Supporting Information

Regio- and stereospecific living polymerization and copolymerization of (*E*)-1,3pentadiene with 1,3-butadiene by half-sandwich scandium catalysts

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Experimental section

General procedures and materials. All the manipulations were performed under pure nitrogen with rigorous exclusion of air and moisture using standard Schlenk techniques and an Mbraun glovebox. Nitrogen was purified by being passed through a Dryclean column (4A molecular sieves, Nikka Seiko Co.) and a Gasclean CC-XR column (Nikka Seiko Co.). Solvent were purified by an Mbraun SPS-800 Solvent Purification System and dried over fresh Na Chips in a glovebox. Half-sandwich scandium dialkyl complexes (1-4) were prepared as described previously (N. Yu, M. Nishiura, X. Li, Z. Xi, Z. Hou, Chem. Asian J., 2008, 3, 1406. X. Li, M. Nishiura, Li. Hu, K. Mori, Z. Hou, J. Am. Chem. Soc., 2009, 131, 13870.). [PhNMe₂H][B(C₆F₅)₄], [Ph₃C][B(C₆F₅)₄] and B(C₆F₅)₃ were purchased from Tosoh Finechem Corporation. (E)-1,3-Pentadiene (Aldrich) and 1,3-butadiene (Aldrich) were dried by stirring with MS-4A (Aldrich) for 24 hours. (Z)-1,3-Pentadiene (TCI) dried over CaH₂, vacuumtransferred and degassed by two freeze-pump-thaw cycles prior to polymerization experiments. The deuterated solvents were obtained from ISOTEC. ¹H, ¹³C NMR and ¹H–¹³C HMQC spectra were recorded on JEOL EX270 (270 MHz for ¹H; 67 MHz for ¹³C) spectrometer and JEOL ECA400WB (400 MHz for ¹H; 100 MHz for ¹³C). All deuterated NMR solvents were stored over molecular sieves under a nitrogen atmosphere. The microstructure of the polypentadiene (PPD) and polybutadiene (PBD) products were determined by ¹H and ¹³C NMR spectroscopy in CDCl₃ (and C₆D₆) solution at 25 °C. Molecular weight and molecular weight distribution for the polymer were measured by gel permeation chromatography (GPC, Tosoh HLC-8220 GPC, Column: Super HZM-H HT \times 3) at 40 °C using THF as an eluent at a flow rate of 0.35 mL/min. The molecular weight was calculated by a standard procedure based on the calibration with standard polystyrene samples. The DSC measurements were performed on a DSC6220 (SII Co.) at a rate of 20 °C/min.

A typical procedure for polymerization of (*E*)-1,3-pentadiene (Table S1, run 1) In the glove box, a predetermined amount of (*E*)-1,3-pentadiene, and 40 µmol of [PhNMe₂H][B(C₆F₅)₄] (13 mg) in toluene were introduced sequentially at room temperature in a 30 mL round-bottom glass flask with a stirring bar. A few minutes later, 40 µmol of complex **4** (13 mg) in toluene was added into the flask. The polymerization was carried out for a predetermined period of time, and was terminated by addition of 2 mL of methanol. The resulting mixture was poured into methanol (100 mL) to precipitate the polymer product. The polymer was collected by filtration and dried under vacuum at 60 °C to a constant weight. The microstructure of the PPD was determined by ¹H and ¹³C NMR spectroscopy in CDCl₃ (and C₆D₆) solution at 25 °C. ¹H NMR (PPD): δ = 0.85–0.94 (–CH₃ of 1,4-PPD unit), 1.10–1.32 (–CH₂– of 1,2-PPD unit), 1.61–1.70 (–CH₃ of 1,2-PPD unit), 1.89–2.18 (–CH– of 1,2-PPD unit and –CH₂– of 1,4-PPD unit), 2.42–2.55 (–CH– of 1,4-PPD unit), 5.11–5.35 ppm (–CH= of 1,2-PPD unit and –CH= of 1,4-PPD unit). ¹³C NMR (PPD): δ = 17.96 (*syndiotactic-trans*-1,2-PPD unit), 20.91 (*isotactic-cis*-1,4-PPD unit), 123.90–127.91 and 135.52–137.35 ppm (1,2-PPD unit and 1,4-PPD unit).

The isomer contents of the polypentadiene products were calculated from the 1 H and 13 C NMR spectra according to the following formulas (eqs 1–4):

Mol 1,4-PD% = { $I_{\rm H5}/(I_{\rm H5} + I_{\rm H5v})$ } × 100 (1)

Mol 1,2-PD% = { $I_{H5v}/(I_{H5} + I_{H5v})$ } × 100 (2)

in which I_{H5} is the integration of the resonance at 0.85–0.94 ppm and I_{H5v} is the integration of the resonance at 1.61–1.70 ppm in the ¹H NMR spectrum (see Figures S4 and S6).

Mol *cis*-1,4-PD% = { $I_{C1}/(I_{C1} + I_{C1t})$ } × 100 (3)

Mol *trans*-1,4-PD% = { $I_{C1t}/(I_{C1} + I_{C1t})$ } × 100 (4)

in which I_{C1} is the integration of the signals at 35.15 ppm assigned as the methylene carbon of the *cis*-1,4-PD unit and I_{C1t} is the integration of the signals at 40.44 ppm assigned as the methylene carbon of the *trans*-1,4-PD unit in the ¹³C NMR spectrum (see Figures S5 and S7).

Preparation of isotactic*cis***-1**,**4**-**polyPD**-*b*-*cis***-1**,**4**-**polyBD diblock copolymer.** In the glove box, a predetermined amount of 1,3-butadiene (BD), and 32.0 mg (40 μ mol) of [PhNMe₂H][B(C₆F₅)₄] in toluene were introduced sequentially at room temperature in a 30 mL round-bottom glass flask with a stirring bar. A few minutes later, 13 mg (40 μ mol) of complex **4** in toluene was added into the flask. After the 1,3-butadiene (BD) was completely consumed, (*E*)-1,3-pentadiene (PD) was added to the reaction mixture. The copolymerization was carried out for a predetermined period of time, and was terminated by addition of 2 mL of methanol. The resulting mixture was poured into methanol (100 mL) to precipitate the polymer product. The polymer was collected by filtration and dried under vacuum at 60 °C to a constant weight.

1 au	cat.	$T_{\rm p}$	[PD]/[Sc]	time	yield	micro	ostructure (%	$(b)^{b}$	$M_{\rm n}^{c}$	$M_{ m w}/M_{ m n}^{\ c}$	$T_{\rm g}^{\ d}$
run	[Sc]	(°C)			(%)	<i>cis</i> -1,4-	trans-1,4-	1,2-	(× 10 ⁴)		(°C)
1	4	25	92	5 min	100	72	15	13	2.61	1.27	-50
2	4	25	130	5 min	100	71	16	13	3.39	1.27	-50
3	4	25	150	5 min	100	70	16	14	4.50	1.27	-52
4	4	25	185	5 min	100	70	17	13	5.12	1.26	-49
5	4	25	260	5 min	99	71	16	13	6.79	1.26	-50
6	4	25	331	5 min	99	71	16	13	10.04	1.26	-50
7	4	0	92	5 min	96	85	8	7	2.56	1.28	-55
8	4	0	150	5 min	95	85	8	7	5.09	1.28	-55
9	4	0	220	5 min	100	85	8	7	9.99	1.30	-55
10	4	-20	150	5 min	100	92	4	4	4.24	1.32	-57
11^e	4	-20	150×2	$10 \min \times 2$	200	92	4	4	9.28	1.37	-57
12	4	-40	150	40 min	98	97	1	2	4.46	1.37	-60
13 ^e	4	-40	150×2	50 min × 2	191	97	1	2	6.21	1.37	-59
14	4	-60	150	4 h	30	> 99 ^f	0	< 1	3.10	1.65	-62 ^g
15	4	-60	150	24 h	55	> 99 ^f	0	< 1	4.45	1.54	-61 ^g

Table S1	. Polymerization	of (<i>E</i>)-1	,3-Pentadiene (P	D) by 4/	[PhMe ₂ NH]	$[B(C_6F_5)_4]^a$
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^{*a*} Conditions: [Sc], 40 µmol; [PhMe₂NH][B(C₆F₅)₄], 40 µmol; toluene, 10 mL. ^{*b*} Determined by ¹H and ¹³C NMR. ^{*c*} Determined by GPC in THF at 40 °C. ^{*d*} Measured by DSC. ^{*e*} After polymerization of 150 equiv. of PD for 10 (run 19) or 50 min (run 21), another 150 equiv. of PD was added and the mixture was stirred for a further 10 (run 20) or 50 min (run 22). ^{*f*}*mm* > 99%. ^{*g*}*T*_m = 45 °C (run 23) and 43 °C (run 24) were also observed.

run	Temp.	[PD]	[BD]	t	yield	BD cont. ^b	cis-1,4/trans-	$M_{\rm n}^{\ d}$	$M_{\rm w}/M_{\rm n}^{\ d}$	$T_{\rm g}^{\ e}$
		/[Sc]	/[Sc]				1,4/1,2°			
	(°C)			(min)	(%)	(mol%)	(%)	(×10 ⁴)		(°C)
1	-20	0	150	5	100	100	86/4/10	2.61	1.14	-103
2^{f}	-20	150	150	10 + 10	200	50	$92/4/4^{g}$	7.17	1.28	-102, -65
3	-40	0	150	120	100	100	87/4/9	1.71	1.22	-105
4 ^{<i>f</i>}	-40	150	150	120 + 120	190	50	97/1/2 ^g	3.48	1.39	-104, -70

Table S2. Polymerization and Block-copolymerization of 1,3-Butadiene (BD) and (*E*)-1,3-Pentadiene (PD) by $4/[PhMe_2NH][B(C_6F_5)_4]$ at -20 °C and -40 °C^{*a*}

^{*a*} Conditions: 40 µmol of Sc, 40 µmol of [PhMe₂NH][B(C₆F₅)₄], 10 mL of toluene. ^{*b*} Calculated from yield. ^{*c*} Determined by ¹H and ¹³C NMR. ^{*d*} Determined by GPC in THF at 45 °C. ^{*e*} Measured by DSC. ^{*f*} After polymerization of 150 equiv. of BD for 10 min or 120 min, another 150 equiv. of PD was added and the mixture was stirred for a further 10 min or 120 min. ^{*g*} Microstructure for PPD. Microstructure for PBD: *cis*-1,4-/*trans*-1,4-/1,2 = 86/4/10.



Figure S1. GPC curves of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at 25 °C (a) (Tables S1, runs 1-6) and 0 °C (b) (Tables S1, runs 7-9).



Figure S2. GPC curves of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -20 °C (a) (Tables S1, runs 10, 11) and -40 °C (b) (Table S1, runs 12, 13).



Figure S3. GPC curves of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -60 °C (Table S1, runs 14, 15).



Figure S4. ¹H NMR spectra of polypentadienes obtained by 1-4/[PhNMe₂H][B(C₆F₅)₄] at 25 °C (Table 1, runs 1-4). (**: H_2O)







Figure S6. ¹H NMR spectra of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ (Table S1, runs 3, 8, 10, 12, 14). (**: H₂O)



Figure S7. ¹³C NMR spectra of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ (Table S1, runs 3, 8, 10, 12, 14). (*: CDCl₃)



Figure S8. ¹³C NMR spectra of the aliphatic region of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ (Table S1, runs 3, 8, 10, 12, 14).



Figure S9. ¹³C NMR spectra (*: C₆D₆) of polypentadiene obtained by **4**/[PhNMe₂H][B(C₆F₅)₄] (Table S1). The peak at 21.641 ppm (-20 °C), 21.631 ppm (-40 °C) and 21.641 ppm (-60 °C) are attributable to "*mm*" triads (see *J. Organomet. Chem.* **1993**, *451*, 67, *mm*; δ = 21.632 ppm, *mr*; δ = 21.490 ppm).



Figure S10. ¹H-¹³C HMQC spectrum of polypentadiene obtained at -40 °C by $4/[PhNMe_2H][B(C_6F_5)_4]$ (Table S1, run 12). (*: CDCl₃)



Figure S11. DSC curves of polypentadienes obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ (Table S1, Table S1, runs 3, 8, 10, 12, 14).







Figure S13. GPC curves of polypentadiene (a) and pentadiene-butadiene diblock copolymer (b) obtained at -20 °C (Table S1 run 10 and Table S2, run 4).



Figure S14. ¹H NMR spectra of pentadiene-butadiene diblock copolymer, polybutadiene and polypentadiene obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -20 °C (Table S1, run 10 and Table S2, runs 1, 2). (**: H₂O)



Figure S15. ¹³C NMR spectra of pentadiene-butadiene diblock copolymer, polybutadiene and polypentadiene obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -20 °C (Table S1, run 10 and Table S2, runs 1, 2). (*: CDCl₃)



Figure S16. DSC curves of pentadiene-butadiene diblock copolymer, polybutadiene and polypentadiene obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -20 °C (Table S1, run 10 and Table S2, runs 1, 2).



Figure S17. GPC curves of polypentadiene (a) and pentadiene-butadiene diblock copolymer (b) obtained at -40 °C (Table S2, runs 3, 4).



Figure S18. ¹H NMR spectra of polybutadiene and pentadiene-butadiene diblock copolymer obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -40 °C (Table S2, runs 3, 4).



Figure S19. ¹³C NMR spectra of polybutadiene and pentadiene-butadiene diblock copolymer obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -40 °C (Table S2, runs 3, 4).



Figure S20. DSC curves of pentadiene-butadiene diblock copolymer, polybutadiene and polypentadiene obtained by $4/[PhNMe_2H][B(C_6F_5)_4]$ at -40 °C (Table S1, run 12 and Table S2, runs 3, 4).

Computational details. All calculations were performed with Gaussian 09 program.¹ The DFT method of B3PW91² was used for geometry optimizations. The 6-31G* basis set was used for C and H atoms, and the Sc atom was treated by the Stuttgart/Dresden effective core potential (ECP) and the associated basis set.³ In the Stuttgart/Dresden ECP used in this study, the most inner 10 electrons of Sc are included in the core. The 11 valence electrons of Sc were treated by the optimized basis set, viz. (8s7p6d1f)/[6s5p3d1f]. Such a basis set contains one f-polarization function with an exponent of 0.27. Each optimized structure was analyzed by harmonic vibrational frequencies at the same level of theory to characterize the minima (Nimag = 0) and the transition states (Nimag = 1). The C₅Me₅ ligand was used as a model of C₅Me₄SiMe₃ to save computational time. Such a strategy was successfully adopted in our previous work.⁴

1. Cationic Active Species The cationic scandium allyl complex $[(\eta^5-C_5Me_5)Sc(\eta^3-C_3H_5)]^+$ has been optimized to model the initial active species (Figure S21). The Sc-C(allyl) distances are 2.321 Å for Sc-C1 and Sc-C3) bonds and 2.408 Å Sc-C2 contacts.



$$[(\eta^{5}-C_{5}Me_{5})Sc(\eta^{3}-C_{3}H_{5})]^{+}$$

Figure S21. Optimized structure (distance in Å) of cationic half-sandwich scandium alkyl species.

2. Chain initiation The computed free energies (in kcal/mol) of stationary points possibly involved in the chain initiation stage are shown in Table S3. (s-*trans*-1,4-*supine* represents insertion of the *E*-1,3-pentadiene with s-*trans* coordination mode into Sc-allyl bond to form a π -allyl intermediate with *supine* form). Structures of four inserted intermediates are shown in Figure S22. Among these pathways, s-*cis*-1,4-*prone* showed a lowest energy barrier (8.3 kcal/mol) to give *cis*-1,4 pentenyl unit which is consistent with the fact that the *cis*-1,4-selectivity was dominant during the polymerization.

Table S3. The computed free energies (ΔG , kcal/mol) relative to the separate active species and the monomer

Insertion modes	ΔG (kcal/mol)						
	Pre-reaction complex (C)	Transition state (TS)	Insertion product (P)				
s-trans-1,4-supine	$-11.5(C_{ss})$	$5.1(\mathbf{TS}_{ss})$	$-18.1(\mathbf{P_{ss}})$				
s-trans-1,4-prone	$-12.2(C_{sp})$	$12.5(\mathbf{TS_{sp}})$	$-12.7(\mathbf{P_{sp}})$				
s-cis-1,4-supine	-7.4(C _{as})	10.0(TS _{as})	$-12.4(\mathbf{P}_{as})$				
s-cis-1,4-prone	-8.1(C _{ap})	$0.2(TS_{ap})$	-15.7(P _{ap})				



Figure S22. Four structures of allyl inserted intermediates, *syn-supine* (P_{ss}), *syn-prone* (P_{sp}), *anti-supine* (P_{as}), and *anti-prone* (P_{ap}) are shown.

3. Isotacticity The critical factor for the formation of tactic 1,4-polymers is the mutual orientation of the butenyl and butadiene part.⁵ Based on the more accessible chain propagation species P_{ap} , the anti(*prone*)-s-*cis*(*prone*) and anti(*prone*)-s-*cis*(*supine*) polymerization processes (Figure S23) have been investigated computationally as indicated by Table S4. The anti(*prone*)-s-*cis*(*prone*) process which gives isotactic *cis*-1,4-poly(PD), is more favorable both kinetically and thermodynamically than anti(*prone*)-s-*cis*(*supine*) process to give syndiotactic *cis*-1,4-poly(PD), probably because of the steric repulsion between Cp* and methyl of pentadiene in 'TS_{ac} (Figure S24).

Table S4. The Computed free energies for insertion of (E)-PD into the Sc-C bond of P_{ap} in a s-cis(supine)and s-cis(prone) manner (energies are relative to the energy sum of P_{ap} and (E)-PD).

Insertion modes	ΔG (kcal/mol)					
	С	TS	Р			
anti-s-cis(prone)	$4.2(C_{ac})$	8.8(TS _{ac})	-8.9(P _{ac})			
anti-s-cis(supine)	$0.6('C_{ac})$	23.8('TS _{ac})	-4.9(' P _{ac})			



anti(prone)-s-cis(supine)

Figure S23. Formation of iso- and syndiotactic-cis-1,4-polymer units.



Figure S24. Structures (distances in Å) for TS_{ac} and 'TS_{ac}.

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Table S5. Optimized artesian coordinates (Å) and zero-point energy corrected energies (a.u.) of the transition

states as well as their imaginary frequencies (cm⁻¹)

Atom		Х	Y	Z
TS -74	19 106501581 <i>i</i> 32	4		
10 ₅₅ , /		•		
Sc	-0.23812600	0.44807600	-0.26966000	
С	-0.41479700	3.11157100	-0.34265300	
Н	-1.33345700	3.27040400	-0.91081100	
Н	-0.12865800	3.98958100	0.23114100	
С	0.65314000	2.56302400	-1.13910600	
С	0.46516900	1.69643600	-2.19079800	
Н	-0.52790500	1.59318000	-2.64699800	
Н	1.29824500	1.32048800	-2.77360700	
С	1.44527800	-0.35533100	1.31222300	
С	2.09933400	-0.22035700	0.05107000	
С	1.50909500	-1.15156900	-0.85402200	
С	0.50462000	-1.87831700	-0.14618400	
С	0.47467800	-1.39586900	1.19829300	
С	3.30553400	0.62848200	-0.22703600	
Н	4.22119700	0.04489100	-0.06482100	
Н	3.36693300	1.49708700	0.43657200	
Н	3.34150300	0.98800300	-1.26096100	
С	1.96655900	-1.45719000	-2.25146100	
Н	2.56394400	-2.37836500	-2.26390800	
Н	2.60006100	-0.66519200	-2.66034800	
Н	1.13343400	-1.61324800	-2.94681000	
С	-0.21876700	-3.07919500	-0.68740900	
Н	0.44645500	-3.95281500	-0.68172500	
Н	-0.54678700	-2.94239800	-1.72410000	
Н	-1.09440800	-3.34730700	-0.08925200	
С	-0.27242200	-2.01838800	2.34170100	
Н	-1.23918500	-2.43989100	2.04677600	
H	-0.45070900	-1.31492000	3.16174900	
H	0.31305400	-2.84576500	2.76394300	
Н	1.67362400	2.74475100	-0.80234200	
C	1.80932100	0.36/5/800	2.57820700	
H	0.9514/500	0.49160300	3.24/64500	
H	2.56868200	-0.193/3000	3.13836100	
H	2.22864600	1.36063400	2.38600100	
C	-2.86016600	-0./3289000	-0.15948400	
C II	-2.6/943000	0.555/6800	0.2503/400	
H	-2.4195/800	-1.55155/00	0.43/04300	
н С	-3.10398/00	1.34118200	-0.33/51000	
C	-1.//349400	0.99429800	1.29200900	
	-1.34/10900	2.34991900	1.31288900	
П U	-0.00/30400	2.03144700	2.11489000	
11 U	-2.00704200	0.21200100	1.1033/300	
п С	-1.4/320900	1 15270500	2.00103200	
с ц	-3.13/43400	1 85104400	-1.2///0900 1 0570000	
н	-3.23007000	-1.63194400	-1.93700000	
Н	-4 12361600	-1.00073300	-0.07792400	
11	-т.12501000	0.27072000	1.05075200	

TS_{sp}, -749.094380112, *i*342

Sc	-0.06524300	0.53246600	-0.25457100	
С	1.53038300	2.64699400	0.17733300	
Н	1.25588300	3.68555300	0.01224600	
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Č	1 73680400	-0 87713400	0 64044300	
Č	1 44608000	-1 35065200	-0 67456100	
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C II	0.56180000	2 53180200	1 8/200/00	
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	-1./0093400	-2.30130400	1.14/80200	
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C	0.4/16/100	-0.84995900	2.9266/400	
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C	-2.91100000	0.72288400	-0.64995500	
C	-2.27601800	1.24883600	0.44144400	
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H	-2.43264800	0.76217900	1.40863600	
C	-1.381/0500	2.38441400	0.38438/00	
C	-0.27/16/00	2.54624400	1.25120400	
H	-0.00669600	3.528//200	1.62075700	
H	-0.12819700	1.78063800	2.02399800	
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	1 1916/200	2.40075500	1 60025800	
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	2 07940000	-1.23636700	-2.45551000	
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Н	0.89008300	-3.33850200	-2.13905600	
Н	-0.06694100	-2.02191400	-2.80518000	
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C C	-2.79903800	0.80343000	-0.14930000	
C II	-2.00990800	-0.32733000	0.14308400	
H	-3.40225600	1.041/9/00	-1.02692400	
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С	3.06924800	1.28779600	-0.68017800
Η	4.09281500	0.98607800	-0.42227000
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С	1.77548900	-1.00382900	-2.53479600
Η	2.62008000	-1.68561100	-2.70003900
Η	2.01516700	-0.07064000	-3.05255000
Η	0.90906600	-1.44971700	-3.03551000
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H	-2.93339200	-2.56528800	-1.24782900
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Н	-0.19830800	-2.93851600	1.05762900	
Η	-0.60129800	-1.15708100	2.57092900	
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Н	3.76497600	2.29376800	-0.25719800
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Н	3.43521400	-0.97597400	-2.30188200
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Η	2.13045600	3.67154300	0.32074300
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