

Supporting Information

**Highly efficient degradation of dye pollutants by Ce-doped
MoO₃ catalyst at room temperature**

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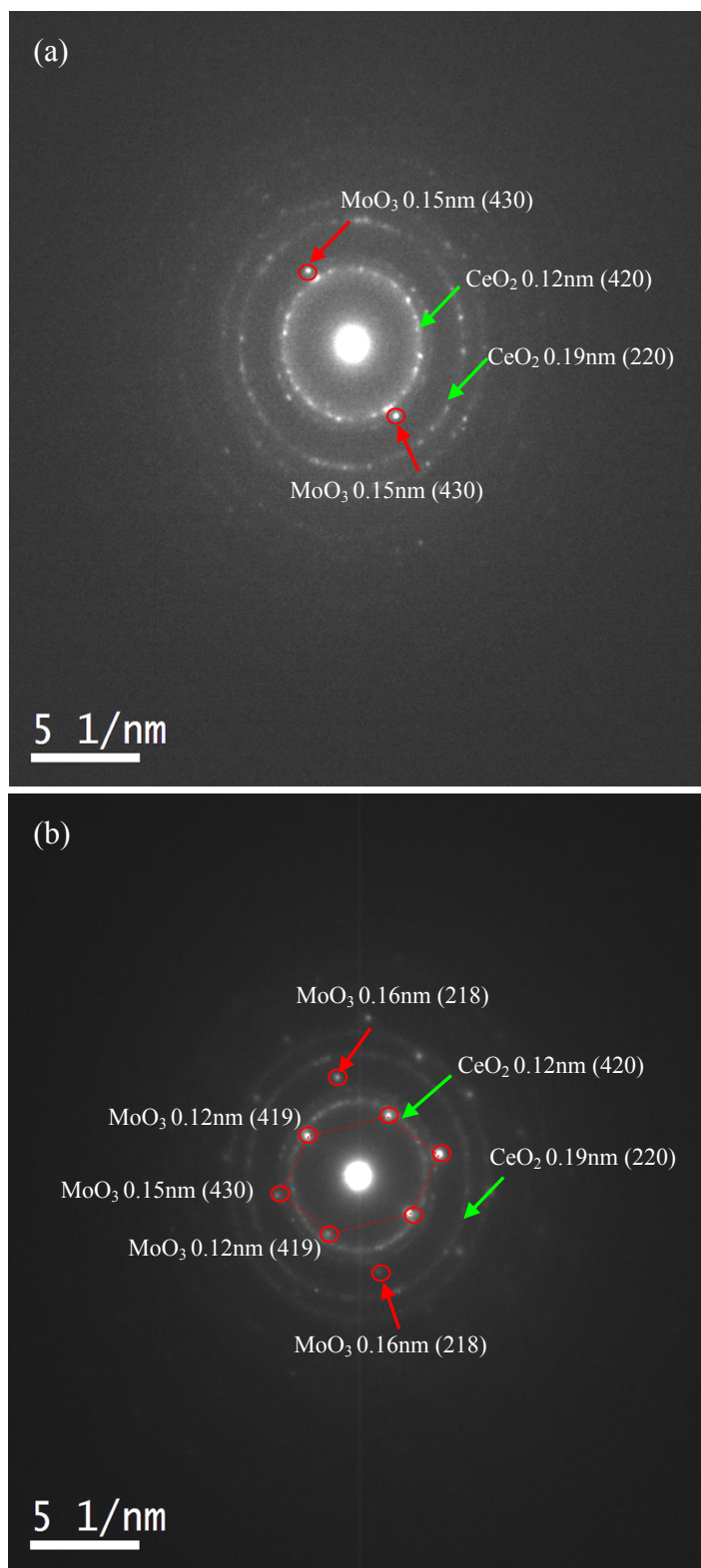


Figure S1. Selected area electron diffraction (SAED) patterns of the typical Ce(*x*)/MoO₃ samples: (a) *x*=5; (b) *x*=40

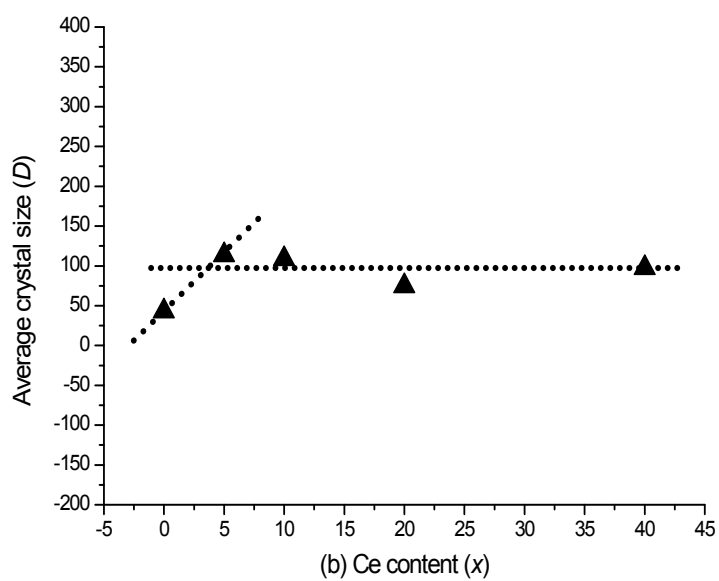
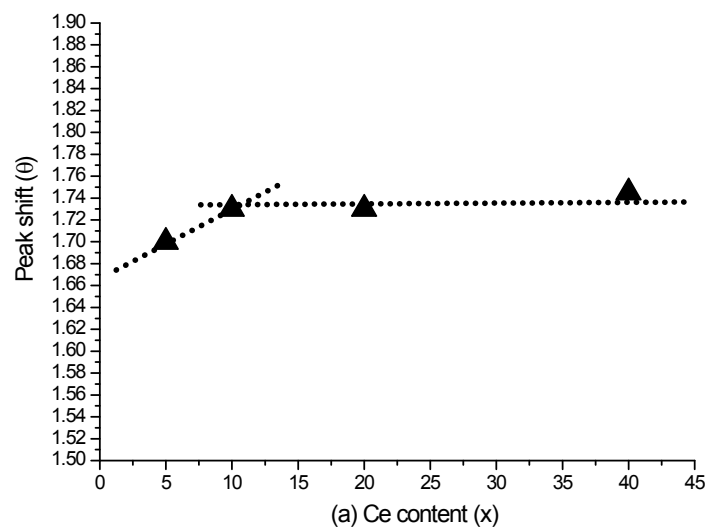


Figure S2. Peak shift degrees (θ) and average crystal sizes with the increase of CeO_2

content: calculated by (210) peak at $2\theta=25.79^\circ$

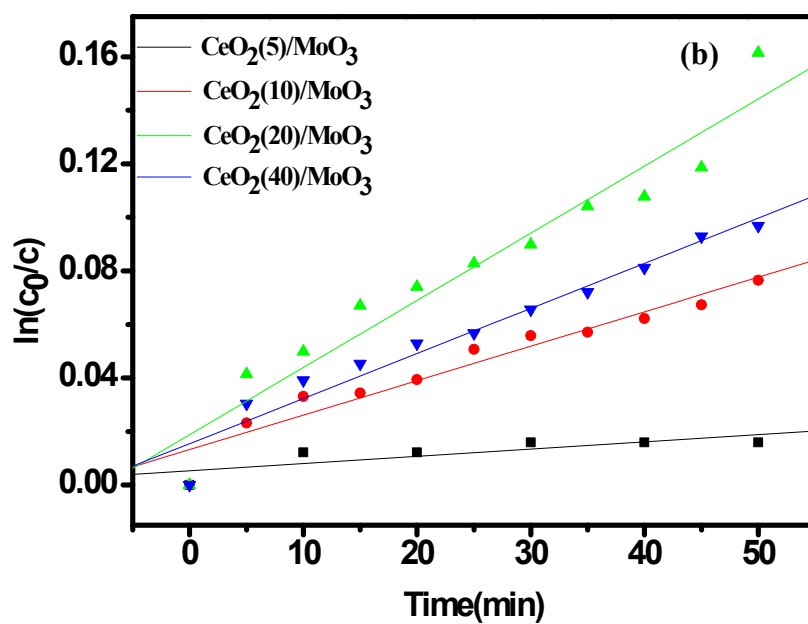
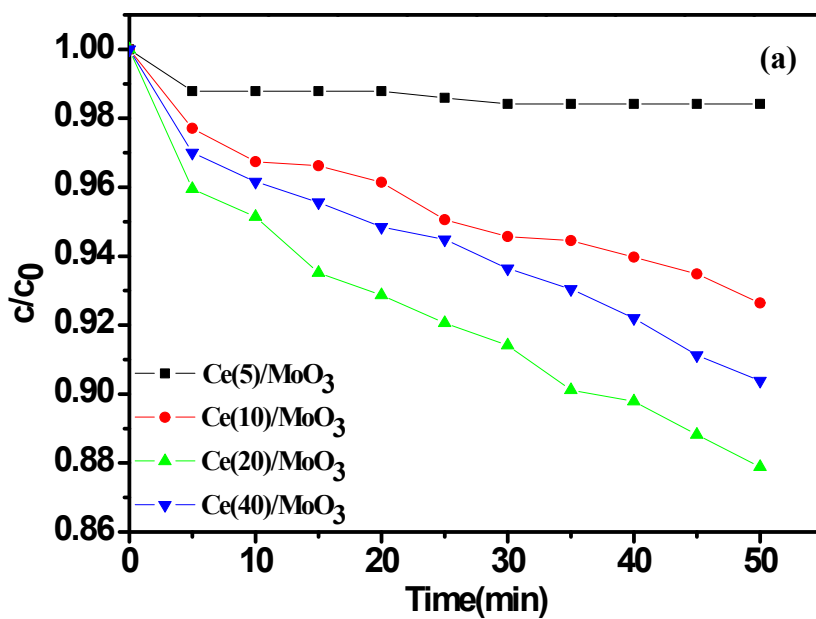


Figure S3. Degradation curves (a) and reaction kinetic curves of MO by Ce(x)/MoO₃ catalysts within 50 min at room temperature and normal atmospheric pressure: 200 mL 15 mgL⁻¹ MO aqueous solution

Table S1 Apparent reaction kinetic constants (k_a) for the degradations of MO dye in single MO solution and MB-MO mixture dye solution by Ce(x)/MoO₃ catalysts

Samples	k_a (Single MO solution) (min ⁻¹)	k_a (MO in MB-MO mixture dyes) (min ⁻¹)
Ce(5)/MoO ₃	2.700×10^{-4}	0.98228×10^{-4}
Ce(10)/MoO ₃	1.29×10^{-3}	1.05×10^{-3}
Ce(20)/MoO ₃	2.51×10^{-3}	2.33×10^{-3}
Ce(40)/MoO ₃	1.68×10^{-3}	1.68×10^{-3}

After reacting for 50 min

The degradation rates follow the order as follows:

Ce(20)/MoO₃ > Ce(40)/MoO₃ > Ce(10)/MoO₃ > Ce(5)/MoO₃

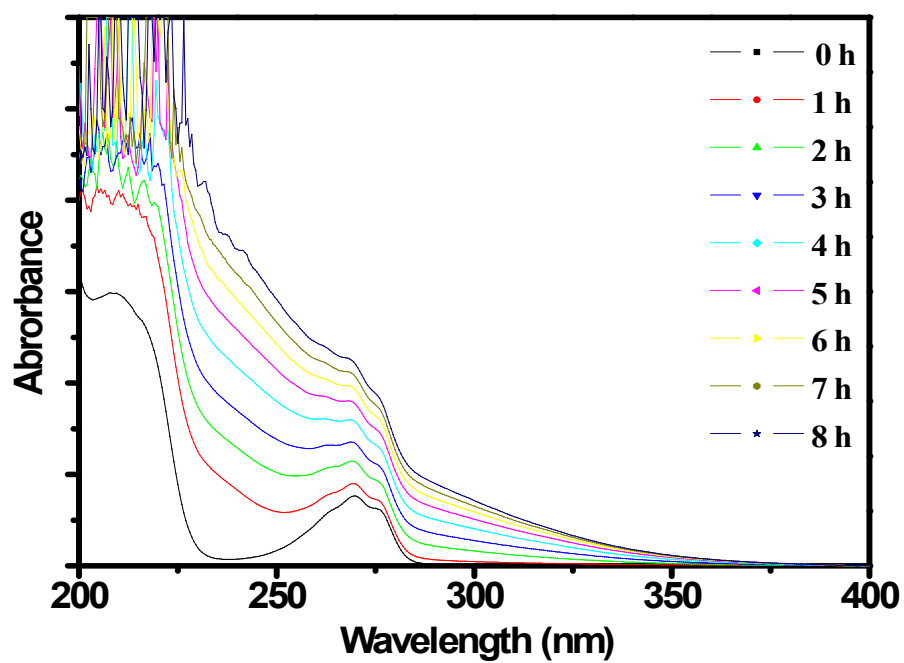


Figure S4. UV-vis absorption spectra of phenol solution over Ce(40)/MoO₃ at different reaction times: 200mL 10 mgL⁻¹ phenol in the ethanol-water solution ($V_{\text{ethanol}}/V_{\text{water}}=2/5$)

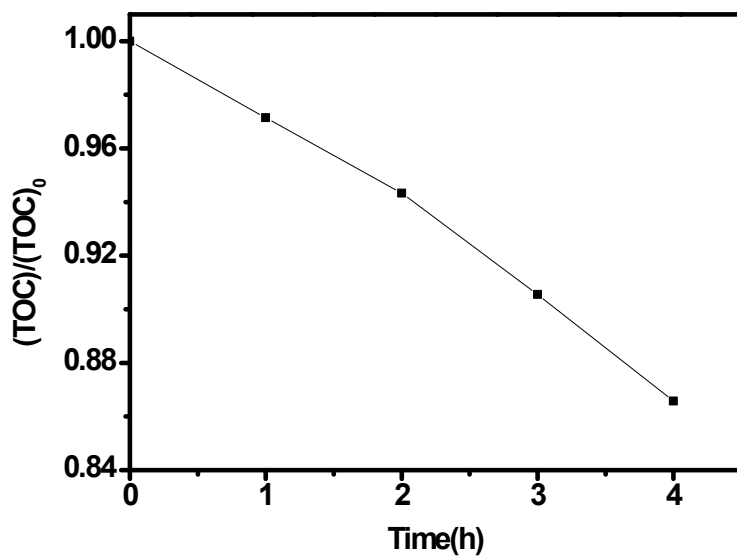


Figure S5. Changes of TOC during the degradation of MB over Ce(5)/MoO₃

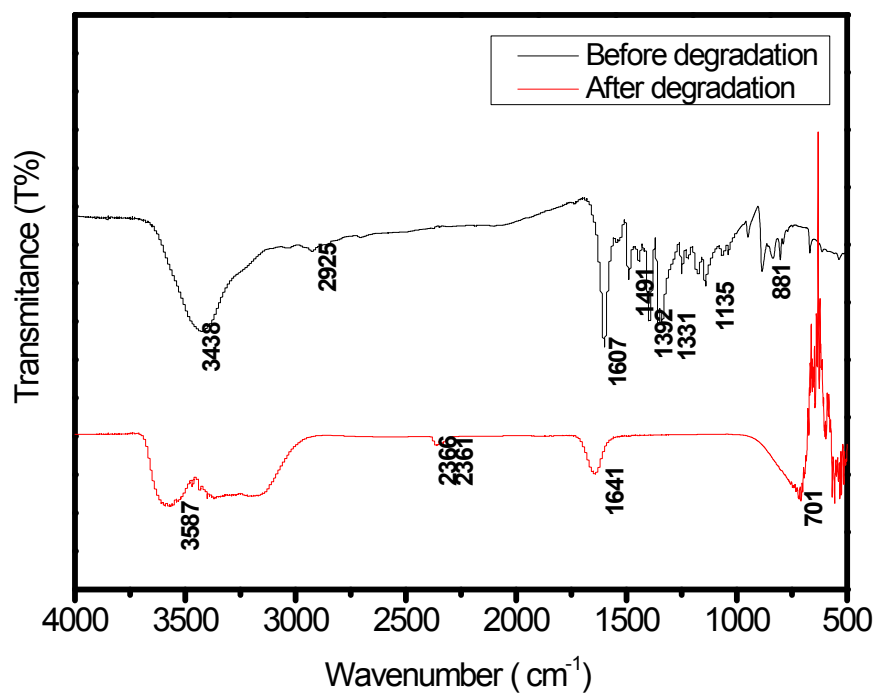


Figure S6. FT-IR spectra of MB before and after degradation

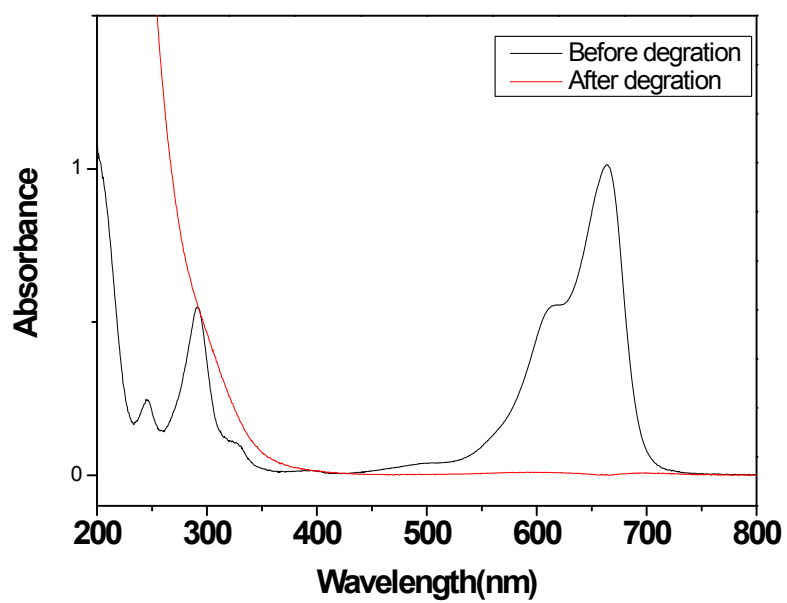
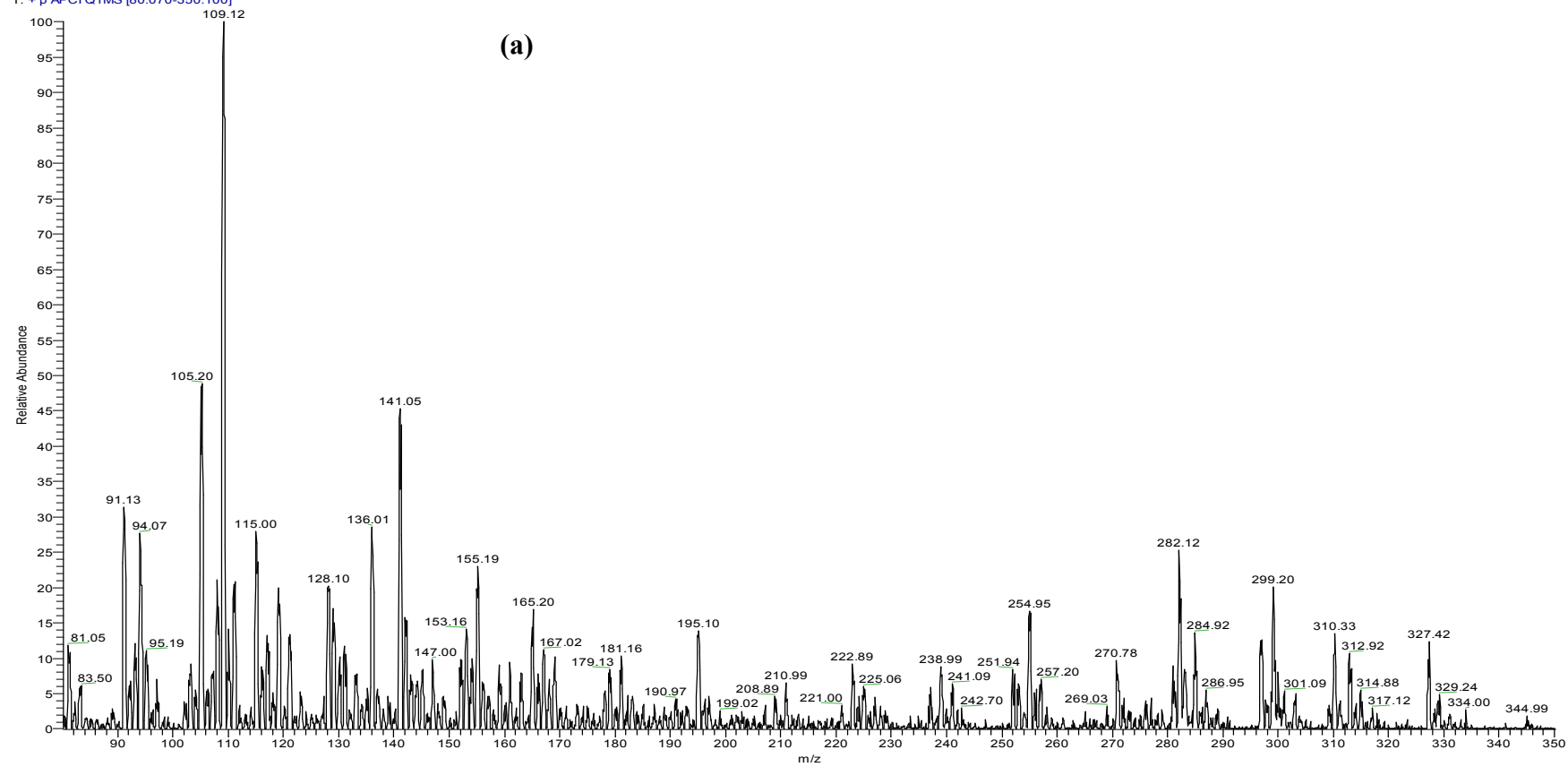


Figure S7. UV-vis absorption spectra of MB solution before and after degradation

B #1 RT: 0.00 AV: 1 NL: 2.81E5
T: + p APCI Q1MS [80.070-350.100]



test_140522173026 #1 RT: 0.01 AV: 1 SM: 15B NL: 1.54E5
T: + p APCI Q1MS [80.070-350.000]

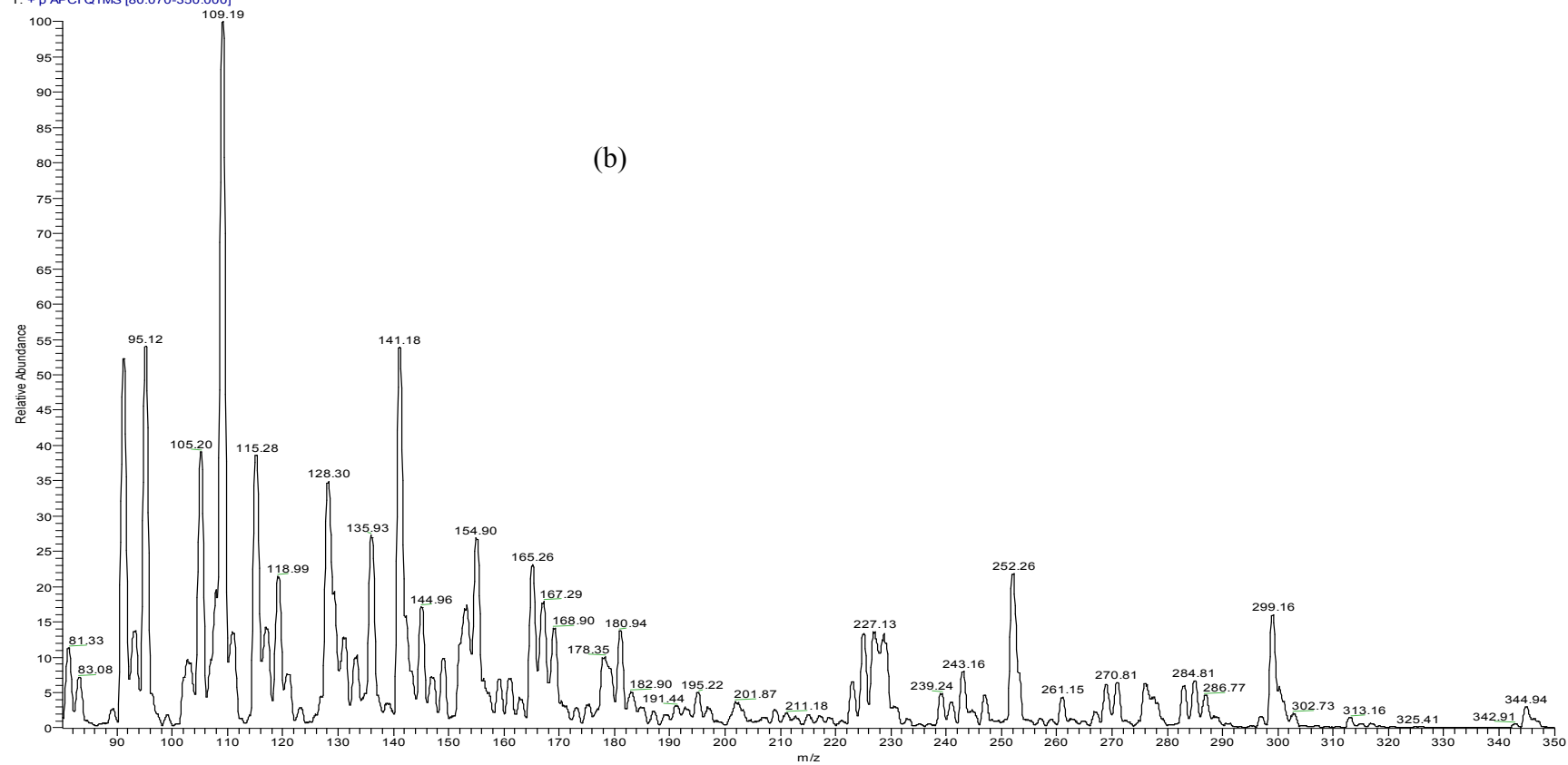


Figure S8. Mass spectra (MS) of MB: (a) background; (b) after degradation by Ce(5)/MoO₃