Comparison of Chemical Interactions with Li^+ and Catalytic Reactivity of Electrochemically Generated $[Fe^IClL]^{2-}$ and $[Co^IL]^-$ Complexes (L = salen or salophen)

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Supporting Information

Figures S1 – S5



Figure S1. Cyclic voltammograms obtained at a glassy carbon electrode (scan rate = 100 mV s^{-1}) for reduction of (A) 0.25 mM [Co^{II}(salen)] and (B) 0.25 mM [Co^{II}(salophen)] in acetonitrile (0.1 M [*n*-Bu₄N][PF₆]) containing (a) 0; (b) 0.125; (c) 0.25; (d) 0.375; (e) 0.5 mM LiClO₄.



Figure S2. Comparison of simulated (o o o) and experimental (—) cyclic voltammograms for reduction of 0.25 mM (A) [$Co^{II}(salen)$] and (B) [$Co^{II}(salophen)$] at a glassy carbon electrode in acetonitrile (0.1 M [n-Bu₄N][PF₆]) containing (a) 0; (b) 0.25; (c) 0.5 mM LiClO₄. Simulation parameters are given in the text and Table 2.



Figure S3. Cyclic voltammograms obtained at a glassy carbon electrode (scan rate = 100 mV s^{-1}) for reduction of 0.25 mM (A) [Co^{II}(salen)] and (B) [Co^{II}(salophen)] in acetonitrile (0.1 M [*n*-Bu₄N][PF₆]) containing (a) 0; (b) 0.125; (c) 0.25; (d) 0.375; (e) 0.5; (f) 1.0 mM benzyl chloride.



Figure S4. Cyclic voltammograms obtained at a glassy carbon electrode (scan rate = 100 mV s^{-1}) for reduction of 0.5 mM (A) [Co^{II}(salen)] and (B) [Co^{II}(salophen)] in acetonitrile (0.1 M [*n*-Bu₄N][PF₆]) containing (a) 0; (b) 0.25; (c) 0.5; (d) 1.0 mM bromobenzene.



Figure S5. Cyclic voltammograms obtained at a glassy carbon electrode (scan rate = 100 mV s^{-1}) for reduction of 0.5 mM (A) [Co^{II}(salen)] and (B) [Co^{II}(salophen)] in acetonitrile (0.1 M LiClO₄) containing 0.25 (b) and 0.5 mM (c) benzyl chloride.