

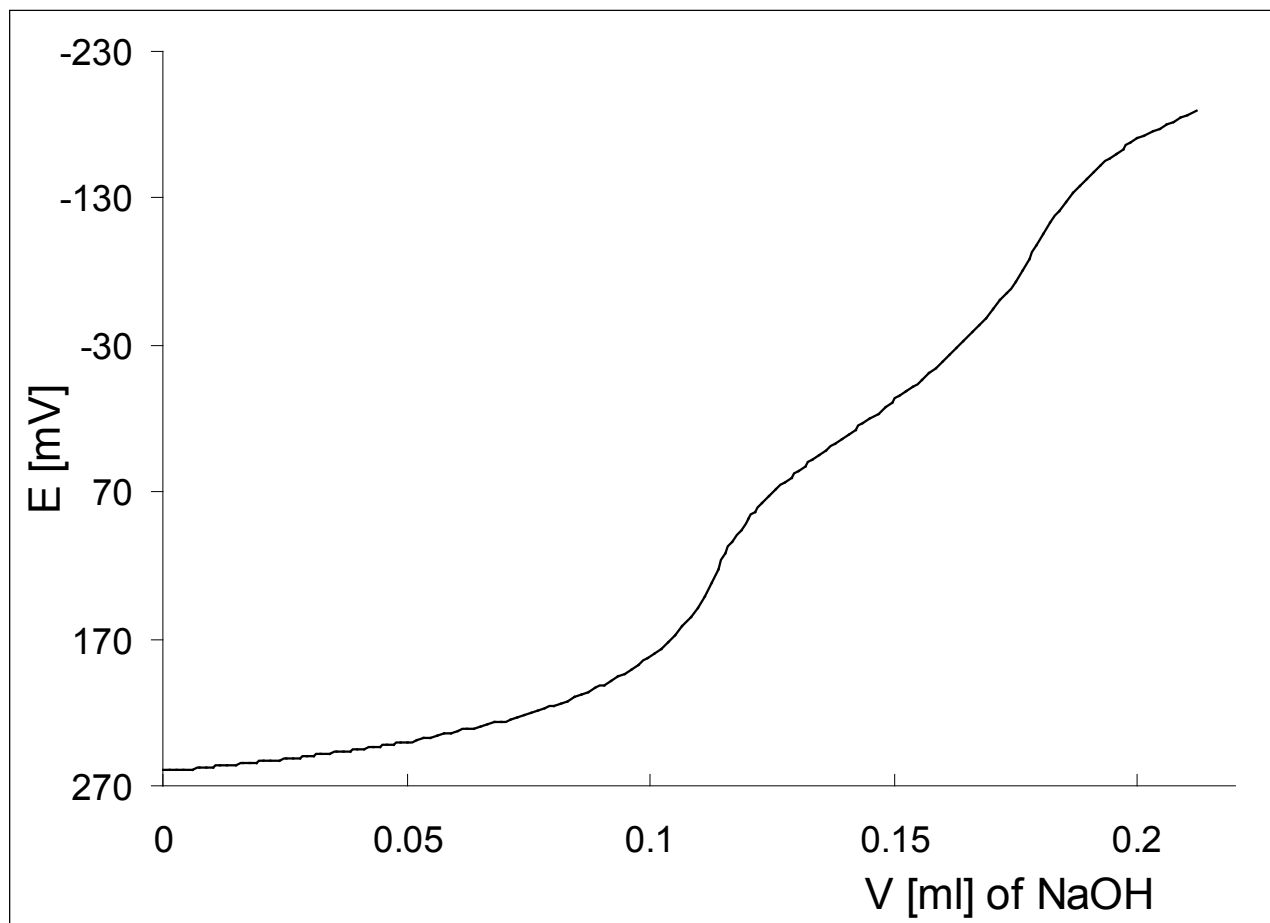


Fig. 1S

Fig 1S. pH-metric titration curve for the (1-2,7-21)NPG ligand. The volume of the ligand solution was 2cm³.

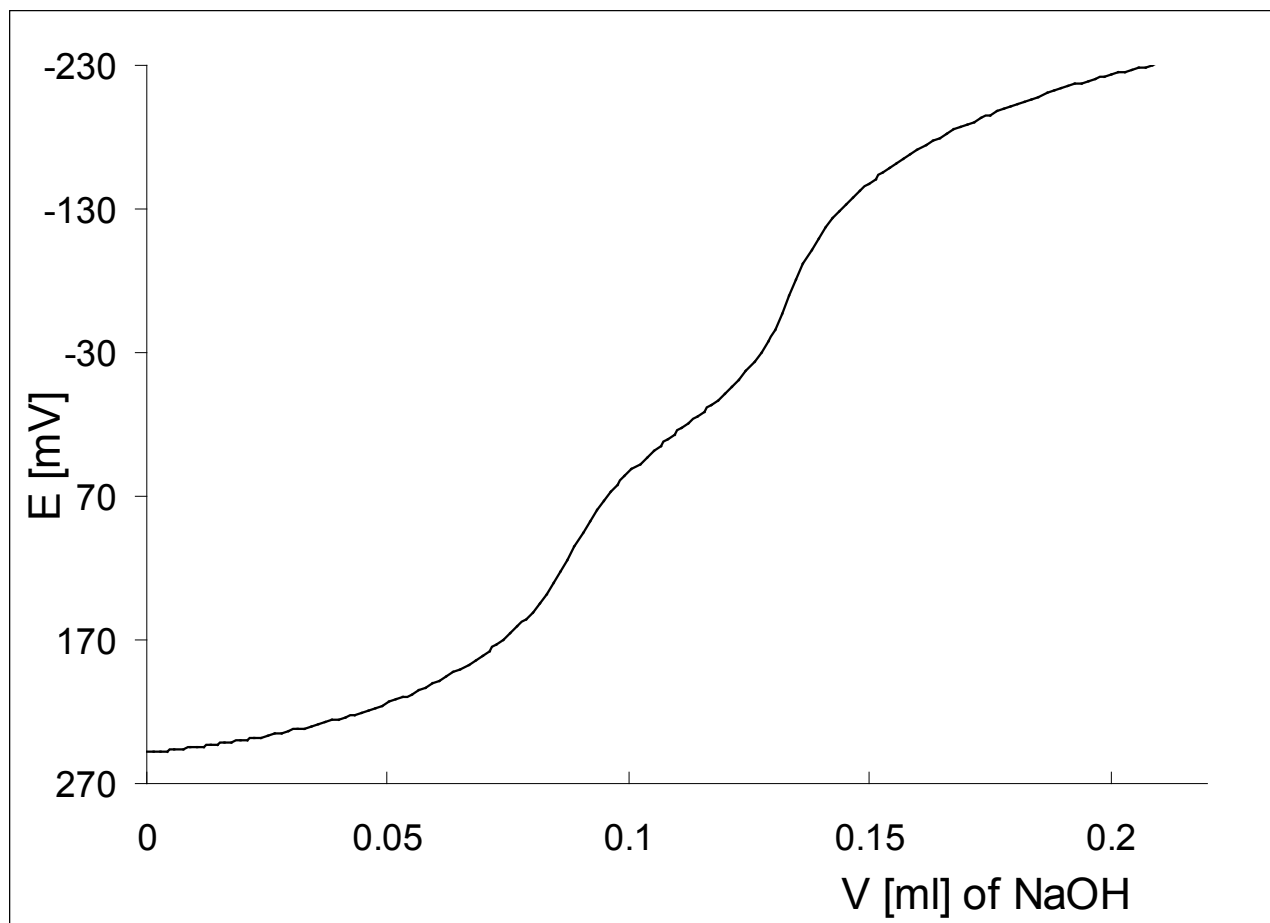


Fig. 2S

Fig 2S. pH-metric titration curve for the Ac-(1-2,7-21)NPG ligand. The volume of the ligand solution was 1.9cm³.

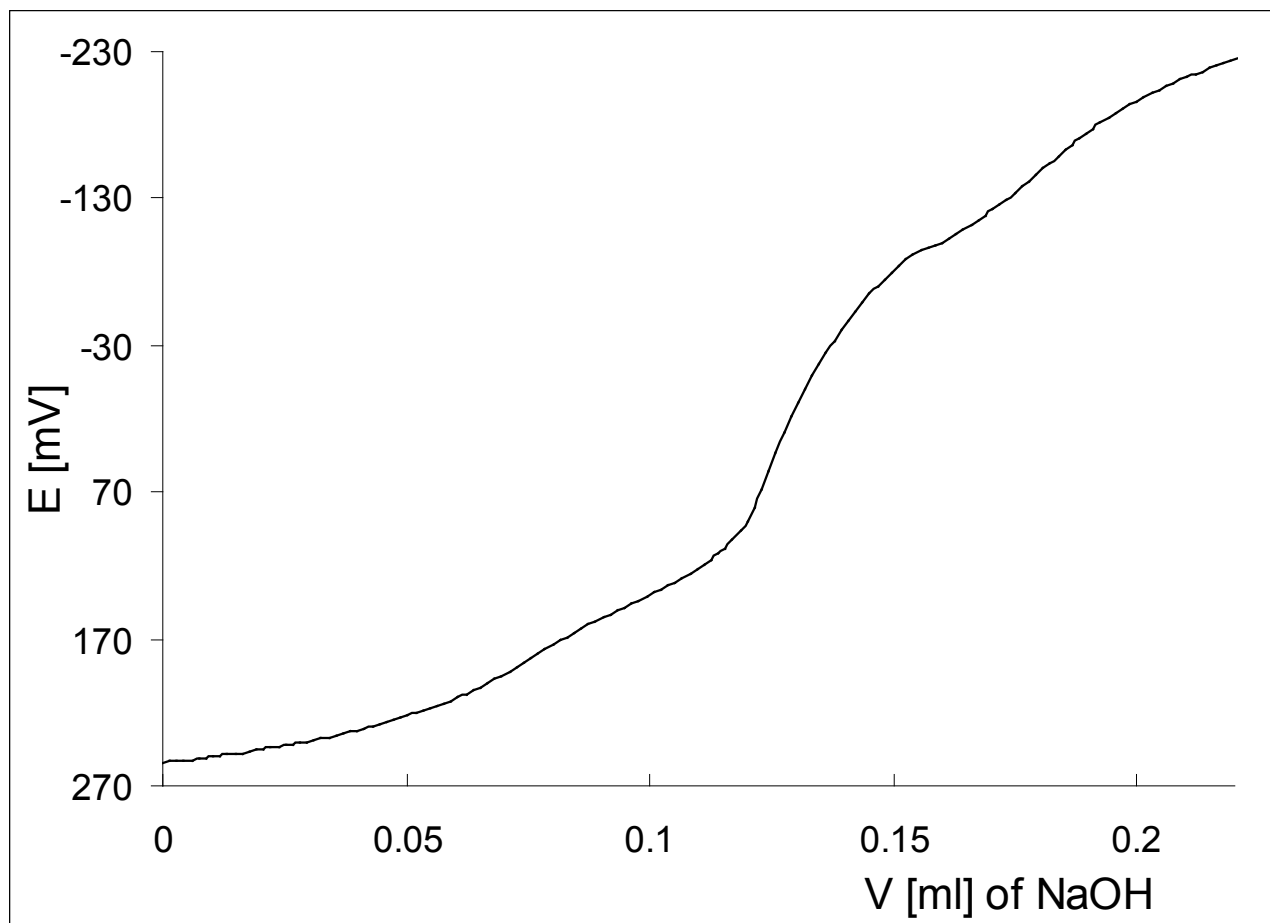


Fig. 3S

Fig 3S. pH-metric titration curve for the Cu(II) - (1-2,7-21)NPG 1:1 molar ratio system. The volume of the ligand solution was 1.5cm³.

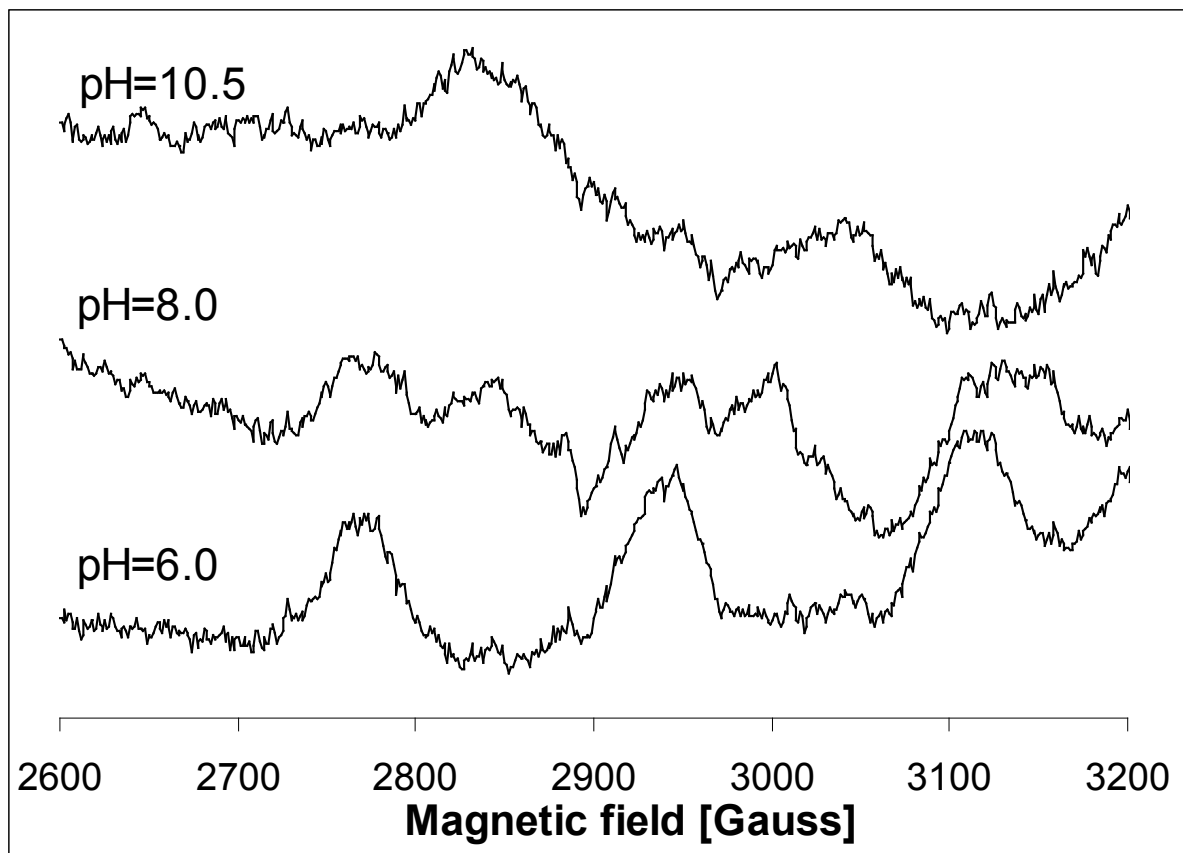


Fig. 4S

Fig 4S. The EPR spectra for the copper(II) complexes formed with (1-2,7-21)NPG at different pH and metal-to-ligand 1:1 molar ratio.

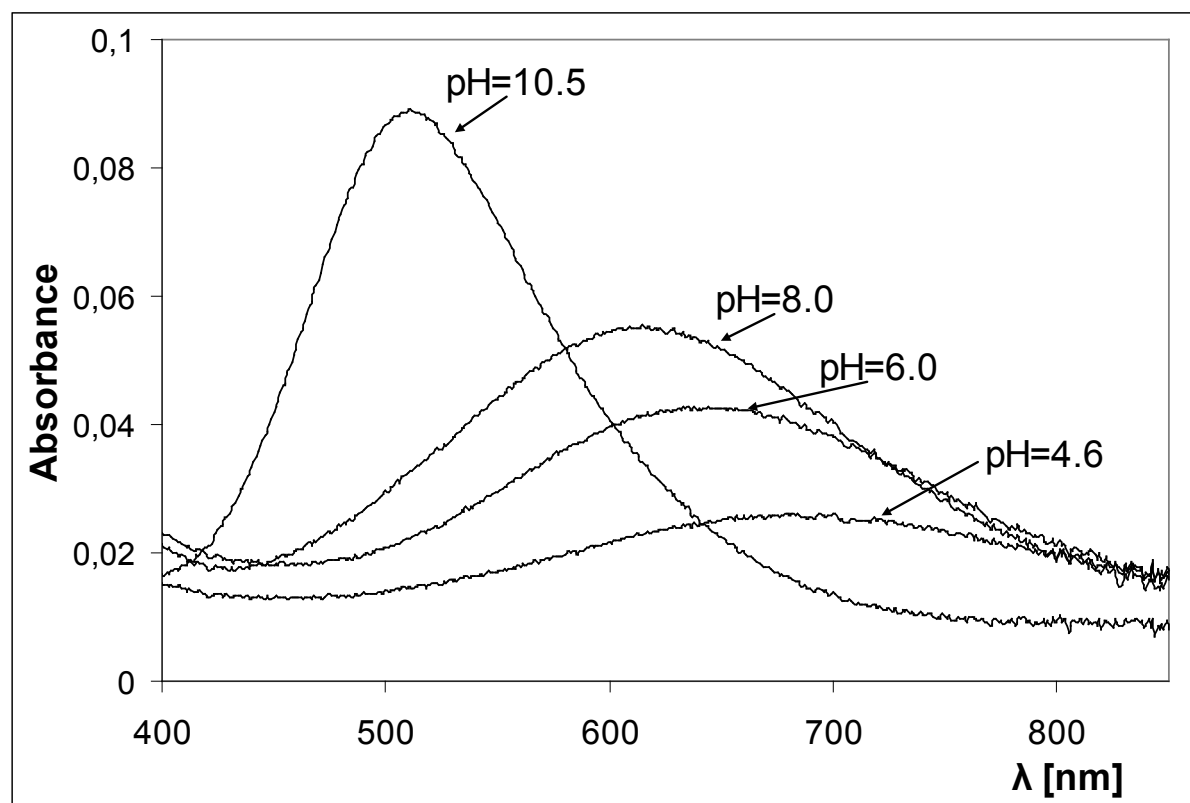


Fig. 5S

Fig 5S. The visible absorption spectra for the copper(II) complexes formed with (1-2,7-21)NPG at different pH and metal-to-ligand 1:1 molar ratio.

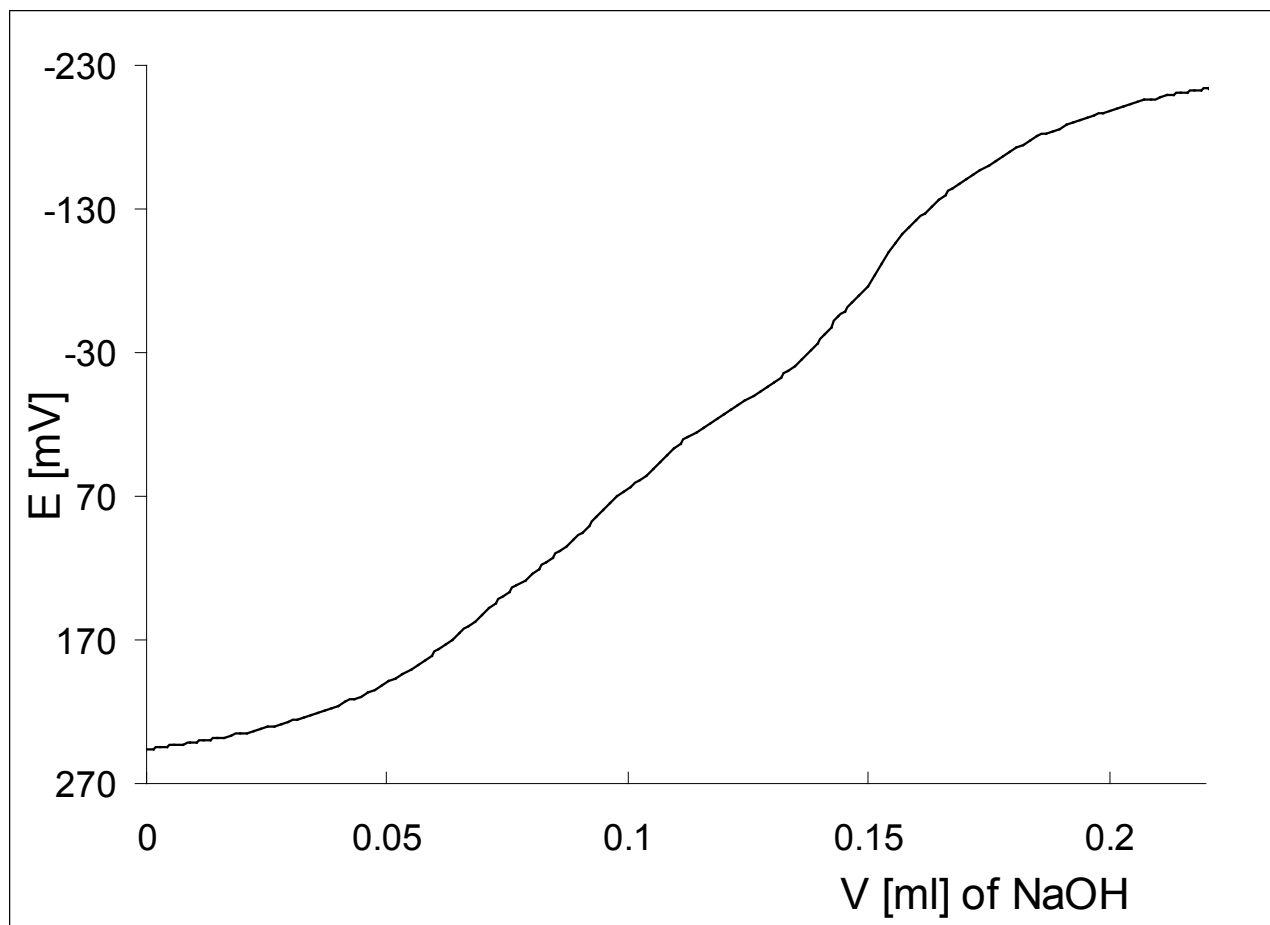


Fig. 6S

Fig 6S. pH-metric titration curve for the Cu(II) – Ac-(1-2,7-21)NPG 1:1 molar ratio system. The volume of the ligand solution was 1.6cm³.

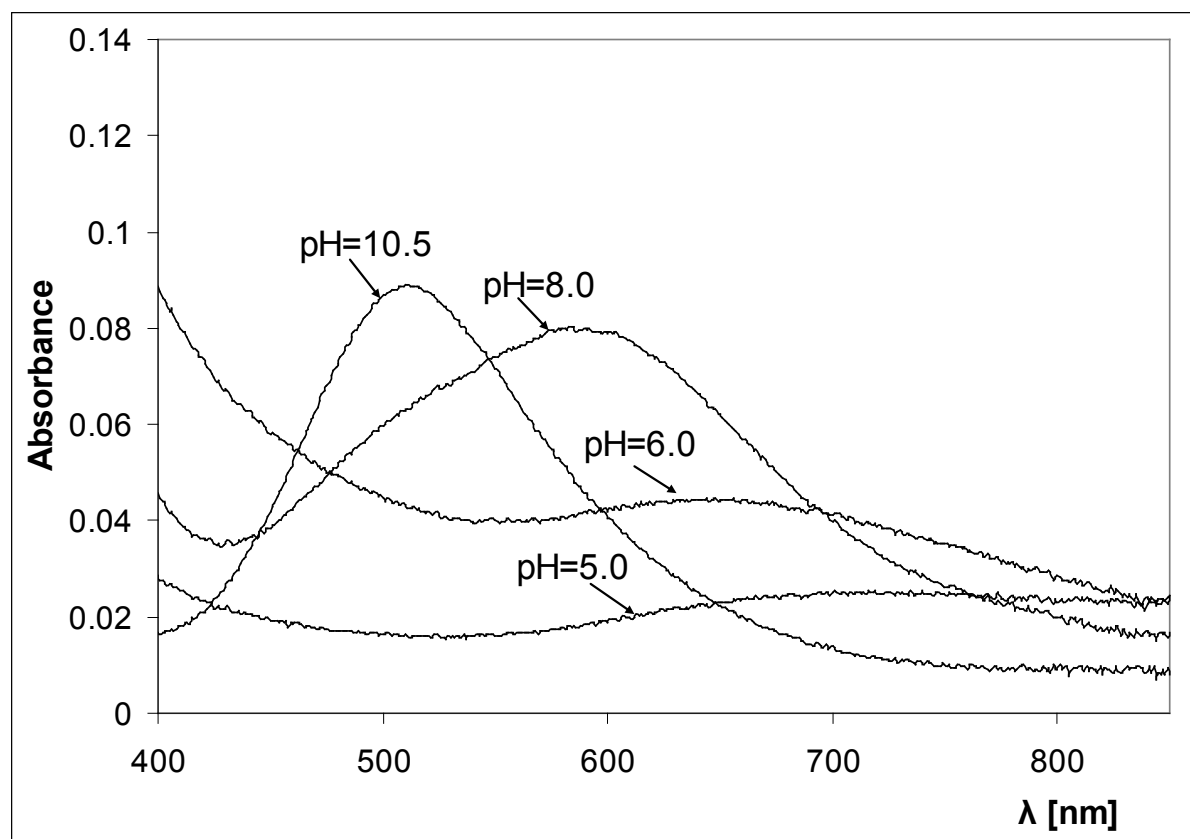


Fig. 7S

Fig 7S. The visible absorption spectra for the copper(II) complexes formed with Ac-(1-2,7-21)NPG at different pH and metal-to-ligand 1:1 molar ratio.

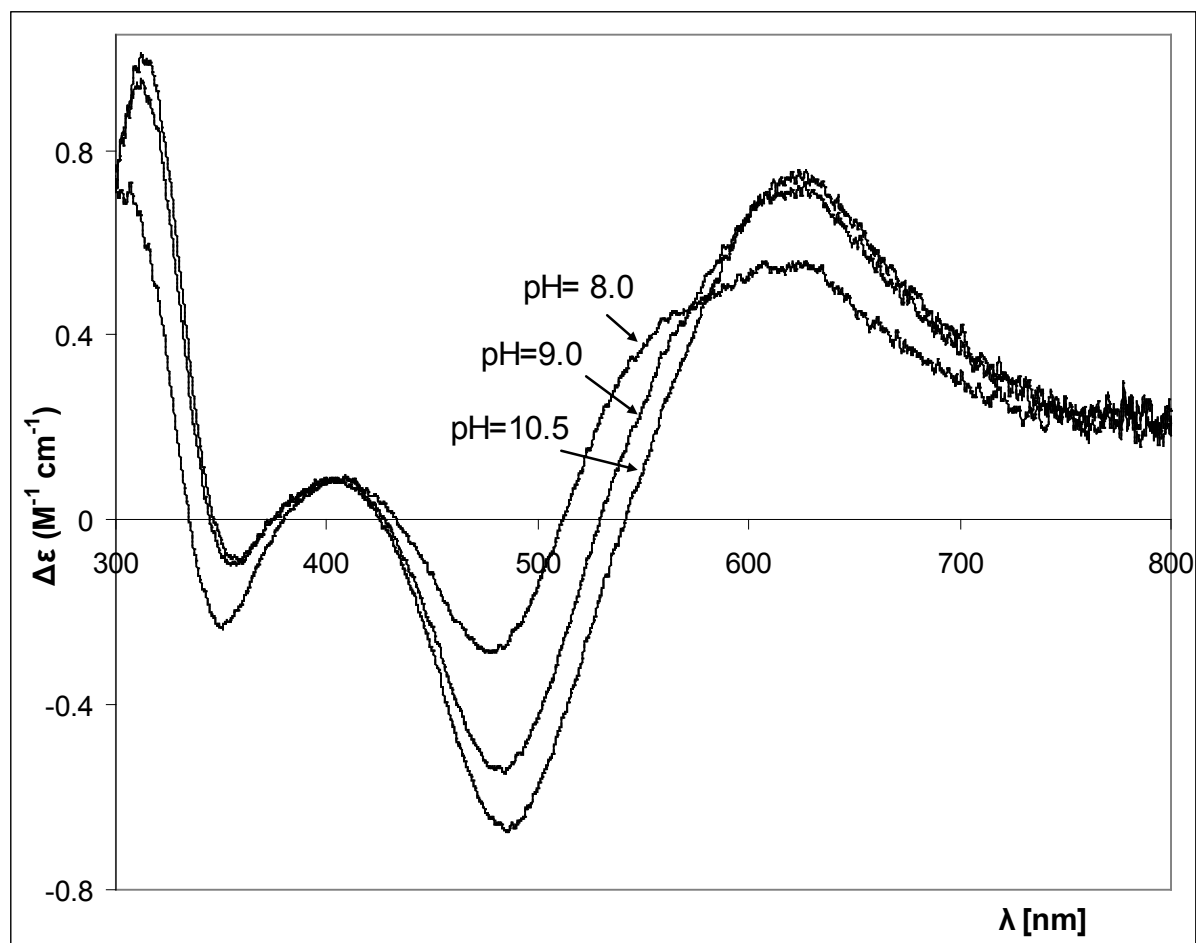


Fig. 8S

Fig. 8S. CD spectra of the species formed in the copper(II)-Ac-(1-2,7-21)NPG system at different pH; metal-to-ligand molar ratio 1:1.