

The first “Kuhn verdazyl” ligand and comparative studies of its PdCl₂ complex with analogous 6-oxoverdazyl ligands

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Supporting information

NMR spectra page S2

Electrochemical data (expanded) page S8

Figure S1. ^1H NMR of **4**.

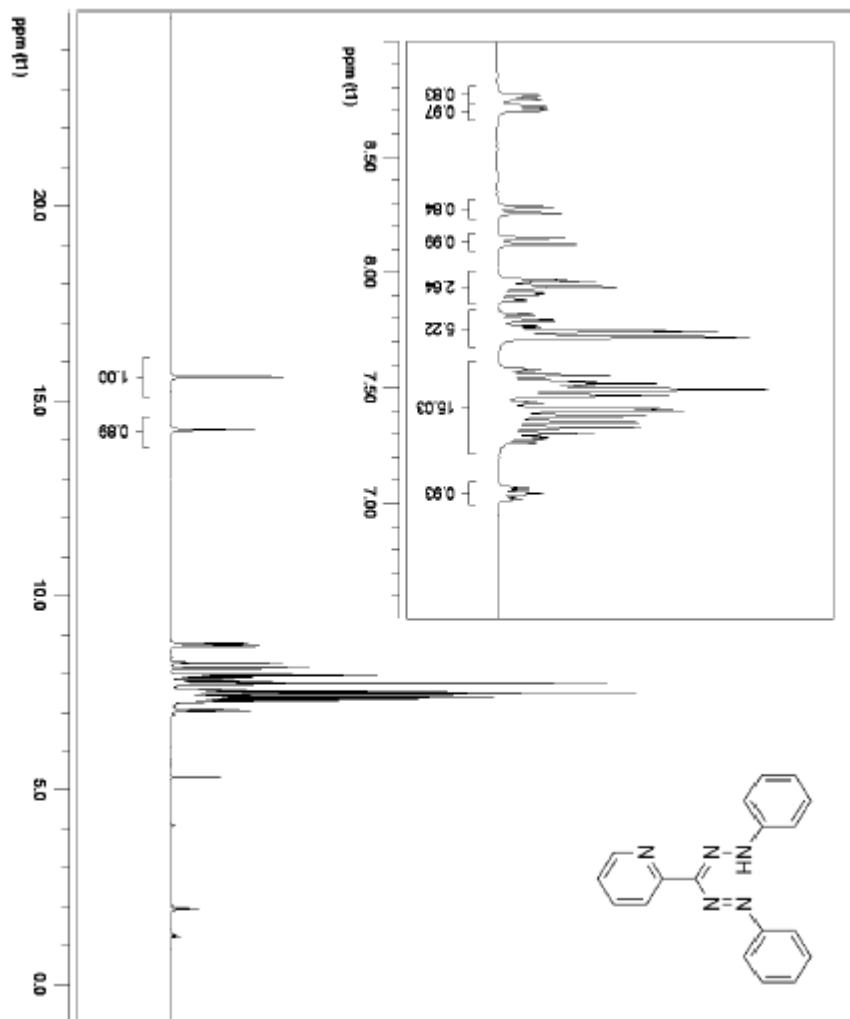


Figure S2. ^{13}C NMR of **4**.

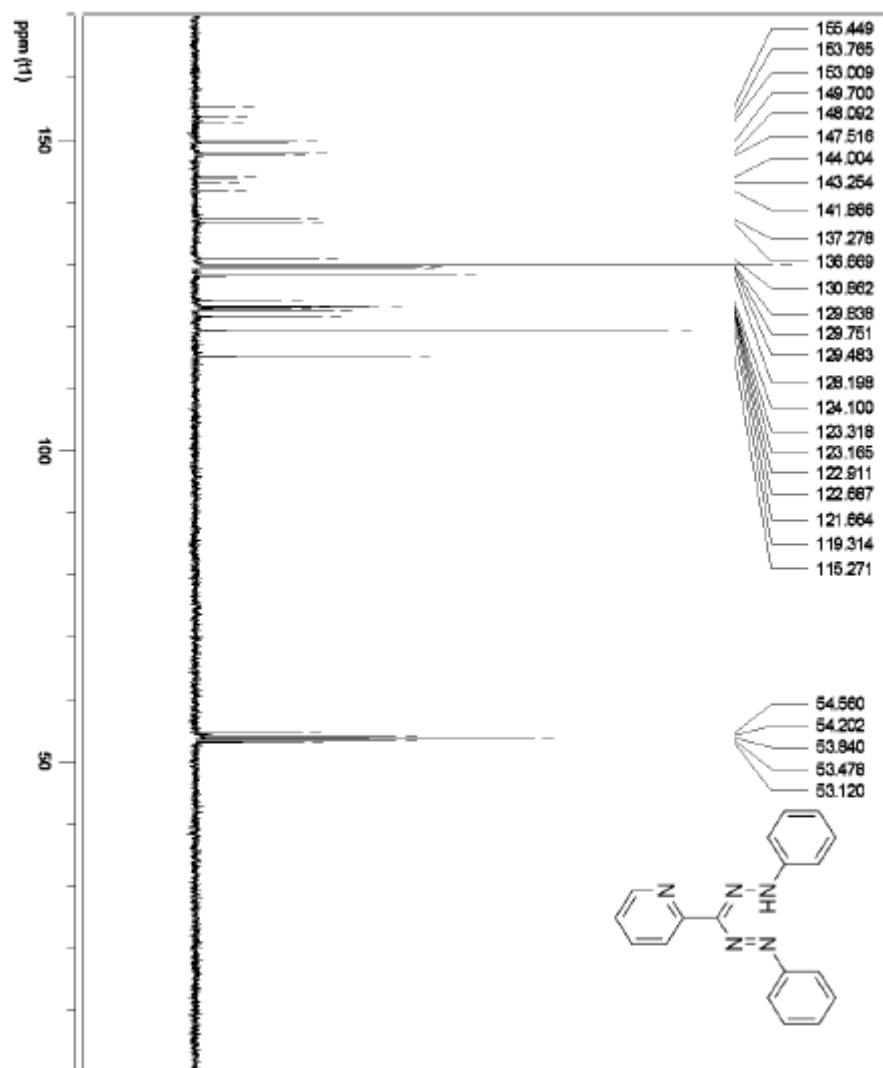


Figure S3. ^1H NMR of **6**.

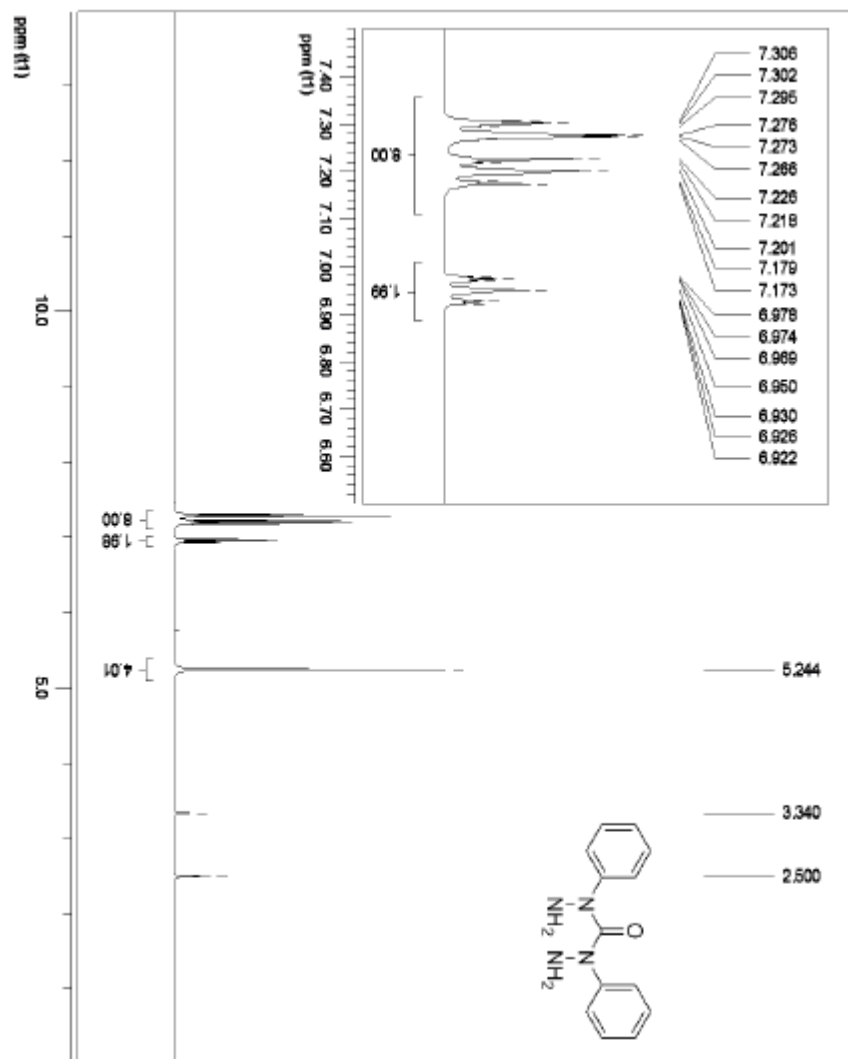


Figure S4. ^{13}C NMR of 6.

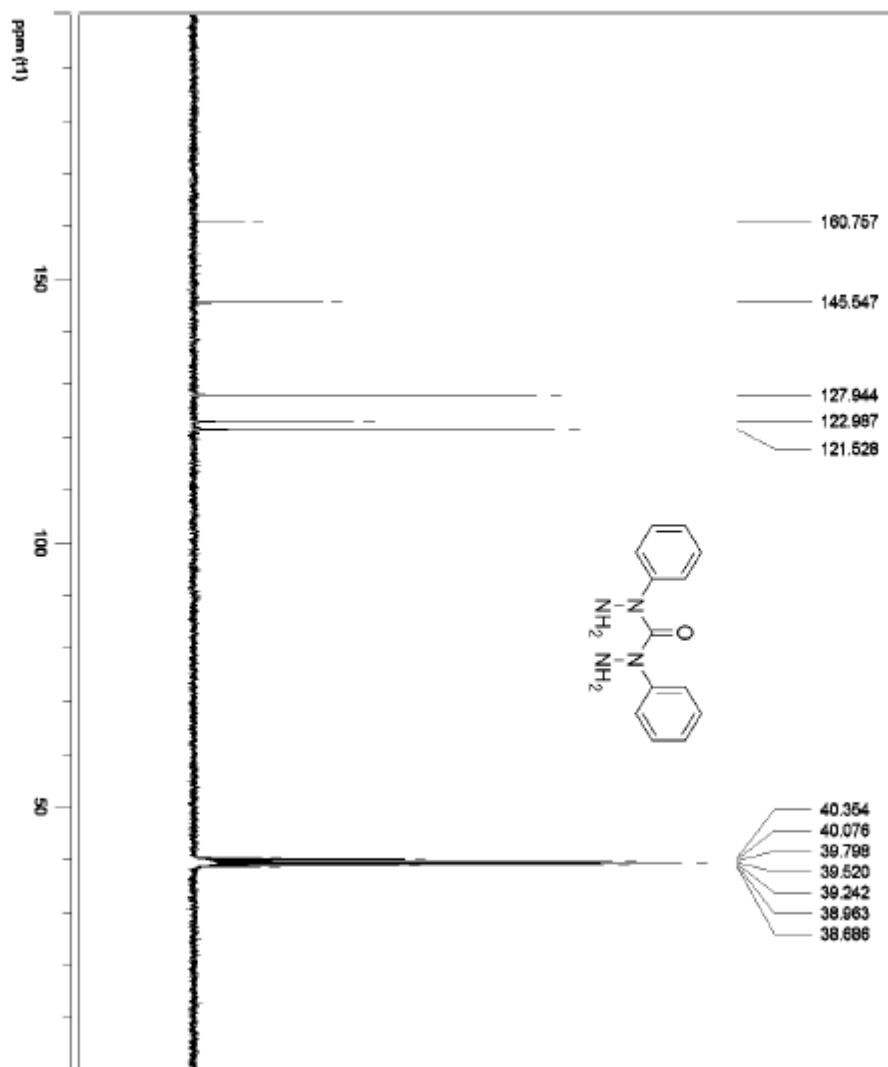


Figure S5. ^1H NMR of **7**.

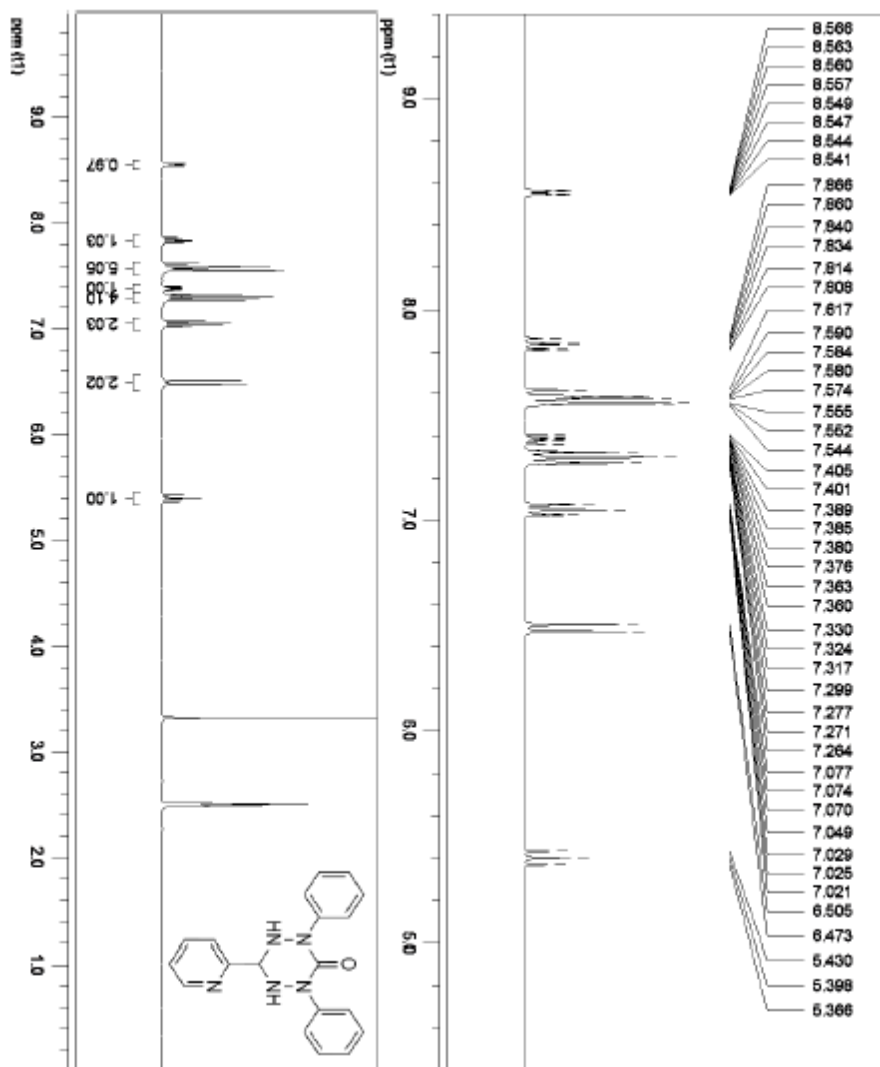


Figure S6. ^{13}C NMR of **7**.

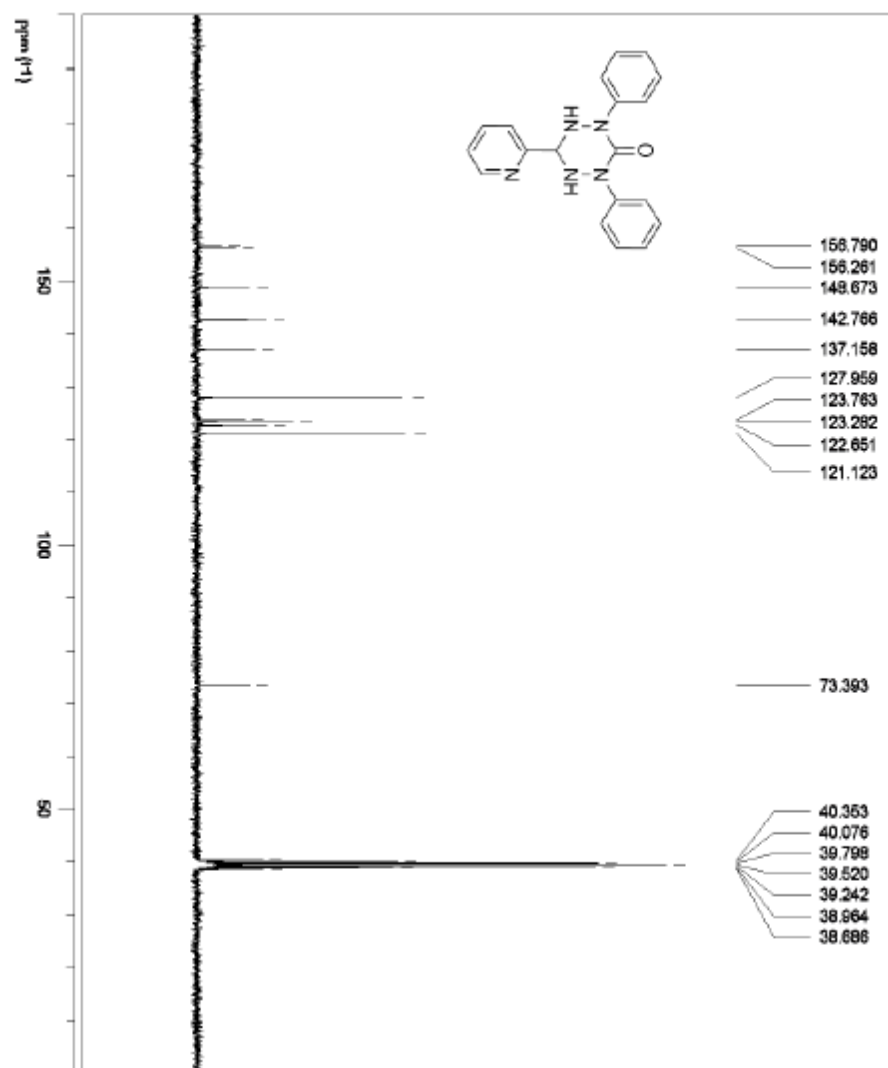


Table S1. Electrochemical data

Table 3. Electrochemical data. Potentials are in V vs. Fc^+/Fc in CH_2Cl_2 solution, cell potentials in V except for ΔE_p in mV

Cpd.	reduction				oxidation				E_{cell}
	$E_{\text{c}}^{\text{p}(1)}$	$E_{\text{c}}^{\text{a}(2)}$	$\Delta E_p^{(3)}$ (mV)	E_{red}°	$E_{\text{c}}^{\text{p}(1)}$	$E_{\text{c}}^{\text{a}(2)}$	$\Delta E_p^{(3)}$	E_{ox}°	
1a	-1.401	-1.341	70	-1.37	+0.171	+0.229	58	+0.20	1.57
1a.PdCl₂	-0.469	-0.409	60	-0.44	-	+0.67	-	+0.67 ^a	1.11 ^b
2a	-1.044	-0.965	79	-1.00	+0.438	+0.511	73	+0.47	1.47
2a.PdCl₂	-0.190	-0.130	60	-0.16	+0.781	+0.842	61	+0.81	0.97
3a	-1.329	-1.265	64	-1.30	-0.256	-0.190	66	-0.22	1.08
3a.PdCl₂	-0.439	-0.379	60	-0.41	+0.386	+0.447	61	+0.42	0.83

(1) cathodic peak potential

(2) anodic peak potential

(3) difference between anodic & cathodic peak potentials

^airreversible process; anodic peak potential given

^b E_{cell} calculated using anodic peak potential