

## Supporting Information

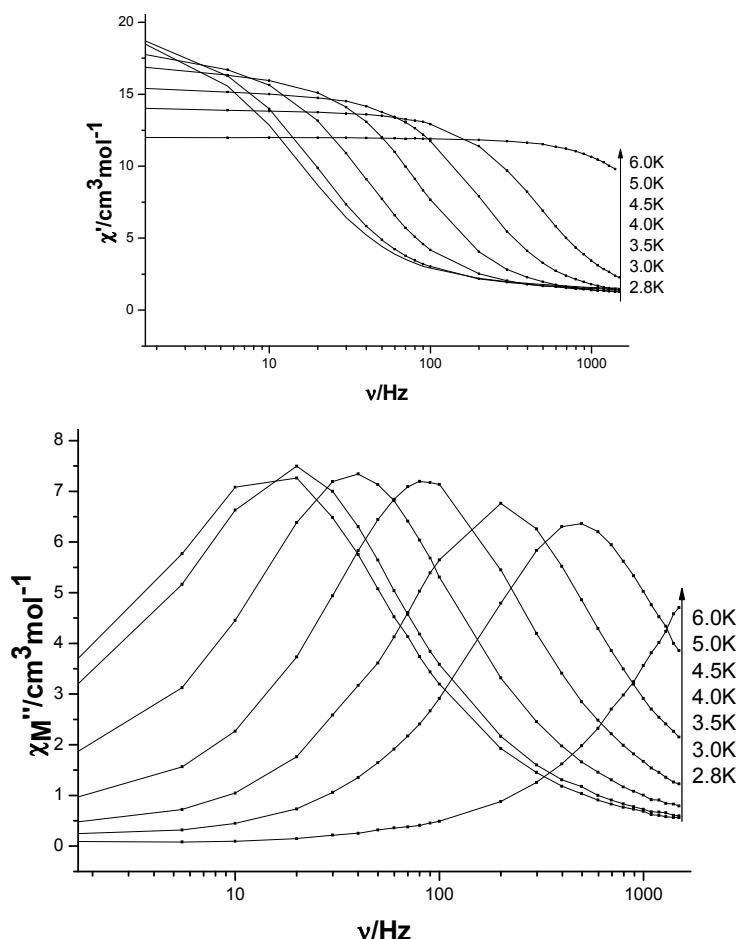
### A series of 1D Dy(III) compound showing slow magnetic relaxation: synthesis, structure, and magnetic studies

Yan Zhu, Feng Luo\*, Yu-mei Song, Hai-xiao Huang, Gong-ming Sun, Xiao-zhao Tian, Zi-Zun Yuan, Zhen-wei Liao, Ming-biao Luo, Shu-juan Liu, Wen-yuan Xu, and Xue-Feng Feng

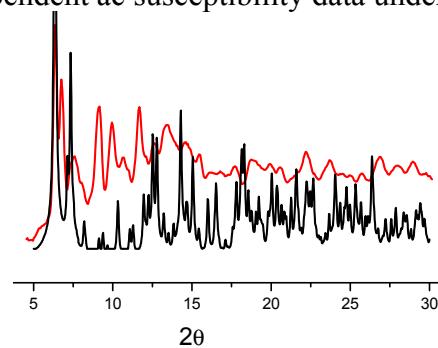
#### IR Studies.

The IR spectrum of these compounds is investigated in KBr pellets. For **1**, the IR bonds of  $3100\text{ cm}^{-1}$  is ascribed to the C-H stretching bond of L1 phenyl ring.  $3353\text{ cm}^{-1}$  indicates that the water molecules afford H-bonds. The absence of IR bonds around  $1700\text{ cm}^{-1}$  implies that in **1** all the L1 ligand is deprotonated. And the typical antisymmetric ( $1642\text{ cm}^{-1}$ ) and symmetric ( $1596, 1517\text{ cm}^{-1}$ ) stretching bands with respective values of  $\Delta=v_a(\text{coo}^-)-v_s(\text{coo}^-)$  clearly indicate the presence of chelate ( $46\text{ cm}^{-1}$ ), bidentate ( $125\text{ cm}^{-1}$ ) coordination modes of the carboxyl groups.<sup>1</sup>  $1508\text{ cm}^{-1}$  is a typical IR bonds of  $-\text{NO}_2$  group. The 1,4- substituent of L1 ligand is manifested by  $875\text{ cm}^{-1}$  IR bond. Similarly, for **2** and **3**, the typical IR bonds of phenyl ring are around  $3085\text{ cm}^{-1}$ ,  $3067\text{ cm}^{-1}$ . The H-bonds from water molecule are also observed, as evidenced by IR bonds of  $3478\text{ cm}^{-1}$  for **2** and  $3340\text{ cm}^{-1}$  for **3**. The above-discussed coordination mode of carboxyl groups are also confirmed by IR spectrum: for **2**, chelate and bidentate coordination mode can be deduced from the typical antisymmetric ( $1659\text{ cm}^{-1}$ ) and symmetric ( $1605, 1530\text{ cm}^{-1}$ ) stretching bands with respective values of  $\Delta=54, 129\text{ cm}^{-1}$ , whereas for **3**,  $1698\text{ cm}^{-1}$  approves monodentate coordinated mode and the corresponding IR bonds of  $1617, 1555, 1487, 1421\text{ cm}^{-1}$  with  $\Delta=62, 130, 196\text{ cm}^{-1}$  suggest chelate, bidentate, and monodentate coordinated mode. The typical IR bonds of  $-\text{NO}_2$  group is  $1510\text{ cm}^{-1}$  for **2** and  $1512\text{ cm}^{-1}$  for **3**. As observed in **1**, the IR bond of 1,4-substituent of L1 ligand in **3** is about  $875\text{ cm}^{-1}$ . By contrast, the IR bond of 1,3-substituent of L2 ligand in **2** is around  $783\text{ cm}^{-1}, 720\text{ cm}^{-1}$ .

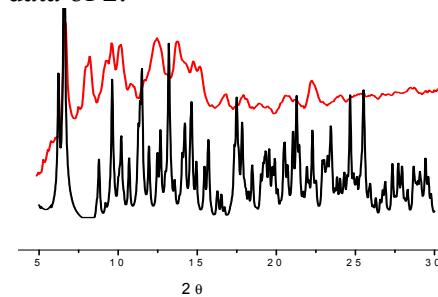
1 M. Kakihana, T. Nagumo, M. Okamoto, H. Kakihana. *J. Phys. Chem.*, 1987, **6128**, 91.



**Fig. S1** The frequency dependent ac susceptibility data under dc=1000 Oe, ac=3 Oe.



**Fig. S2** The experimental/red XRD pattern of bulk samples and simulated/black XRD pattern from singel crystal data of **2**.



**Fig. S3** The experimental/red XRD pattern of bulk samples and simulated/black XRD pattern from singel crystal data of **3**.