Supporting Information for

Magnetic Nanoparticles Supported Ionic Liquid Catalysts for CO₂ Cycloaddition Reactions

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General procedure for the synthesis of catalysts:

Synthesis of precursor 1-3:

The precursor **1-3** were prepared based on the procedure of Shi et al.¹ The compound **4** (2 mmol) and imidazole (1 mmol) were mixed in 20 mL dry toluene and heated to 110 $^{\circ}$ C overnight. The mixture was cooled to room temperature, and washed by Et₂O at least 5 times, and dried in vacuum to give the pure product.





1: ¹H NMR (300 MHz, CDCl₃): δ 0.61-0.66 (2H, m), 1.99-2.03 (2H, m), 3.53 (9H, s), 4.14 (3H, s), 4.31-4.36 (2H, m), 7.51-7.79 (2H, m), 10.45 (1H, m); ¹³C NMR (CDCl₃, 75 MHz): δ 5.8, 24.0, 36.4, 50.5, 51.5, 121.9, 123.8, 137.5;

2: ¹H NMR (300 MHz, CDCl₃): δ 0.31-0.36 (2H, m), 0.66 (3H, t), 1.04-1.12 (2H, m), 1.61-1.75 (4H, m), 3.20 (9H, s), 4.05-4.12 (4H, m), 7.31-7.50 (2H, m), 10.26 (1H, m); ¹³C NMR (CDCl₃, 75 MHz): δ 5.7, 13.2, 19.2, 23.9, 31.9, 49.4, 50.2, 51.3, 121.9, 122.3, 136.9;

3: ¹H NMR (300 MHz, CDCl₃): δ 0.57-0.66 (2H, m), 0.79-0.86 (3H, m), 1.31 (6H, m), 1.93-2.02 (4H, m), 3.53 (9H, s), 4.36-4.38 (4H, m), 7.55-7.74 (2H, m), 10.63 (1H, m); ¹³C NMR (CDCl₃, 75 MHz): δ 5.7, 13.7, 22.1, 23.9, 25.7, 30.0, 30.9, 49.7, 50.2, 51.4, 122.0, 122.2, 137.1.

General procedure for the synthesis of PVP-stabilized magnetite nanoparticles:

Magnetite (Fe₃O₄) particles were prepared in a co-precipitation step based on the procedure of Massart et al² and T. J. Yoon, et al³. FeCl₃:6H₂O (22.0 g, 81.4 mmol) and FeCl₂:4H₂O (8.0 g, 40.7 mmol) were dissolved in 120mL deionized water under argon with vigorous stirring at 85 °C. the pH value of the solution was adjusted to 9 by concentrated NH₃:H₂O. After 4 hours, the magnetite precipitates were washed to pH=7 by deionized water. The black precipitate was collected with a permanent magnet under the reaction flask, and the supernatant was decanted. The sediment was redispersed in 100 mL of deionized water. The PVP aqueous solution (8.8 mL, 25.6 g/L) was added, and stirred for 1 day at room temperature. The PVP-stabilized magnetite nanoparticles were separated by addition of aqueous acetone (H₂O/acetone = 1/10, v/v) and centrifugation at 4000 rpm for 10 min. The supernatant solution was removed, and the precipitated particles were washed by ethanol twice. The particles were dried in vacuum.

General procedure for the synthesis of SiO₂-coated magnetite nanoparticles:

SiO₂-coated magnetite nanoparticles were prepared according to the procedure of Hyeon et al⁴. PVP stabilized magnetite nanoparticles (2.0 g) were dispersed in 400 mL ethanol. $NH_3.H_2O$ (12 mL) and TEOS (4.0 mL) was added successively. After stirring for 24 hours, the black precipitate was collected with a permanent magnet, and rinsed with ethanol three times. The product was dried and

stored in vacuum. The content of nitrogen is less than 0.3 % which is determined by elemental analysis.

Procedure for the synthesis of the MNP-1:

1.0 g SiO₂-coated magnetite nanoparticles were dispersed in 30 mL CHCl₃ by sonication for 1 hour. 0.28 g of **1** (1 mmol) was then added, and the reaction mixture was refluxed for 2 days under argon. After being cooled to room temperature, the products were sedimented on magnet and rinsed thrice with 100 mL CHCl₃. The obtained particles were dried in vacuum. The loading of the catalyst is determined to be 0.60 mmol/g by elemental analysis. IR spectrum (KBr, cm⁻¹): 3343, 2938, 1632, 1079 cm⁻¹.

Procedure for the synthesis of the MNP-2:

1.0 g SiO₂-coated magnetite nanoparticles were dispersed in 30 mL CHCl₃ by sonication for 1 hour. 0.32 g of **2** (1 mmol) was then added, and the reaction mixture was refluxed for 2 days under argon. After being cooled to room temperature, the products were sedimented on magnet and rinsed thrice with 100 mL CHCl₃. The obtained particles were dried in vacuum. The loading of the catalyst is determined to be 0.23 mmol/g by elemental analysis. IR spectrum (KBr, cm⁻¹): 3425, 2938, 2861, 1630, 1084 cm⁻¹.

Procedure for the synthesis of the MNP-3:

1.0 g SiO₂-coated magnetite nanoparticles were dispersed in 30 mL CHCl₃ by sonication for 1 hour. 0.35 g of **3** (1 mmol) was then added, and the reaction mixture was refluxed for 2 days under argon. After being cooled to room temperature, the products were sedimented on magnet and rinsed thrice with 100 mL CHCl₃. The obtained particles were dried in vacuum. The loading of the catalyst is determined to be 0.27 mmol/g by elemental analysis. IR spectrum (KBr, cm⁻¹): 3343, 2929, 2857, 1631, 1083 cm⁻¹.





Table S1 XRD of magnetite nanoparticles

Sample	d (nm)					
Prepared Fe ₃ O ₄	0.296	0.252	0.209	0.170	0.161	0.148
Standard Fe ₃ O ₄	0.296	0.253	0.209	0.171	0.161	0.148

A dry powder sample of Fe_3O_4 was used to analyze the XRD pattern of the nanoparticles. The observed diffraction pattern coincides with the JCPDS database for magnetite. The interlayer spacings (d), calculated using the Bragg equation, agree well with the data for standard magnetite (Table S1).



Figure S2 Magnetic curve of silica-coated magnetite (black) and Catalyst 1 (blue)

The room-temperature magnetization curves of the magnetic nanoparticles before (black line) and after (blue line) functionalization prove that the magnetic nanoparticles are superparamagnetic. The saturation magnetization of samples changed from 11.7 emu/g to 4.2 emu/g because of the functionalization by precursor **1**.

Figure S3 TEM images (all the samples were prepared in MeOH).



A: MNP; the size of the nanoparticles is about 8-10nm. The shape of the nanoparticles is spherical.B: SiO₂-MNP; the thickness of the SiO₂ layer is about 1-2 nm.

Figure S4. Catalyst recycling^[a]



[a]Epichlorohydrin (10 mmol), **MNP-1** (1 mol%), at 140 $^{\circ}$ C, 4 h, 1 MPa CO₂ pressure, and the results are determined by GC.

Reference:

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NMR spectra:















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IR spectra of the compound:



