

Supplementary Information

Theoretical Evaluation of Pristine, Single B and N-Doped, and BN Co-Doped Graphenylene as Metal-Free Cathode Catalysts for Nonaqueous Li-O₂ Batteries

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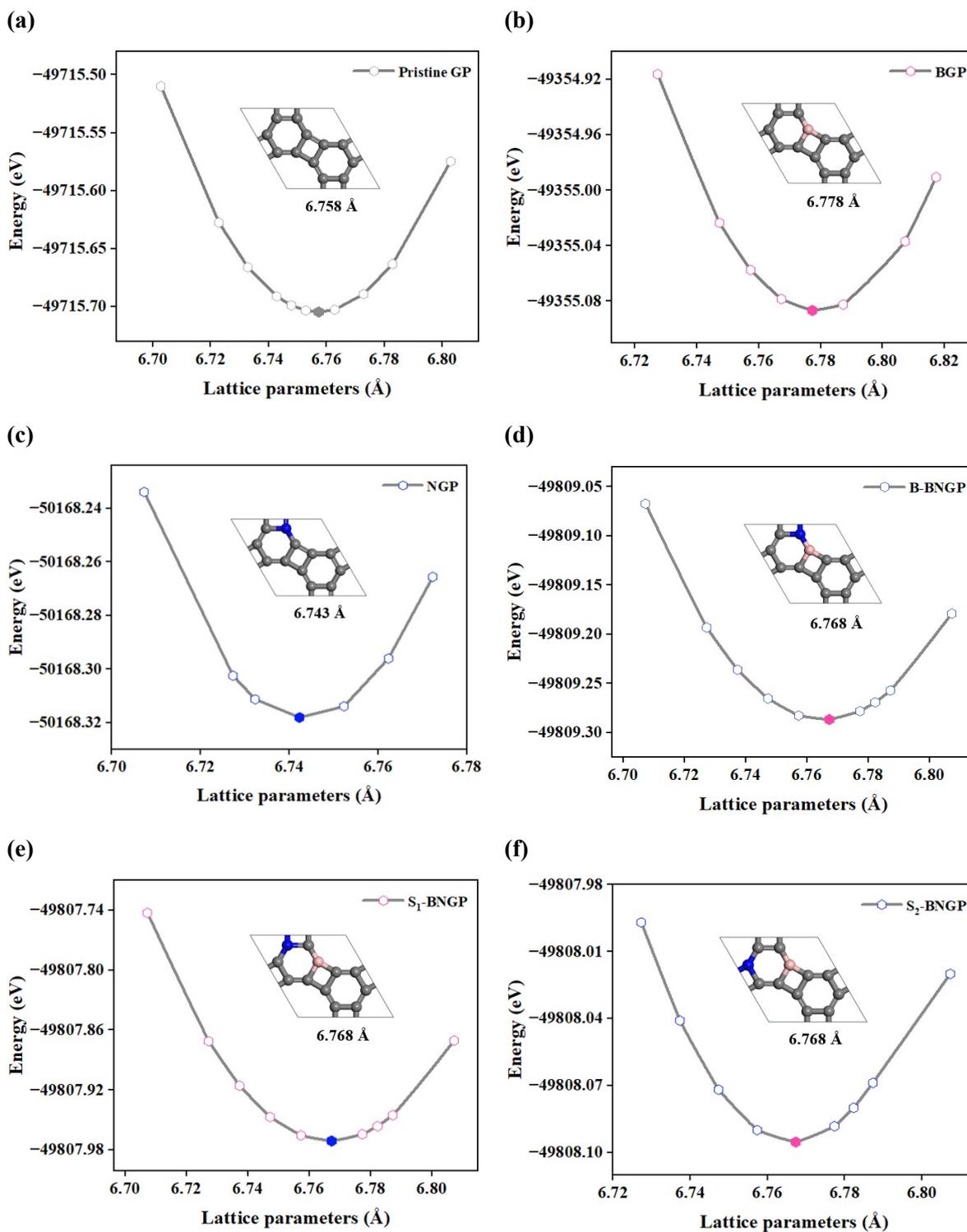


Figure S1. Optimized lattice constants (Å) of (a) pristine GP, (b) BGP, (c) NGP, (d) B-BNGP, (e) S₁-BNGP, and (f) S₂-BNGP.

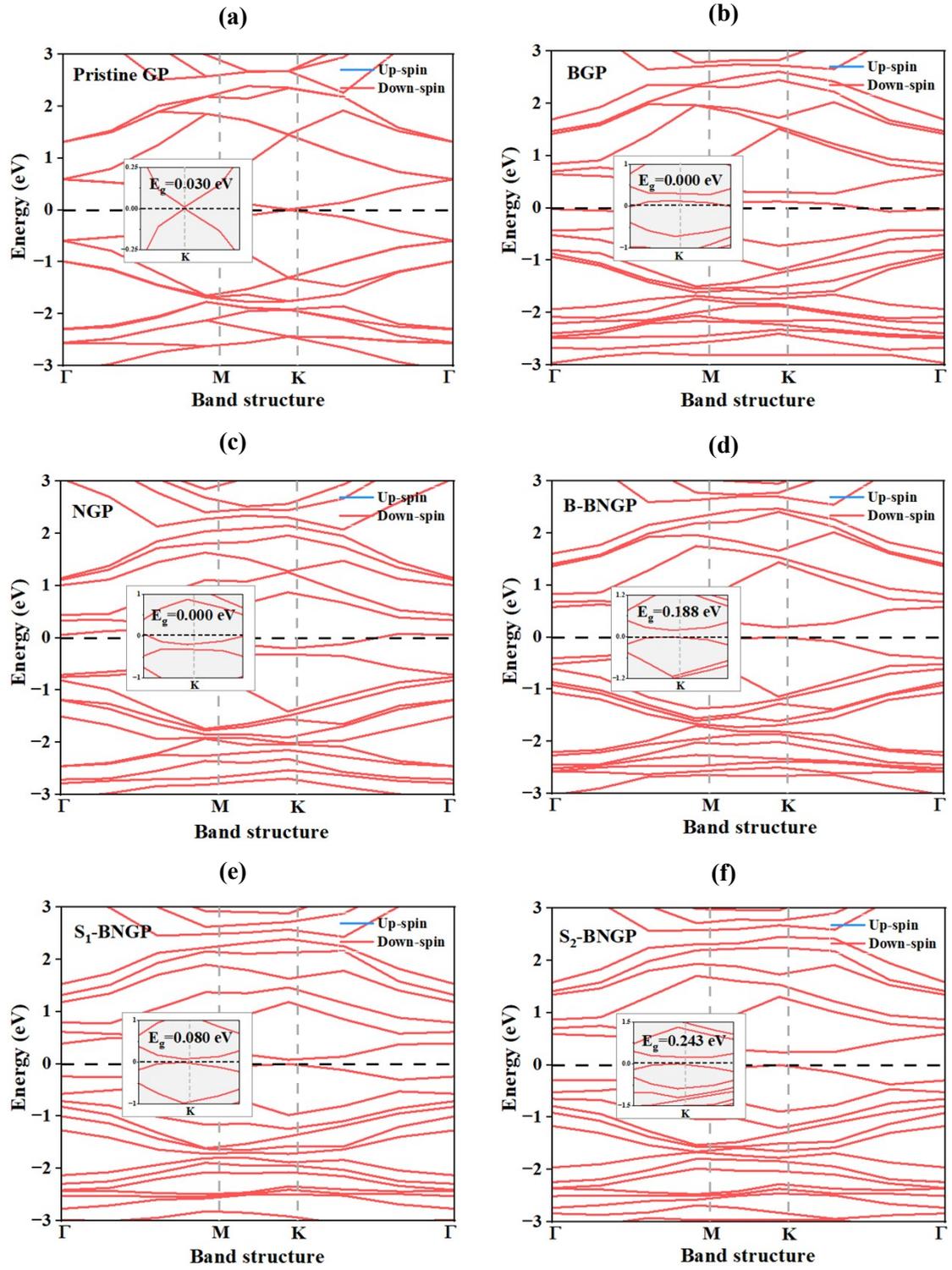


Fig S2. Electronic band structures of (a) pristine GP, (b) BGP, (c) NGP, (d) B-BNGP, (e) S_1 -BNGP, and (f) S_2 -BNGP; the Fermi level is indicated by a black dashed line set at zero energy. Blue and red lines represent spin-up and spin-down states, respectively, and the insets show magnified views around the K point and the band gaps.

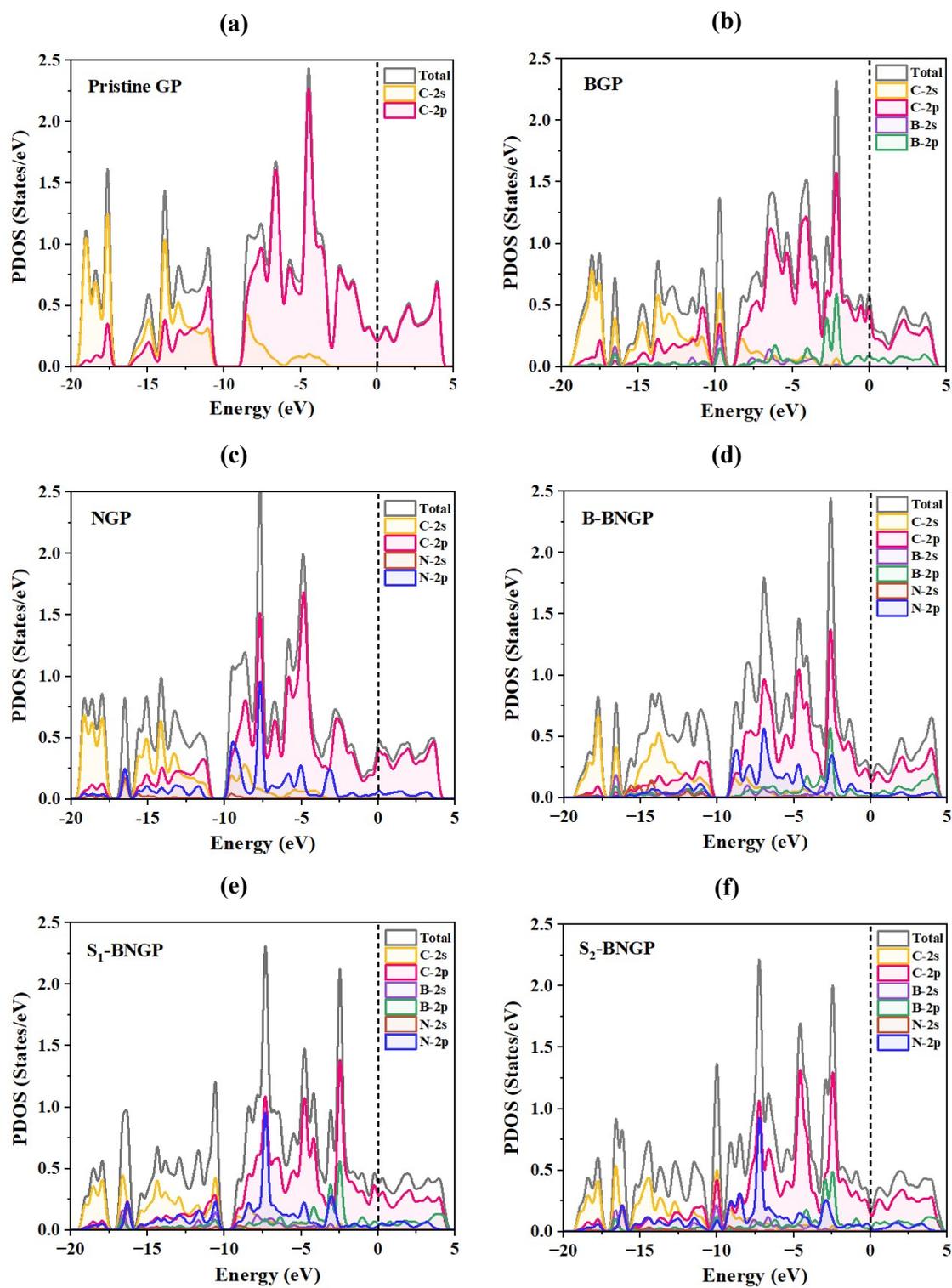


Fig S3. PDOS images of (a) pristine GP, (b) BGP, (c) NGP, (d) B-BNGP, (e) S_1 -BNGP, and (f) S_2 -BNGP; the Fermi level is indicated by a black dashed line set at zero energy.

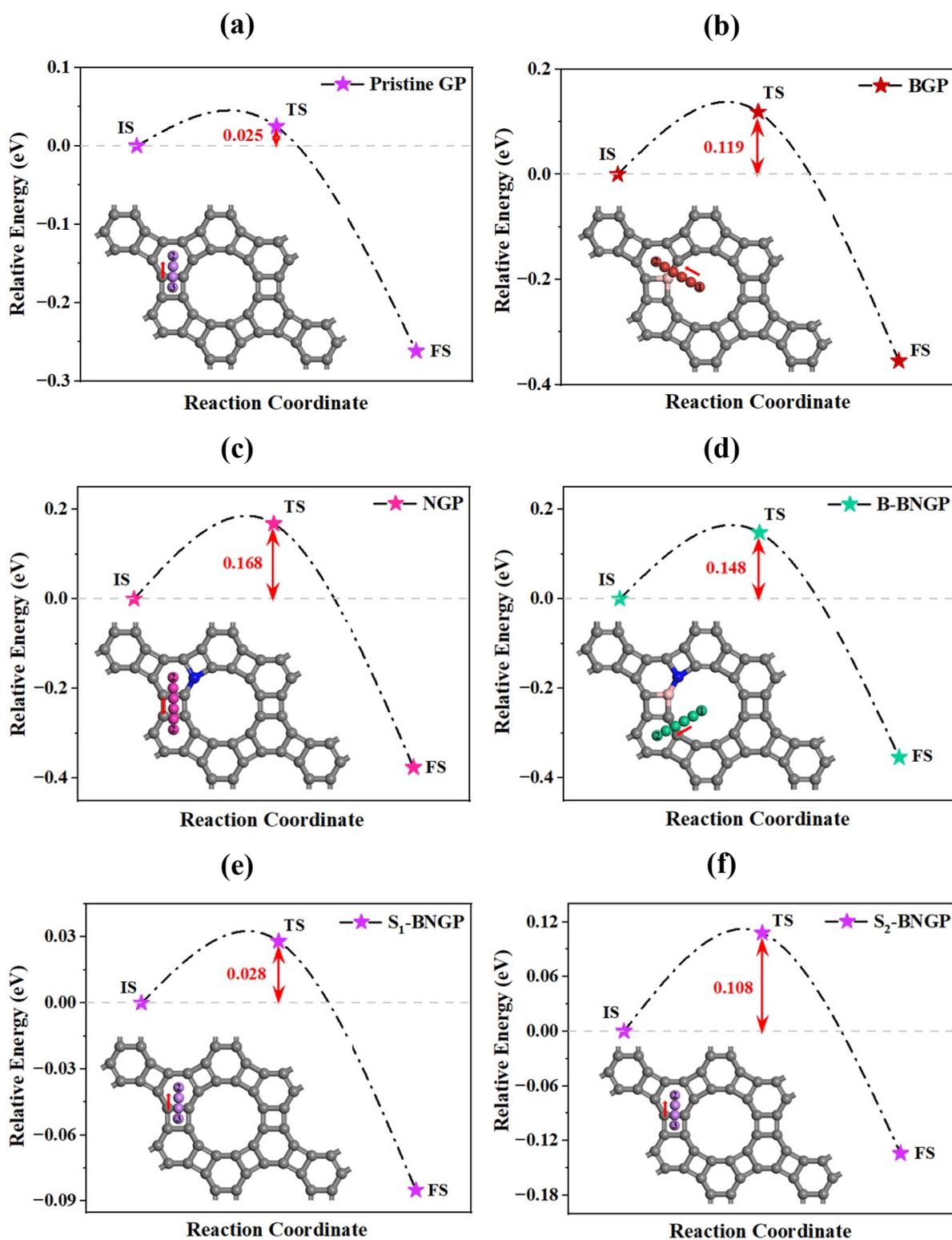


Fig S4. Relative energy profiles and diffusion barriers (eV) for Li atom migration on (a) pristine GP, (b) BGP, (c) NGP, (d) B-BNGP, (e) S₁-BNGP, and (f) S₂-BNGP surfaces; IS, TS, and FS denote the initial, transition, and final states, respectively; C, B, and N atoms are shown as gray, pink, and blue spheres. Migration paths H₃→H₂, H₂→H₂, and H₁→H₂ are marked by purple, pink, and red-green balls, respectively.

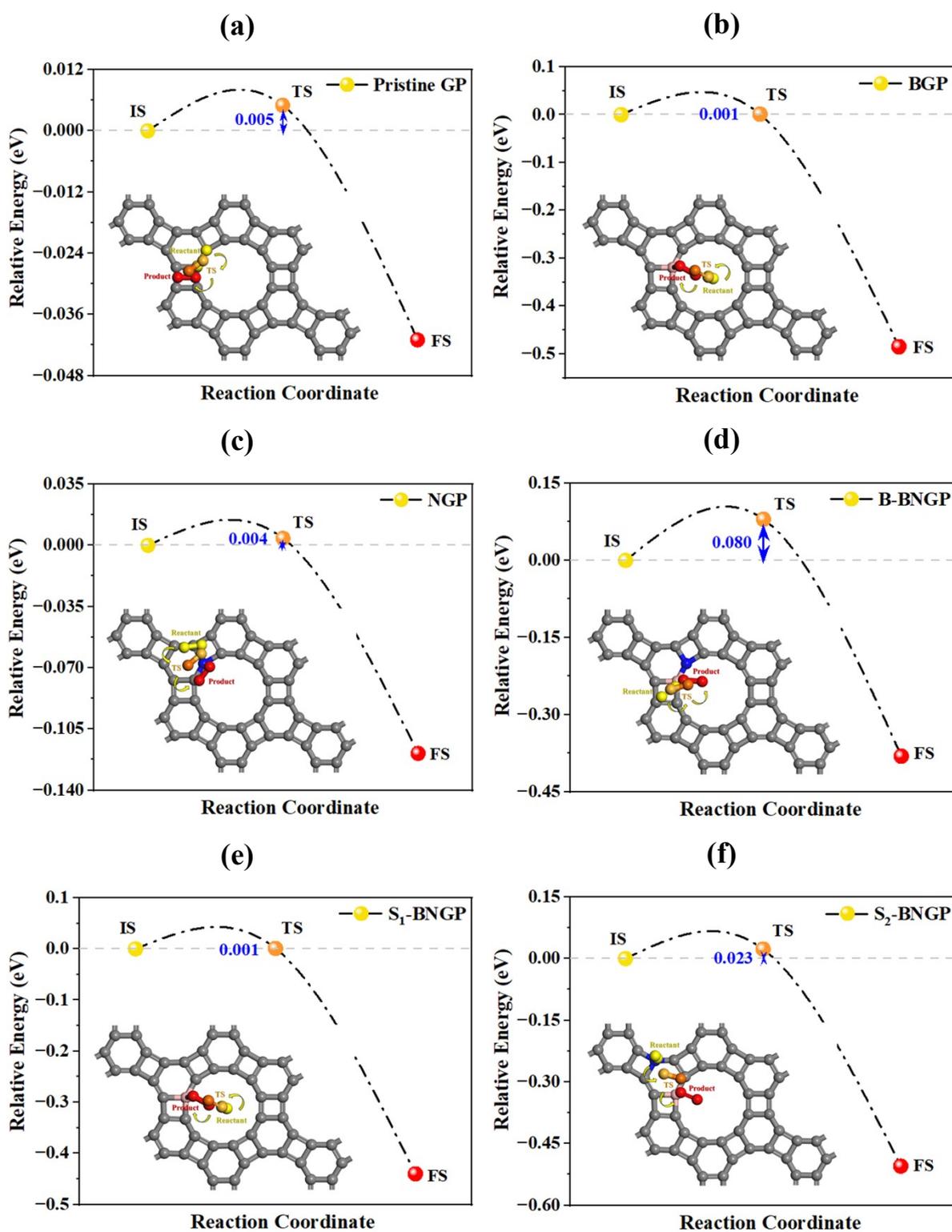


Fig S5. Relative energy profiles and diffusion barriers (eV) for O₂ molecule migration on (a) pristine GP, (b) BGP, (c) NGP, (d) B-BNGP, (e) S₁-BNGP, and (f) S₂-BNGP surfaces; C, B, and N atoms are shown as gray, pink, and blue spheres, respectively; initial (IS), transition (TS), and final (FS) states are indicated by yellow, orange, and red spheres, respectively.

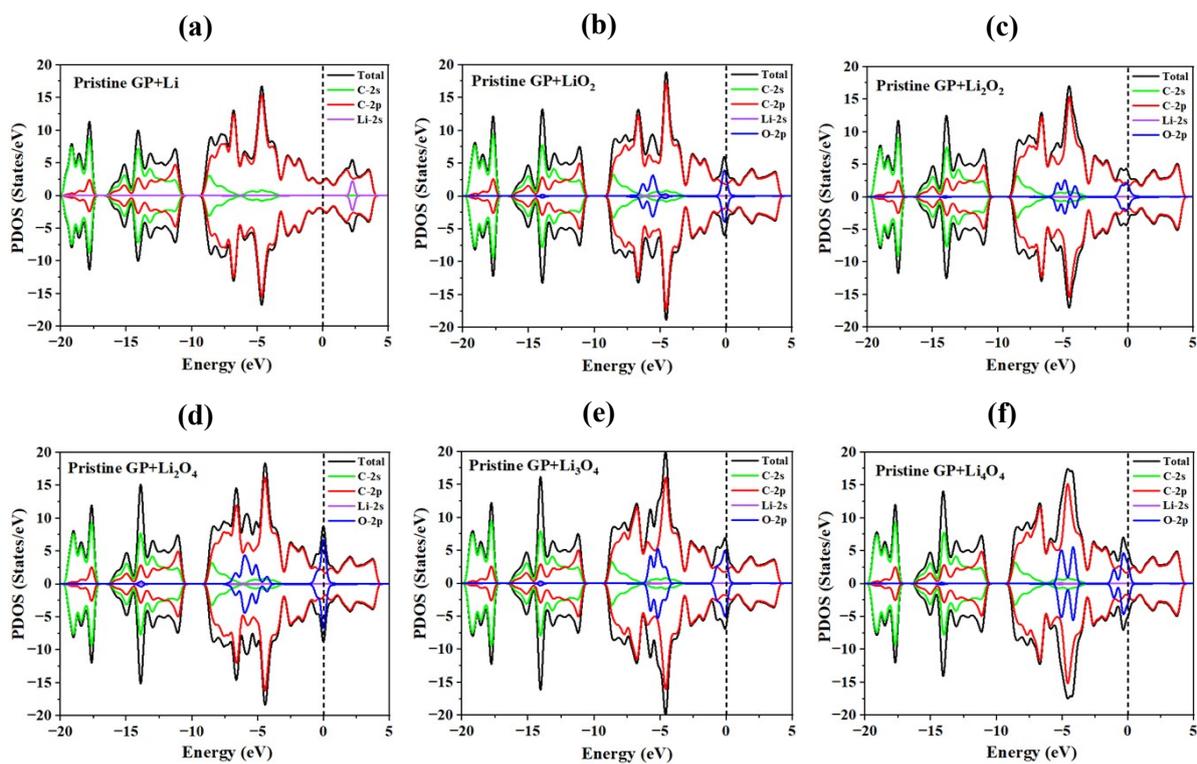


Fig S6. PDOS images of (a) Li/O₂, (b) LiO₂, (c) Li₂O₂, (d) Li₂O₄, (e) Li₃O₄, and (f) Li₄O₄ adsorbed on the pristine GP, BGP, NGP, S₁-BNGP, and S₂-BNGP structures, respectively; the Fermi level is indicated by a black dashed line set at zero energy.

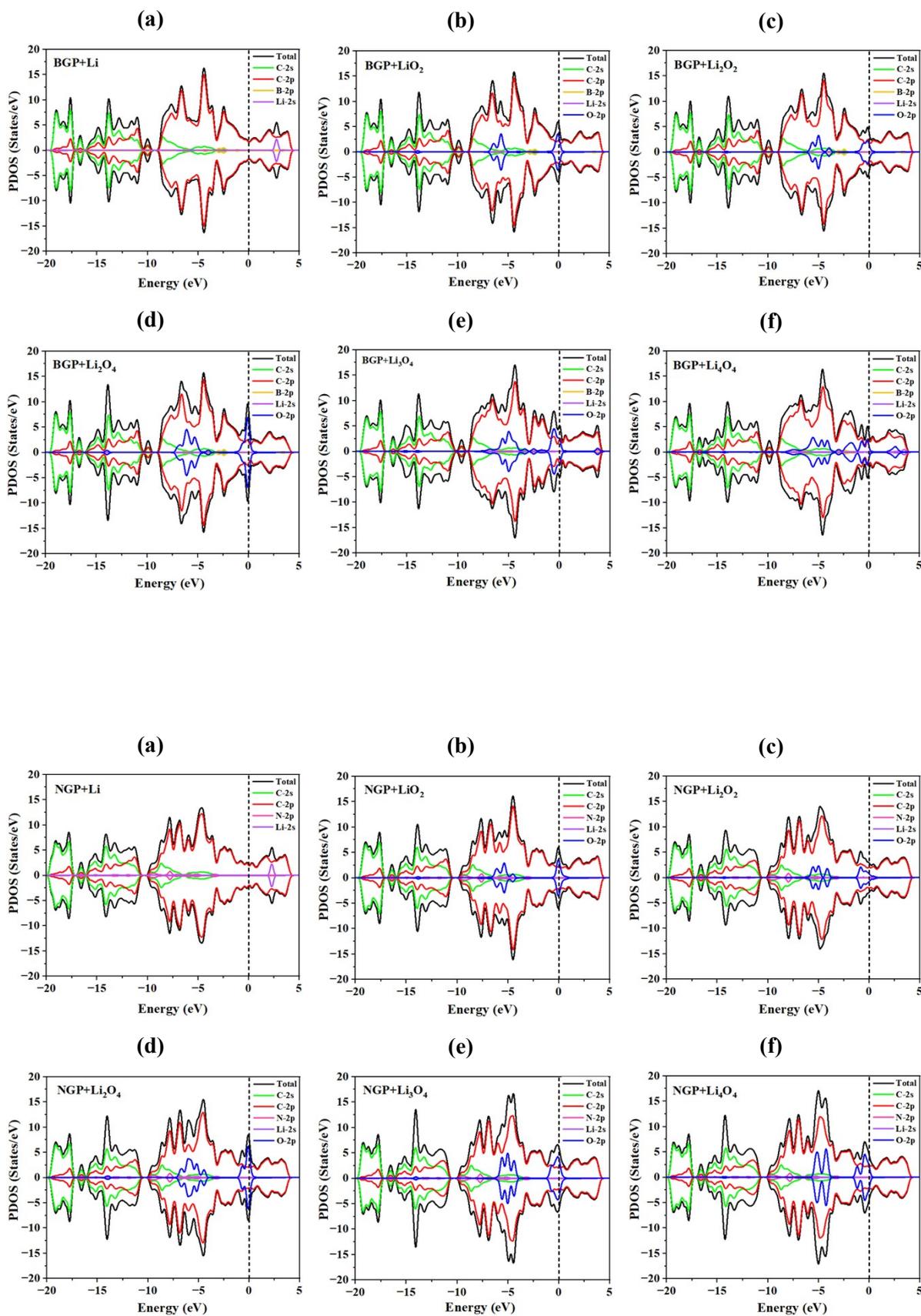


Fig S6. (continued).

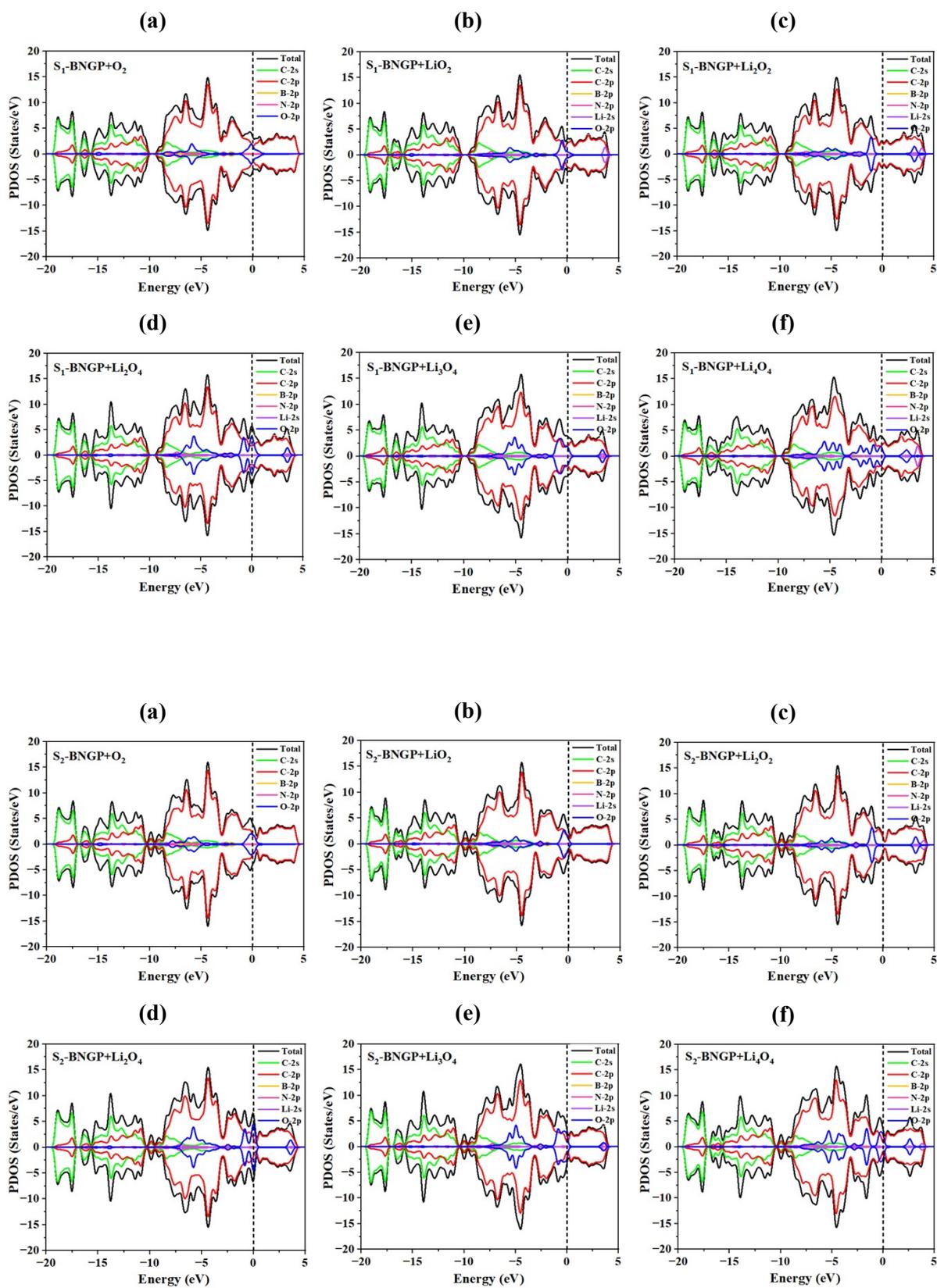


Fig S6. (continued).

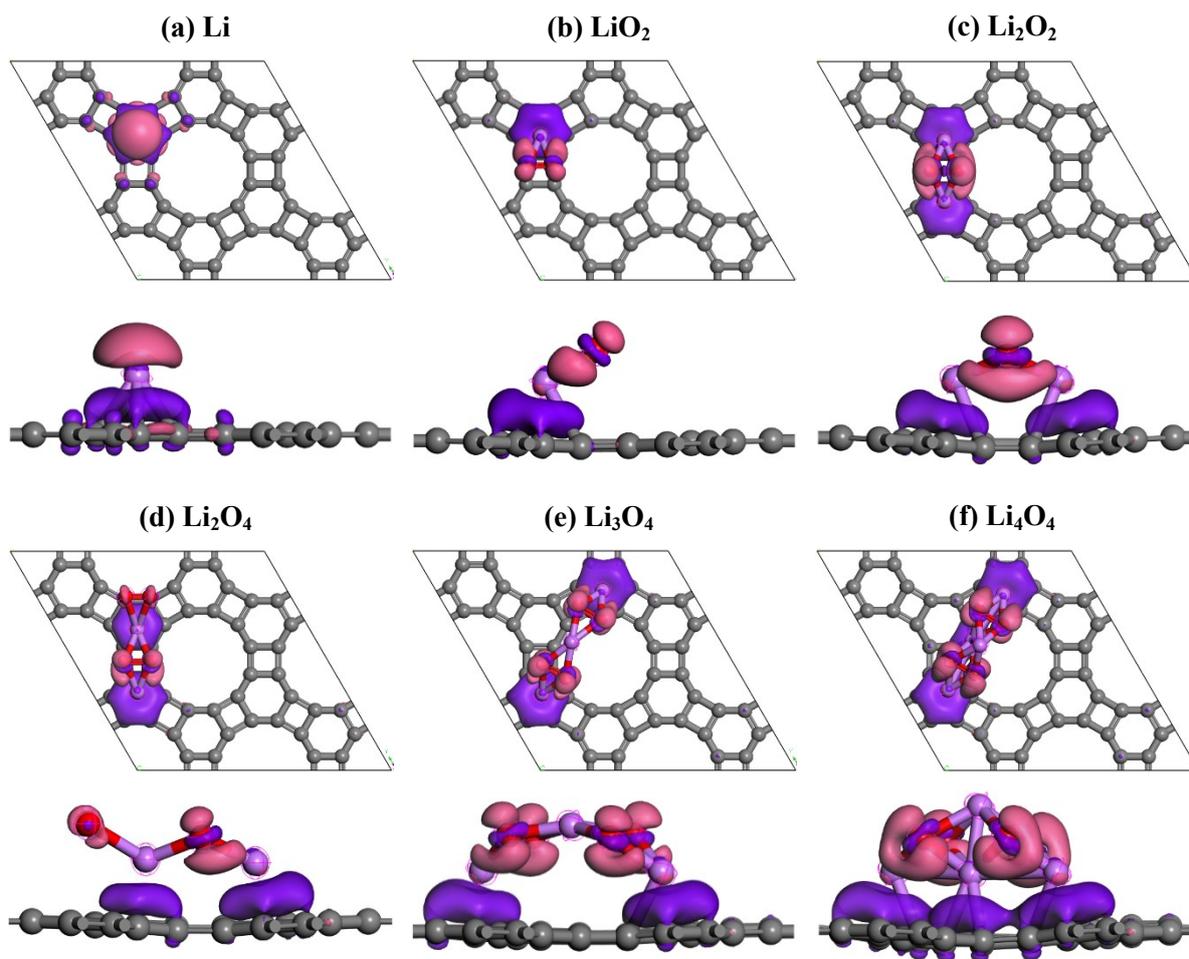


Fig S7. CDD images of (a) Li/O₂, (b) LiO₂, (c) Li₂O₂, (d) Li₂O₄, (e) Li₃O₄, and (f) Li₄O₄ adsorbed on the pristine GP, BGP, NGP, S₁-BNGP, and S₂-BNGP structures, respectively; purple and pink regions represent charge accumulation and depletion, respectively.

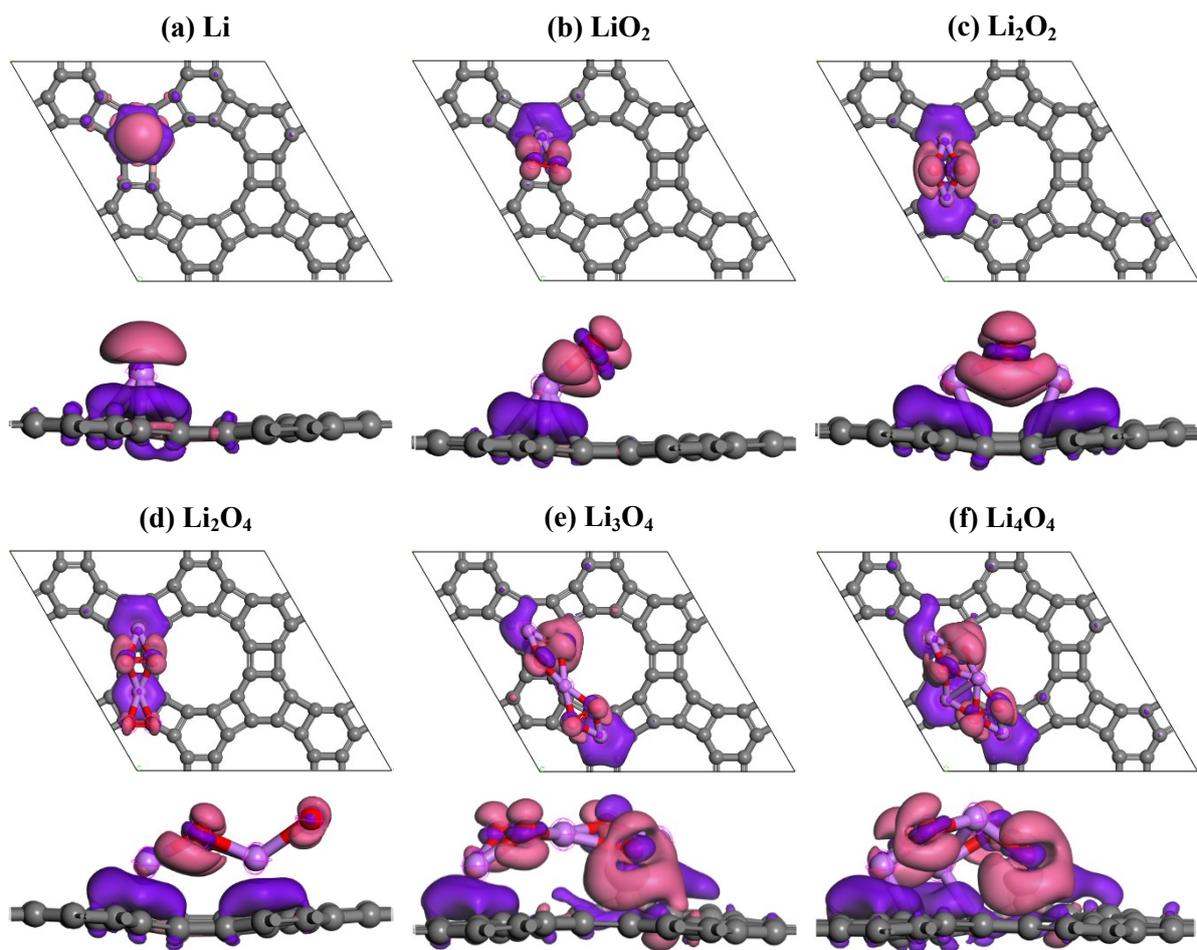


Fig S7. (continued).

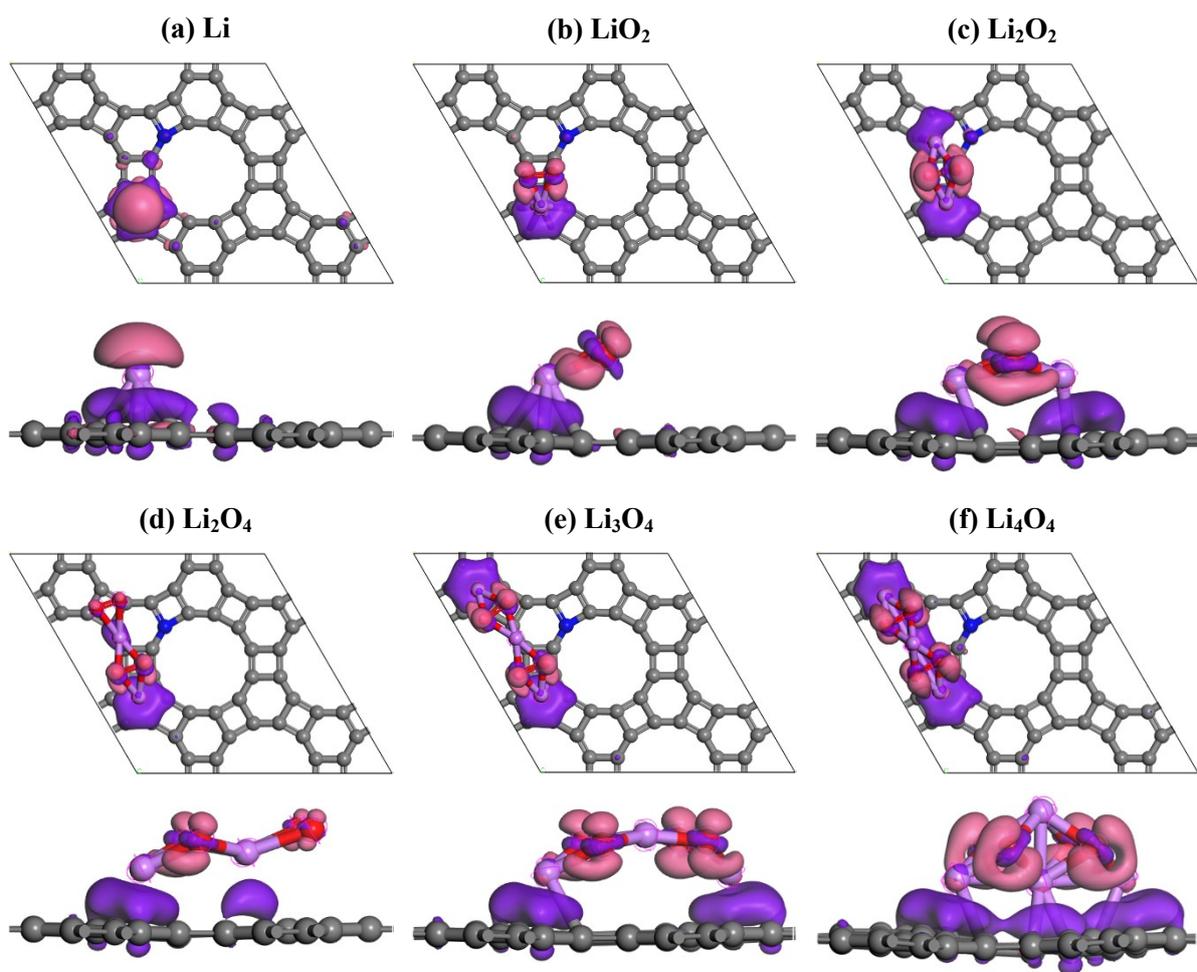


Fig S7. (continued).

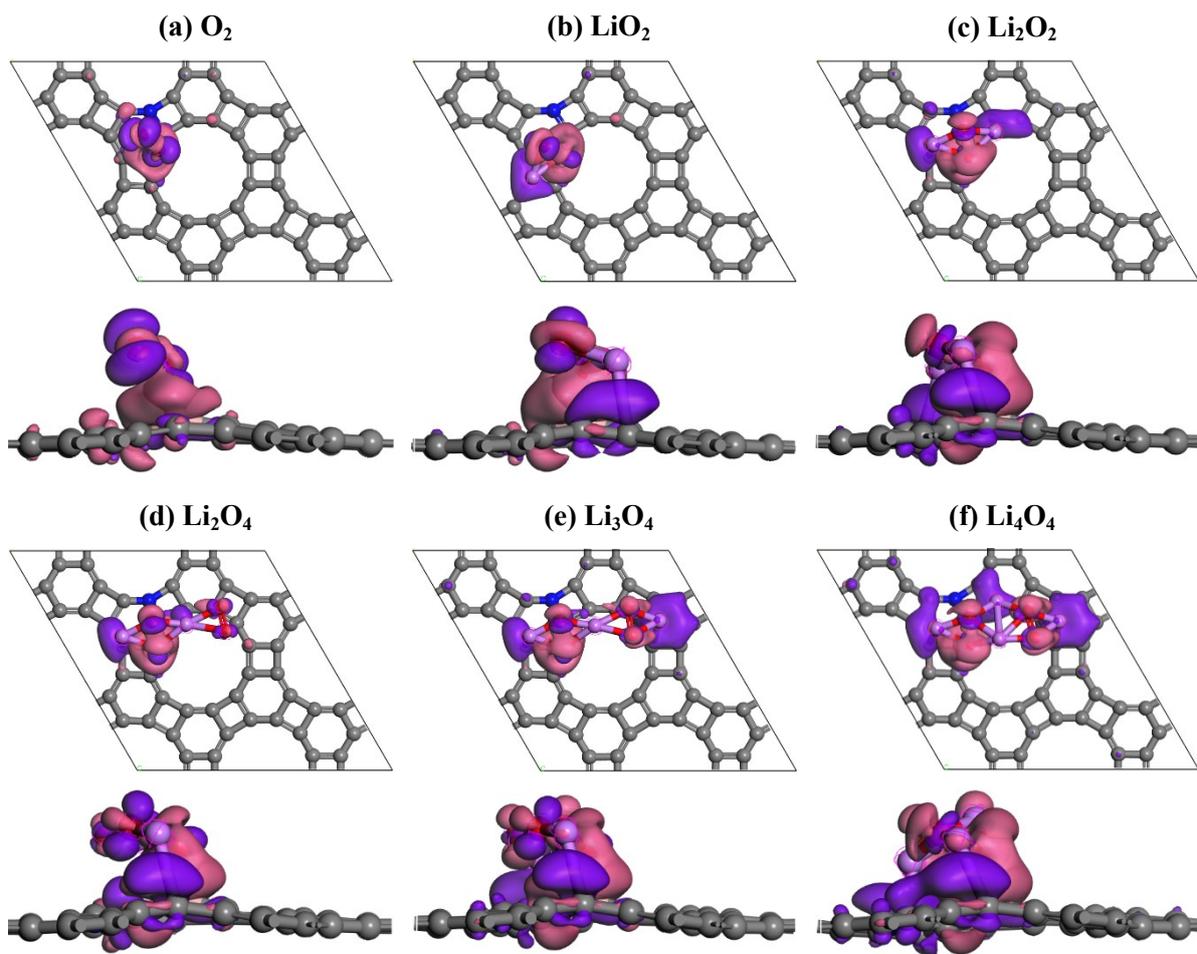


Fig S7. (continued).

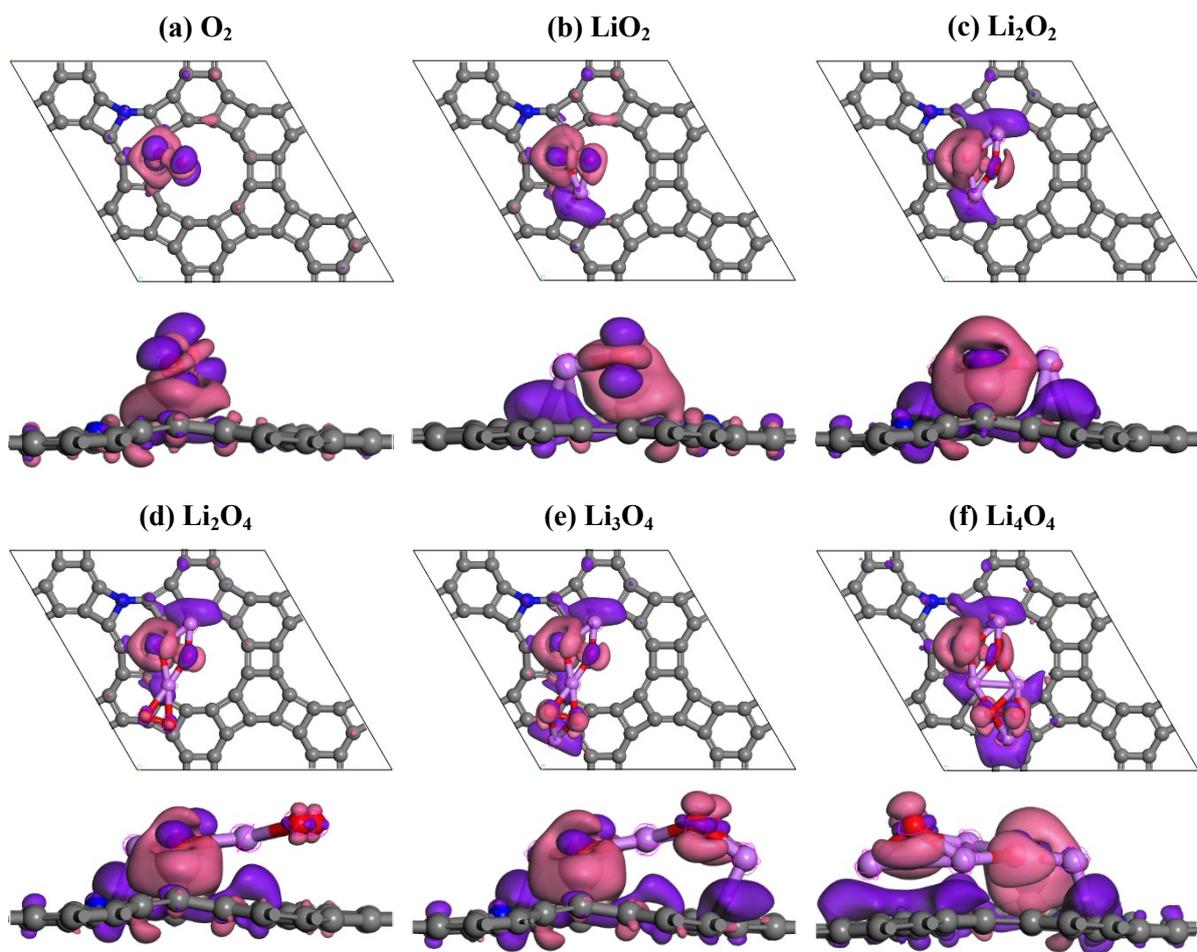


Fig S7. (continued).

Table S1. Optimized lattice constants (Å) of pristine, B-doped, N-doped, and BN co-doped structures.

	Pristine GP	BGP	NGP	B-BNGP	S ₁ -BNGP	S ₂ -BNGP
a	6.758	6.778	6.743	6.768	6.768	6.768
b	6.758	6.778	6.743	6.768	6.768	6.768

Table S2. Bond lengths of C-C, B-C, N-C, and B-N in pristine, B-doped, N-doped, and BN co-doped structures.

System	B1 (Å)	B2 (Å)	B3 (Å)
Pristine GP	1.366 (C-C)	1.471 (C-C)	1.470 (C-C)
BGP	1.468 (B-C)	1.525 (B-C)	1.549 (B-C)
NGP	1.344 (N-C)	1.452 (N-C)	1.459 (N-C)
B-BNGP	1.401 (B-N)	1.542 (B-C) 1.416 (N-C)	1.548 (B-C) 1.460 (N-C)
S ₁ -BNGP	1.460 (B-C) 1.343 (N-C)	1.532 (B-C) 1.458 (N-C)	1.542 (B-C) 1.458 (N-C)
S ₂ -BNGP	1.463 (B-C) 1.350 (N-C)	1.530 (B-C) 1.447 (N-C)	1.553 (B-C) 1.465 (N-C)

Table S3. Adsorption of Li at various sites and distances in different GP-based systems: initial position (IP), initial distance (ID), final position (FP), final distance (FD), and adsorption energy (E_{ads}).

Substrate	Adsorbate	IP	ID (Å)	FP	FD (Å)	E_{ads} (eV)
Pristine GP	Li	H2	0.000	H2	0.003	-0.426
		H2	0.500	H2	1.794	-0.807
		H2	1.000	H2	1.787	-0.807
		H2	1.500	H2	1.802	-0.807
		H2	2.000	H2	1.805	-0.808
		H2	2.500	H2	1.807	-0.806
		H1	1.805	H1	0.099	-0.333
		H3	1.805	H3	1.976	-0.547
		B1	1.805	H2	1.824	-0.807
		B2	1.805	H3	1.971	-0.547
		B3	1.805	H2	1.809	-0.806
		T	1.805	H2	1.781	-0.806
BGP	Li	H1	1.805	H1	0.400	-0.961
		H2	1.805	H2	1.837	-1.291
		H3	1.805	H2	1.669	-1.290
		B1	1.805	H2	1.651	-1.291
		B2	1.805	H2	1.660	-1.291
		B3	1.805	H2	1.657	-1.291
		T1 (Top C)	1.805	H2	1.647	-1.292
		T2 (Top B)	1.805	H2	1.664	-1.291
NGP	Li	H1	1.805	H1	0.281	-0.206
		H2	1.805	H2	1.832	-0.352
		H3	1.805	H2	1.790	-0.726
		B1	1.805	H2	1.811	-0.352
		B2	1.805	H2	1.784	-0.725
		B3	1.805	H2	1.806	-0.725
		T1 (Top C)	1.805	H3	1.946	-0.176
T2 (Top N)	1.805	H3	1.893	-0.181		
B-BNGP	Li	H1	1.805	H1	0.178	-0.501
		H2	1.805	H2	1.785	-0.746
		H3	1.805	H2	1.692	-0.847
		B1	1.805	H2	1.730	-0.749
		B2	1.805	H2	1.696	-0.848
		B3	1.805	H2	1.701	-0.847
		T1 (Top C)	1.805	H2	1.724	-0.749
		T2 (Top B)	1.805	H2	1.694	-0.847
T3 (Top N)	1.805	H2	1.723	-0.749		
S ₁ -BNGP	Li	H1	1.805	H1	0.419	-0.749
		H2	1.805	H2	1.777	-0.809
		H3	1.805	H3	1.798	-0.732
		B1	1.805	H2	1.757	-0.815
		B2	1.805	H3	1.810	-0.733
		B3	1.805	H3	1.793	-0.735
		T1 (Top C)	1.805	H2	1.768	-0.815
		T2 (Top B)	1.805	H3	1.727	-0.735
T3 (Top N)	1.805	H2	1.767	-0.816		
S ₂ -BNGP	Li	H1	1.805	H1	0.338	-0.706
		H2	1.805	H2	1.773	-0.718
		H3	1.805	H3	1.820	-0.655
		B1	1.805	H2	1.784	-0.723
		B2	1.805	H3	1.830	-0.656
		B3	1.805	H3	1.811	-0.656
		T1 (Top C)	1.805	H3	1.987	-0.590
		T2 (Top B)	1.805	H3	1.732	-0.653
T3 (Top N)	1.805	H2	1.768	-0.724		

Table S4. Adsorption of O₂ at various sites, distances, and orientations in different GP-based systems: initial position (IP), initial orientation (IO), initial distance (ID), final position (FP), final orientation (FO), final distance (FD), and adsorption energy (E_{ads}).

Substrate	Adsorbate	IP	IO ^a	ID (Å)	FP	FO	FD (Å)	E _{ads} (eV)
Pristine GP	O ₂	H2	H	1.500	H2	H	2.815	-0.345
		H2	V	1.500	H2	V	3.419	-0.270
		H2	H	2.000	H2	H	2.809	-0.341
		H2	V	2.000	H2	V	3.439	-0.271
		H2	H	2.500	H2	H	2.806	-0.342
		H2	V	2.500	H2	V	3.432	-0.271
		H2	H	3.000	H2	H	2.793	-0.341
		H2	V	3.000	H2	V	3.425	-0.271
		H1	H	2.815	H1	H	2.281	-0.348
		H1	V	2.815	H1	V	2.239	-0.347
		H3	H	2.815	H3	H	2.797	-0.359
		H3	V	2.815	H3	V	3.442	-0.274
		B1	H	2.815	B1	H	2.939	-0.315
		B1	V	2.815	B1	O	3.299	-0.274
		B2	H	2.815	B2	H	2.876	-0.329
		B2	V	2.815	B2	O	3.399	-0.276
		B3	H	2.815	B3	H	2.810	-0.350
		B3	V	2.815	B3	H	2.856	-0.339
T	H	2.815	B3	H	2.776	-0.357		
T	V	2.815	T	O	3.309	-0.275		
BGP	O ₂	H1	H	2.815	Top B	O	2.343	-0.859
		H1	V	2.815	H1	O	2.387	-0.281
		H2	H	2.815	Top B	O	2.569	-0.855
		H2	V	2.815	Top B	O	2.362	-0.859
		H3	H	2.815	Top B	O	2.565	-0.855
		H3	V	2.815	Top B	O	2.567	-0.854
		B1	H	2.815	Top B	O	2.568	-0.854
		B1	V	2.815	Top B	O	2.364	-0.860
		B2	H	2.815	Top B	O	2.352	-0.859
		B2	V	2.815	Top B	O	2.349	-0.859
		B3	H	2.815	Top B	O	2.563	-0.856
		B3	V	2.815	Top B	O	2.567	-0.856
		T1 (Top C)	H	2.815	Top B	O	2.567	-0.855
		T1 (Top C)	V	2.815	Top B	O	2.573	-0.855
T2 (Top B)	H	2.815	Top B	O	2.564	-0.856		
T2 (Top B)	V	2.815	Top B	O	2.564	-0.855		
NGP	O ₂	H1	H	2.815	H1	O	2.031	-0.424
		H1	V	2.815	H1	V	2.212	-0.410
		H2	H	2.815	H3	O	2.626	-0.464
		H2	V	2.815	H2	O	3.413	-0.312
		H3	H	2.815	H3	O	2.615	-0.464
		H3	V	2.815	B1	O	2.532	-0.512
		B1	H	2.815	B1	O	2.530	-0.512
		B1	V	2.815	B1	O	3.311	-0.336
		B2	H	2.815	B2	O	2.664	-0.430
		B2	V	2.815	B1	O	2.513	-0.512
		B3	H	2.815	H3	O	2.662	-0.463
		B3	V	2.815	B1	O	2.518	-0.513
		T1 (Top C)	H	2.815	Top C	O	2.925	-0.362
		T1 (Top C)	V	2.815	Top C	O	3.230	-0.328
T2 (Top N)	H	2.815	B1	O	2.538	-0.512		
T2 (Top N)	V	2.815	Top N	O	3.373	-0.328		
B-BNGP	O ₂	H1	H	2.815	H1	O	2.139	-0.355
		H1	V	2.815	H1	V	2.288	-0.330

		H2	H	2.815	B1	O	2.469	-0.750
		H2	V	2.815	H2	O	3.359	-0.261
		H3	H	2.815	H3	O	2.634	-0.347
		H3	V	2.815	H3	O	2.208	-0.269
		B1	H	2.815	B1	O	2.471	-0.750
		B1	V	2.815	B1	O	3.315	-0.271
		B2	H	2.815	B2	O	2.512	-0.668
		B2	V	2.815	H1	O	2.238	-0.751
		B3	H	2.815	Top B	O	2.465	-0.746
		B3	V	2.815	Top B	O	2.474	-0.750
		T1 (Top C)	H	2.815	B2	O	2.791	-0.336
		T1 (Top C)	V	2.815	B3	O	3.263	-0.267
		T2 (Top B)	H	2.815	Top B	O	2.472	-0.751
		T2 (Top B)	V	2.815	Top B	O	2.231	-0.752
		T3 (Top N)	H	2.815	Top B	H	2.235	-0.751
		T3 (Top N)	V	2.815	Top N	O	3.382	-0.261
		H1	H	2.815	Top B	O	2.219	-0.825
		H1	V	2.815	H1	O	2.253	-0.349
		H2	H	2.815	Top B	O	2.411	-0.837
		H2	V	2.815	Top B	O	2.232	-0.825
		H3	H	2.815	Top B	O	2.408	-0.836
		H3	V	2.815	Top B	O	2.410	-0.836
		B1	H	2.815	Top B	O	2.248	-0.826
		B1	V	2.815	Top B	O	2.412	-0.836
		B2	H	2.815	Top B	O	2.246	-0.826
		B2	V	2.815	Top B	O	2.234	-0.825
		B3	H	2.815	Top B	O	2.420	-0.836
		B3	V	2.815	Top B	O	2.404	-0.836
		T1 (Top C)	H	2.815	Top C	O	2.621	-0.476
		T1 (Top C)	V	2.815	Top B	O	2.216	-0.825
		T2 (Top B)	H	2.815	Top B	O	2.423	-0.836
		T2 (Top B)	V	2.815	Top B	O	2.410	-0.836
		T3 (Top N)	H	2.815	Top B	O	2.409	-0.836
		T3 (Top N)	V	2.815	Top N	O	3.359	-0.305
		H1	H	2.815	Top B	O	2.253	-0.821
		H1	V	2.815	H1	O	2.315	-0.343
		H2	H	2.815	Top B	O	2.474	-0.781
		H2	V	2.815	Top B	O	2.253	-0.821
		H3	H	2.815	Top B	O	2.473	-0.785
		H3	V	2.815	Top B	O	2.411	-0.705
		B1	H	2.815	Top B	O	2.469	-0.785
		B1	V	2.815	Top B	O	2.481	-0.785
		B2	H	2.815	Top B	O	2.253	-0.822
		B2	V	2.815	Top B	O	2.256	-0.821
		B3	H	2.815	Top B	O	2.472	-0.784
		B3	V	2.815	Top B	O	2.475	-0.785
		T1 (Top C)	H	2.815	Top C	O	2.624	-0.444
		T1 (Top C)	V	2.815	Top B	O	2.479	-0.785
		T2 (Top B)	H	2.815	Top B	O	2.466	-0.783
		T2 (Top B)	V	2.815	Top B	O	2.469	-0.785
		T3 (Top N)	H	2.815	B1	O	2.611	-0.400
		T3 (Top N)	V	2.815	Top N	O	3.431	-0.288

^a H, V, and O represent horizontal, vertical, and oblique orientations, respectively.

Table S5. Calculated distances between Li_xO_{2y} intermediates and C atoms (\AA) on pristine, B-doped, N-doped, and BN co-doped structures.

Li_xO_{2y}	Pristine GP	BGP	NGP	B-BNGP	S ₁ -BNGP	S ₂ -BNGP
Li	1.805	1.647	1.790	1.696	1.767	1.768
O ₂	2.797	2.364	2.518	2.231	2.411	2.253
LiO ₂	1.939	1.876	1.926	1.774	2.095	1.909
Li ₂ O ₂	1.992	1.887	2.035	1.807	2.075	1.934
Li ₃ O ₂	1.542	1.551	2.260	1.294	1.921	1.507
Li ₄ O ₂	1.542	1.402	1.541	1.282	2.112	2.003
Li ₂ O ₄	1.846	1.892	1.976	1.842	2.128	1.982
Li ₃ O ₄ (i)	1.967	1.529	1.965	1.859	1.849	1.899
Li ₃ O ₄ (ii)	1.979	1.951	2.029	1.440	1.971	1.545
Li ₄ O ₄ (i)	2.194	1.482	2.039	1.987	1.976	1.883
Li ₄ O ₄ (ii)	1.933	1.558	1.715	1.246	2.143	1.957

Table S6. Adsorption energies (E_{ads} , eV) of Li_xO_{2y} intermediates on pristine, B-doped, N-doped, and BN co-doped structures.

Li_xO_{2y}	Pristine GP	BGP	NGP	B-BNGP	S ₁ -BNGP	S ₂ -BNGP
Li	-0.808	-1.292	-0.726	-0.848	-0.816	-0.724
O ₂	-0.359	-0.860	-0.513	-0.752	-0.837	-0.822
LiO ₂	-2.358	-2.627	-2.930	-2.392	-2.738	-2.400
Li ₂ O ₂	-4.691	-5.018	-4.537	-4.956	-5.222	-5.417
Li ₃ O ₂	-6.031	-6.462	-5.830	-6.028	-6.254	-6.497
Li ₄ O ₂	-7.222	-8.123	-6.851	-7.396	-9.423	-9.788
Li ₂ O ₄	-6.114	-6.501	-6.038	-6.148	-6.294	-6.548
Li ₃ O ₄ (i)	-8.567	-8.727	-8.464	-7.872	-8.457	-8.219
Li ₃ O ₄ (ii)	-8.125	-8.127	-7.542	-7.131	-8.385	-7.830
Li ₄ O ₄ (i)	-10.390	-11.316	-10.332	-10.492	-10.848	-11.018
Li ₄ O ₄ (ii)	-9.365	-9.569	-9.359	-9.372	-10.727	-10.971

Table S7. Variation of the Hirshfeld charge transferred for Li_xO_y intermediates formed on pristine, B-doped, N-doped, and BN co-doped structures during the reduction process. Negative and positive values of the transferred charge indicate an increase and decrease in the number of electrons, respectively.

Structure	Li_xO_y	Transferred charge ($ e $) for the adsorbed O and Li atoms									
		Li ₁	Li ₂	Li ₃	Li ₄	O ₁	O ₂	O ₃	O ₄	Total	
Pristine GP	LiO ₂	0.278					-0.141	-0.141			-0.004
	Li ₂ O ₂	0.283	0.283				-0.238	-0.237			0.091
	Li ₃ O ₂	0.278	0.279	0.357			-0.314	-0.287			0.313
	Li ₄ O ₂	0.270	0.270	0.358	0.355		-0.327	-0.326			0.600
	Li ₂ O ₄	0.168	0.290				-0.180	-0.180	-0.049	-0.049	0.000
	Li ₃ O ₄ (i)	0.286	0.259	0.260			-0.203	-0.203	-0.204	-0.203	-0.008
	Li ₄ O ₄ (i)	0.350	0.263	0.264	0.165		-0.255	-0.258	-0.253	-0.253	0.023
BGP	LiO ₂	0.267					-0.128	-0.112			0.027
	Li ₂ O ₂	0.265	0.277				-0.223	-0.208			0.111
	Li ₃ O ₂	0.264	0.272	0.345			-0.297	-0.265			0.319
	Li ₄ O ₂	0.252	0.266	0.349	0.343		-0.321	-0.314			0.575
	Li ₂ O ₄	0.275	0.166				-0.164	-0.158	-0.039	-0.040	0.040
	Li ₃ O ₄ (i)	0.357	0.242	0.266			-0.309	-0.145	-0.198	-0.196	0.017
	Li ₄ O ₄ (i)	0.371	0.323	0.255	0.237		-0.301	-0.139	-0.258	-0.263	0.225
NGP	LiO ₂	0.286					-0.158	-0.152			-0.024
	Li ₂ O ₂	0.289	0.305				-0.260	-0.246			0.088
	Li ₃ O ₂	0.292	0.304	0.372			-0.322	-0.344			0.302
	Li ₄ O ₂	0.306	0.296	0.346	0.344		-0.317	-0.336			0.639
	Li ₂ O ₄	0.285	0.203				-0.200	-0.199	-0.064	-0.070	-0.045
	Li ₃ O ₄ (i)	0.270	0.257	0.266			-0.214	-0.212	-0.211	-0.212	-0.056
	Li ₄ O ₄ (i)	0.264	0.347	0.263	0.166		-0.260	-0.260	-0.257	-0.259	0.004
B-BNGP	LiO ₂	0.380					-0.260	-0.094			0.026
	Li ₂ O ₂	0.376	0.348				-0.345	-0.158			0.221
	Li ₃ O ₂	0.402	0.350	0.328			-0.326	-0.153			0.601
	Li ₄ O ₂	0.281	0.261	0.318	0.423		-0.637	-0.411			0.235
	Li ₂ O ₄	0.386	0.233				-0.326	-0.149	-0.102	-0.107	-0.065
	Li ₃ O ₄ (i)	0.378	0.260	0.272			-0.330	-0.148	-0.204	-0.209	0.019
	Li ₄ O ₄ (i)	0.371	0.321	0.269	0.247		-0.307	-0.137	-0.269	-0.270	0.225
S ₁ -BNGP	LiO ₂	0.277					-0.133	-0.147			-0.003
	Li ₂ O ₂	0.284	0.287				-0.225	-0.254			0.092
	Li ₃ O ₂	0.269	0.279	0.355			-0.307	-0.294			0.302
	Li ₄ O ₂	0.260	0.276	0.363	0.347		-0.318	-0.315			0.613
	Li ₂ O ₄	0.286	0.173				-0.192	-0.202	-0.080	-0.079	-0.094
	Li ₃ O ₄ (i)	0.267	0.262	0.263			-0.200	-0.210	-0.207	-0.204	-0.029
	Li ₄ O ₄ (i)	0.261	0.349	0.262	0.167		-0.250	-0.255	-0.252	-0.253	0.029
S ₂ -BNGP	LiO ₂	0.355					-0.098	-0.264			-0.007
	Li ₂ O ₂	0.363	0.347				-0.157	-0.336			0.217
	Li ₃ O ₂	0.366	0.335	0.402			-0.167	-0.332			0.604
	Li ₄ O ₂	0.280	0.258	0.400	0.319		-0.410	-0.633			0.214
	Li ₂ O ₄	0.226	0.351				-0.139	-0.315	-0.095	-0.094	-0.066
	Li ₃ O ₄ (i)	0.214	0.346	0.330			-0.136	-0.309	-0.209	-0.205	0.031
	Li ₄ O ₄ (i)	0.233	0.357	0.260	0.284		-0.134	-0.295	-0.267	-0.272	0.166

Table S8. Calculated discharge, equilibrium, and charge potentials ($U_{DC}/U_{eq}/U_C$) (in V), and ORR, OER, and total overpotentials ($\eta_{ORR}/\eta_{OER}/\eta_{TOT}$) (in V) for pristine, B-doped, N-doped, and BN co-doped structures.

Structure	U_{DC}	U_{eq}	U_C	η_{ORR}	η_{OER}	η_{TOT}
Pristine GP	2.163	2.563	3.526	0.400	0.963	1.363
BGP	2.364	2.790	3.720	0.426	0.930	1.356
NGP	2.096	2.550	3.324	0.454	0.774	1.228
B-BNGP	2.304	2.584	2.877	0.280	0.293	0.574
S ₁ -BNGP	2.255	2.680	3.338	0.425	0.658	1.083
S ₂ -BNGP	2.432	2.718	3.013	0.286	0.295	0.581

To evaluate the effect of the O₂ energy correction on the calculated thermodynamic quantities, the formation free energies (ΔG_f) of discharge species and the equilibrium potential (U_{eq}) for one representative structure in this study (B-BNGP) were compared with and without applying the O₂ correction ($E_{CV} = 0.86$ eV) (Table S9). Applying this correction results in a uniform shift in the Gibbs free energies of formation and makes the ΔG_f values of Li-O₂ species more negative, indicating an improved thermodynamic description of these systems within the DFT framework. Furthermore, the calculated U_{eq} values obtained by including this correction become closer to the experimentally reported value for Li-O₂ batteries (approximately 2.96 V). These results demonstrate that although the O₂ correction improves the absolute accuracy of the thermodynamic quantities, the main trends and key conclusions of this study remain robust and reliable with respect to this correction. Similar behavior was also observed for the other investigated structures.

Table S9. Effect of O₂ energy correction on ΔG_f (eV) and U_{eq} (V) for the representative B-BNGP structure.

	Without O ₂ correction	With O ₂ correction
ΔG_f (Li ₂ O ₂)	-3.981	-4.841
ΔG_f (Li ₄ O ₄)	-8.615	-10.335
U_{eq}	2.154	2.584