

Supplementary Information for:

**Labilized Ligands Under Ultra-Confinement: A Solid-State ^2H NMR
Investigation of $\text{Ln}(\text{BH}_4)_2(\text{THF})_2/\text{HY}_{30}$**

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Table of Contents

1. Experimental Details	3
1.1. Synthesis	3
1.1.1. Sc(BH ₄) ₃ (THF) ₂	3
1.1.2. Ln(BH ₄) ₂ (THF) ₂ /Y ₃₀	3
1.2. Solid-State ² H NMR Spectroscopy	3
1.3. Density Functional Theory	3
1.3.1. Cluster Model Calculations	3
1.3.2. Plane Wave Calculations	4
2. Density Functional Theory	4
2.1. THF Rotation	4
2.2. THF Dissociation	6
2.3. Intramolecular THF Exchange	6
3. Lineshape Simulations	7
4. Supplementary Figures	10
5. Geometry-Optimized Coordinates	11
5.1. Sc(BH ₄) ₂ (THF) ₂ /Y ₃₀	11
5.2. Y(BH ₄) ₂ (THF) ₂ /Y ₃₀	15
5.3. La(BH ₄) ₂ (THF) ₂ /Y ₃₀	18
6. Supplementary References	22

1. Experimental Details

1.1. Synthesis

1.1.1. $\text{Sc}(\text{BH}_4)_3(\text{THF})_2$

$\text{Sc}(\text{BH}_4)_3(\text{THF})_2$ was synthesized using literature methods.^{1,2} Briefly, ScCl_3 (1.516 g, 10.01 mmol) and excess sodium borohydride (1.515 g, 40.05 mmol) were heated at reflux in dry THF (100 mL) for 5 days. The solution was filtered through a cannula wrapped with filter paper to remove the residual NaBH_4 and NaCl . The filtrate was concentrated to saturated solution by evaporating under reduced pressure and then cooled at $-30\text{ }^\circ\text{C}$ for 2 days, providing 1.905 g (81% yield) $\text{Sc}(\text{BH}_4)_3(\text{THF})_3$. The colorless powders were stored at $-30\text{ }^\circ\text{C}$ freezer in a N_2 glovebox. $\text{Y}(\text{BH}_4)_3(\text{THF})_3$ and $\text{La}(\text{BH}_4)_3(\text{THF})_3$ complexes were prepared via the same protocol. To deuterate the compound, 0.100 g of $\text{Sc}(\text{BH}_4)_3(\text{THF})_2$ was stirred with 0.2 mL $\text{THF-}d_8$ at room temperature for 30 h and then was evacuated under vacuum (10^{-3} mbar) for 12 h.

1.1.2. $\text{Ln}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$

0.020 g of $\text{Sc}(\text{BH}_4)_3(\text{THF})_3$ (0.075 mmol) was dissolved in toluene (10 mL) and added to 0.300 g of zeolite HY_{30} that had been capped by Ph_3SiCl (namely $\text{Ph}_3\text{Si-HY}_{30}$).^{2,3} The mixtures were stirred at room temperature for 20 h. The supernatant was decanted and the solids were washed with toluene (3×10 mL). The obtained solids were then dried under vacuum (10^{-3} mbar) at room temperature for 24 h, giving $\text{Sc}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$. The material was subsequently deuterated by taking 0.100 g $\text{Sc}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$ and stirring it with 0.2 mL $\text{THF-}d_8$ at room temperature for 30 h. The material was then evacuated under vacuum (10^{-3} mbar) for 12 h. The procedure was repeated to prepare the Y and La analogues. Metal loadings were measured using inductively-coupled plasma optical emission spectroscopy (ICP-OES) and equaled 0.16, 0.069, and 0.076 mmol/g for Sc, Y, and La, respectively.

1.2. Solid-State ^2H NMR Spectroscopy

All solid-state NMR experiments were performed using an Agilent DD2 400 MHz NMR spectrometer equipped with a Chemagnetics 5 mm magic-angle spinning (MAS) probe operated as a static probe. Samples were flame-sealed into 5 mm NMR tubes to prevent their degradation and inserted directly into the coil of the probe. Samples were cooled using a nitrogen heat exchanger in a temperature range of 20 to $-120\text{ }^\circ\text{C}$. The ^2H NMR spectra were acquired using a quadrupolar echo sequence with 2 μs excitation and refocusing pulses and a 30 μs delay. The number of scans was varied for each sample and temperature to obtain a reasonable signal-to-noise ratio and varied between 1024 and 10240 scans. The recycle delay was set to 2 s.

1.3. Density Functional Theory

1.3.1. Cluster Model Calculations

Density functional theory (DFT) calculations were performed on a large cluster model of the micropore in zeolite Y consisting of 37 interconnected silicon tetrahedral sites that were

hydride-terminated. The desired $\text{Ln}(\text{BH}_4)_2(\text{THF})_2$ was coordinated to an oxide site bonded to an aluminum tetrahedral site, to mirror the experimental chemical environment as closely as possible. Geometry optimizations froze the positions of the silicon atoms to those from the crystal structure to prevent the pore from relaxing in response to the grafted complex, which could influence the computed steric effects. The geometry-optimized structure for the La analogue is depicted in **Figure S1** and the coordinates are also given for the three supported complexes below. Calculations were all performed at the PBE-D3/TZP^{4,5} level of theory using the Amsterdam Density Functional (ADF) program.⁶

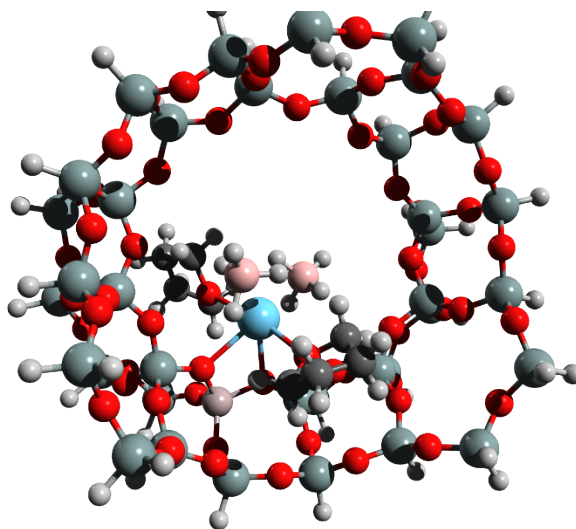


Figure S1. Geometry-optimized structure of $\text{La}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$.

1.3.2. Plane Wave Calculations

The activation energy barrier for intramolecular THF exchange was investigated using the nudged elastic band (NEB) method^{7,8} as implemented in the *neb.x* module of the Quantum ESPRESSO (v7.1).^{9,10} Calculations used the PBE exchange-correlation functional.⁴

Pseudopotentials were taken from the SSSP library.¹¹ Kinetic energy cutoffs for wavefunctions and charge density were set to 544 eV and 2177 eV, respectively. Brillouin zone integration was performed using Gamma-point sampling due to the large size of the simulation cell. To determine the energy barrier, the climbing-image NEB (CI-NEB) scheme^{9,10} was used with 11 images to identify the saddle point. Calculations were considered converged when the forces on the climbing image fell below 0.05 eV/Å.

2. Density Functional Theory

2.1. THF Rotation

We performed potential energy surface scans as a function of the O-Ln-O-C dihedral angle in a Ln-bonded THF to rotate the ligand around its metal-oxygen bond. The scan was performed

in 10° increments and yielded rotation barriers of 17.4, 17.5, and 20.6 kJ/mol for the La, Y, and Sc complexes, respectively. The full list of relative energies are listed below in **Table S1**.

Table S1. Relative bonding energy as a function of THF rotation.

$\varphi_{\text{O-Ln-O-C}} / ^\circ$	Sc $\Delta E / \text{kJ mol}^{-1}$	Y $\Delta E / \text{kJ mol}^{-1}$	La $\Delta E / \text{kJ mol}^{-1}$
0	7.358016	2.163648	2.926199
10	9.178669	3.518879	3.493018
20	10.52416	4.742126	4.001288
30	11.39735	5.882406	4.783766
40	13.03558	8.063173	6.374793
50	15.42216	10.33693	8.277545
60	18.23869	13.78096	11.58764
70	22.12795	17.59972	15.39079
80	20.88793	15.62939	17.32951
90	18.93954	14.35673	12.66518
100	18.28894	13.49179	11.65638
110	17.18983	12.51395	11.40864
120	16.20217	11.87319	10.90417
130	15.51203	11.45443	10.62133
140	15.12086	10.23895	9.342868
150	12.83021	8.735616	8.079215
160	9.545189	6.978500	6.706525
170	8.736955	4.874241	4.931135
180	7.095729	2.109432	2.833807
190	5.688671	0.260712	1.583728
200	5.167010	0.000000	1.861086
210	5.832286	1.667508	3.660131
220	8.155538	2.806371	4.914831
230	12.83835	6.617284	8.614055
240	10.92686	12.39139	13.28603
250	17.89549	17.47121	17.38837
260	18.97982	16.55504	14.36889
270	14.86133	14.80315	13.28692
280	12.00751	12.78800	12.31753
290	10.70296	11.56031	11.27587
300	10.46902	10.47325	10.51426
310	8.401626	7.415305	5.155721
320	5.804088	3.769142	1.267408
330	2.408686	0.830708	0.000000
340	0.159630	0.135423	0.414750
350	0.000000	0.482409	1.346540

2.2. THF Dissociation

To estimate the free energy change associated with THF dissociation, geometry optimizations were performed on both $\text{Ln}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$ and $\text{Ln}(\text{BH}_4)_2(\text{THF})/\text{Y}_{30}$ ($\text{Ln} = \text{Sc}, \text{Y},$ or La). Given that all silicon atoms' positions needed to be frozen, we could not estimate the free energy of the solid species because vibrational framework contributions to entropy have been neglected, however, we did estimate the entropy contributions from the THF translational and rotational degrees of freedom for the released THF as these are likely important. The translational entropy contribution was calculated as 0.1621 kJ/mol K, and the rotational entropy contribution as 0.1043 kJ/mol K, using the ideal gas rigid-rotor approximation. The bonding energies are listed in **Table S2** for the six complexes. Dissociation enthalpies equal 118, 111, and 87 kJ/mol for the La, Y, and Sc analogues, respectively, which are reduced to 40.3, 32.6, and 8.7 kJ/mol at 20 °C when considering the new degrees of freedom of the free THF.

Table S2. Bonding energies of $\text{Ln}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$ and $\text{Ln}(\text{BH}_4)_2(\text{THF})/\text{Y}_{30}$

Model	$E / \text{kJ mol}^{-1}$
$\text{La}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$	-99845.35406
$\text{Y}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$	-99785.62787
$\text{Sc}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$	-99713.09967
$\text{La}(\text{BH}_4)_2(\text{THF})/\text{Y}_{30}$	-92802.43295
$\text{Y}(\text{BH}_4)_2(\text{THF})/\text{Y}_{30}$	-92750.38491
$\text{Sc}(\text{BH}_4)_2(\text{THF})/\text{Y}_{30}$	-92701.73143
THF	-6924.567608

2.3. Intramolecular THF Exchange

The energy barrier for the intramolecular THF exchange in $\text{La}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$ was estimated using a minimum energy path (MEP) search via the climbing-image nudged elastic band (CI-NEB) method. The energy profile reaches its maximum when the two THF ligands pass each other in a sterically constrained configuration (**Figure S2**), resulting in a calculated energy barrier of approximately 150 kJ/mol. The conformational rearrangement requires the two BH_4 ligands to move in synchronization with the THF ligands; however, this collective motion is severely hindered by steric repulsion with the zeolite walls, which significantly increases the overall energy of the transition state.

In configurations where steric hindrance becomes excessive, the system alleviates the energy strain through the dissociation of one THF ligand from the La center, which occurs well before the system reaches the highest energy intermediate. Along this part of the MEP, a pronounced elongation of one La–O is observed, reaching 3.0 Å, indicating that one of the THF ligand is already in a nearly dissociated state. While the current DFT-based MEP calculation focuses on electronic energy at 0 K, in a real-world scenario, the dissociation of a ligand provides a significant entropic gain. This additional entropic contribution lowers the free energy barrier for dissociation, making it even more favorable than the positional exchange. Consequently, it is

expected that the ligands will dissociate at an earlier stage than predicted by the static energy profile, effectively pre-empting the intramolecular THF exchange.

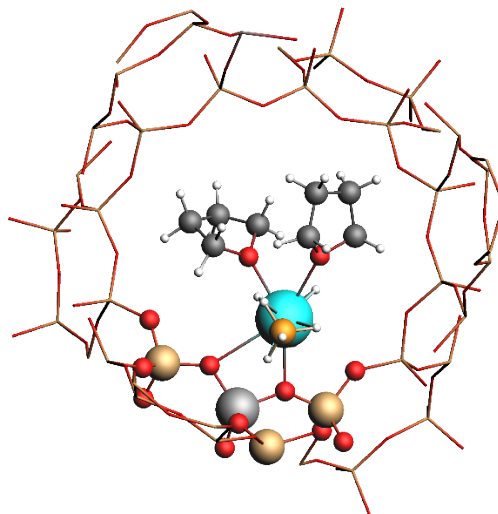


Figure. S2 Highest energy geometry of $\text{La}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$ as it performs intramolecular THF exchange.

3. Lineshape Simulations

All lineshape simulations were performed using an open-source program developed in our lab that is available in the following repository.¹² They assumed a quadrupolar coupling constant of 163 kHz and no axial asymmetry, consistent with the expected C_Q for C-D bonds in a cyclic ether environment. Polar angles were calculated from a given structure using the approach described by Bryce.¹³ Briefly, the polar angles for a given internuclear vector were calculated as:

$$\beta = \cos^{-1}\left(\frac{z}{\sqrt{x^2 + y^2 + z^2}}\right)$$

$$\gamma = \begin{cases} \cos^{-1}\left(\frac{x}{\sqrt{x^2 + y^2}}\right) & \text{if } y > 0 \\ 60^\circ - \cos^{-1}\left(\frac{x}{\sqrt{x^2 + y^2}}\right) & \text{if } y \leq 0 \end{cases}$$

Calculations were performed using a 3-angle powder average comprising a 678-angle REPULSION grid¹⁴ interleaved with a 15-step γ angle average. 256 time increments were considered with a spectral width of 400 kHz and 5 kHz of Lorentzian broadening. The full set of polar angles describing a single molecule are listed in **Table S3**. Additional tetrahedral averaging was performed to emulate the isotropic motion seen in the grafted species.

Table S3. Polar angles describing the rotation-related C-H bonds in $\text{Ln}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$

Site ^a	$\beta / ^\circ$	$\gamma / ^\circ$
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α1	94	101
α1	141	287
α2	26	289
α2	150	264
β1	61	135
β1	84	252
β2	43	99
β2	134	336

^a α and β refer to the respective positions of the two chemically distinct deuterons on THF. The numbers refer to ligands 1 and 2 on either side of the molecule.

Exchange rates were calculated on-the-fly for a given temperature and energy barrier using the Eyring relationship below. T refers to the temperature, k_B to the Boltzmann constant, h to the Planck constant, and ΔG^\ddagger to the free energy barrier.

$$k_{ex} = \frac{k_B T}{h} \exp\left(\frac{-\Delta G^\ddagger}{RT}\right)$$

Component-based fits were also performed for the $\text{Sc}(\text{BH}_4)_3(\text{THF})_3$ complex using the fast-motion limit lineshapes associated to libration, intramolecular exchange, and a combination of intramolecular exchange and THF rotation. THF libration was approximated with a lineshape having an averaged C_Q of 145 kHz and a 0.14 asymmetry parameter, which corresponds to a discrete case with a 30° libration angle. As the temperature was increased, the librating component was observed to narrow, suggesting an increase in librational amplitude. This was accounted for using a temperature-dependent order parameter ($\langle S \rangle$). The temperature dependent weights are listed in **Table S4** and are plotted in **Figure S3**.

Table S4. Weights used in the fit shown in Figure 3 for $\text{Sc}(\text{BH}_4)_3(\text{THF})_2$

T / K	static	librating	exchange	exchange and rotation	$\langle S \rangle_{\text{libration}}$
-120	56	39	6	0	1.00
-100	33	53	13	0	1.00
-90	15	62	23	0	1.00
-80	0	67	27	7	1.00
-70	0	61	30	7	0.95
-60	0	56	11	29	0.90
-50	0	52	0	45	0.90
-40	0	48	0	48	0.88
-30	0	44	0	52	0.85
-20	0	43	0	53	0.83
-10	0	36	0	60	0.75
0	0	31	0	65	0.72
10	0	18	0	82	0.72

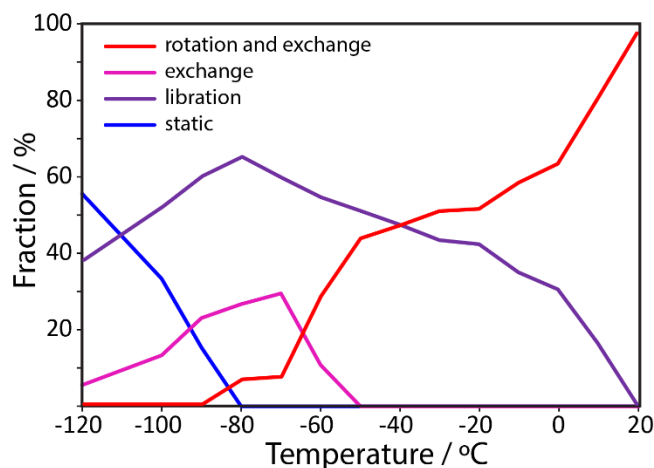


Figure S3. Temperature dependence of the weights used in the fit shown in Figure 3 for $\text{Sc}(\text{BH}_4)_3(\text{THF})_2$.

Simulations of the variable temperature NMR data for the HY_{30} -supported complexes undertook a time-domain approach as only this method can simulate the narrowing of the isotropic resonance from intermolecular exchange. Nevertheless, these simulations incorporate a component from a purely rotating THF group that is visible as a baseline feature in the spectra and thought to originate from silanol-grafted sites that are visible by ^{11}B NMR.

4. Supplementary Figures

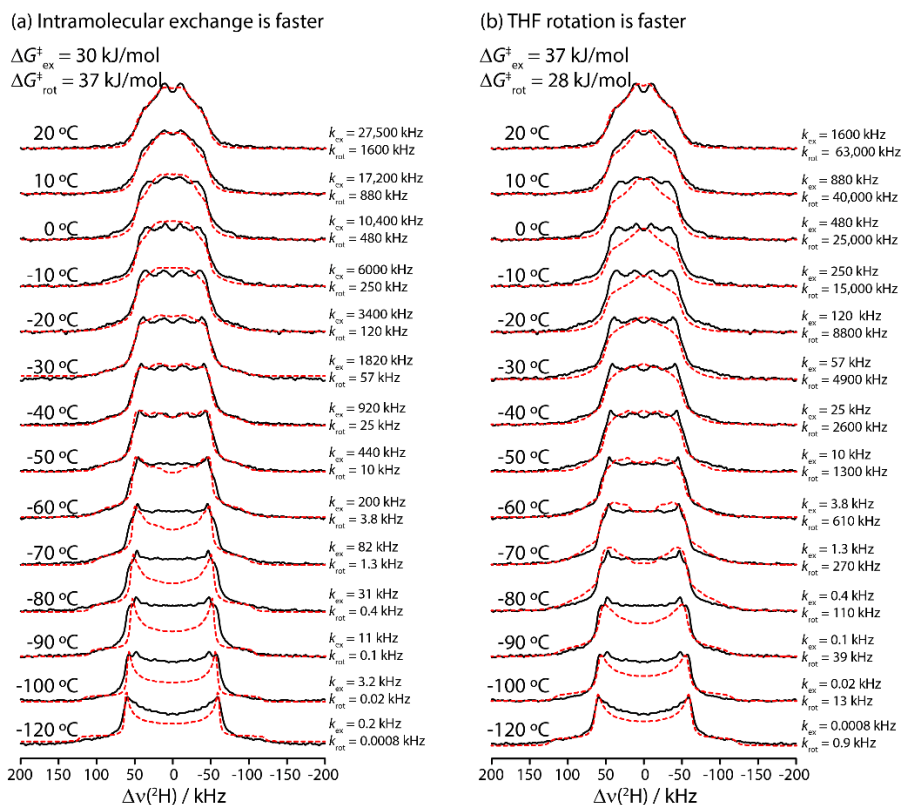


Figure S4. Variable temperature ^2H NMR spectra of $\text{Sc}(\text{BH}_4)_3(\text{THF})_2$ with time-domain simulations assuming constant ΔG_{\ddagger} and that either (a) intramolecular exchange or (b) THF rotation is faster. The intermediate lineshapes have a cusp in the center and do not match the experimental result in (b) suggesting that THF rotation has the higher barrier.

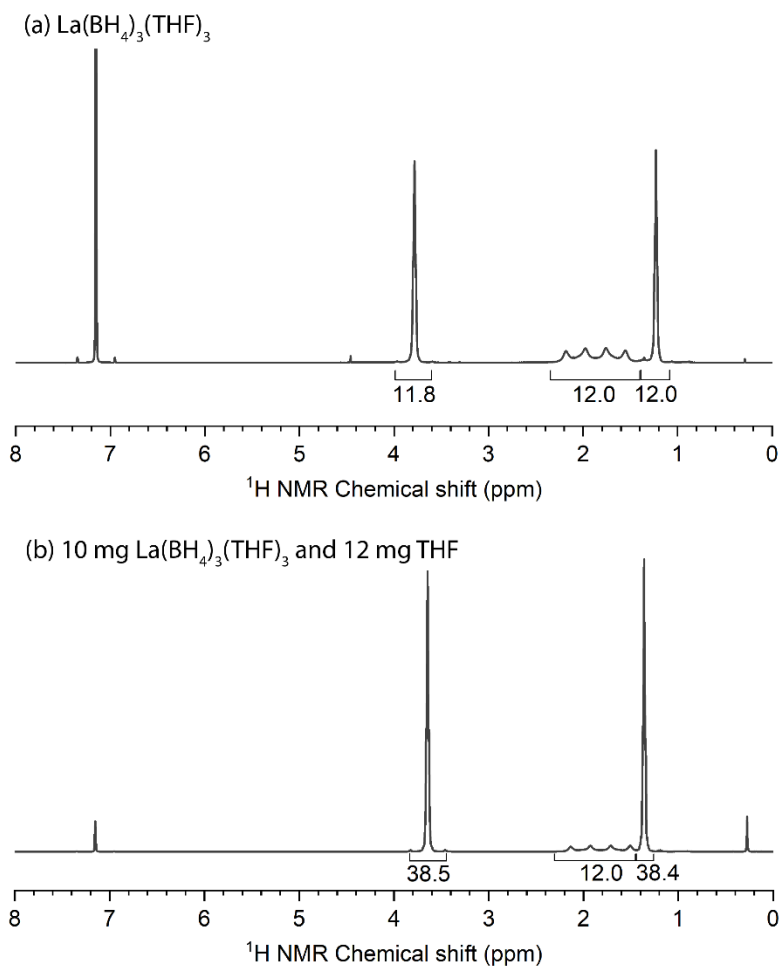


Figure S5. Liquid-state ^1H NMR spectra acquired for C_6D_6 solutions of $\text{La}(\text{BH}_4)_3(\text{THF})_3$ acquired without (a) and with (b) the addition of excess THF. As is evident from the spectrum and integrals, the excess THF is in rapid chemical exchange with the La-bound THF.

5. Geometry-Optimized Coordinates

5.1. $\text{Sc}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$

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Sc 18.82325582801248 7.882464776531349 6.368538960153203
O 19.75020945829827 6.400044498581743 8.270558188712583
O 17.32500447640999 6.18331184580245 6.646229724860877
B 17.79948924507838 8.373394752749633 4.376676039017444
B 18.14587056465124 9.471422717708549 7.876485094275379
O 20.75205317999711 9.014652040565661 6.114441744959668
H 19.22816674017682 8.988733031264287 8.213834997127426
H 17.37764440658761 8.504086148507819 7.747424933790826
H 17.03463332175281 8.377037189322321 5.356016793320886
H 18.69844346972966 9.165701918422 4.679270979972886
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C 15.49686960079504 4.796967902687604 5.956028079254798
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Si 16.08298507918421 5.274972788485146 12.19765473006501
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5.2. $Y(BH_4)_2(THF)_2/Y_{30}$

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5.3. $\text{La}(\text{BH}_4)_2(\text{THF})_2/\text{Y}_{30}$

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