

Supporting information for

Influence of Aliovalent Substitution on Structure and Dynamics in Sodium Halide $\text{Na}_{3-2x}\text{Y}_{1-x}\text{Nb}_x\text{Cl}_6$ Solid Electrolytes

Brian B. Phan ^a, Tso Shuen ^b, Dmitry Vrublevskiy ^a, Qingyu Yan ^b and Vladimir K. Michaelis ^{a,*}

^a Department of Chemistry, University of Alberta, Edmonton, Alberta, Canada, T6G 2G2

^b School of Materials Science and Engineering, Nanyang Technological University, Singapore 639798

Table of Contents

Figure S1	2
Table S1	2
Figure S2	3
Figure S3	3
Table S2	4
Figure S4	5
Figure S5	6
Figure S6	7
Figure S7	8

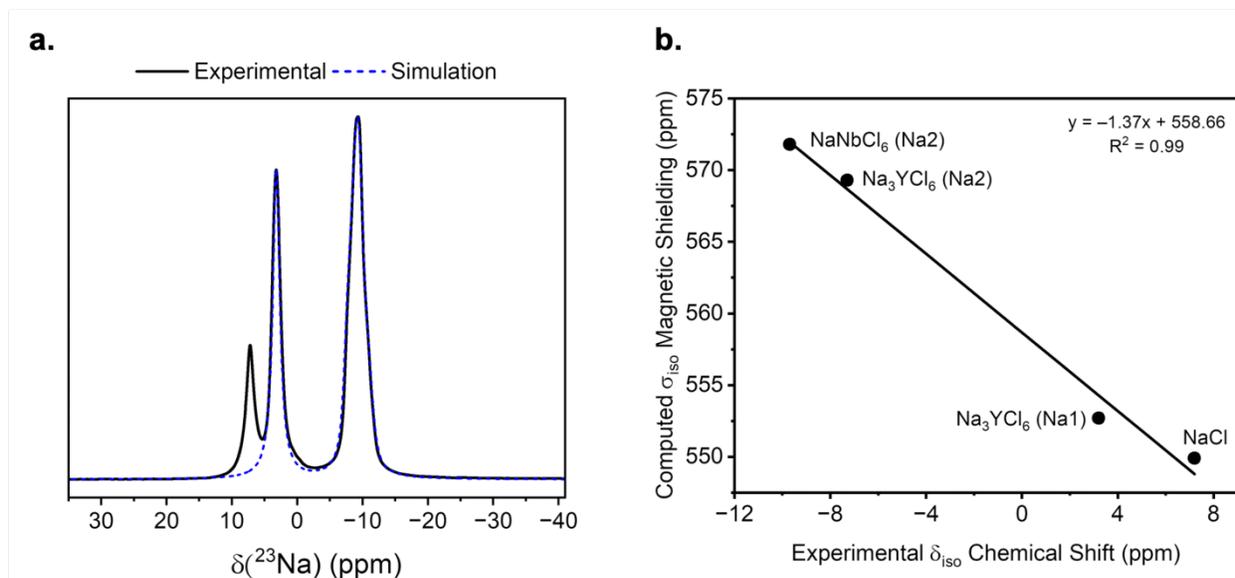


Figure S1. (a) ^{23}Na MAS NMR ($B_0 = 11.7$ T, $\nu_{\text{rot}} = 10$ kHz) of Na_3YCl_6 simulations (dashed blue line) of octahedral and prismatic sites overlaid on top of experimental (black) data. (b) Plot of ^{23}Na magnetic shielding values computed from CASTEP against experimental isotropic chemical shifts of NaCl , Na_3YCl_6 , and NaNbCl_6 .

Table S1. Experimental and GIPAW DFT computed ^{23}Na NMR shielding parameters for Na_3YCl_6 and NaNbCl_6 .

Sample	δ_{iso} (ppm)		C_Q (MHz)		η	
	Experimental	Computed ^a	Experimental	Computed	Experimental	Computed
Na_3YCl_6	3.2(3)	5.9	0.0(2)	0.9	0.1(6)	0.45
	-7.3(3)	-10.6	1.0(3)	1.3	0.6(9)	0.65
NaNbCl_6	-9.7(3)	-13.1	0.0(3)	0.9	0.0(1)	0.6

^a ^{23}Na magnetic shielding tensors were converted into chemical shift values using $\delta(\text{ppm}) \approx \sigma_{\text{ref}}(\text{ppm}) - \sigma_{\text{DFT}}(\text{ppm})$, where $\sigma_{\text{ref}} = 558.6$ ppm was used based on the relationship of NaCl and the end member.

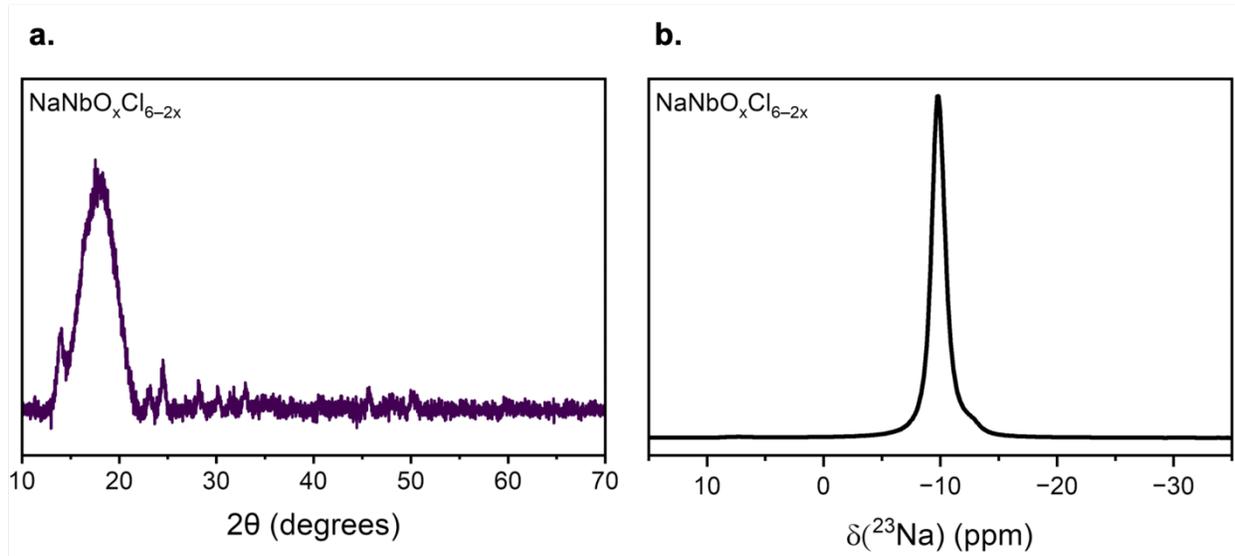


Figure S2. (a) PXRD and (b) ^{23}Na MAS NMR ($B_0 = 18.8$ T, $v_{\text{rot}} = 10$ kHz) of $\text{NaNbO}_x\text{Cl}_{6-2x}$.

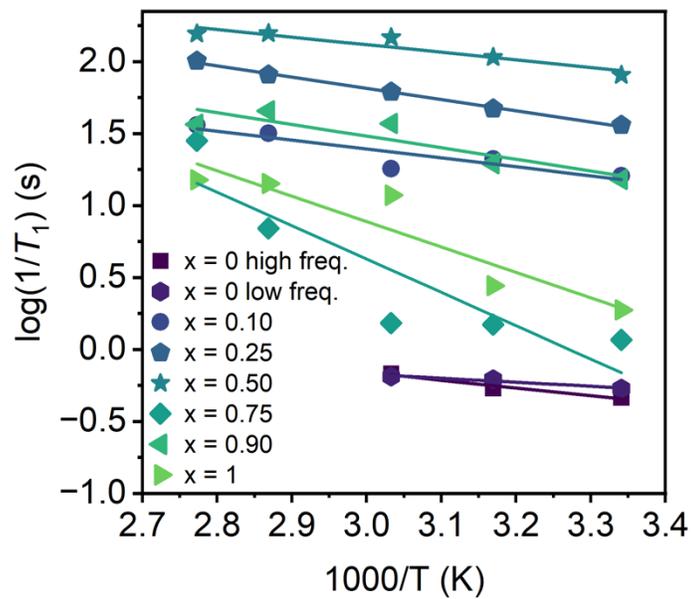


Figure S3. Arrhenius plot of local Na-ion hopping rates derived from spin-lattice relaxation rates as a function of $1000/T$.

Table S2. Activation energies extracted from ^{23}Na VT NMR of $\text{Na}_{3-2x}\text{Y}_{1-x}\text{Nb}_x\text{Cl}_6$ series.

Sample	E_a VT NMR (eV)
Na_3YCl_6	0.04(6) @ 3.2 ppm 0.02(4) @ -9.1 ppm
$\text{Na}_{2.8}\text{Y}_{0.90}\text{Nb}_{0.10}\text{Cl}_6$	0.05(3)
$\text{Na}_{2.5}\text{Y}_{0.75}\text{Nb}_{0.25}\text{Cl}_6$	0.06(6)
$\text{Na}_2\text{Y}_{0.50}\text{Nb}_{0.50}\text{Cl}_6$	0.04(5)
$\text{Na}_{1.5}\text{Y}_{0.25}\text{Nb}_{0.75}\text{Cl}_6$	0.19(9)
$\text{Na}_{1.2}\text{Y}_{0.10}\text{Nb}_{0.90}\text{Cl}_6$	0.06(9)
NaNbCl_6	0.15(1)

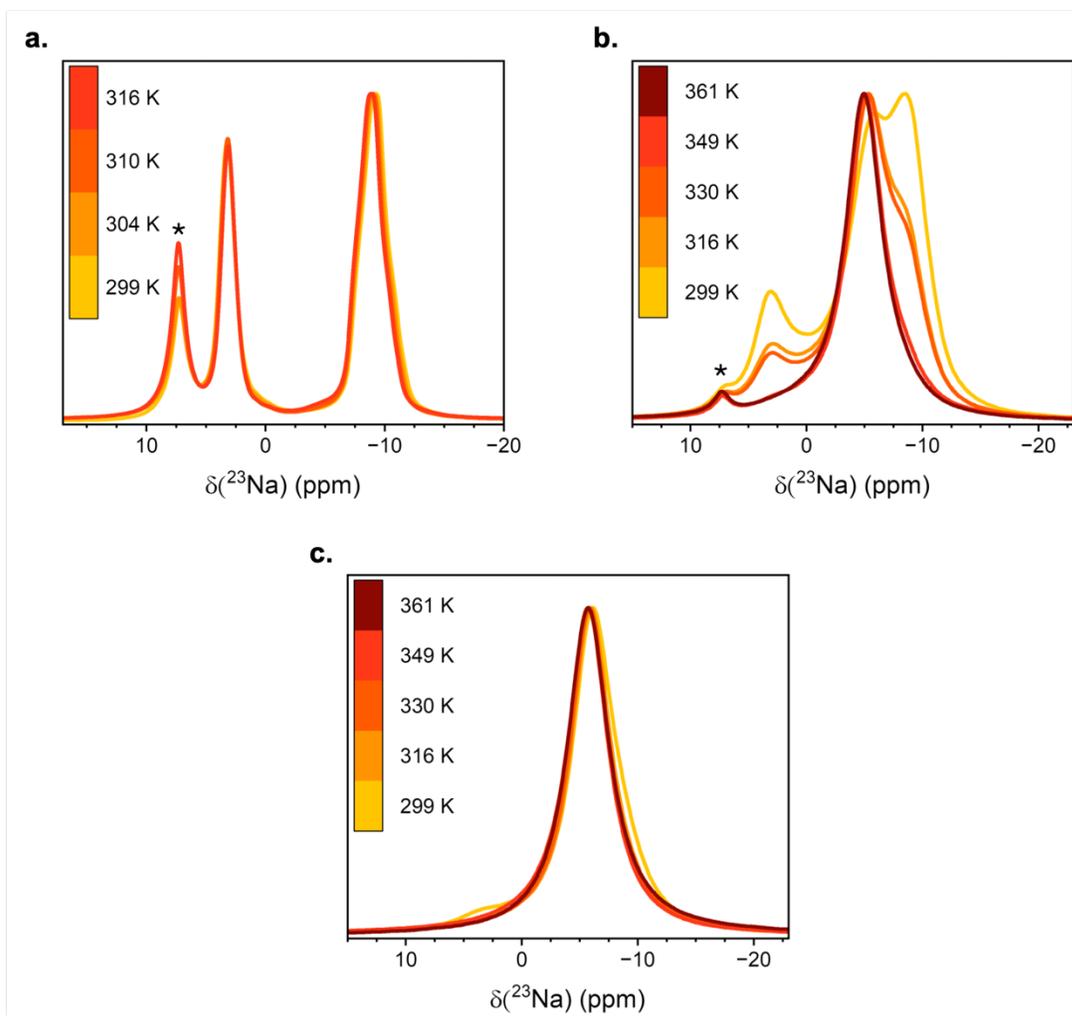


Figure S4. VT ^{23}Na MAS NMR spectra ($B_0 = 11.7$ T, $\nu_{\text{rot}} = 10$ kHz) of (a) YNb0 (b) YNb10 and (c) YNb25. Asterisk (*) indicates NaCl impurity.

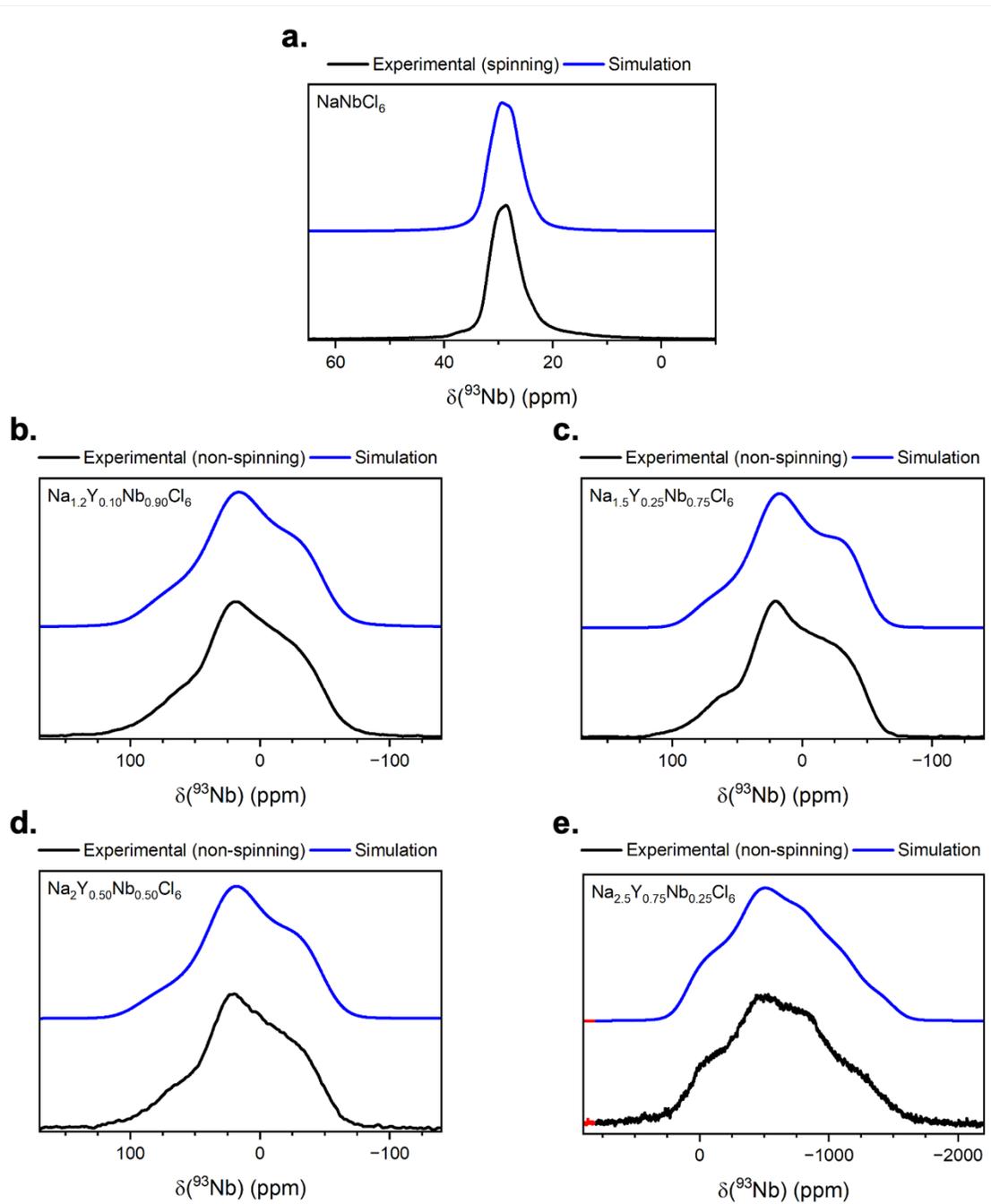


Figure S5. ^{93}Nb NMR spectra of (a) NaNbCl_6 acquired at $B_0 = 18.8$ T ($\nu_{\text{rot}} = 22$ kHz). Non-spinning ^{93}Nb NMR spectra ($B_0 = 11.7$ T) of (b) YNb90, (c) YNb75, (d) YNb50, and (e) YNb25 ($B_0 = 18.8$ T).

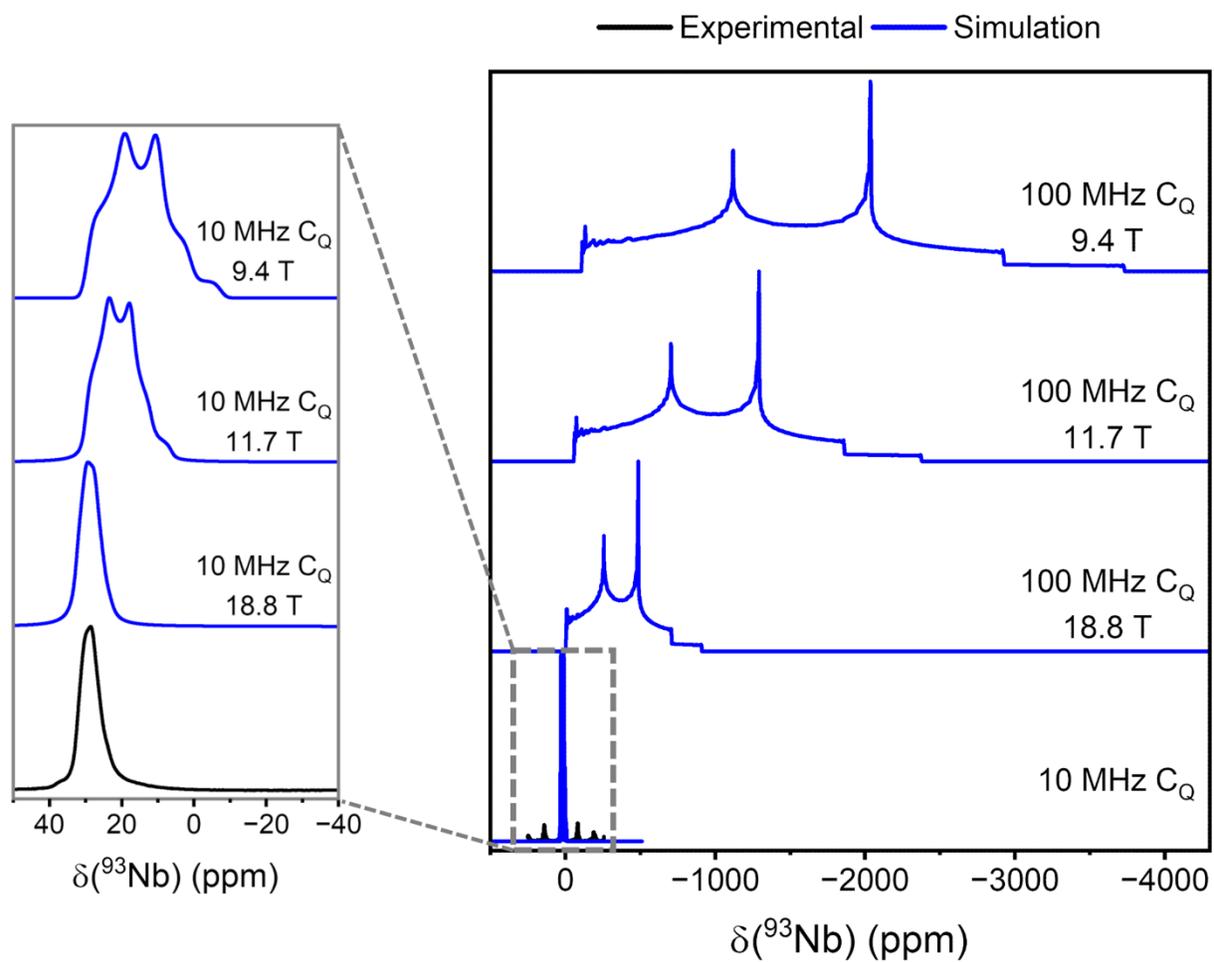


Figure S6. ^{93}Nb MAS NMR of NaNbCl_6 (experimental, $\nu_{\text{rot}} = 22$ kHz; simulation, $\nu_{\text{rot}} = \infty$) simulated at 9.4 T, 11.7 T, and 18.8 T with either $C_Q = 10$ or 100 MHz. For consistency, $\eta = 0.47$ and $\delta_{\text{iso}} = 30$ ppm were used across the various simulations.

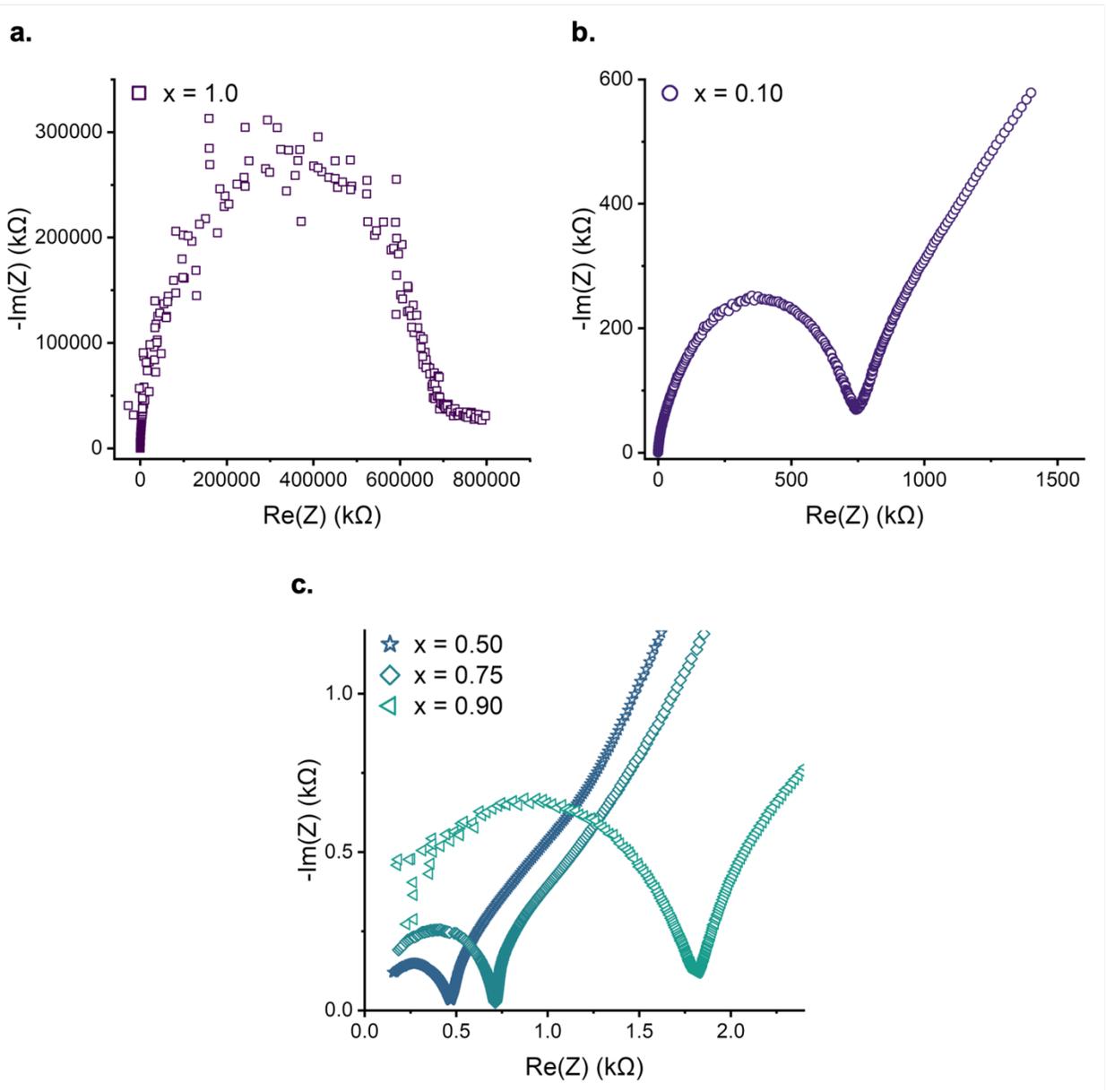


Figure S7. Nyquist plot of room temperature electrochemical impedance measurements for (a) YNb0, (b) YNb10, and (c) YNb50, YNb75, and YNb90.