

Supporting Information for

Minimisation of E-Factor in the Synthesis of *N*-Hydroxylamines: The Role of Silver (I)-based Coordination Polymers in Green Chemistry

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E-Factor, ^{S1-S6} defined as the mass ratio of waste to desired product. Therefore, as the E-Factor increases the amount of waste increases and, lead to greater negative environmental impact. The ideal E-Factor is 0. Put quite simply, it is kilograms (of raw materials) in, minus kilograms of desired product, divided by kilograms.⁷

Following the important role of environmental acceptability in organic synthesis, our work based on invention of a method to minimize the E-Factor. Calculation of this factor and comparison with literature review show this is achieved.

For entry e:

$$E = \frac{0.208 \text{ g (UHP)} + 1 \text{ mL} \left(0.71 \frac{\text{g}}{\text{mL}} \right) \text{ (ether)} + 0.358 \text{ g (entry e)} - 0.81 * 0.390 \text{ g}}{0.81 * 0.390 \text{ g}} = 3.04$$

**Same calculation for other entries lead to E-Factors = 2.36-4.29
(entries a:4.29, b:4.21, c:4.00, d:3.15, e:3.04, f:2.36)**

- Ref 5 (*J. Org. Chem.*, 1990, **55**, 1981-1983)

$$E = \frac{41 \text{ g (amine)} + (2 + 1.2) \left(0.79 \frac{\text{g}}{\text{mL}} \right) \text{ (acetone)} + 0.074 \text{ g (dioxirane)} - 0.76(0.335 \text{ g})}{0.76 * 0.335 \text{ g}} = 170.2$$

S1-R. A. Sheldon, *Chem. Ind. (London)*, **1992**, 903–906

S2-R. A. Sheldon, in *Precision Process Technology*, M.P. C. Weijnen, A. A.H. Drakenburg (Eds.), Kluwer, Dordrecht, 1993, pp. 125–138.

S3-R. A. Sheldon, in *Industrial Environmental Chemistry*, D.T. Sawyer, A. E. Martell (Eds.), Plenum, New York, 1992, pp. 99–119.

S4-R. A. Sheldon, *J. Chem. Technol. Biotechnol.*, 1997, **68**, 381–388.

S5-R. A. Sheldon, *J. Mol. Catal. A: Chemical*, 1996, **107**, 75–83.

S6-R. A. Sheldon, *Pure Appl. Chem.*, 2000, **72**, 1233–1246.

S7-R. A. Sheldon, I. Arends and U. Hanefeld, *Green Chemistry and Catalysis*, 2007, p. 2.

- **Ref 6** (*J. Org. Chem.*, 1999, **64**, 7451-7458)

For step 1:

$$\frac{2.6 \text{ g (amine)} + 40 \text{ mL} \left(0.86 \frac{\text{g}}{\text{mL}} \right) (\text{toluene}) + 2.12 \text{ g (benzaldehyde)} - 4 \text{ g (imine)}}{4 \text{ g}} = 8.78$$

For step 2:

$$\frac{4 \text{ g (imine)} + (20 + 50) \text{ mL} \left(1.33 \frac{\text{g}}{\text{mL}} \right) (\text{CH}_2\text{Cl}_2) + 4.1 \text{ g (mCPBA)} - 3.5 \text{ g (oxaziridine)}}{3.5 \text{ g}} = 27.88$$

For step 3:

$$\frac{0.47 \text{ g (oxaziridine)} + (30) \text{ mL} \left(1.33 \frac{\text{g}}{\text{mL}} \right) (\text{CH}_2\text{Cl}_2) + 8 \text{ mL} \left(1.49 \frac{\text{g}}{\text{mL}} \right) (\text{TFA}) - 0.3 \text{ g}}{0.3 \text{ g}} = 173$$

Therefore net E-Factor = **173+27.88+8.78 = 209.66**

- **Ref 7** (*J. Org. Chem.*, 2000, **65**, 5937-5941)

$$\begin{aligned} E &= \frac{0.388 \text{ g (amine)} + 15 \text{ mL} \left(0.87 \frac{\text{g}}{\text{mL}} \right) (\text{benzene}) + 1.39 \text{ g (oxone)} + 150 \text{ mL} \left(0.79 \frac{\text{g}}{\text{mL}} \right) (\text{MeOH}) - 0.331 \text{ g}}{0.331 \text{ g}} \\ &= 401.83 \end{aligned}$$

- **Ref 9** (*Chem. Commun.*, 2000, 5, 409-410)

$$\frac{0.0068 \text{ g (BINAP)} + 2.5 \text{ mL} \left(0.89 \frac{\text{g}}{\text{mL}} \right) (\text{THF}) + 0.0026 \text{ g (NBu}_4\text{BH}_4\text{)} + 0.092 \text{ (N. oxide)} - 0.076 \text{ g}}{0.076 \text{ g}} = 29.6$$

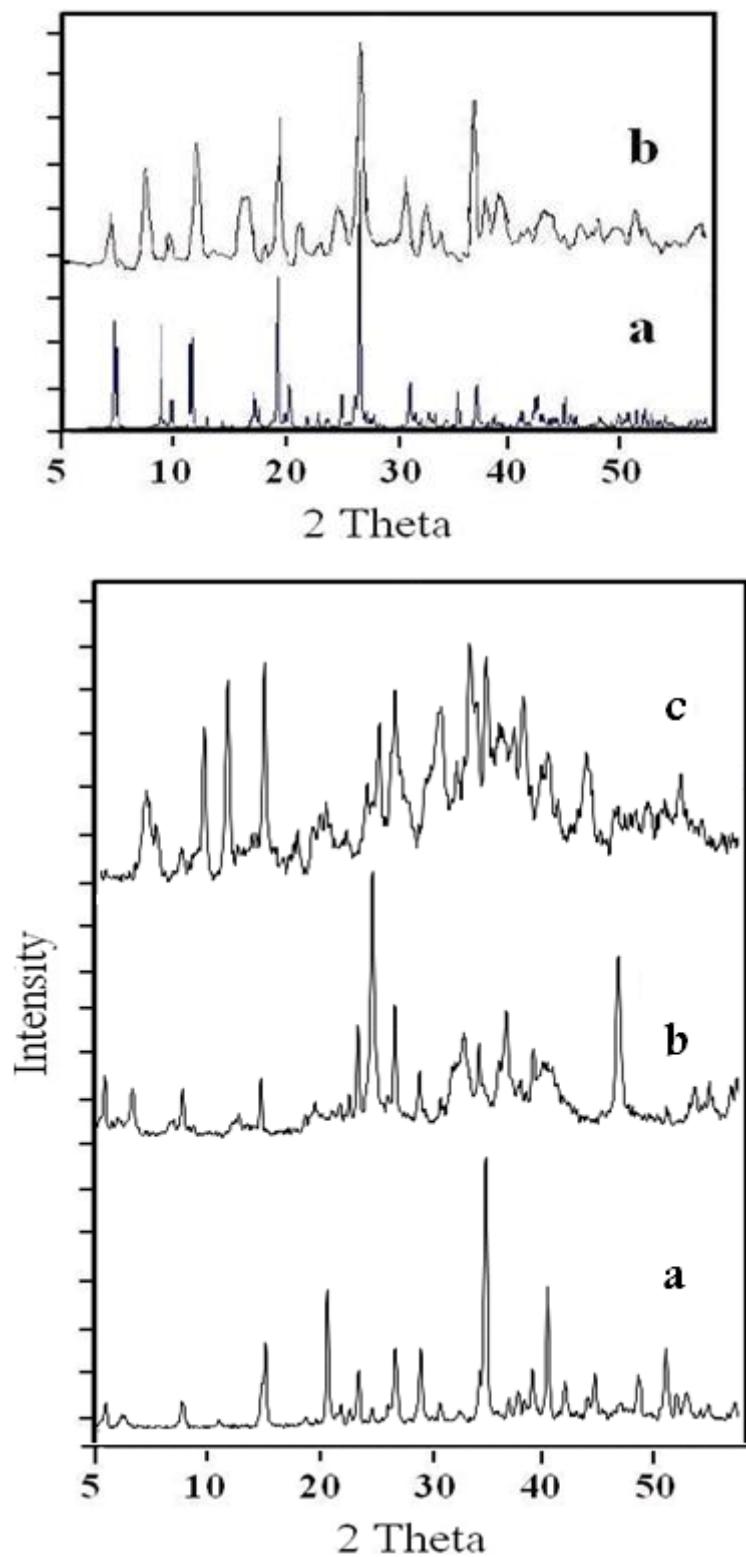
E-Factor = 29.6 (only in hydrogenation step, 2 to 3, without regarding transformation to nitrone)

- **Ref 10** (*J. Org. Chem.* 1994, **59**, 4858-4861)

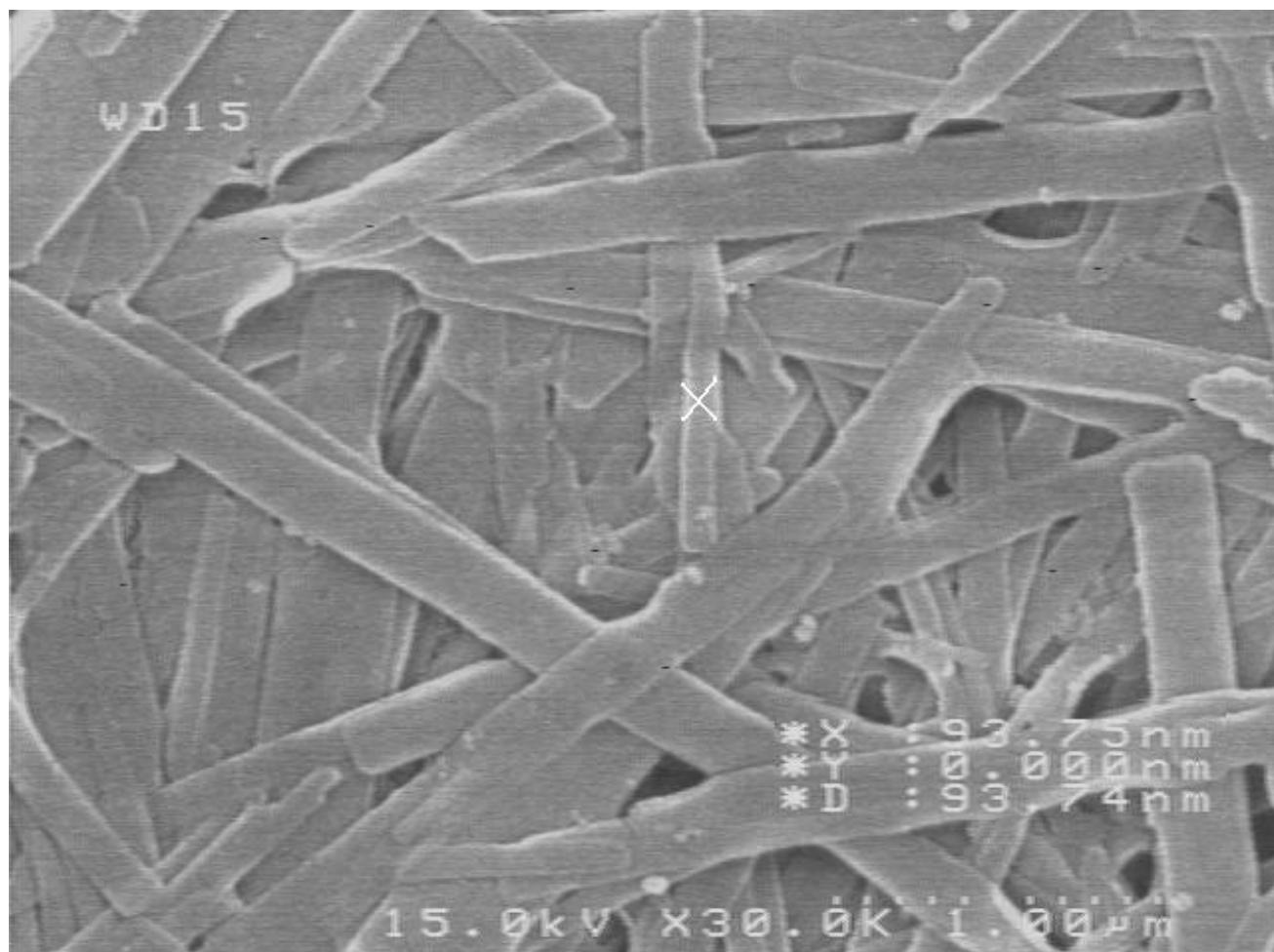
$$\frac{0.183 \text{ g (imine)} + 20 \text{ mL} \left(0.79 \frac{\text{g}}{\text{mL}} \right) (\text{acetone}) + 0.074 \text{ g (dioxirane)} - 0.107 \text{ g}}{0.107 \text{ g}} = 149$$

E-Factor = 149 (until synthesis of nitrone 11 and without regarding its aqueous hydrolysis step by HCl to N-hydroxylamine 12)

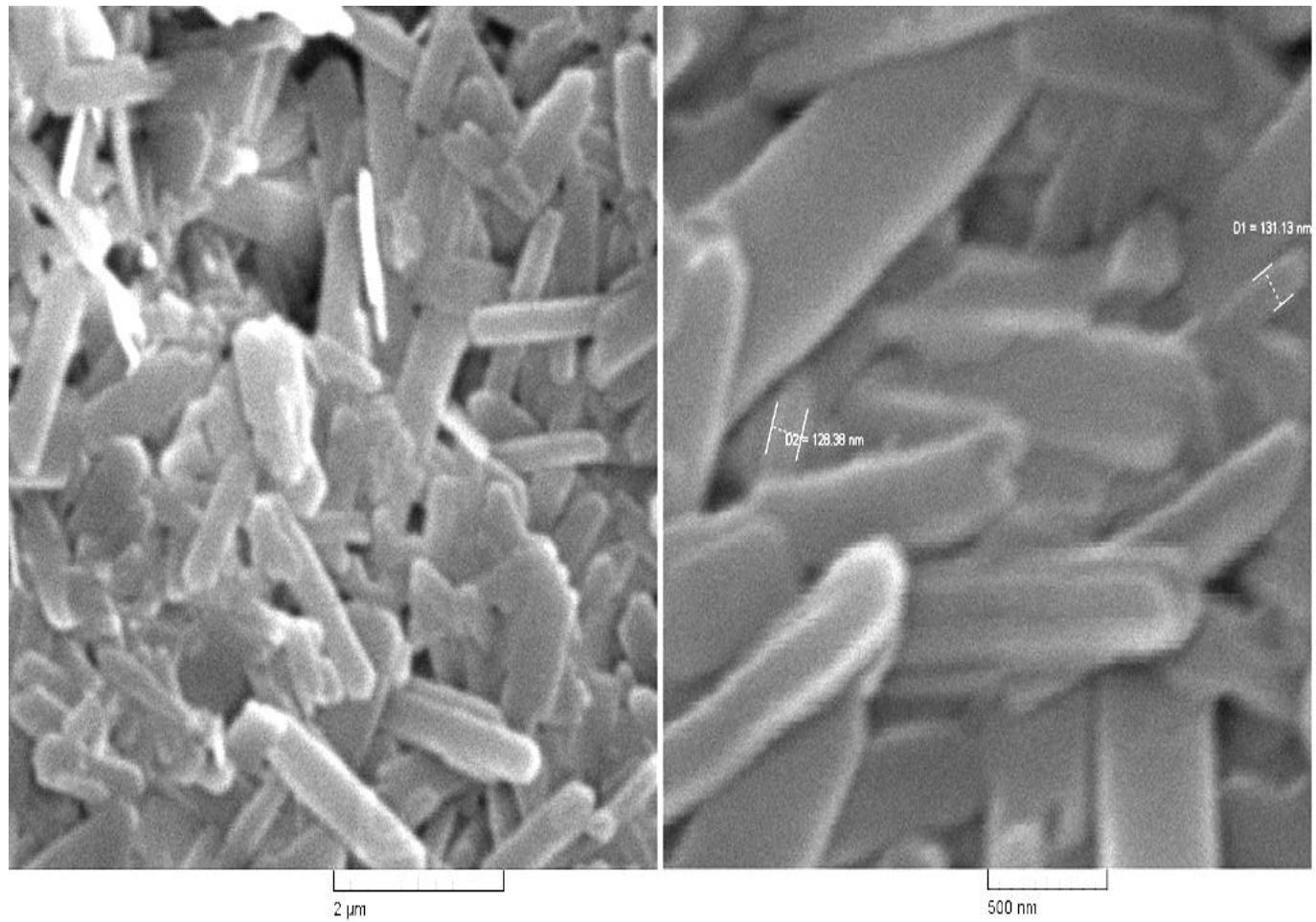
1. XRD patterns of $[\text{Ag}(\mu\text{-bpfb})(\text{NO}_3)]_n$ (top) and $[\text{Ag}(\mu\text{-bpfb})(\text{X})]_n$ ($\text{X}^- = \text{NO}_3^-$, SCN^- , N_3^- and ClO_4^-) 4a-d nano-structures (down)



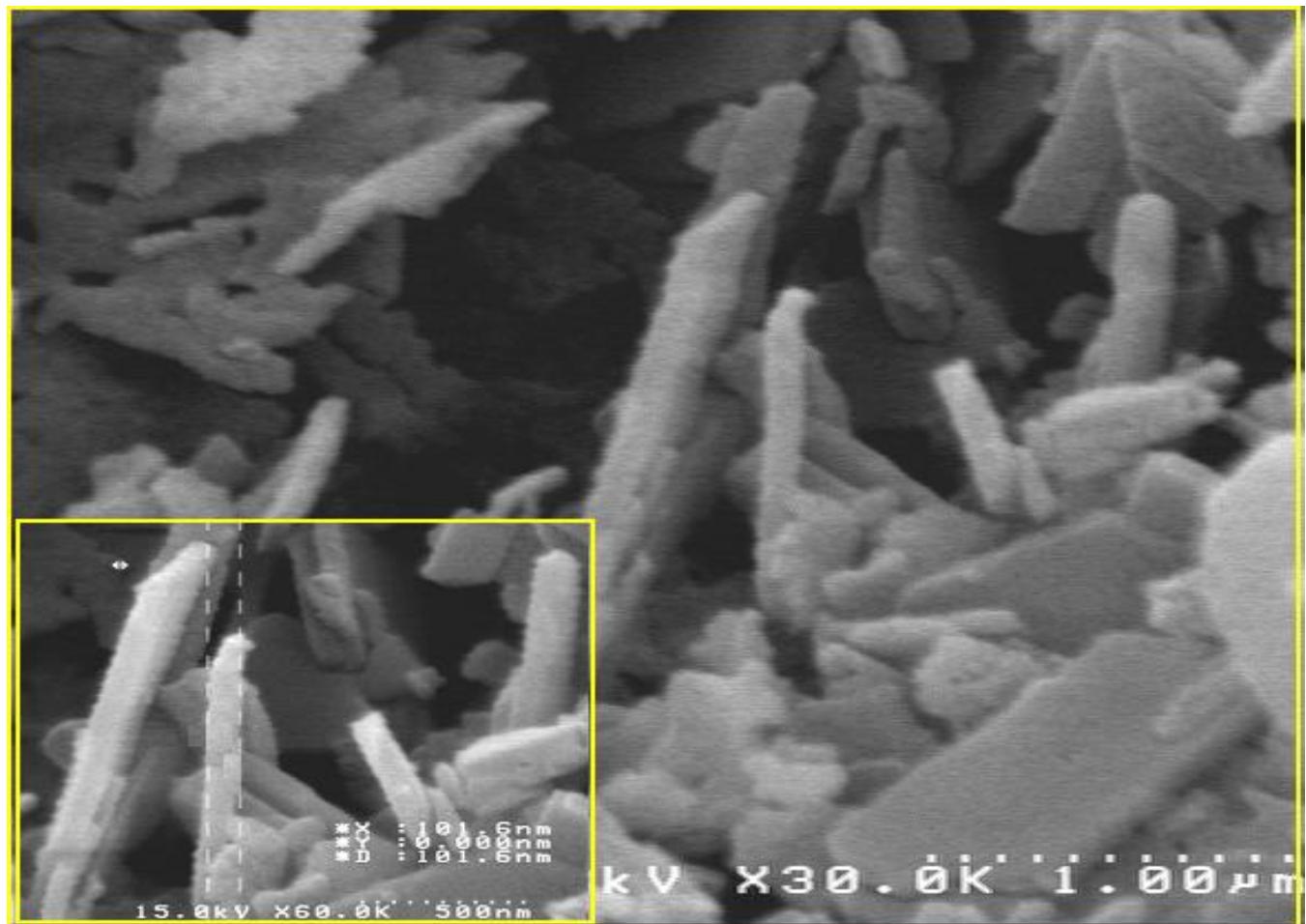
2. SEM images



$[\text{Ag}(\mu\text{-bpfb})(\text{NO}_3)]_n$ (**4a**)



$[\text{Ag}(\mu\text{-bpfb})(\text{SCN})]_n$ (**4b**)

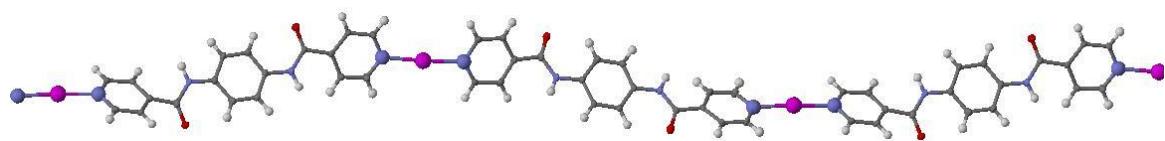


[Ag(μ -bpfb)(N₃)]_n (**4c**)

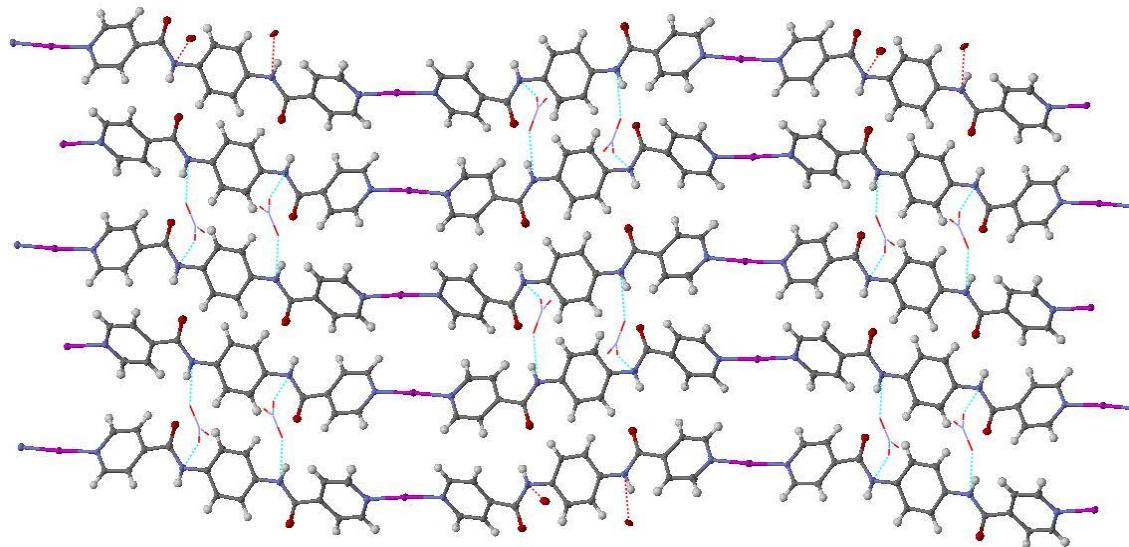


$[\text{Ag}(\mu\text{-bpfb})(\text{ClO}_4)]_n$ (**4d**)

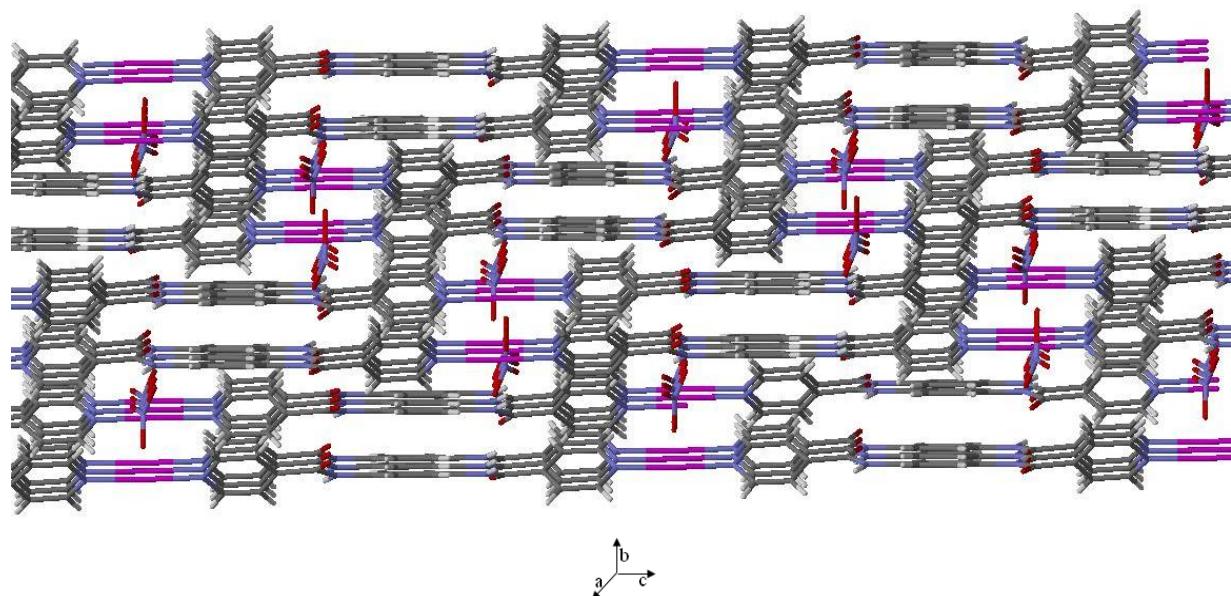
3. X-ray crystallographic results



A fragment of the one-dimensional chain in compound **4a** along crystallographic c axis [001]
(Ag = violet, O = red, C = gray, N = blue).

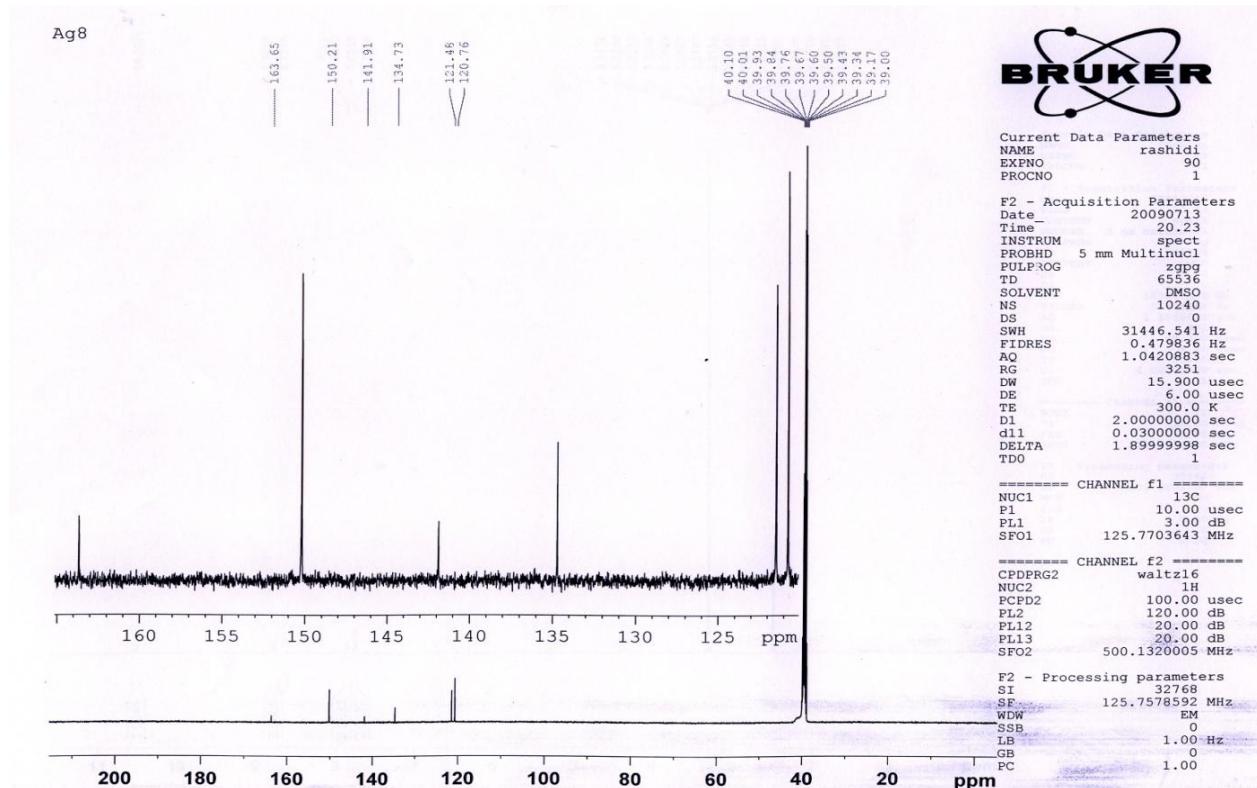
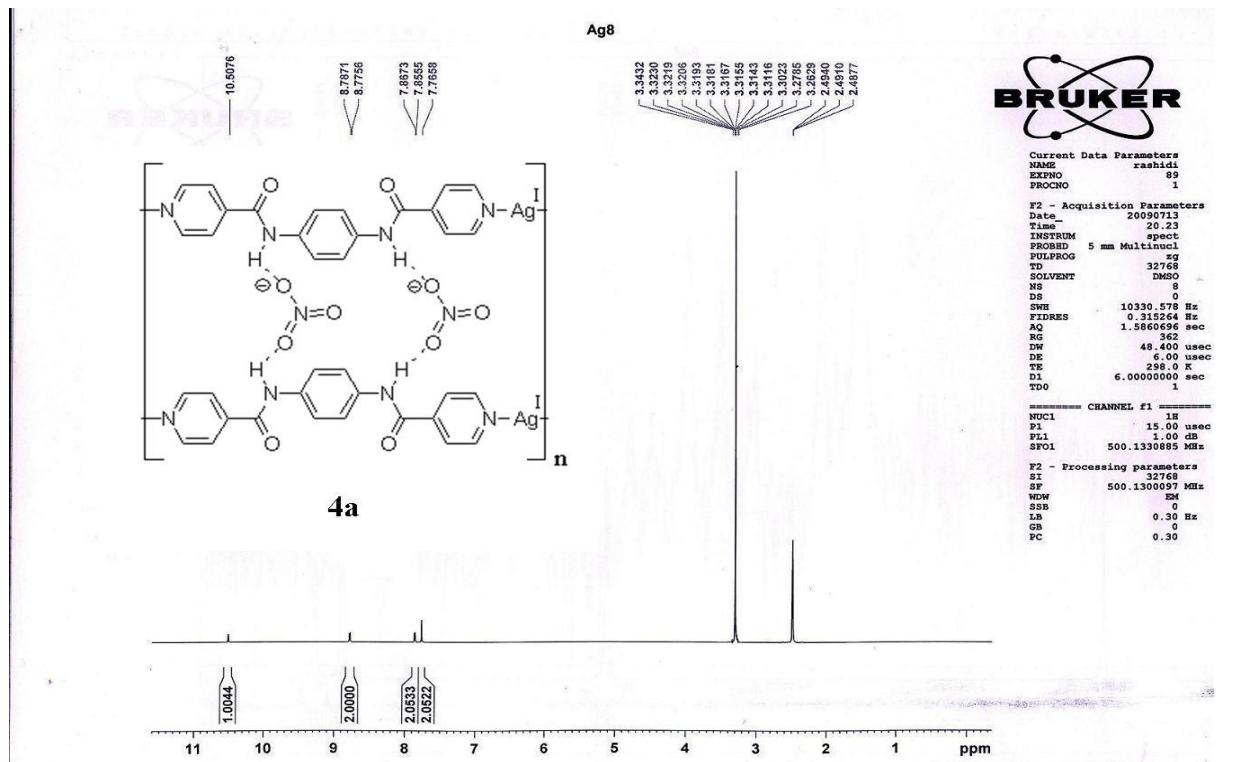


Hydrogen bonds form the 2D coordination polymer (Ag = violet, O = red, C = gray, N = blue).

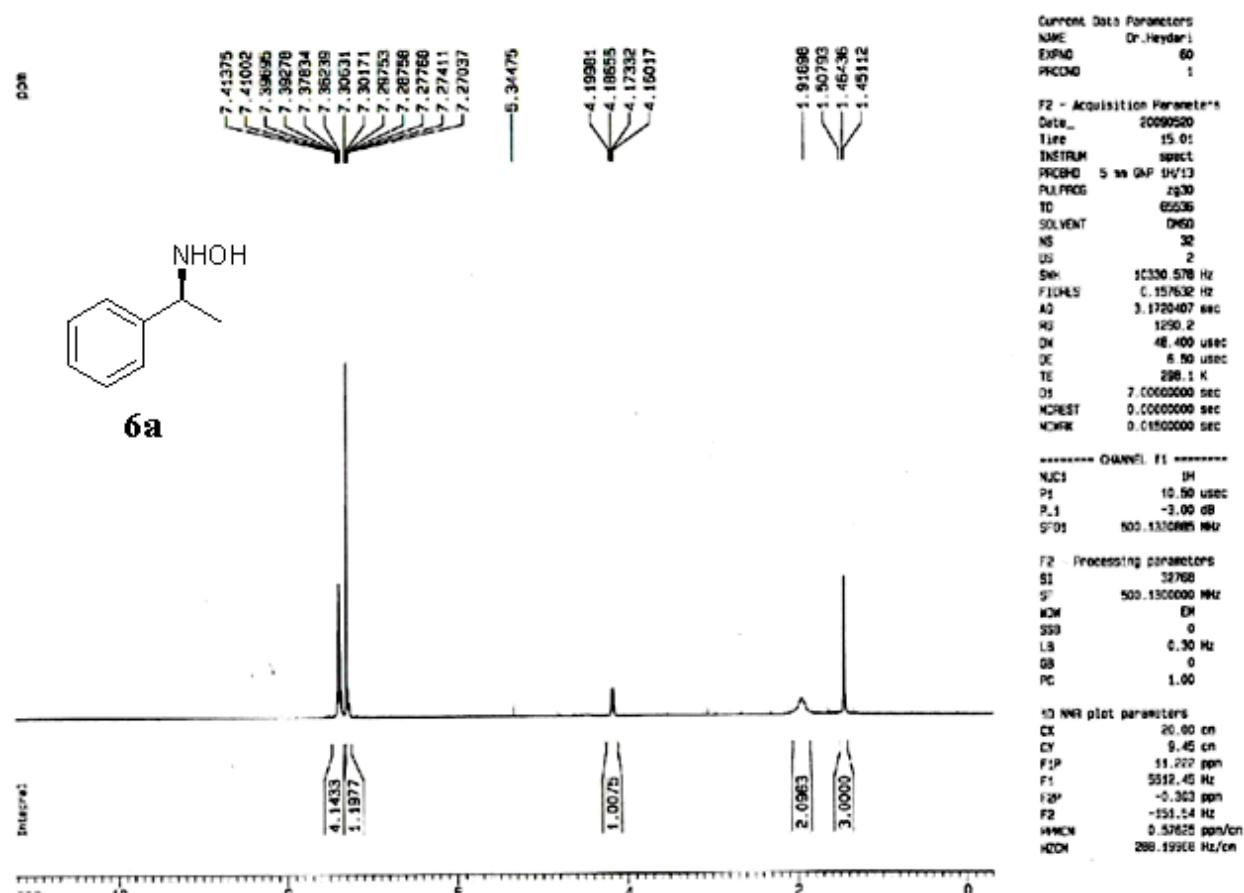


The π - π stacking interaction in supermolecule structure (Ag = violet, O = red, C = gray, N = blue).

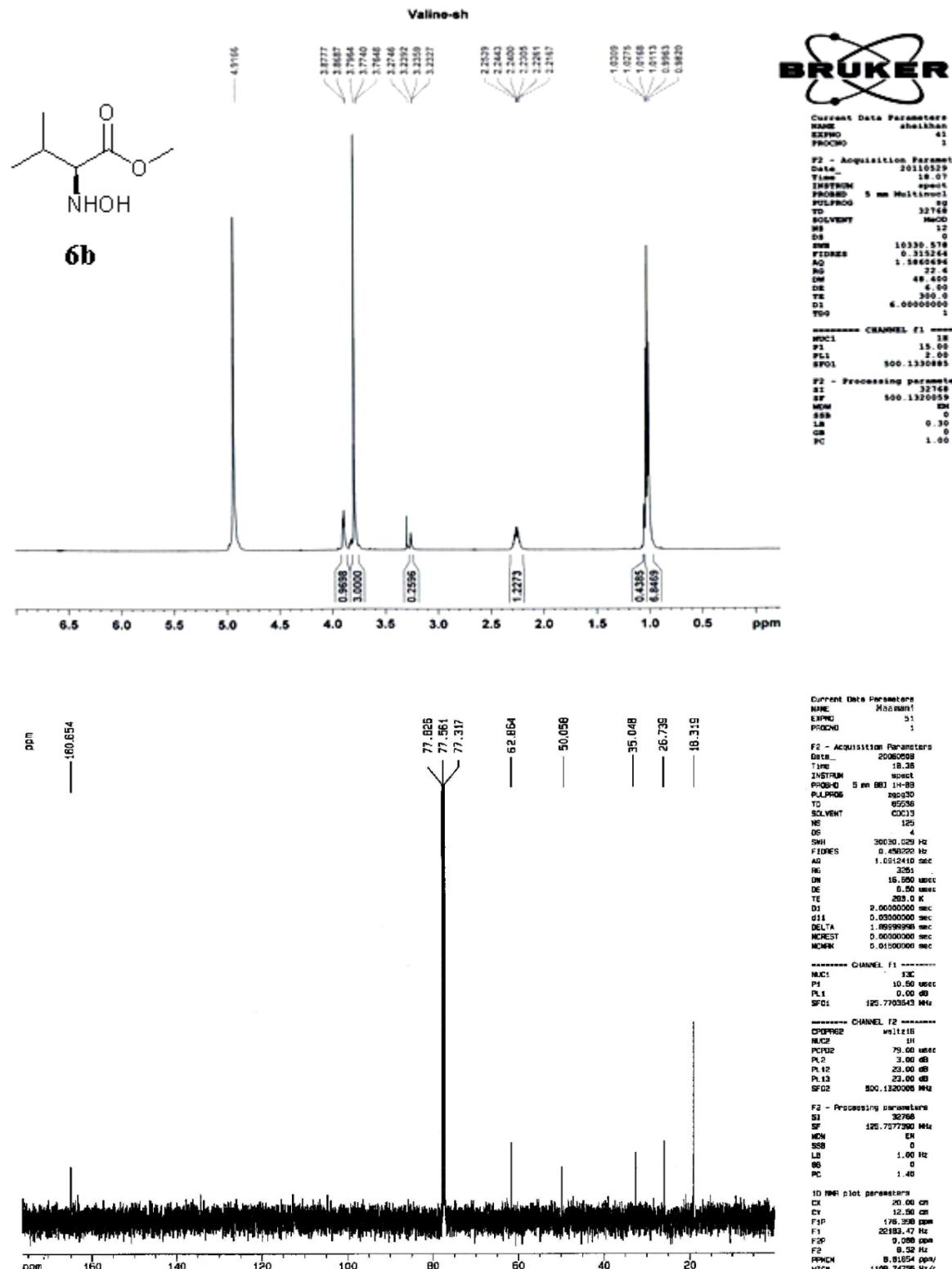
4. ^1H and ^{13}C -NMR of $[\text{Ag}(\mu\text{-bpfb})(\text{NO}_3)]_n$ (4a)



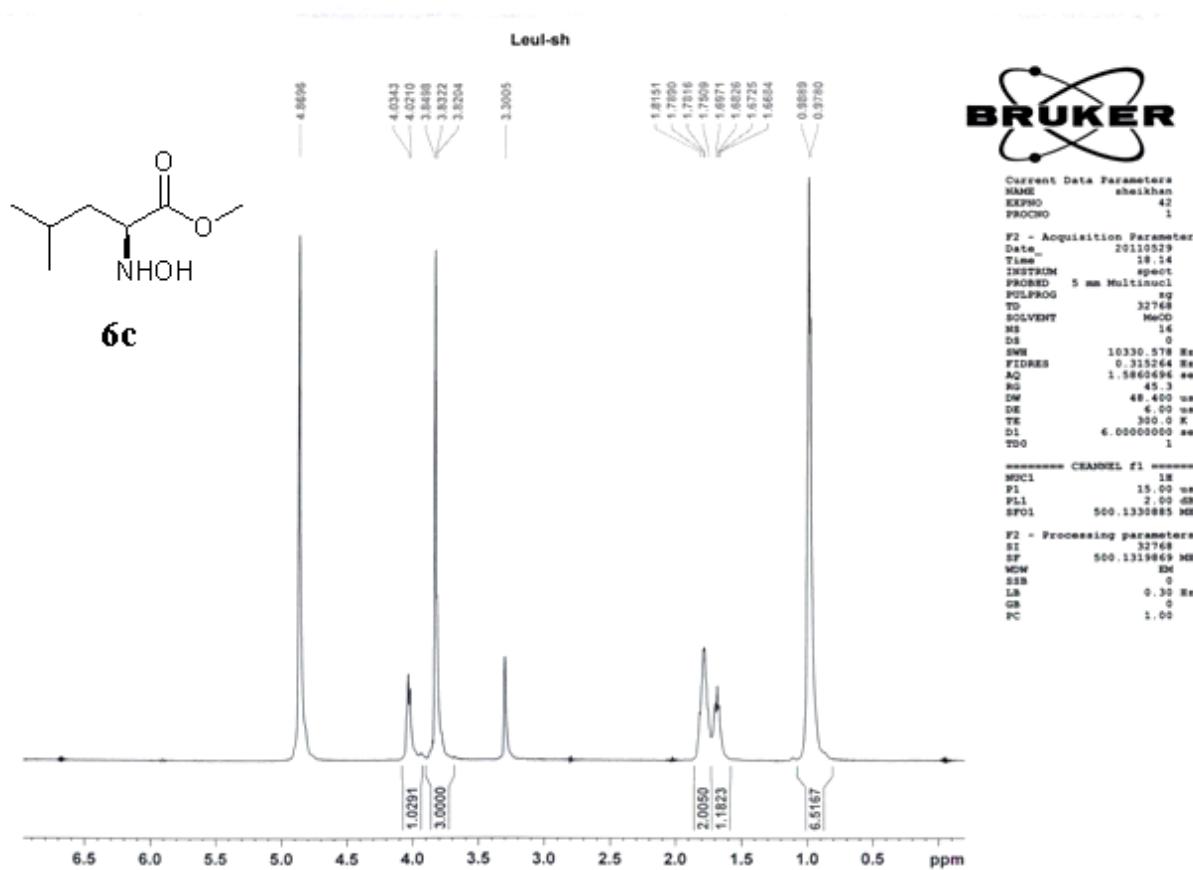
5. ^1H NMR of *N*-monoalkylated hydroxylamines derivatives



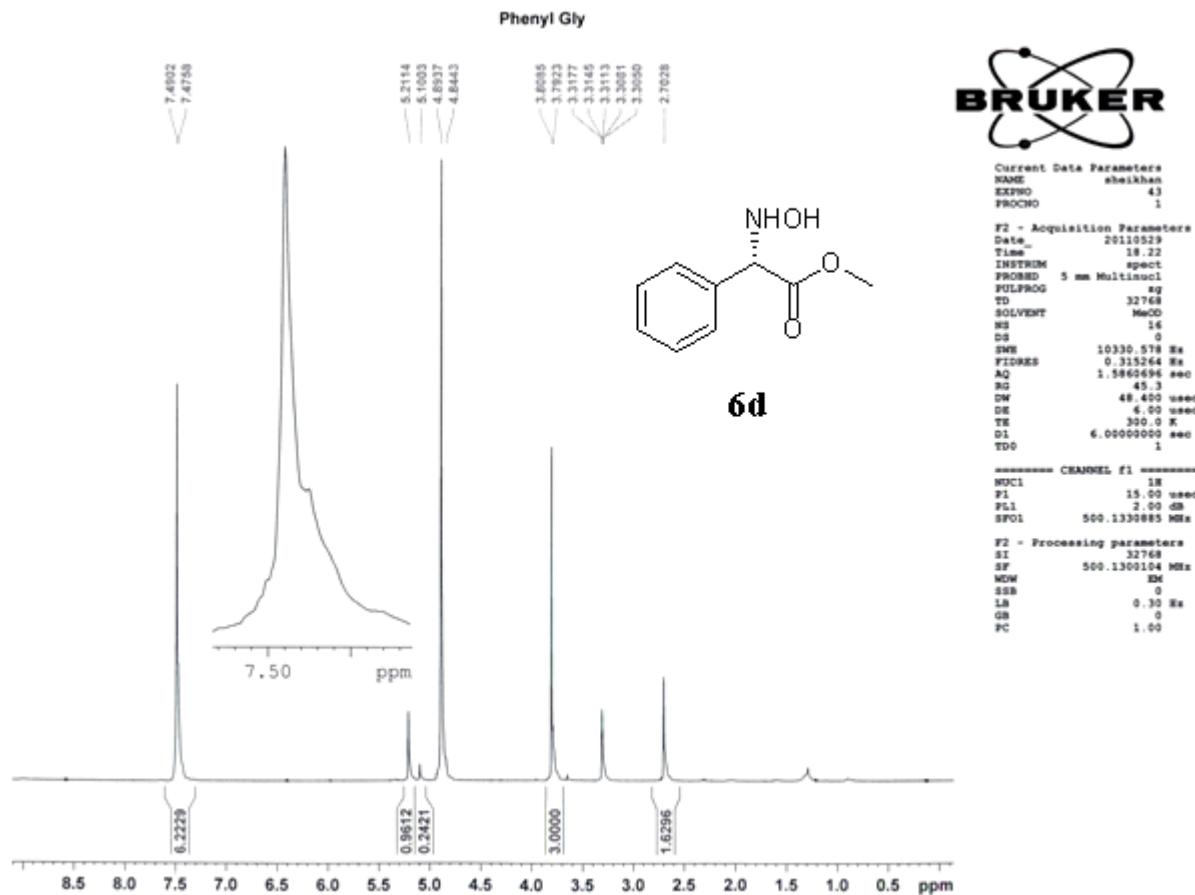
White crystals, mp 69–71 °C [ref 20]



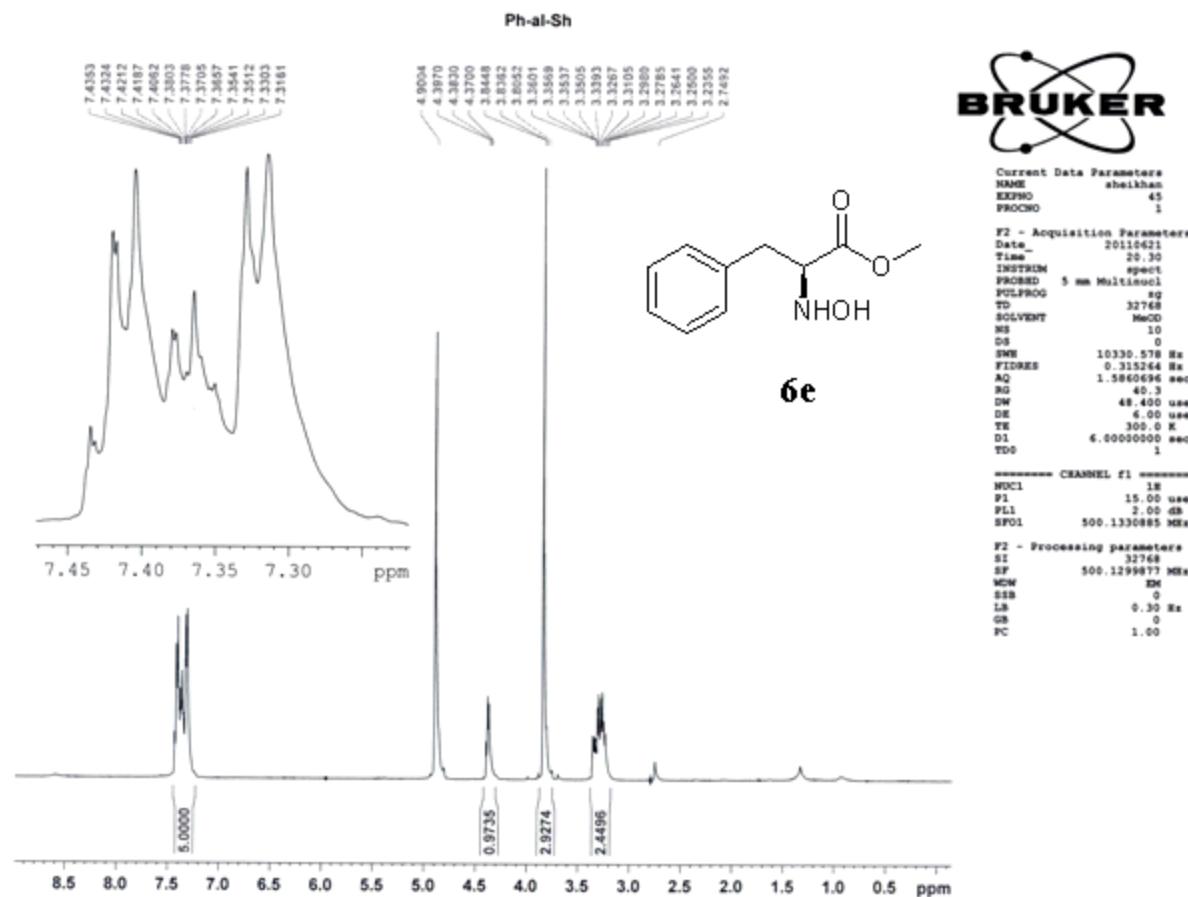
White crystals, mp 158–159 °C [ref 5]



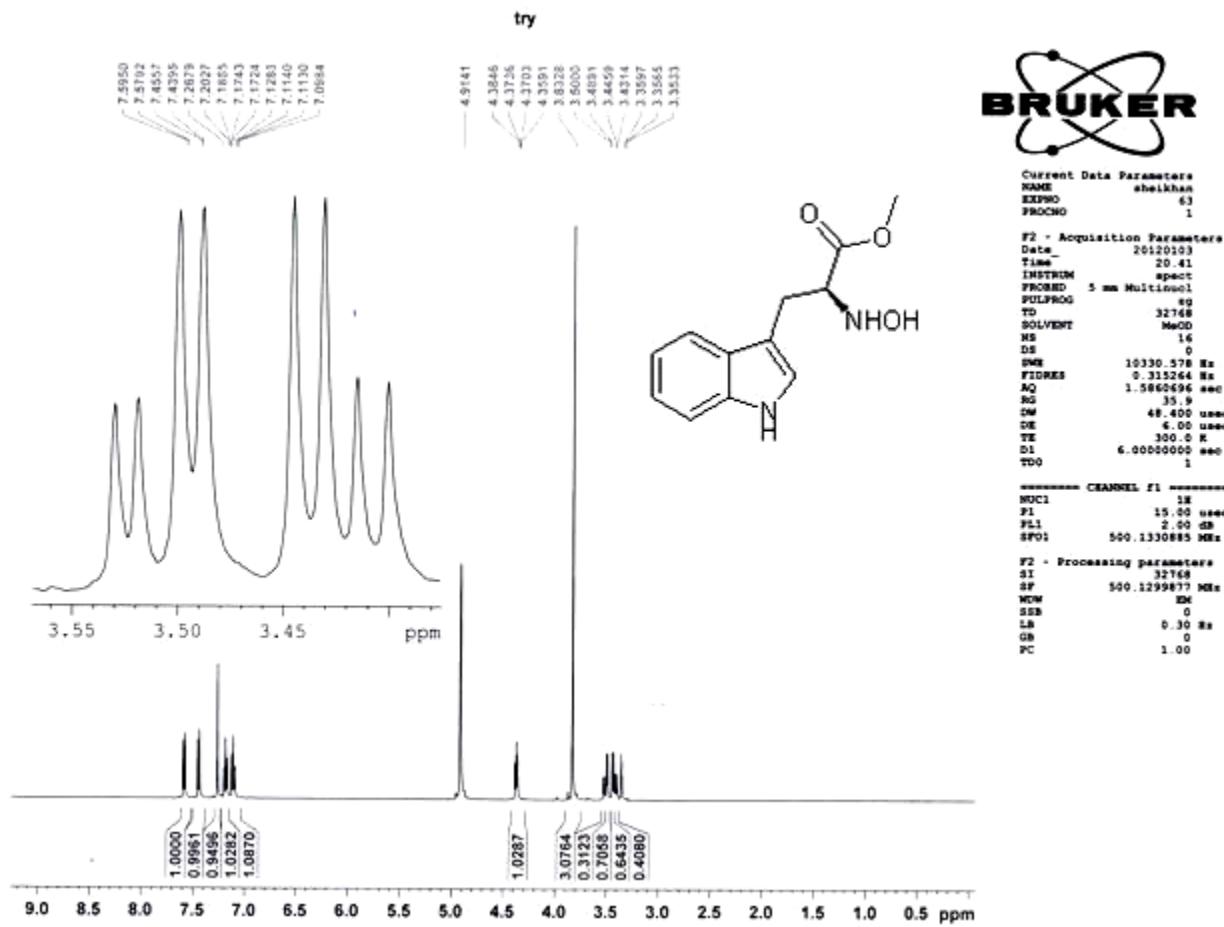
White crystals, mp 65–67 °C [ref 19]



Pale yellow crystals, mp 140–142 °C [ref 4]



White crystals, mp 69–71 °C [ref 5,6]



Pale yellow crystals, mp 82 °C [ref 21]