

Supplementary Information

Steric hinderance-induced formation of zwitterionic carbonates from alkanolamines and CO₂: highly efficient CO₂ absorbents

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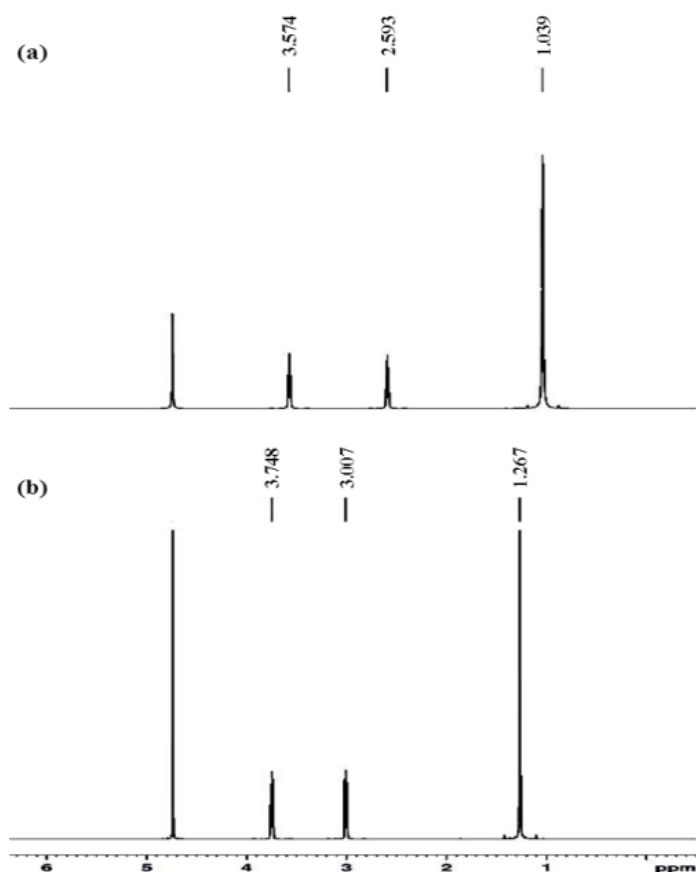


Fig. S-1 ¹H NMR spectra of TBA and TBAE-CO₂ in D₂O.

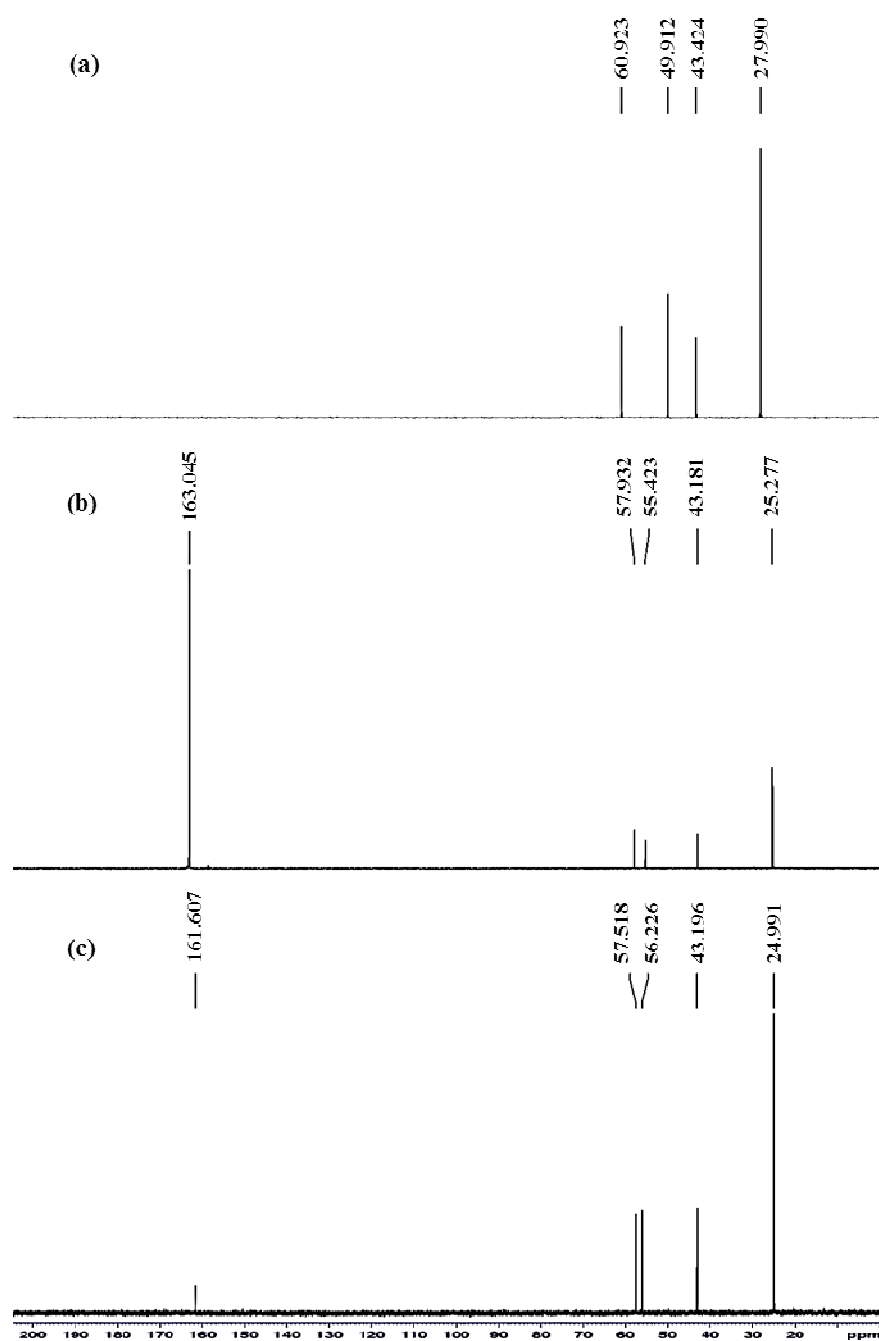


Fig. S-2 ^{13}C NMR spectra of TBA and TBA- CO_2 adducts in D_2O : (a) TBAE, (b) TBAE- $^{13}\text{CO}_2$, (c) TBAE- $^{12}\text{CO}_2$.

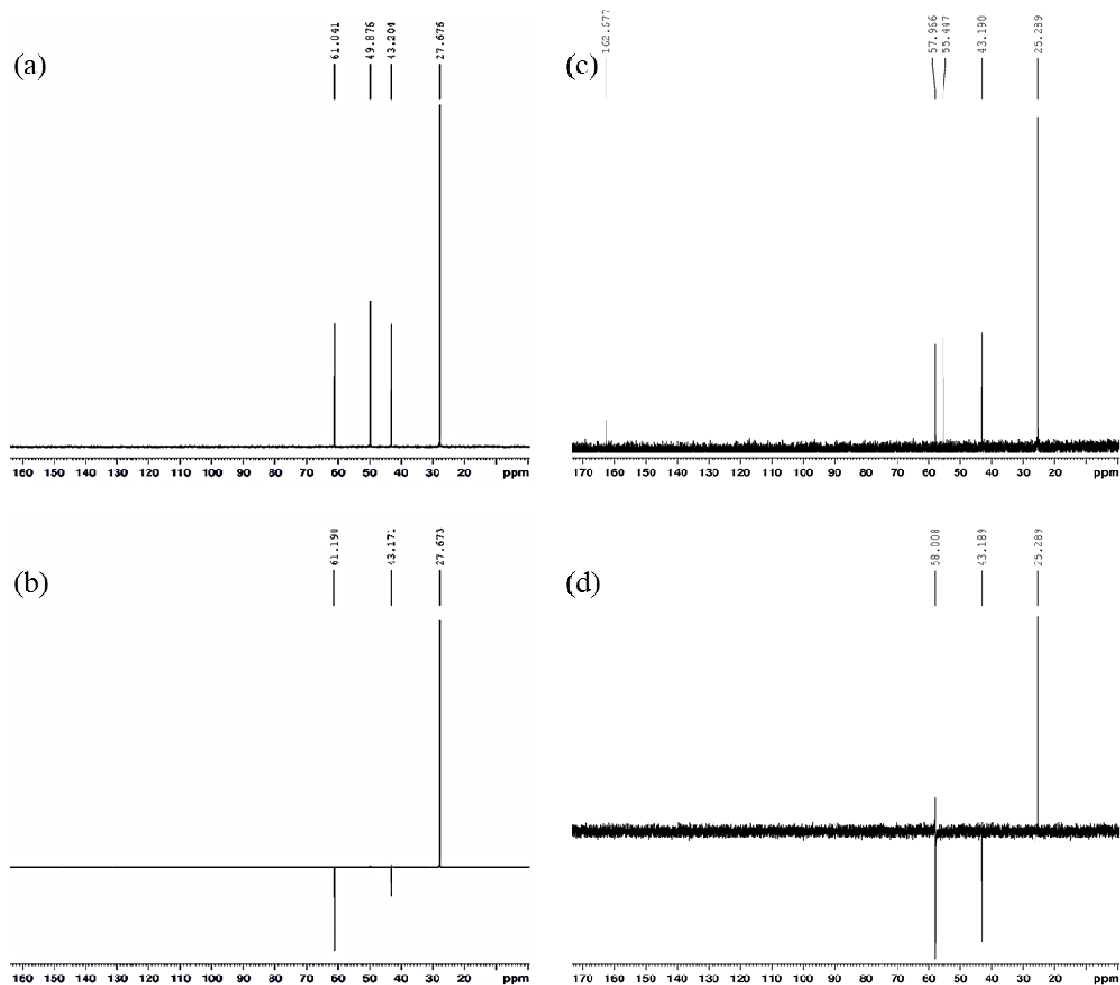


Fig. S-3 DEPT-135 spectra of TBA and TBA- CO_2 adducts in D_2O : (a) ^{13}C NMR of TBAE, (b) DEPT-135 of TBAE, (c) ^{13}C NMR of TBAE- CO_2 , (d) DEPT-135 of TBAE- CO_2 .

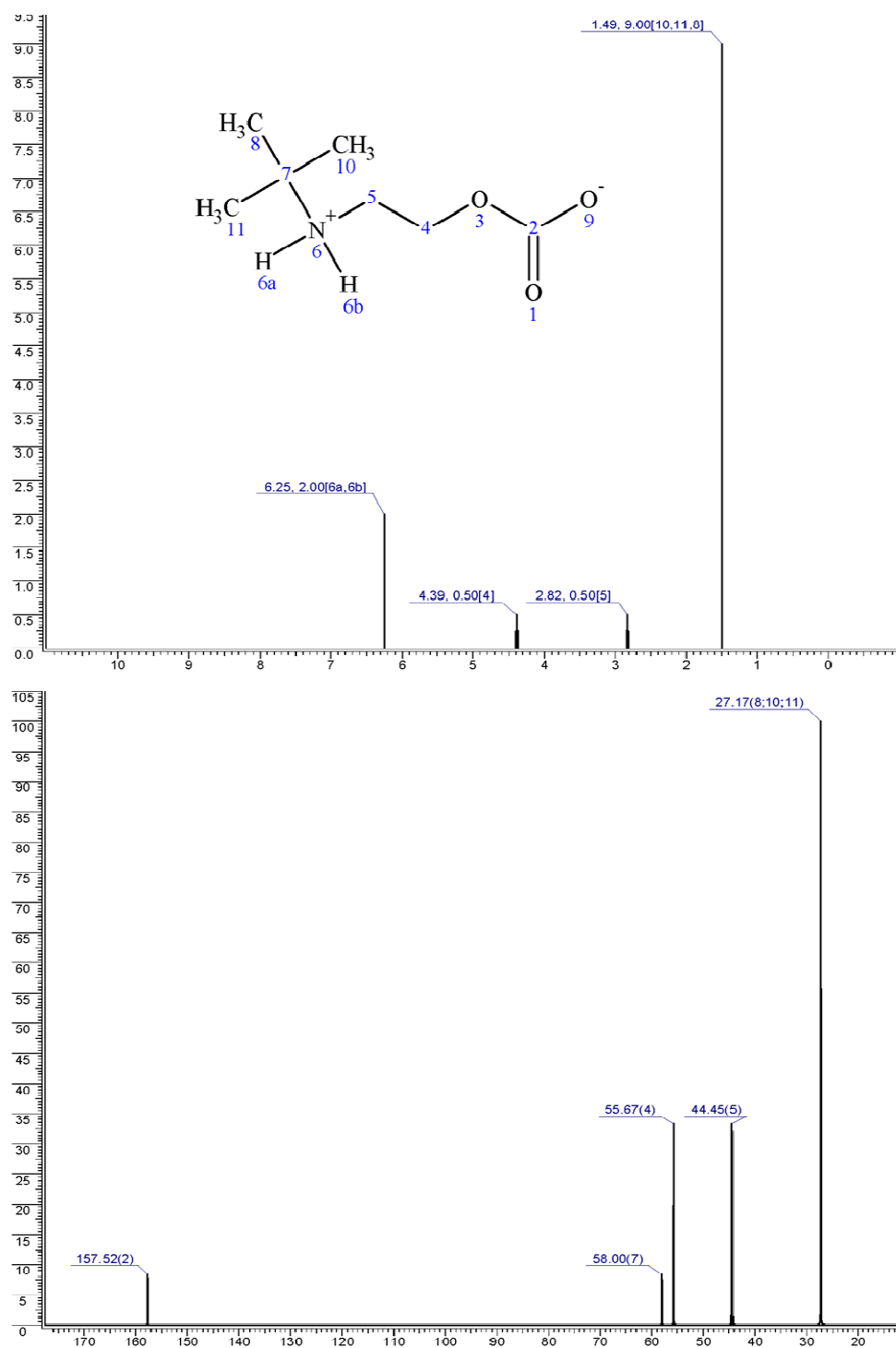


Fig. S-4 Predicted NMR spectra of $t\text{-C}_4\text{H}_9\text{NH}_2^+\text{CH}_2\text{CH}_2\text{OCO}_2^-$ using ACD/NMR DB (v.6.12) software: (top) ^1H NMR, (bottom) ^{13}C NMR.

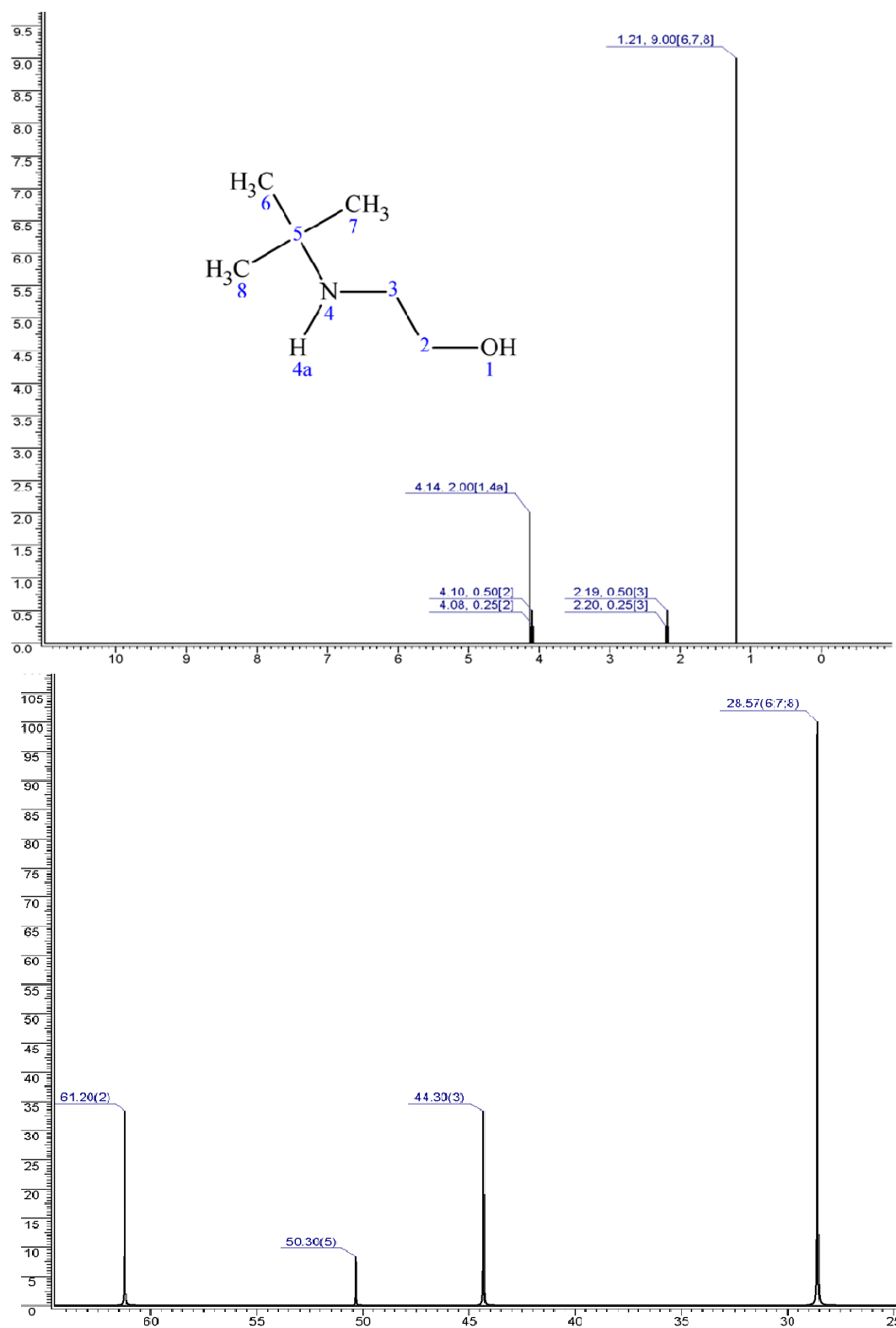


Fig. S-5 Predicted NMR spectra of TBAE using ACD/NMR DB (v.6.12) software: (top) ^1H NMR, (bottom) ^{13}C NMR.



Fig. S-6 A photograph of a high pressure FT-IR cell.

Fig. S-7 Schematic diagram of CO₂ absorption-desorption apparatus

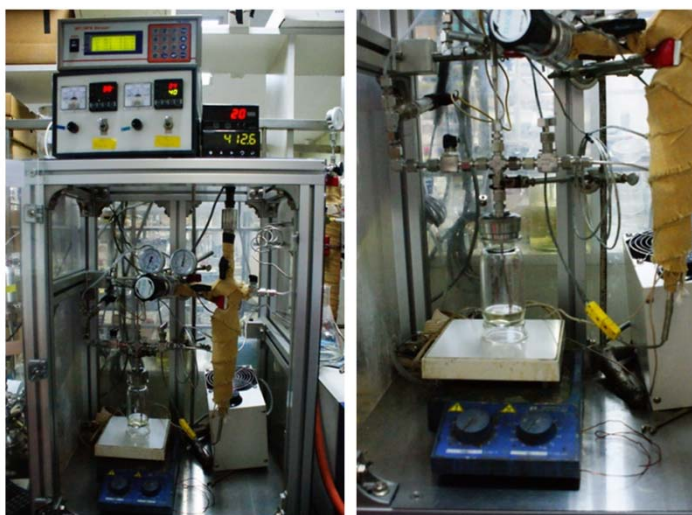


Fig. S-8 A CO₂ absorption apparatus used for the CO₂ absorption capacity of TBAE solution in diethyl ether and for the synthesis of TBAE-CO₂ adduct.

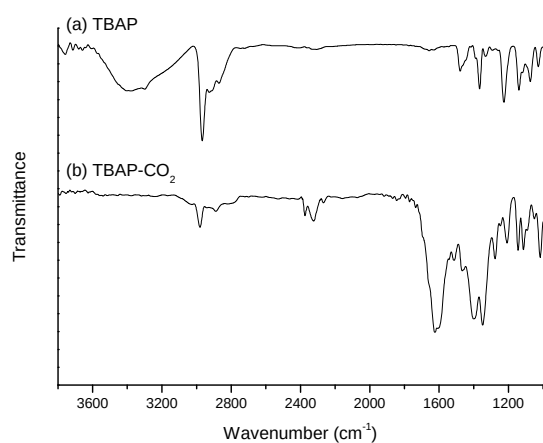


Fig. S-9 IR spectra showing the interactions of neat TBAP with CO₂.

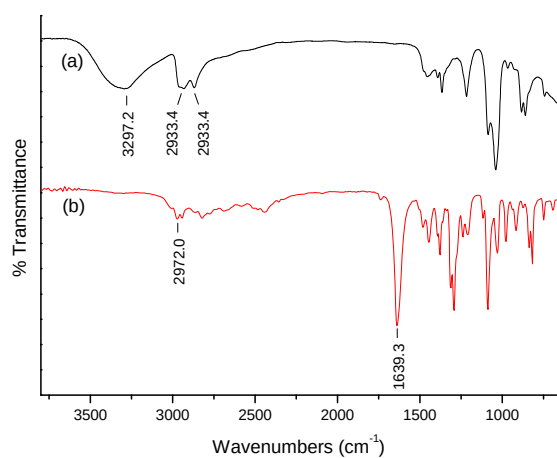
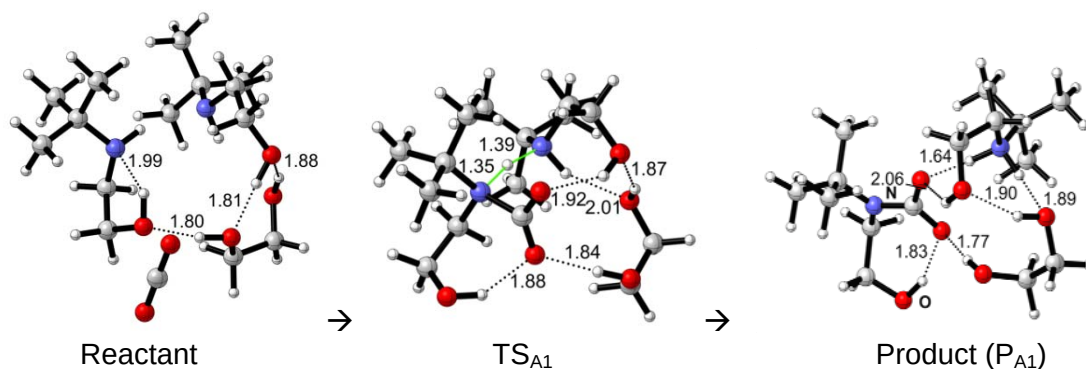


Fig. S-10 FT-IR spectrum of the neat TBAE and TBAE-CO₂ adduct isolated from the reaction of the TBAE with CO₂ in diethyl ether.

A: Carbamate pathway



B: Carbonate pathway

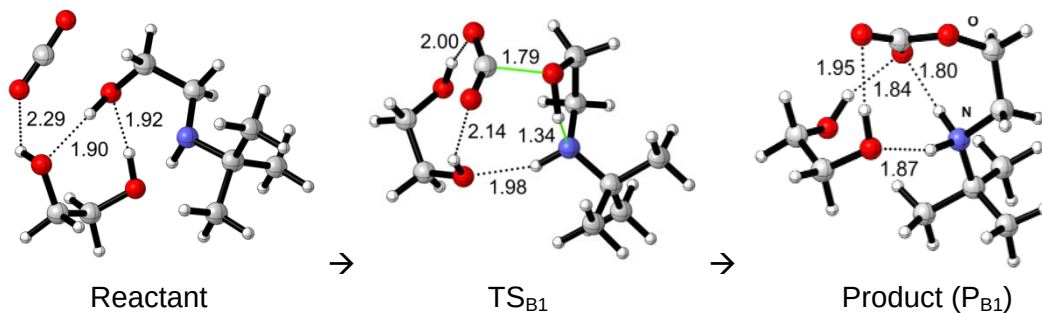
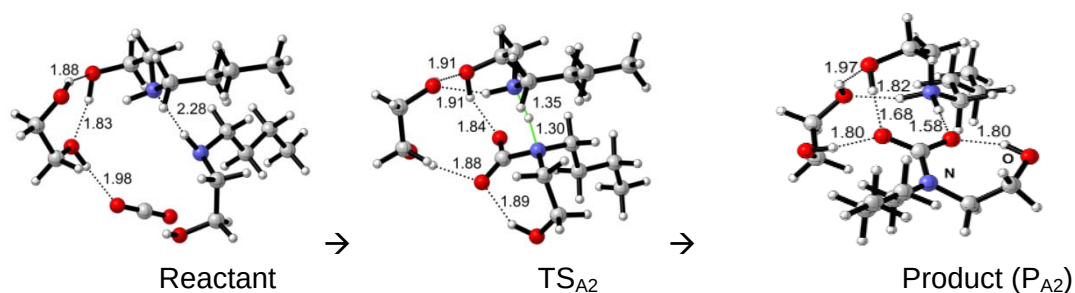


Fig. S-11 Optimized structures showing the interactions of TBAE with CO₂. Pathway A: (TS_{A1}) transition state ($\Delta G^{\ddagger}_{\text{TS}_{A1}} = 30.1$ kcal/mol), (P_{A1}) TBAE-CO₂ adduct ($\Delta G_{\text{P}_{A1}} = 1.4$ kcal/mol); Pathway B: (TS_{B1}) transition state ($\Delta G^{\ddagger}_{\text{TS}_{B1}} = 16.4$ kcal/mol), (P_{B1}) TBAE-CO₂ adduct ($\Delta G_{\text{P}_{B1}} = -2.8$ kcal/mol).

A: Carbamate pathway



B: Carbonate pathway

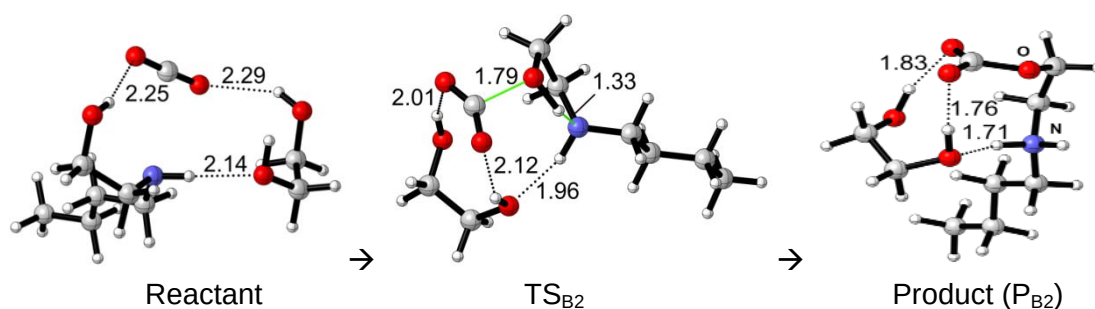


Fig. S-12 Optimized structures showing the interactions of BAE with CO₂. Pathway A: (a) transition state ($\Delta G^\ddagger_{\text{TS}_{A2}} = 3.9$ kcal/mol), (b) BAE-CO₂ adduct ($\Delta G_{\text{P}_{A2}} = -20$ kcal/mol); Pathway B: (c) transition state ($\Delta G^\ddagger_{\text{TS}_{B2}} = 18.7$ kcal/mol), (d) BAE-CO₂ adduct ($\Delta G_{\text{P}_{B2}} = 2.5$ kcal/mol).

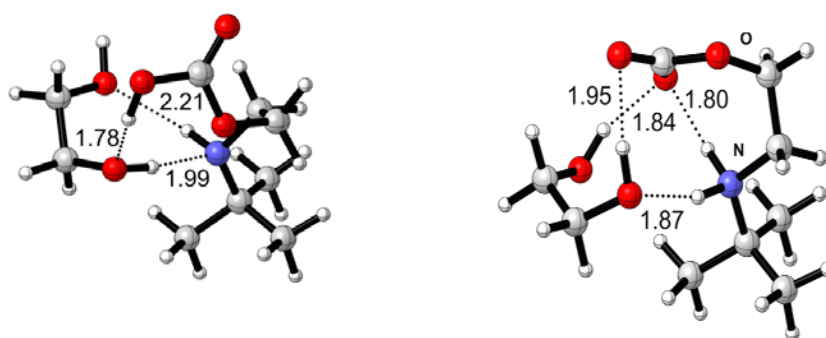
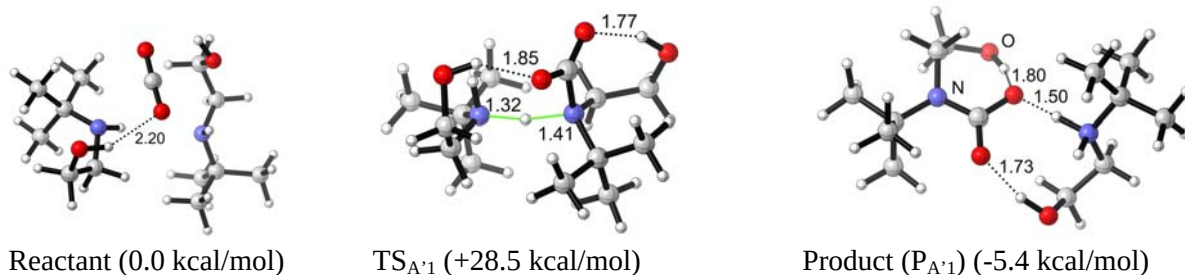
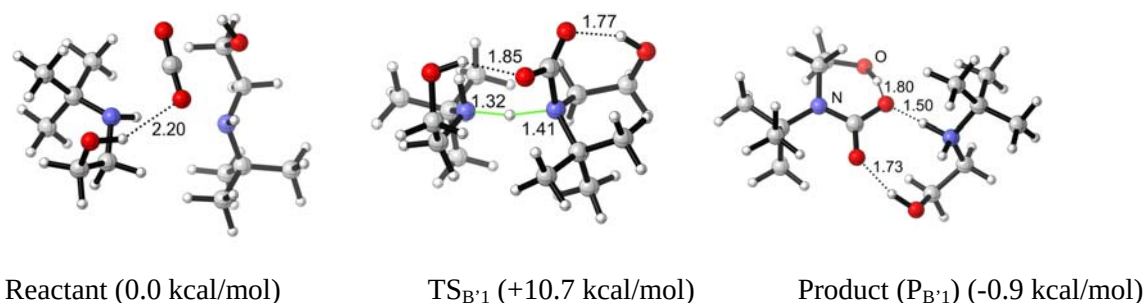


Fig. S-13 Optimized structures of the intermediate carbonic acid species (left, $\Delta G = +4.1$ kcal/mol) and carbonate (right, $\Delta G = -2.8$ kcal/mol) formed from the interaction of TBAE with CO₂.

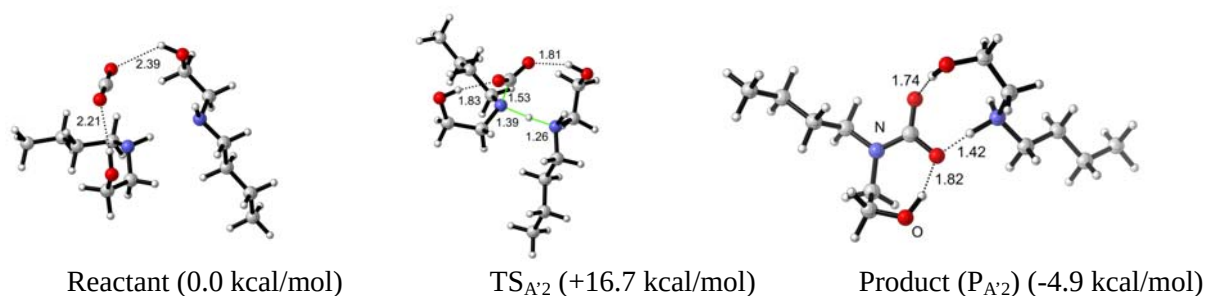
A'1: Carbamate pathway for TBAE in the absence of EG



B'1: Carbonate pathway for TBAE in the absence of EG



A'2: Carbamate pathway for BAE in the absence of EG



B'2: Carbonate pathway for BAE in the absence of EG

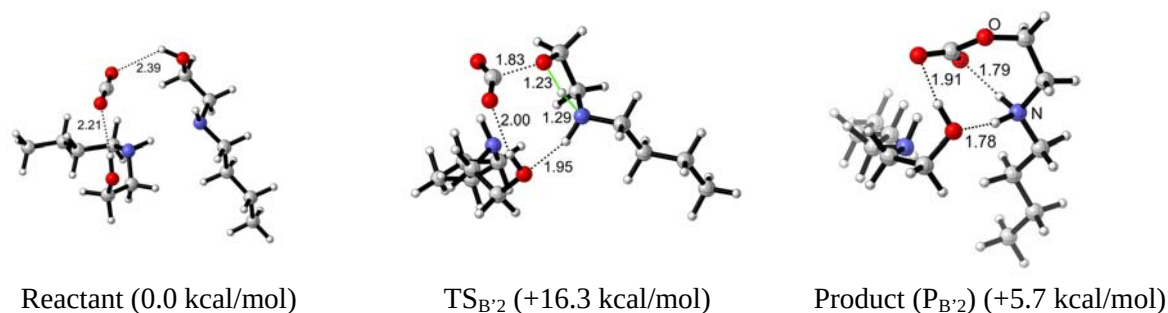


Fig. S-14 Optimized structures showing the interactions of CO₂ with TBAE and BAE in the absence of EG. Pathway A': Carbamate pathway, Pathway B': Carbonate pathway.

Reference [23]

Gaussian 09 (Revision A.1), M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda,

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