

Supporting Information

A Ball-Milling Strategy for the Synthesis of Benzothiazole Benzimidazole and Benzoxazole Derivatives under Solvent-Free Conditions

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General

All chemicals and solvents were used as received without any purification. The Fritsch “Pulverisette 7 premium line” (Fritsch GmbH, Idar-oberstein, Germany) planetary ball mill was employed as a reaction chamber with 80 mL grinding cups (tungsten carbide) and milling balls (20×5 mm; tungsten carbide). Prior to the reaction, all vessels and milling balls were cleaned several times with acetone followed by aqua regia to avoid any contamination. ¹H (400 MHz) and ¹³C (100 MHz) NMR spectra were recorded on a JNM-ECS400 JEOL spectrophotometer in DMSO-*d*₆ or CDCl₃ solvent using TMS as an internal reference. The splitting patterns are indicated as s (singlet), bs (broad singlet), d (doublet), dd (doublet of doublet), t (triplet), and m (multiplet). Fourier transform infrared (FT-IR) spectra of the entitled compounds were recorded in the range of 4000-600 cm⁻¹ using a Bruker Tensor 27 spectrophotometer. For solid samples, KBr pellets were prepared for IR measurement, and neat sample was used for liquids. The PANalytical X’PERT PRO diffractometer was employed for analysis of the crystal structure of ZnO powder, operated at 45 KV, 40 mA using Ni-filtered Cu K_α radiation with a scan speed of 10°/min for 2θ in a range from 10 to 75. For XRD analysis, air-dried fine powder samples were prepared in a uniform thin layer on a zero background sample holder. The morphology and elemental (EDX) composition were visualized by scanning electron microscopy (SEM, JEOL JSM-6610LV) using a voltage of 15 KV. A dilute solution of ZnO (100 μM) in ethanol was placed on carbon tape and was dried at room temperature under vacuum. The Metrohm Microtrac Ultra Nanotracer Particle Size Analyzer (Dynamic Light Scattering) was used to measure the hydrodynamic diameter of ZnO particles using a DMSO/water (5:5, v/v) system as the dispersion medium. N₂ absorption measurements were carried out on an Autosorb-IQ Quantachrome Instruments volumetric adsorption analyzer. The sample was degassed at 200 °C for 3 h in the degassing port. The Brunauer-Emmett-Teller (BET) model from a linear part of the BET plot (P/P₀: 0.05-0.3) was employed for the surface area calculation.

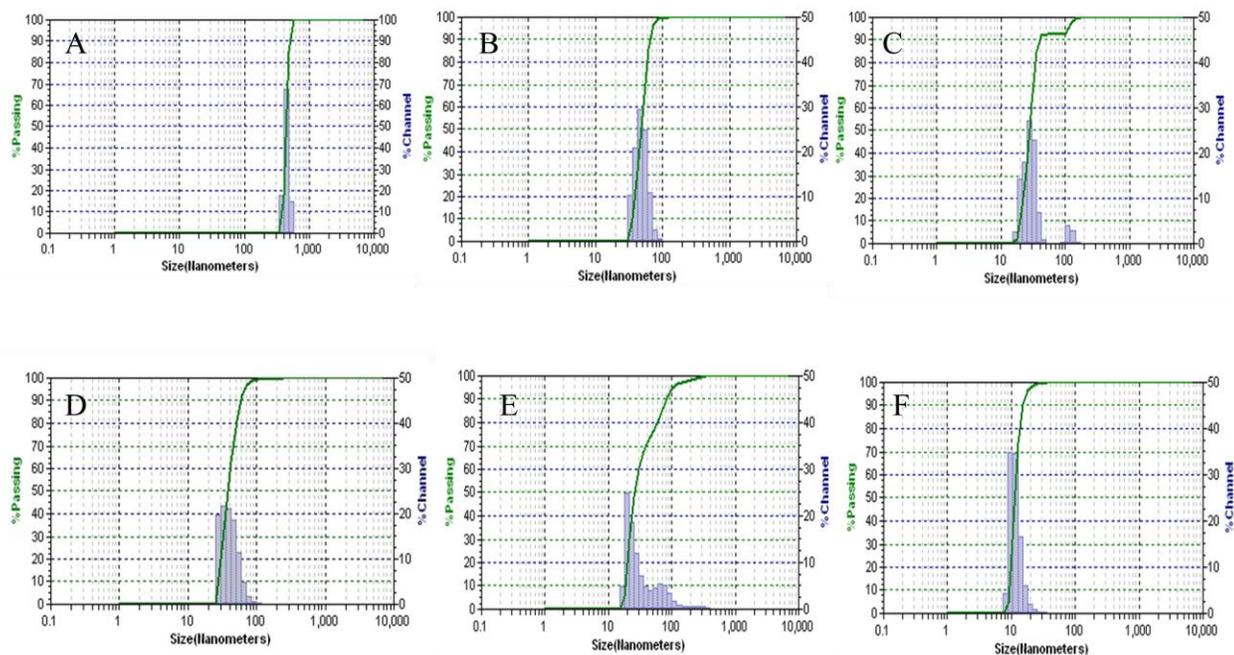


Figure S1. A dynamic light scattering (DLS) study of six different catalysts indicating the size of the ZnO nanoparticles (nm) (A) NPs-1, (B) NPs-2, (C) NPs-3, (D) NPs-4, (E) NPs-5, and (F) NPs-6.

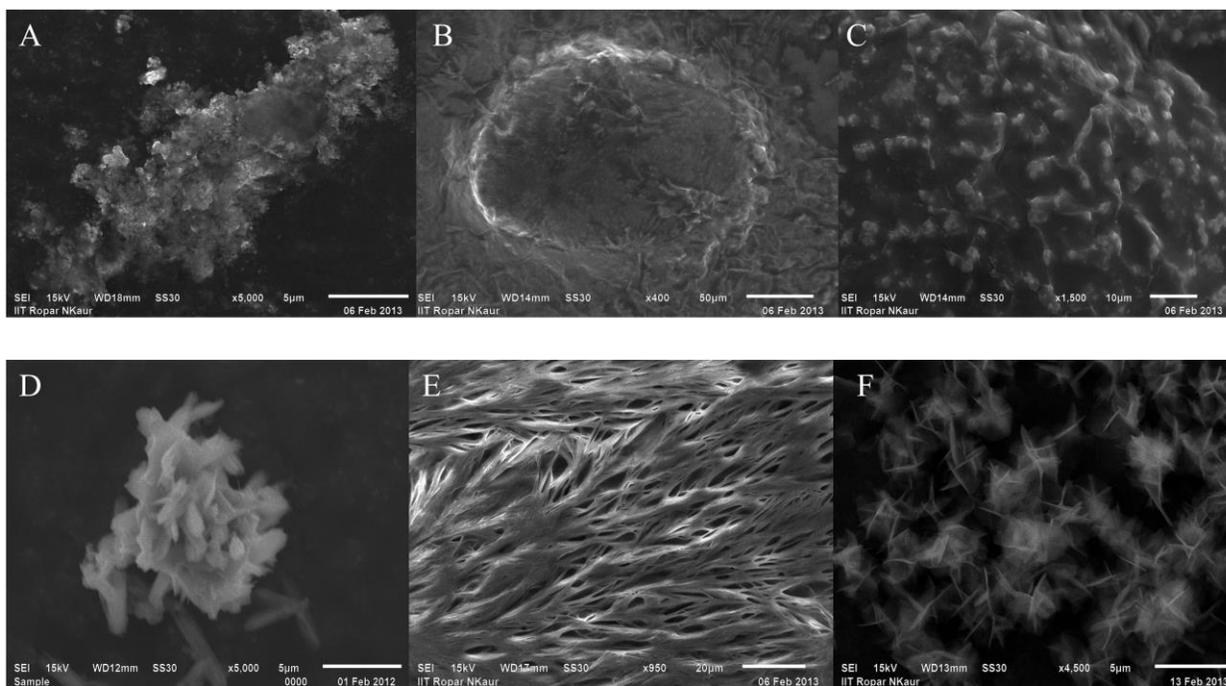


Figure S2. SEM images of six different catalysts showing the size of ZnO nanoparticles (nm) (A) NPs-1, (B) NPs-2, (C) NPs-3, (D) NPs-4, (E) NPs-5, and (F) NPs-6.

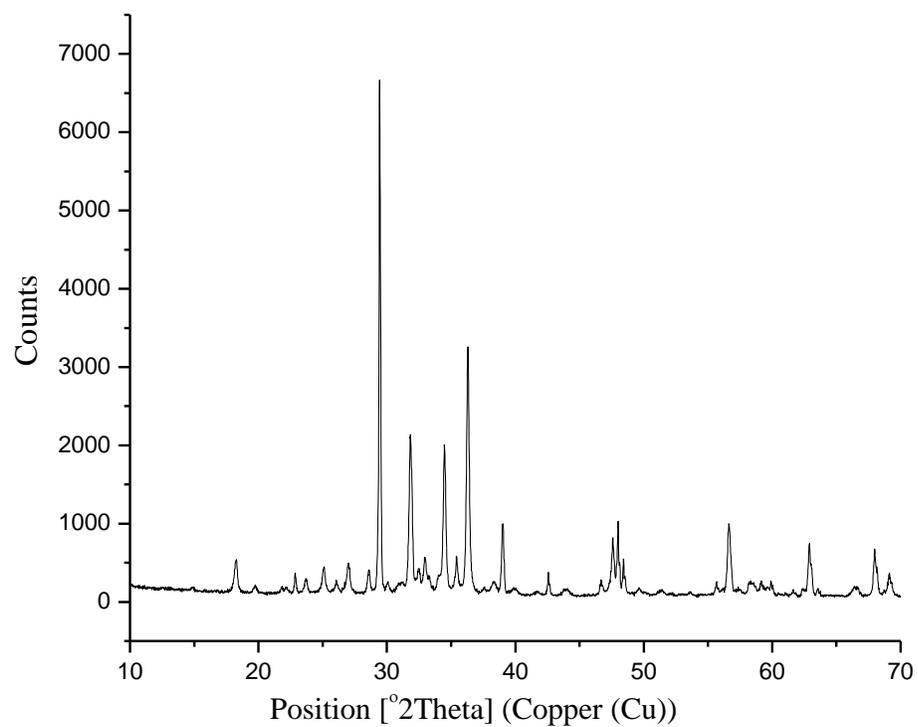


Figure S3. XRD pattern of NPs-6.

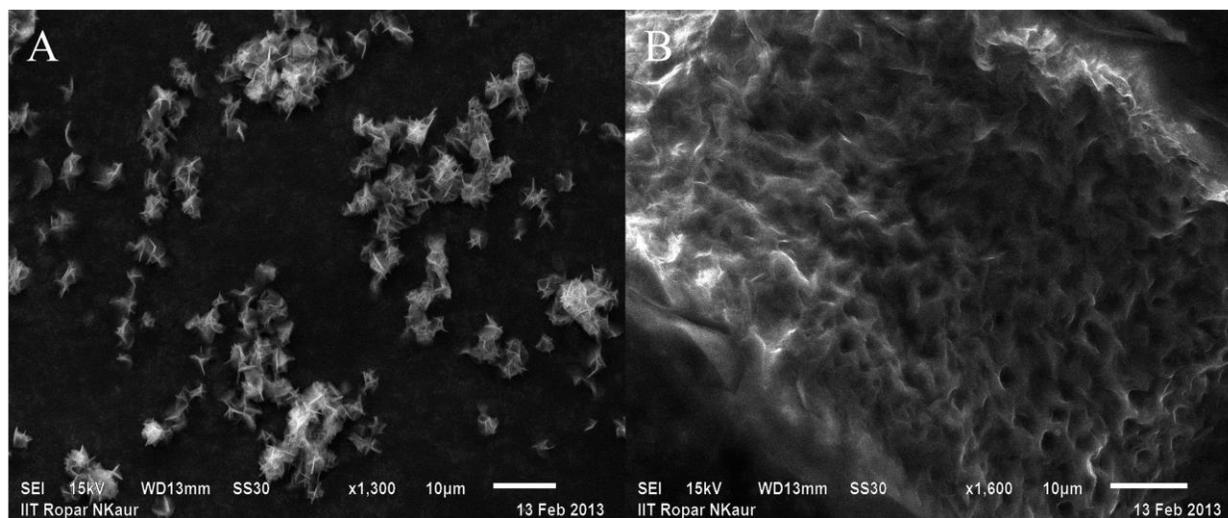


Figure S4. SEM image of catalyst NPs-6 (A) before reaction and (B) after the fifth reaction.

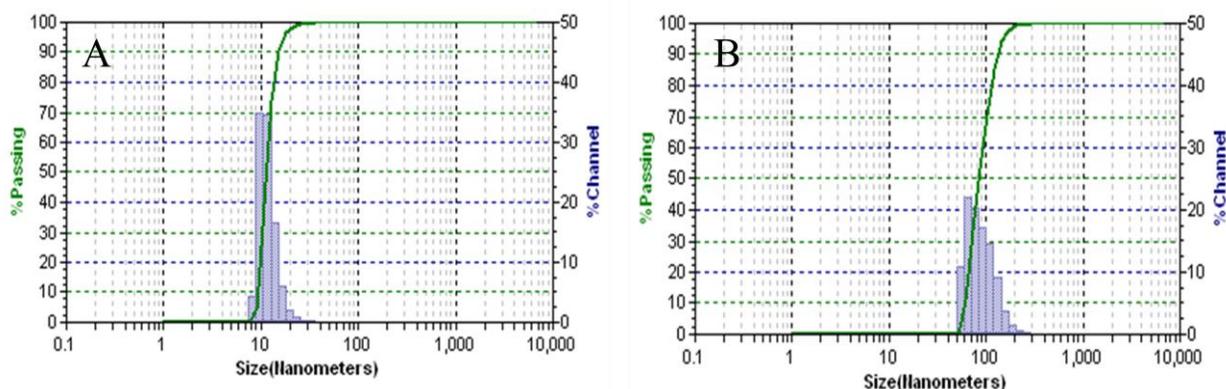


Figure S5. Size measurement of catalyst NPs-6 (A) before reaction and (B) after the fifth reaction.

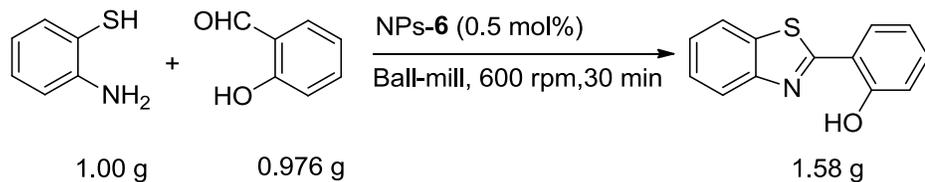
Table S1. Ecoscale calculation for the reaction of 2-aminothiophenol with 2-hydroxybenzaldehyde.^a

	Detail of parameters	Penalty Points ^b
1.	Yield 87%	6.5
2.	Cost of reactants to obtain 10 mmol of product	
	2-aminothiophenol (1.25 g)	0
	2-hydroxybenzaldehyde (1.22 g)	0
	ZnO nanoparticles (0.5 mol%, 81 mg)	0
	methanol (10 mL)	0
3.	Safety	
	2-aminothiophenol	5 (T)
	2-hydroxybenzaldehyde	5 (T)
	ZnO	5 (N)
	methanol	10 (F, T)
4.	Technical setup	
	ball-mill	2
5.	Temperature/Time	
	room temperature, < 1h	0
	annealing for ZnO, Heating > 1h	3
6.	Work-up and Purification	
	adding solvent	0
	simple filtration	0

^aThe reaction was performed on a 10 mmol scale.

^bThe total of all penalties was 36.5, which gave a score of 63.5 (100 - 36.5), which is indicative of an acceptable synthesis.

Table S2. Calculation of E-factor and mass intensities for the reaction of 2-aminothiophenol with 2-hydroxybenzaldehyde.^a



Total amount of reactants: 1.00 g + 0.976 g = 1.98 g

Amount of final product: 1.58 g

Amount of waste: (1.98 g - 1.58 g) = 0.40 g

E-Factor = Amount of waste/Amount of product = 0.40 g/1.58 g = 0.25

Mass intensity (MI) = 1.25

^aThe calculation was performed for 8 mmol scale.

