenous solution. The solution was transferred to a Reactor containing (R,R)-diaminocyclohexane (1.86 g, 16.3 mmol) and imidazole hydrochloride (1.70 g,16.5 mmol), rinsed with 6 mL CH₃CN. The resulting mixture was then heated at 82 °C for 12 h. After the solution was cooled to 60 °C, a slurry was obtained. 13 mL H₂O was added over 30 min at 60 °C. After 1.0 h, the obtained slurry was cooled to 20 °C (Low temperature reduces the solubility of samples in solution) over 0.5 h. After 0.5 h, the solid was collected by filtration and then washed with CH₃CN/H₂O (2:1, 30 mL) and H₂O (20 mL) successively. After being dried under vacuum at 50 °C, the **L4** was obtained as a pale-yellow solid (4.60 g, 40%)

¹H NMR (600 MHz, CDCl₃) δ 7.57 (m, *J* = 7.7, 3.7 Hz, 2H), 7.32 – 7.26 (m, 14H), 7.24 – 7.18 (m, 8H), 6.91 (m, *J* = 7.6, 3.9 Hz, 2H), 6.34 (d, *J* = 6.5 Hz, 2H), 3.77 (s, 2H), 1.89 – 1.82 (m, 2H), 1.64 (d, *J* = 9.5 Hz, 4H), 1.21 (t, *J* = 10.1 Hz, 2H), 1.07 – 0.94 (m, 2H); ¹³C{H} NMR (150 MHz, CDCl₃) δ 169.3, 140.9, 140.7, 137.6, 137.6, 136.6, 136.5, 134.3, 133.9, 133.8, 130.2, 128.8, 128.6, 128.5, 128.5, 128.4, 128.4, 127.5, 127.5, 53.9, 32.0, 24.6; ³¹P NMR (243 MHz, CDCl₃) δ -9.69.

The spectroscopic data correspond to those reported in the literature.³

2.5 Preparation of Mn1



 $Mn(CO)_5Br$ (0.54 g, 2.0 mmol) and L1 (1.32 g, 2.0 mmol) were loaded in a 50 mL Schlenk tube under nitrogen and the contents stirred in dry toluene (10 mL) for 10 h at 90 °C. After cooling to room temperature, the resulting red suspension was cooled to room temperature (Low temperature reduces the solubility of samples in solution) and the precipitate was collected and washed successively *n*-hexane (2 × 5 mL), then dried in vacuo to afford **Mn1** as a red powder (1.50 g, 84% yield). A single crystal of **Mn1** suitable for the X-ray determination was grown by slow diffusion of *n*-hexane into a dichloromethane solution of **Mn1**. ¹H NMR (400 MHz, DMSO-*d*₆) δ 8.81 – 8.54 (m, 2H), 7.99 – 7.90 (m, 2H), 7.73 – 7.67 (m, 3H), 7.64 – 7.55 (m, 10H), 7.50 – 7.43 (m, 6H), 7.17 – 7.07 (m, 5H), 6.54 (s, 2H), 1.94 (s, 2H), 1.70 (s, 2H), 1.45 (s, 2H), 1.24 (s, 2H); ¹³C{H} NMR (100 MHz, DMSO-*d*₆) δ 168.8, 143.5, 139.8, 133.9, 132.3, 131.8, 130.2, 129.7, 71.2, 31.4, 23.6, 19.0; ³¹P NMR (162 MHz, CDCl₃) δ 63.45; IR (ATR, cm⁻¹, KBr): 1612 (s, *v*_{C=N}), 1882 (s, *v*_{CO}), 1950 (s, *v*_{CO}); Anal. Calcd for [C₄₆H₄₀MnN₂O₂P₂Br (*M*w: 848.11)]: C, 65.03; H, 4.75; N, 3.30. Found: C, 64.97; H, 4.78; N, 3.28%.

2.6 Preparation of Mn2



 $Mn(CO)_5Br$ (406.0 mg, 1.49 mmol) and L2 (0.990 mg, 1.49 mmol) were loaded in a 50 mL Schlenk tube under nitrogen and the contents stirred in dry toluene (20 mL) for 12 h at 110 °C. After cooling to room temperature, the reaction mixture was concentrated (Remove low boiling point solvents) under reduced pressure. The residue was dissolved in dichloromethane (2 mL) and *n*-hexane (25 mL) added to form a precipitate (Inert solvents reduce the solubility of samples in solution). The solution above the precipitate was carefully removed by pipette. The precipitate was also washed with *n*-hexane (2 × 20 mL) and then dried to afford **Mn2** as a pale-yellow solid (958.0 mg, 78%).

¹H NMR (400 MHz, DMSO- d_6) δ 7.84 (t, J = 7.2 Hz, 4H), 7.78 (d, J = 7.0 Hz, 2H), 7.61 (d, J = 6.5 Hz, 6H), 7.51 (t, J = 6.8 Hz, 12H), 7.33 – 7.27 (m, 2H), 7.26 – 7.13 (m, 2H), 4.42 (d, J = 15.5 Hz, 2H), 4.21 (d, J = 15.5 Hz, 2H), 3.40 (brs, 2H, NH), 2.60 (s, 2H), 1.87 (d, J = 12.0 Hz, 2H), 1.53 (d, J = 9.0 Hz, 2H), 1.33 – 1.23 (m, 2H), 0.50 (d, J = 10.3 Hz, 2H); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 225.9, 225.6, 140.0, 136.1, 135.4, 133.5, 133.0, 132.3, 132.1, 131.9, 131.3, 130.9, 129.9, 129.7, 129.4, 128.7, 125.3, 60.3, 51.4, 29.4, 23.3; ³¹P NMR (162 MHz, CDCl₃) δ 61.93; IR (ATR, cm⁻¹, KBr): 1622 (s, v_{NH}), 1852 (s, v_{CO}), 1930 (s, v_{CO}), 3233 (s, v_{NH}); Anal. Calcd for [C₄₆H₄₄MnN₂O₂P₂Br (*M*w: 852.14)]: C, 64.72; H, 5.20; N, 3.28. Found: C, 64.67; H, 5.32; N, 3.24%.

3. NMR Spectra for ligands and complexes

Figure S1 ¹H, ¹³C and ³¹P NMR spectra for L1 in CDCl₃



Figure S2 ¹H, ¹³C and ³¹P NMR spectra for L2 in CDCl₃



140 120 100 80 60 40 20 0 -20 -40 -60 -80 -100 -120 -140 -160 -180 -200 -220 -240 f1 (pm)

Figure S3 ¹H and ¹³C NMR spectra for L3 in CDCl₃



-10 130 120

Figure S4 ¹H, ¹³C and ³¹P NMR spectra for L4 in CDCl₃





Figure S5 ¹H, ¹³C and ³¹P NMR spectra for Mn1 in DMSO-*d*₆









Figure S8 FT-IR spectrum of L2



Figure S9 FT-IR spectrum of Mn1



Figure S10 FT-IR spectrum of Mn2



4. Catalytic study

4.1 Optimizing reaction conditions for the ATH of acetophenone (a1) to (S)-1-phenylethan-1-ol (b1)

Under a nitrogen atmosphere, a mixture of acetophenone (**a1**) (0.25 mmol), the manganese complex (**Mn1** – **Mn2**) (5 μ mol) and base (namely *t*-BuOK, *t*-BuONa, *i*-PrONa, NaOMe KOH, NaOH, K₂CO₃, Na₂CO₃) (0.05 mmol) was dissolved in dry and degassed EtOH (5 mL) and then stirred and heated to the desired temperature (35 - 85 °C). At the specified reaction time (4 – 14 h), 0.1 mL of the reaction mixture was sampled and immediately diluted with 0.5 mL of EtOH, dodecane introduced, before being analyzed by GC. The composition of the reaction mixture was confirmed by running GC of a mixture of pure ketone, alcohol and dodecane. The enantiomeric excess (ee) of the alcohol was determined by chiral HPLC or GC analysis.

 Table S1 Screening of base^a

	0 2 mol% Mn1 20 mol% Base 55 °C, 12 h 5 mL EtOH a1		Ph-P-H-P Ph CO P Mn1	$ Br \stackrel{\Theta}{\longrightarrow} \\ \overrightarrow{Ph} \\ h $
Run	Base	Tem. (°C)	Yield (%) ^b	Ee (%) ^b
1	t-BuOK	55	99	65
2	<i>t</i> -BuONa	55	99	65
3	<i>t</i> -BuOLi	55	76	67
4	K ₃ PO ₄	55	31	67
5	CH₃ONa	55	30	68
6	КОН	55	97	73
7	NaOH	55	99	75
8	LiOH∙H₂O	55	63	68
9	K ₂ CO ₃	55	98	85
10	DBU	55	8	63
11	CsCO₃	55	41	59

^{*a*} Conditions: 0.25 mmol substrate, 5 μ mol **Mn1** (2 mol%), 0.05 mmol base (20 mol%), 5 mL EtOH, 50 bar H₂, 55 °C, 12 h; ^{*b*} Yields (%) and enantioselectivities (ee, %) were determined by GC (*n*-dodecane was used as an internal standard) and chiral-phase GC using a CP-Chirasil-Dex CB column, respectively.

	0 2 a1	2 mol% Mn1 20 mol% K ₂ CO ₃ - 55 °C, 4 - 14 h 5 mL EtOH b1	OH N COT Ph ⁻ P ⁻ Ph ⁻ CO Ph ⁻ CO Ph ⁻ CO Ph ⁻ CO Ph ⁻ CO Mn	Br ^O P-Ph Ph 1
Run	Time (h)	Tem. (°C)	Yield (%) ^b	Ee (%) ^b
1	4	55	48	73
2	6	55	61	79
3	8	55	76	62
4	10	55	79	73
5	12	55	97	82
6	14	55	99	77

Table S2 Optimizing the time for the AH of a1^a

^{*a*} Conditions: 0.25 mmol substrate, 0.05 mmol K₂CO₃, 5 μ mol **Mn1** (2 mol%), 0.05 mmol K₂CO₃ (20 mol%), 5 mL EtOH, 50 bar H₂, 55 °C, 4 - 14 h; ^{*b*} Yields (%) and enantioselectivities (ee, %) were determined by GC (*n*-dodecane was used as an internal standard) and chiral-phase GC using a CP-Chirasil-Dex CB column, respectively.

Table S3 Optimizing the temperature for the AH of **a1**^{*a*}

a1	2 mol% Mn1 20 mol% K ₂ CO ₃ 35 - 85 °C, 12 h 5 mL EtOH	OH OH N N N N N N N N N N N N N	$Br \stackrel{\Theta}{\longrightarrow} Br \stackrel{O}{\longrightarrow} Br \stackrel{O}{\longrightarrow$
Run	Tem. (°C)	Yield (%) ^b	Ee (%) ^b
1	35	9	50
2	45	11	50
3	55	97	82
4	65	99	71
5	75	99	58
6	85	99	50

^{*a*} Conditions: 0.25 mmol substrate, 0.05 mmol K_2CO_3 , 5 µmol **Mn1** (2 mol%), 5 mL EtOH, 50 bar H₂, 35 - 85 °C, 12 h;

^b Yields (%) and enantioselectivities (ee, %) were determined by GC (*n*-dodecane was used as an internal standard) and chiral-phase GC using a CP-Chirasil-Dex CB column, respectively.

4.2 General procedure for the asymmetric hydrogenation of ketones

Under a nitrogen atmosphere, a mixture of substrate (**a1** - **a15**) (0.25 mmol), the manganese complex (**Mn1**) (5 μ mol, 2 mol%) and K₂CO₃ (0.05 mmol) was dissolved in dry and degassed EtOH (5 mL) and then stirred and heated to the desired

temperature (55 °C). At the specified reaction time (12 h), 0.1 mL of the reaction mixture was sampled and immediately diluted with 0.5 mL of EtOH, dodecane introduced, before being analyzed by GC. The composition of the reaction mixture was confirmed by running GC of a mixture of pure ketone, alcohol and dodecane. The enantiomeric excess (ee) of the alcohol was determined by chiral GC analysis.

5. Characterization of the chiral alcohol products

The reaction mixture was purified by flash gel chromatography (using a gradient of PE/EA) to give the desired product. The enantiomeric excess (ee) of the alcohol was determined by chiral HPLC.

5.1 ¹H and ¹³ C{H} NMR spectroscopic data for the isolated chiral alcohol products.⁴
 (S)-1-phenylethan-1-ol (b1)



Colorless oil, 93% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.37 – 7.29 (m, 4H, Ar-H), 7.24 – 7.19 (m, 1H, Ar-H), 5.17 (s, 1H, OH), 4.73 (q, J = 6.5 Hz, 1H, CH), 1.34 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 147.8, 128.4, 126.9, 125.7, 68.5, 26.4.

(S)-1-(2-fluorophenyl)ethan-1-ol (**b2**)





Colorless oil, 17% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.52 (td, J = 7.7, 1.9 Hz, 1H, Ar-H), 7.26 – 7.10 (m, 2H, Ar-H), 7.05 (m, J = 10.9, 8.0, 1.3 Hz, 1H, Ar-H), 5.29 (d, J = 4.5 Hz, 1H, OH), 5.04 – 4.95 (m, 1H, CH), 1.32 (d, J = 6.6 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 160.5, 158.1, 134.5, 134.4, 128.7, 128.6, 127.4, 127.3, 124.7, 124.6, 115.3, 115.1, 62.5, 62.5, 25.1.

(S)-1-(2-chlorophenyl)ethan-1-ol (b3)



Colorless oil, 50% isolated yield.¹H NMR (400 MHz, DMSO- d_6) δ 7.61 (m, J = 8.0, 1.8 Hz, 1H, Ar-H), 7.31 (m, J = 7.1, 6.3, 1.4 Hz, 2H, Ar-H), 7.23 – 7.15 (m, 1H, Ar-H), 5.37 (d, J = 4.3 Hz, 1H, OH), 5.03 (qd, J = 6.3, 4.2 Hz, 1H, CH), 1.29 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 145.0, 130.7, 129.2, 128.5, 127.7, 127.3, 65.4, 24.7.

(S)-1-(2-bromophenyl)ethan-1-ol (b4)



White solid, Mp: 54-56 °C, 38% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.66 (m, J = 7.8, 1.8 Hz, 1H, Ar-H), 7.55 (m, J = 8.0, 1.2 Hz, 1H, Ar-H), 7.41 (td, J = 7.5, 1.2 Hz, 1H, Ar-H), 7.19 (td, J = 7.6, 1.8 Hz, 1H, Ar-H), 5.44 (d, J = 4.0 Hz, 1H, OH), 5.05 – 4.97 (m, 1H, CH), 1.34 (d, J = 6.4 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 146.5, 132.5, 129.0, 128.3, 127.6, 121.1, 67.8, 24.8.

(S)-1-(o-tolyl)ethan-1-ol (**b5**)





Colorless oil, 14% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.51 – 7.39 (m, 1H, Ar-H), 7.20 – 7.12 (m, 1H, Ar-H), 7.10 – 7.02 (m, 2H, Ar-H), 5.05 (d, *J* = 4.0 Hz, 1H, OH), 4.92 (m, *J* = 6.4, 4.2 Hz, 1H, CH), 2.26 (s, 3H, CH₃), 1.29 (d, *J* = 6.5 Hz, 3H, CH₃);

¹³C{H} NMR (100 MHz, DMSO-*d*₆) δ 145.8, 133.9, 130.2, 126.7, 126.2, 125.2, 65.3, 25.0, 19.0.

(S)-1-(3-fluorophenyl)ethan-1-ol (b6)



Colorless oil, 99% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.29 (td, J = 8.1, 6.0 Hz, 1H, Ar-H), 7.19 – 7.06 (m, 2H, Ar-H), 7.02 – 6.90 (m, 1H, Ar-H), 5.29 (d, J = 4.4 Hz, 1H, OH), 4.72 (qd, J = 6.4, 4.4 Hz, 1H, CH), 1.30 (d, J = 6.6 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 163.9, 161.5, 151.0, 130.3, 130.2, 121.7, 113.6, 113.4, 112.4, 112.2, 68.0, 68.0, 26.1.

(S)-1-(3-chlorophenyl)ethan-1-ol (**b7**)



Colorless oil, 66% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.40 (t, J = 1.9 Hz, 1H, Ar-H), 7.35 – 7.24 (m, 3H, Ar-H), 5.30 (s, 1H, OH), 4.74 (q, J = 6.4 Hz, 1H, CH), 1.33 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 150.5, 133.3, 130.5, 130.3, 126.8, 125.6, 124.4, 124.3, 67.9, 26.2.

(S)-1-(3-bromophenyl)ethan-1-ol (b8)



Colorless oil, 73% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.50 (t, J = 1.9 Hz, 1H, Ar-H), 7.31 (dt, J = 7.7, 1.6 Hz, 1H, Ar-H), 7.26 (dt, J = 7.9, 1.5 Hz, 1H, Ar-H), 7.18 (t, J = 7.8 Hz, 1H, Ar-H), 5.26 (d, J = 4.4 Hz, 1H, OH), 4.68 (qd, J = 6.4, 4.3 Hz, 1H, CH), 1.27 (d, J = 6.6 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 150.7, 130.6, 129.7, 128.6, 128.4, 124.8, 122.0, 68.0, 26.2.

(S)-1-(m-tolyl)ethan-1-ol (b9)



Colorless oil, 42% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.18 – 7.01 (m, 3H, Ar-H), 6.95 (dt, *J* = 7.4, 1.7 Hz, 1H, Ar-H), 5.07 (d, *J* = 4.2 Hz, 1H, OH), 4.65 (qd, *J* = 6.4, 4.1 Hz, 1H, CH), 2.23 (d, *J* = 1.3 Hz, 3H, CH₃), 1.28 (d, *J* = 6.6 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 147.8, 147.8, 137.3, 128.3, 127.5, 126.4, 122.8, 68.7, 26.4, 21.5.

(S)-1-(4-fluorophenyl)ethan-1-ol (b10)



Colorless oil, 50% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.40 – 7.31 (m, 2H, Ar-H), 7.15 – 7.06 (m, 2H, Ar-H), 5.21 (d, J = 4.2 Hz, 1H, OH), 4.73 (qd, J = 6.4, 4.2 Hz, 1H, CH), 1.31 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 162.6, 160.2, 143.9, 127.6, 115.0, 67.9, 26.3.

(S)-1-(4-chlorophenyl)ethan-1-ol (**b11**)



Colorless oil, 95% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.36 (d, J = 1.2 Hz, 4H, Ar-H), 5.28 (d, J = 2.9 Hz, 1H, OH), 4.77 – 4.70 (m, 1H, CH), 1.32 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 146.8, 131.4, 128.3, 127.6, 67.9, 26.2.

(S)-1-(4-bromophenyl)ethan-1-ol (b12)



colorless oil, 90% isolated yield.¹H NMR (400 MHz, DMSO- d_6) δ 7.57 – 7.47 (m, 2H, Ar-H), 7.36 – 7.28 (m, 2H, Ar-H), 5.30 (s, 1H, OH), 4.73 (q, J = 6.4 Hz, 1H, CH), 1.33 (d, J = 6.5 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 147.2, 131.3, 128.0, 119.8, 67.9, 26.2.

(S)-1-(p-tolyl)ethan-1-ol (b13)



Colorless oil, 59% isolated yield. ¹H NMR (400 MHz, DMSO- d_6) δ 7.29 – 7.22 (m, 2H, Ar-H), 7.14 (d, J = 7.8 Hz, 2H, Ar-H), 5.10 (d, J = 3.9 Hz, 1H, OH), 4.71 (qd, J = 6.3, 3.4 Hz, 1H, CH), 2.30 (s, 3H, CH₃), 1.33 (d, J = 6.4 Hz, 3H, CH₃); ¹³C{H} NMR (100 MHz, DMSO- d_6) δ 144.8, 135.8, 128.9, 125.7, 125.6, 68.3, 26.4, 21.1.

5.2 NMR spectra for chiral alcohol products.

Figure S11 ¹H and ¹³C NMR spectra of b1; recorded in DMSO- d_6 at ambient temperature





Figure S12 ¹H and ¹³C NMR spectra of b2 in DMSO- d_6



Figure S13 ¹H and ¹³C NMR spectra of b3 in DMSO- d_6

Figure S14 ¹H and ¹³C NMR spectra of b4; recorded in DMSO- d_6 at ambient temperature



f1 (ppm) ò . 90 . 50

Figure S15 ¹H and ¹³C NMR spectra of b5; recorded in DMSO- d_6 at ambient temperature





Figure S16 ¹H and ¹³C NMR spectra of b6; recorded in DMSO- d_6 at ambient temperature



Figure S17 ¹H and ¹³C NMR spectra of b7; recorded in DMSO-*d*₆ at ambient temperature

f1 (ppm)



. 90 . 80 . 40 . 30 . 10 f1 (ppm)



Figure S19 ¹H and ¹³C NMR spectra of b9; recorded in DMSO- d_6 at ambient temperature

f1 (ppm)



Figure S20 ¹H and ¹³C NMR spectra of b10; recorded in DMSO- d_6 at ambient temperature

110 100 f1 (ppm)



Figure S21 ¹H and ¹³C NMR spectra of b11; recorded in DMSO-*d*₆ at ambient temperature

. 190 . 40 . 30 f1 (ppm)





. 190 . 40 . 30 f1 (ppm)



f1 (ppm)

6. GC spectra of the chiral alcohol products

(S)-1-(2-fluorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(2-chlorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(2-bromophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(o-tolyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(3-fluorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(3-chlorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(3-bromophenyl)ethan-1-ol CP-Chirasil-Dex CB column Column $25m \times 0.25 \ mm \times 0.25um$ Injector Temp °C 240 Detector Temp °C 260 Injection volume 0.8 uL 120 °C, withdraw time 2 min, then 5 °C /min to Column Temp 180 °C keeping for 30 min 24.446 LmV OH 23.584 22.722 21.860 -22.126 20.998 ÑĂ 20.136 19.274 18.412 17.550 16.688 15.826 0.26 4.96 7.31 16.71 19.06 21.41 2.61 9.66 12.01 14.36 23.76 [Unit: min] Peak area (AU *S) Entry Time (min) Height (AU) Content (%) 1 22.126 873.7 8730.9 50.6353 2 22.501 822.9 8511.8 49.3647 219.406 LmV 189.623 159.840 OH 130.057 100.274 - 22. 302 70.491 15.625 ΞÌ 40.708 10.925 -18.858 -48.641 -78.424 4.40 3.56 5.72 7.88 14.36 16.52 1B. 68 20.84 23.00 10.04 12.20 [Unit: min] Entry Time (min) Height (AU) Peak area (AU *S) Content (%) 211798.4 98398.5 21.9046 1 15.625 2 21.999 56082.7 6828.7 12.4847 3 22.302 24173.8 294732.2 65.6107

(S)-1-(m-tolyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(4-fluorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(4-chlorophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(4-bromophenyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



(S)-1-(p-tolyl)ethan-1-ol Column

Injector Temp °C Detector Temp °C Injection volume Column Temp



7. X-ray structure determination

The single crystal X-ray diffraction studies of Mn1 were carried out on a Rigaku Sealed Tube CCD (Saturn 724⁺) diffractometer using graphite-mono-chromated Cu-K α radiation (λ = 1.54178) at 170.00 K. The cell parameters were obtained by global refinement of the positions of all collected reflections. Intensities were corrected for Lorentz and polarization effects and empirical absorption. The structures were solved by direct methods and refined by full-matrix least-squares on F². All nonhydrogen atoms were refined anisotropically and all hydrogen atoms were placed in calculated positions. Structure solution was performed using SHELXT (2015) and structure refinement performed using SHELXL.⁵ Crystal data and processing parameters for Mn1 are summarized in Tables S4.



Chart S1 ORTEP representations of Mn1. Thermal ellipsoids are shown at 30% probability, and hydrogen atoms have been omitted for clarity.

Table S4 Crystal data and structure refinement for				
Identification code	Mn1			
CCDC No.	2408177			
Empirical formula	$C_{46}H_{40}MnN_2O_2P_2$			
Formula weight	769.68			
Temperature/K	169.99(10)			
Crystal system	trigonal			
Space group	P3 ₁ 21			
a/Å	17.5937(3)			
b/Å	17.5937(3)			
c/Å	29.4019(7)			
α/°	90			
β/°	90			
γ/°	120			

Table S4 Cry	stal data	and structure	refinement	for Mn	1
able 34 City	ystai uata	and shutter	rennement		

Volume/ų	7881.7(3)
Z	6
$\rho_{calc}g/cm^3$	0.973
µ/mm⁻¹	2.853
F(000)	2406.0
Crystal size/mm ³	$0.15 \times 0.1 \times 0.08$
Radiation	Cu Kα (λ = 1.54178)
20 range for data collection/°	5.8 to 155.754
Index ranges	$-21 \leq h \leq 22, -21 \leq k \leq 19, -36 \leq l \leq 35$
Reflections collected	34069
Independent reflections	10455 [R _{int} = 0.0627, R _{sigma} = 0.0589]
Data/restraints/parameters	10455/108/443
Goodness-of-fit on F ²	1.055
Final R indexes [I>=2σ (I)]	$R_1 = 0.0674$, $wR_2 = 0.1839$
Final R indexes [all data]	R ₁ = 0.0796, wR ₂ = 0.1940
Largest diff. peak/hole / e Å $^{\text{-}3}$	0.65/-0.40
Flack parameter	0.102(9)

Table S5 Selected bond distances and bond angles for Mn-1

Selected bond distances (Å)		Selected bond angles	(°)
Mn1-P2	2.269(2)	P2-Mn1-P1	174.28(8)
Mn1-N1	2.081(6)	N1-Mn1-P1	103.62(16)
Mn1-N2	2.078(6)	N1-Mn1-P2	81.88(16)
Mn1-C20	1.795(7)	N2-Mn1-P1	81.37(17)
Mn1-C21	1.767(9)	N2-Mn1-P2	101.43(17)
P1-C34	1.820(7)	N2-Mn1-N1	79.0(2)
P1-C41	1.834(4)	C20-Mn1-P1	90.4(3)
P1-C35	1.826(8)	C20-Mn1-P2	87.6(3)
P2-C13	1.831(8)	C21-Mn1-P1	84.2(3)
P2-C7	1.839(4)	C21-Mn1-N1	170.2(3)
P2-C1	1.822(5)	C21-Mn1-C20	91.2(3)

8. DFT studies

Computational methods

All density functional theory (DFT) calculations were performed to understand the mechanism of the enantioselectivity of ketone hydrogenation catalyzed by manganese (Mn) catalysts. B3LYP-D3⁶ method combined with a mixed basis sets of SDD (and its ECP)⁷ for Mn, K and 6-31G(d)⁸ for the other atoms (denoted as BS1) were used to fully optimize all structures in the gas-phase. Then, vibrational

frequency calculations on these optimized geometries were carried out at the same level of theory to confirm no imaginary frequency for all local minimum, and one appropriate imaginary frequency for each transition state. The single-point energy calculations are also performed using the SDD/6-311+G(d,p) mixed basis set (denoted as BS2) in PCM solvent model⁹ using ethanol solvent to obtain more accurate electronic energies. To examine the effect of DFT functional, a few other common and reliable M06-L, PBE0-D3, and DB97X-D methods¹⁰ were also used for the single-point energy calculations on the key intermediates and transition states in dichloromethane solution. In addition, (relative) distortion/interaction energy analysis, non-covalent interactions (NCIs) plot analysis (isovalue=0.50) on the key structures were performed¹¹. All DFT calculations were carried out by Gaussian16 program¹². All 3D images of the optimized structures were shown by CYLview¹³. All free energies are corrected with Grimme's quasi-harmonic approximation for vibrational entropy correction at all frequencies below than 100 cm⁻¹ by GoodVibes using default settings¹⁴. The steric maps were generated by SambVca 2.0 program¹⁵. The distance of the coordination point from the center of the sphere is defined as 2.3 Angstrom, and the value of mesh spacing for numbering integration is defined as 0.050. Other default parameters are used.



Figure S24 Relative corrected free energies (in kcal/mol) for TSS_{Mn1} and TSS_{Mn2} calculated by PCM B3LYP-D3/BS2//B3LYP-D3/BS1 in ethanol solution.



Figure S25 Steric map of Mn-ligand part and substrate for TSR_{Mn1} with the computed buried volumes (in %).



Figure S26 Steric map of Mn-ligand part and substrate for TSS_{Mn1} with the computed buried volumes (in %).



Figure S27 The noncovalent interaction (NCI) plots for TSR_{Mn2} and TSS_{Mn2} (red: strong

repulsion; green: weak attraction; blue: strong attraction).

optimized structures in gas-phase.				
	⊡G _{soln} (B3LYP-D3)	⊡G _{soln} (M06-L)	⊡G _{soln} (PBE0-D3)	⊡G _{soln} (⊡B97X-D)
Mn1H	0.0	0.0	0.0	0.0
TSR _{Mn1}	15.5	25.6	18.7	21.6
TSS _{Mn1}	14.1	24.8	18.0	20.4
??G	1.4	0.8	0.8	1.3
Mn2H	0.0	0.0	0.0	0.0
TSR _{Mn2}	18.2	25.7	19.1	18.5
TSS _{Mn2}	17.5	25.0	18.6	18.4
??G	0.7	0.6	0.5	0.1

Table S6 The relative corrected free energies (in kcal/mol) of transition states in ethanol solution evaluated by the other PCM DFTs/BS2 methods based on the B3LYP-D3/BS1-optimized structures in gas-phase.

Table S7 The absolute electronic and corrected free energies (in Hartree) of the optimized structures of intermediates and transition states for reaction catalyzed by the Mn catalyst by the B3LYP-D3 method in gas-phase. The free energies (G-qh) are corrected with the quasi-harmonic entropy correction.

	Egas	G _{gas} -qh
Sub	-384.90722	-384.80067
Mn1H	-2740.32620	-2739.63865
TSR _{Mn1}	-3125.24547	-3124.42356
TSS _{Mn1}	-3125.24414	-3124.42118
Mn2H	-2713.80407	-2713.07353
TSR _{Mn2}	-3098.71225	-3097.84949
TSS _{Mn2}	-3098.71312	-3097.85047

Table S8 The absolute electronic energies (in Hartree) for Mn complexes and transition states evaluated by the other PCM (ethanol) DFTs/BS2 methods based on the B3LYP-D3/BS1-optimized structures in gas-phase.

	E _{soln} (B3LYP-D3)	E _{soln} (M06-L)	E _{soln} (PBE0-D3)	E _{soln} (⊡B97X-D)
Sub	-385.02010	-384.94953	-384.55286	-384.86927
Mn1H	-2740.93513	-2740.57474	-2738.23350	-2740.11136
TSR _{Mn1}	-3125.95831	-3125.51123	-3122.78429	-3124.97398
TSS _{Mn1}	-3125.96161	-3125.51353	-3122.78654	-3124.97704
Mn2H	-2714.40040	-2714.01585	-2711.70536	-2713.57707
TSR _{Mn2}	-3099.41711	-3098.95014	-3096.25345	-3098.44245
TSS _{Mn2}	-3099.41818	-3098.95106	-3096.25410	-3098.44254

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10 Cartesian coordinates XYZs.

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Mn1H.xyz
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Mn	-0.033852	-0.409004	-0.480256
Р	-2.270973	-0.422195	-0.019773
Р	2.225769	-0.579040	-0.244897
Ν	0.145427	1.406720	-1.577647
Ν	-0.036839	0.901939	1.095463
0	-0.055269	-3.156521	0.558462
С	-3.127027	-2.395559	3.550762
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С	-3.672832	-3.599946	-2.240205
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Η	-0.165525	4.120586	-2.127179
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Н	-3.009547	1.607849	2.080875
С	4.921250	2.722179	-2.127417
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С	5.480779	1.789213	-1.252806
Н	6.535166	1.836101	-0.992987
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