Supplementary Matterial

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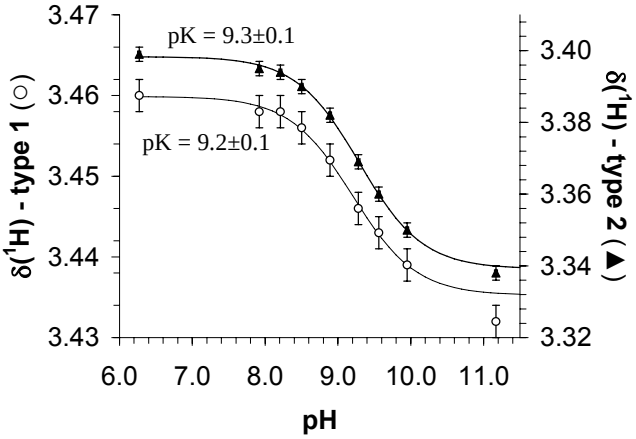

Figure SI1

Fig. SI1 The pH dependence of the ^{1}H chemical shifts of V^V-trihyat complexes formed in solution. $C_V = 4.0 \times 10^{-3}$ M, L : M = 1, I = 0.20 M KCl. The samples contain 10% D_2O. Type 1 hydrogen marked with ○ (left axis) and type 2 with ▲ (right axis).

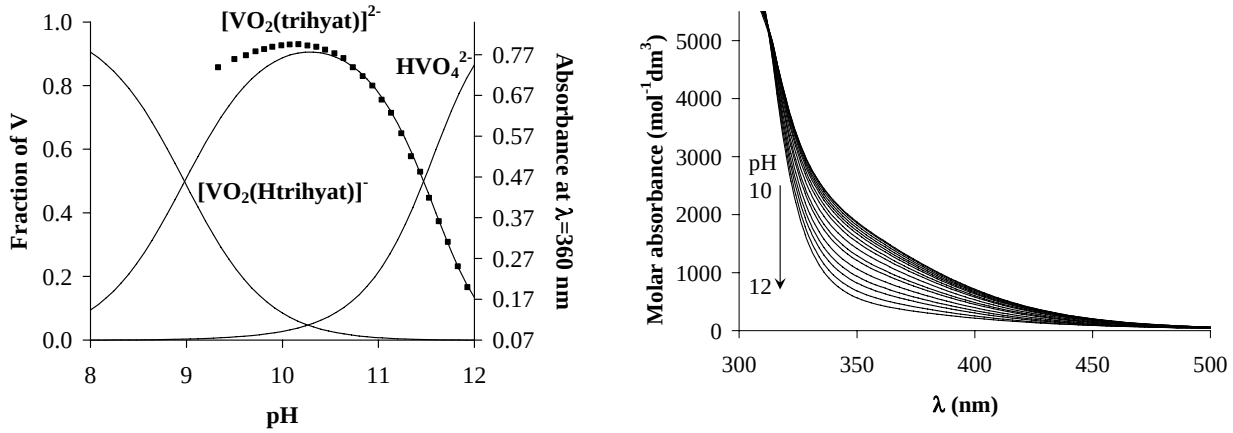

Figure SI2

Fig. SI2 A) Concentration distribution curves of V^{V}-trihyat complexes formed in solutions with c_V= 2.6×10^{-4} M and L:M=~1, calculated by using the stability constants listed in Table 1 together with the observed absorbance at □ = 360 nm. B) The molar absorbances at the same conditions, in the pH range 10-12.