Electronic Supporting Information

A Mononuclear Ruthenium Complex Showing Multiple Proton-Coupled Electron Transfer toward Multi-electron Transfer Reactions

Masaya Okamura,^{*a,b*} Masaki Yoshida,^{*a,c*} Reiko Kuga,^{*a*} Ken Sakai,^{*c,d,e*} Mio Kondo^{*a*} and Shigeyuki Masaoka^{**a,b,f*}

^{*a*} Institute for Molecular Science, Higashiyama 5-1, Myodaiji, Okazaki 444-8787 Japan. ^{*b*} The Graduate University for Advanced Studies[SOKENDAI], Shonan Village, Hayama, Kanagawa 240-0193 Japan

^c Department of Chemistry, Faculty of Science, Kyushu University, Hakozaki 6-10-1, Higashi-ku, Fukuoka 812-8581, Japan.

^d International Institute for Carbon-Neutral Energy Research (WPI-I²CNER), Kyushu University, Motooka 744, Nishi-ku, Fukuoka 819-0395, Japan.

^e International Research Center for Molecular Systems (IRCMS), Kyushu University, Motooka 744, Nishi-ku, Fukuoka 819-0395, Japan.

^f PRESTO, Japan Science and Technology Agency (JST), Honcho 4-1-8, Kawaguchi, Saitama 332-0012, Japan.

E-mail: masaoka@ims.ac.jp

Fax: (+81) 564-59-5589



Fig. S1. Hydrogen-bonded one-dimensional network of the cationic moieties of 1 in the crystal.



Fig. S2. Hydrogen-bonded dimeric structure of the complex **2**' in the crystal. Thermal ellipsoids are displayed at the 50% probability level.



Fig. S3. Time course of absorption spectra after dissolution of 1 (0.5 mM) in water.



Fig. S4. UV-visible spectrum of an acetonitrile solution of 1 (0.5 mM).



Fig. S5. Cyclic voltammogram of **1** (0.5 mM) in 0.1 M TBAP/acetonitrile solution under Ar atomosphere (WE: GC, CE: Pt wire, RE: SCE; Scan rate: 100 mV/s).



Fig. S6. pH-dependence of cyclic voltammograms of $[Ru(trpy)(H_2bim)(OH_2)](PF_6)_2$ (0.5 mM) in aqueous solution under various pH conditions (WE : GC, CE : Pt wire, RE : SCE ; Scan rate :100 mV/s)



Fig. S7. Isodensity surface plots of selected frontier molecular orbitals of $[Ru(trpy)(Hbim)(OH_2)]^+$ based on the optimized ground-state geometry.



Fig. S8. Isodensity surface plots of selected frontier molecular orbitals of [Ru(trpy)(Hbim)(OH)]⁰ based on the optimized ground-state geometry.

Complex	Wavelength (nm)	Oscillator strength	Transition	CI coef (> 0.3)
[Ru(trpy)(H ₂ bim)(OH ₂)] ²⁺	466	0.0469	HOMO-2→LUMO+1	0.68751
	442	0.0912	HOMO-2→LUMO	0.34681
			HOMO-1→LUMO+1	0.39629
			HOMO→LUMO+1	0.36473
	364	0.0606	HOMO-1→LUMO+2	0.42111
			HOMO→LUMO+4	0.39707
$[Ru(trpy)(Hbim)(OH_2)]^+$	500	0.0486	HOMO-2→LUMO+1	0.68547
	477	0.0652	HOMO-2→LUMO	0.44248
			HOMO→LUMO+1	0.33187
	389	0.0381	HOMO→LUMO+2	0.65208
[Ru(trpy)(Hbim)(OH)] ⁰	660	0.056	HOMO→LUMO+1	0.60367
	414	0.117	HOMO-1→LUMO+2	0.66706

Table S1. Wavelength and oscillator strengths of both isomers at MLCT band from TD-DFT calculated absorption spectra.



Fig. S9 Cyclic voltammograms (0.5 mM) of 2 in an aqueous $1 \text{ N H}_2\text{SO}_4$ solution (red dotted line) and blank solution (black line) under Ar atmosphere. (WE : GC, CE : Pt wire, RE : SCE ; Scan rate : 10 mV/s)