Supplementary Information for

"Spectral and Redox Properties of Zinc Porphyrin Core Dendrimers with Triarylamines as Dendron"

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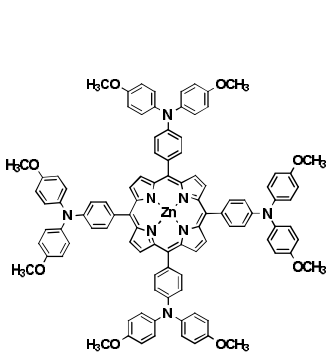

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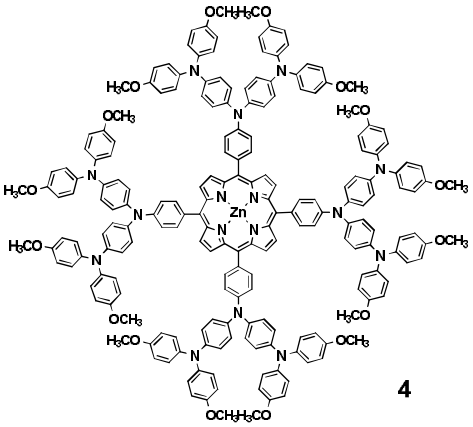

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A：^{1}H NMR spectra

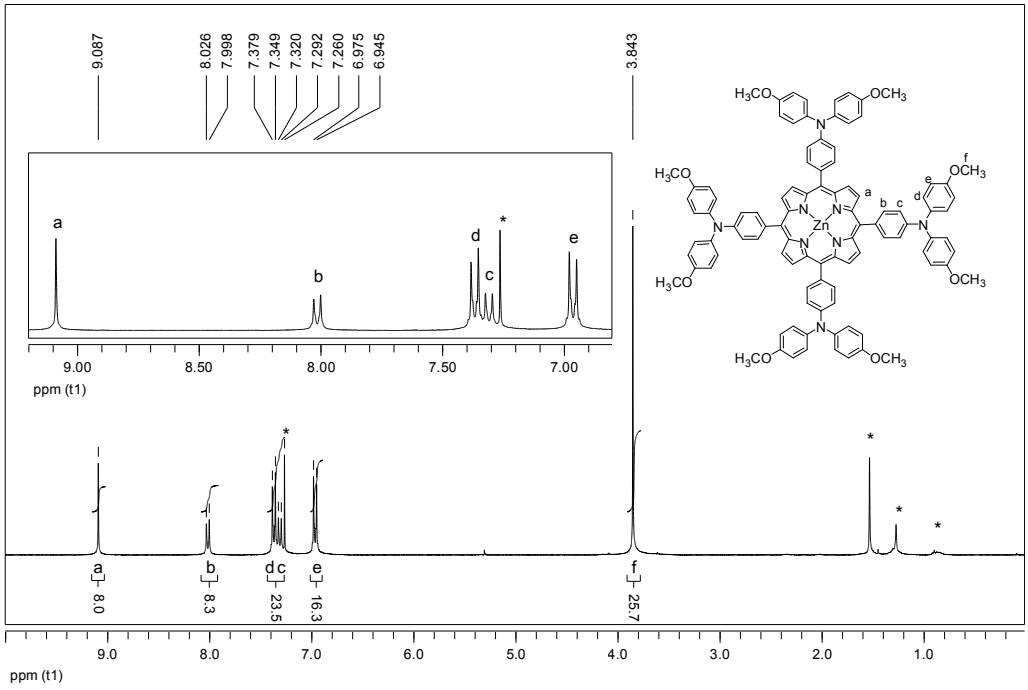

Fig. 1 ^{1}H NMR of **3** in $CDCl_3$.

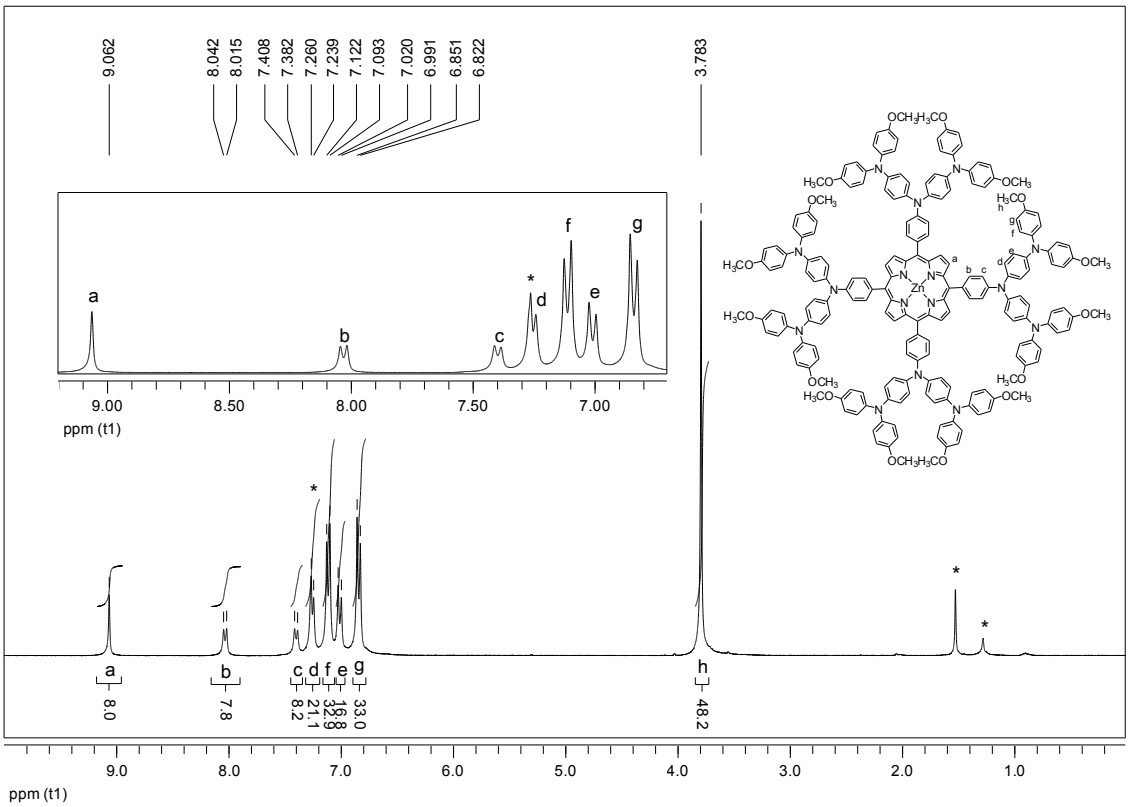

Fig. 2 ^{1}H NMR of **4** in $CDCl_3$.

B. Excitation spectra

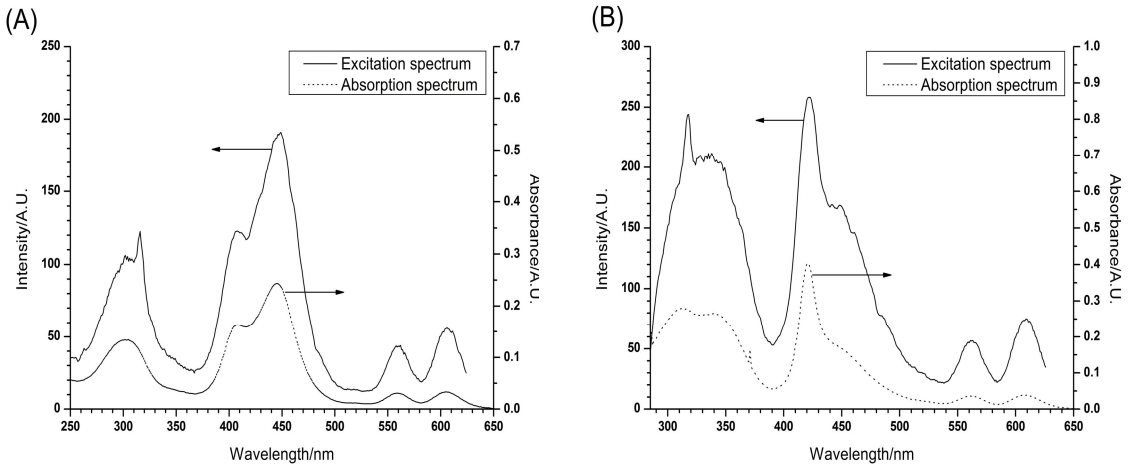

Fig. 3 The absorption and excitation spectra of (A) 2.5×10^{-5} M **3** in CH_2Cl_2; (B) 2.0×10^{-6} M **4** in toluene. Excitation experiments were monitored at emission wavelength (A) 635; (B) 632 nm.

C. Tables of calculated numbers of electron transferred from spectroelectrochemistry

Table 1 The numbers of electron transferred (n) in Ox 1 of **4** was obtained from the plot of $\log[(A_O-A)/(A-A_R)]$[a] vs. applied potential (E). Data is obtained from Fig. 6(D) in article.

Absorption peak (nm)	$E = -\frac{0.059}{n}\log\frac{A_O - A}{A - A_R} + E^0$	n	R square
312	y = -0.061x + 0.423	0.97	$R^2 = 0.993$
426	y = -0.047x + 0.411	1.26	$R^2 = 0.961$
552	y = -0.051x + 0.444	1.16	$R^2 = 0.971$
610	y = -0.056x + 0.407	1.05	$R^2 = 0.914$
774	y = -0.057x + 0.435	1.04	$R^2 = 0.990$
1354	y = -0.060x + 0.430	0.98	$R^2 = 0.998$

[a] A_O, A_R and A represent the absorbance of a peak at oxidative state, reductive state and an applied potential respectively.

Table 2 The numbers of electron transferred (n) in Ox 2 of **4** was obtained from the plot of $\log[(A_O-A)/(A-A_R)]$[a] vs. applied potential (E). Data is obtained from Fig. 6(E) in article.

Absorption peak (nm)	$E = -\frac{0.059}{n}\log\frac{A_O - A}{A - A_R} + E^0$	n	R square
298	y = -0.062x + 0.703	0.95	$R^2 = 0.900$
426	y = -0.099x + 0.735	0.60	$R^2 = 0.965$
556	y = -0.060x + 0.721	0.98	$R^2 = 0.981$
592	y = -0.053x + 0.721	1.11	$R^2 = 0.939$
1172	y = -0.068x + 0.716	0.87	$R^2 = 0.986$
1500	y = -0.064x + 0.710	0.92	$R^2 = 0.961$

[a] A_O, A_R and A represent the absorbance of a peak at oxidative state, reductive state and an applied potential respectively.