

SUPPORTING INFORMATION

SnO₂ quantum dots decoration of CuO nanoparticles with enhanced NO₂ and H₂ gas sensing response via *p-n* heterojunction interfaces

Valentina Paolucci¹, Thirugnanam Natarajan¹, Vittorio Ricci¹, Fabiola Ferrante¹, Carlo Cantalini¹

¹*Department of Industrial and Information Engineering and Economics, University of L'Aquila, I-67100 L'Aquila, Italy*

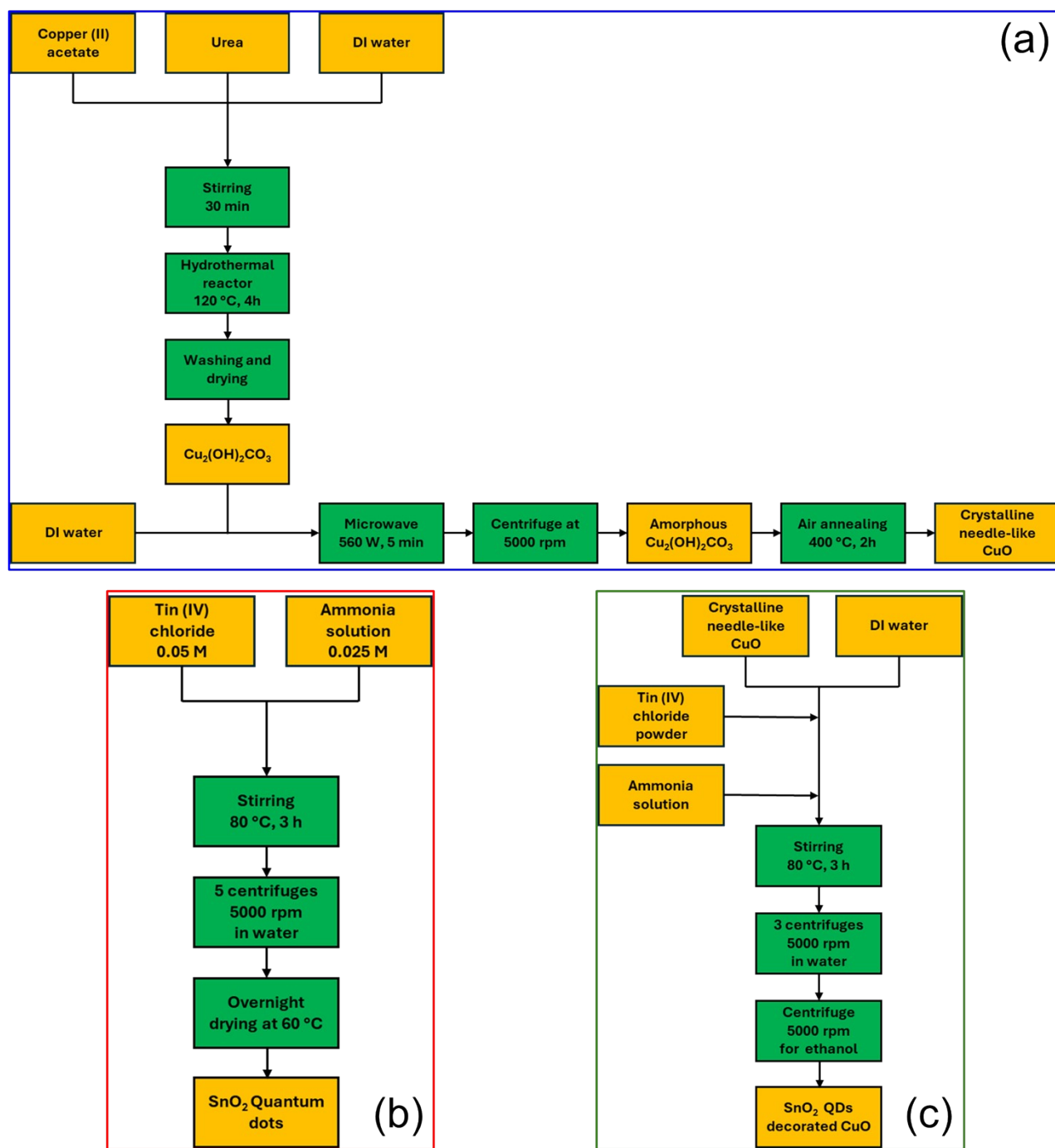


Figure S1. Flowcharts of the synthesis processes of: (a) Needle-like CuO nanoparticles; (b) SnO₂ quantum dots; (c) SnO₂ quantum dots-decorated CuO.

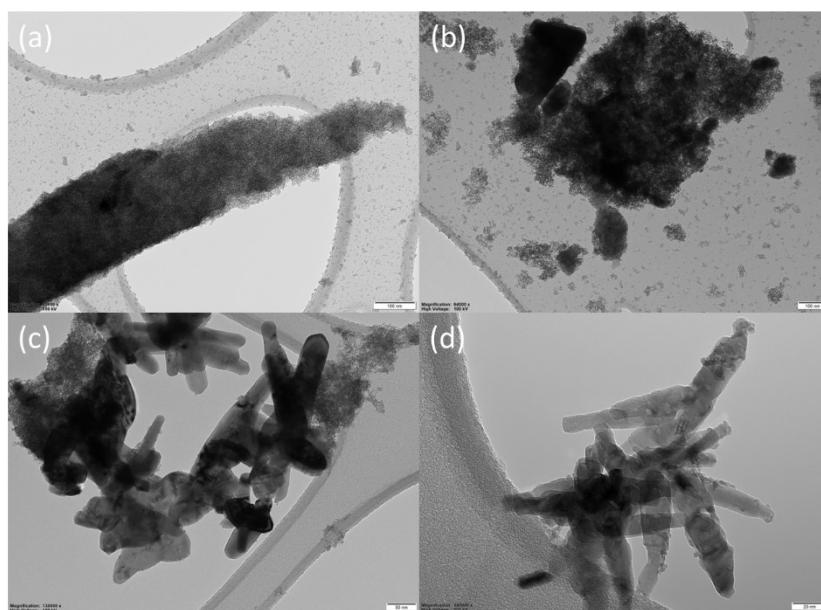


Figure S2. Optimization of the synthesis procedure. (a,b) Equimolar CuO and SnCl₄: SnO₂ QDs formed agglomerated clusters homogeneously dispersed in solution, with no preferential anchoring to CuO. (c) Reduced Sn precursor concentration: partial improvement in the flakes/dots ratio, although nucleation is still independent of CuO. (d) Optimized route: addition of SnCl₄ directly to the CuO dispersion, followed by ammonia, yielding selective nucleation on CuO and the desired decoration with SnO₂ dots.

The synthesis was first attempted by combining equimolar amounts of CuO and SnCl₄. In this case, SnCl₄ was initially dissolved in the 0.05 M ammonia solution and then mixed with the CuO dispersion (**Figure S1a-b**). The resulting SnO₂ quantum dots were agglomerated and these clusters homogeneously dispersed in substrate, rather than anchored to CuO, showing that nucleation occurred independently of the CuO surface. In addition, the amount of SnCl₄ used was found to be excessive. A second trial (**Figure S1c**), performed by reducing the Sn precursor concentration (1:10 as respect to CuO), led to a partial improvement in the flakes/dots ratio, although nucleation was still not correlated with CuO. The final optimization step was achieved by modifying the order of reagent addition: SnCl₄ (2.6 mg) was introduced directly into the CuO dispersion (25 mg in 20 ml DI water), and ammonia was added only afterwards. This simple adjustment successfully directed nucleation onto the CuO surface, leading to controlled decoration rather than uncontrolled nucleation in solution, therefore yielding the desired SnO₂ decoration (**Figure S1d**).

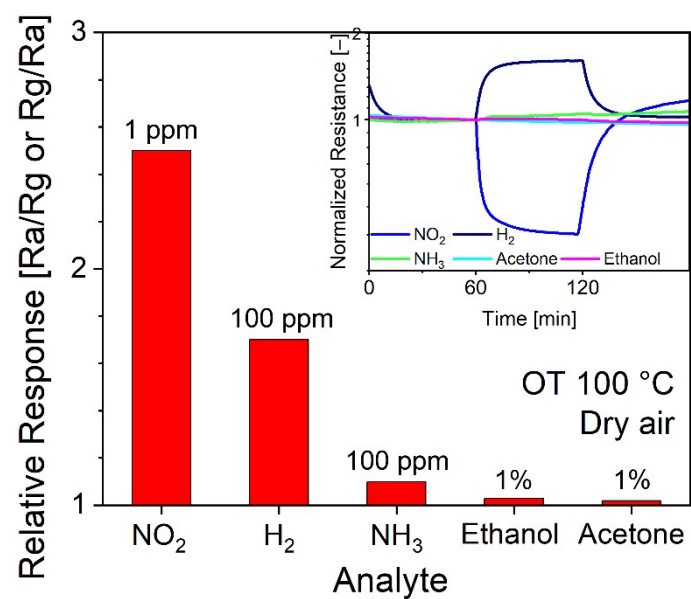


Figure S3. Selectivity responses to different concentrations of various gases and vapours (NO₂, H₂, NH₃, Ethanol and Acetone) measured at 100°C OT and dry air atmosphere.