

Electronic Supplementary Information

Highly Efficiency Triarylene Conjugation Contains substitution groups for Dye-Sensitized Solar Cells

Yuan Jay Chang^a and Tahsin J. Chow^{b*}

^a*Department of Chemistry, National Taiwan University, Taipei 106, Taiwan;*

^b*Institute of Chemistry, Academia Sinica, Taipei 115, Taiwan*

Contents

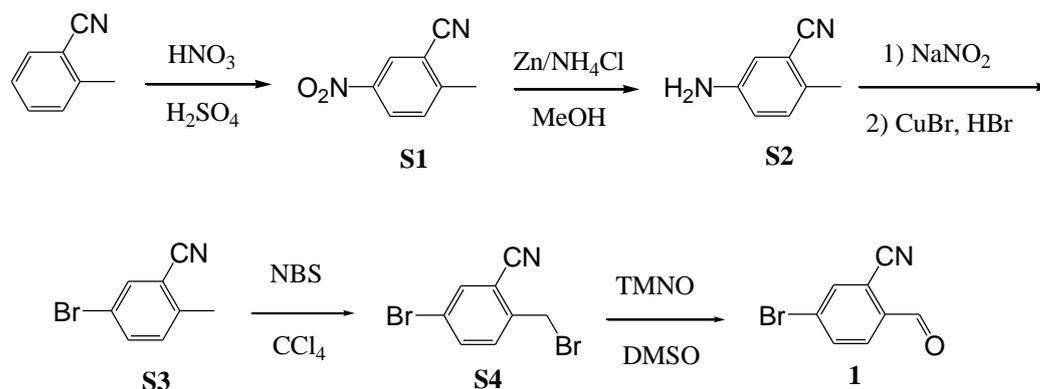
1. The synthesis of compound 1	2-4
2. ¹ H and ¹³ C NMR spectra	5-21
3. Infrared spectra	22
4. Theoretical calculation	23-25
5. Loading amount	26
6. HOMO-LUMO level	26
7. EIS spectra	27

1. The synthesis of compound 1

General Information

All reactions and manipulations were carried out under a nitrogen atmosphere. Solvents were distilled freshly according to standard procedures. ^1H and ^{13}C NMR spectra were recorded on Bruker (AV 400/AV 500 MHz) spectrometer in CDCl_3 , THF-d_8 , and DMSO-d_6 as a solvent. Chemical shifts are reported in scale downfield from the peak for tetramethylsilane. Absorption spectra were recorded on a Jasco-550 spectrophotometer. Emission spectra and photoluminescence quantum yield were obtained from a Hitachi F-4500 spectrofluorimeter. The emission spectra in solutions were measured in spectral grade solvent by a 90° angle detection, and those of thin film were recorded by front-face detection. The redox potentials were measured by using cyclic voltammetry on CHI 620 analyzer. All measurements were carried out in THF solutions containing 0.1 M tetrabutylammonium hexafluorophosphate (TBAPF_6) as supporting electrolyte at ambient condition after purging 10 minutes with N_2 . The conventional three electrode configuration was employed, which consists of a glassy carbon working electrode, a platinum counter electrode, and a Ag/Ag^+ reference electrode calibrated with ferrocene/-ferrocenium (Fc/Fc^+) as an internal reference. Mass spectra were recorded on a VG70-250S mass spectrometer. Elementary analyses were performed on a Perkin-Elmer 2400 CHN analyzer.

All chemicals were purchased from ACROS, Merck, Lancaster, TCI, Sigma-Aldrich, separately, and purified before use. Chromatographic separations were carried out by using silica gel from Merk, Kieselgel si 60 (40 - 63 μm).



Scheme S1. The synthetic route of 1.

2-Methyl-5-nitrobenzonitrile (S1)

A mixture of fuming nitric acid (10 mL) and sulphuric acid (20 mL) in a two-necked flask was stirred at room temperature for 1 h. To it was added *o*-tolunitrile (9 mL, 76.8 mmol) slowly in an ice-bath, then the mixture was stirred for 1 hour at room temperature. The resulted solution was poured into ice-water, and the white solids were collected by filtration. Compound **S1** was purified by recrystallization from ethanol in 90% yield (11.2 g, 69.12 mmol). Spectroscopic data of **S1**: δ_{H} (400 MHz, CDCl_3) 8.46 (d, 1H, $J = 2.4$ Hz), 8.33 (dd, 1H, $J = 2.4, 8.4$ Hz), 7.53 (d, 1H, $J = 8.8$ Hz), 2.68 (s, 3H); δ_{C} (100 MHz, CDCl_3) 149.2, 146.1, 131.5, 127.5, 127.2, 115.9, 114.2, 20.8; m/z (EI) 162.0425 (M^+ , $\text{C}_8\text{H}_6\text{N}_2\text{O}_2$ requires 162.0429).

5-Amino-2-methylbenzonitrile (S2)

A mixture of compound **S1** (5 g, 30.8 mmol) and Zn powder (7.5 g, 114.7 mmol) in methanol (200 mL) was placed in a three-necked flask and was stirred at room temperature for 4 h under a nitrogen atmosphere. The reaction was quenched by adding 0.1 N $\text{NaOH}_{(\text{aq})}$, then was extracted with ethyl acetate. The organic layer was dried over anhydrous MgSO_4 and evaporated under vacuum. The products were purified by silica gel column chromatograph eluted with dichloromethane. Yellow solids of **S2** were obtained in 85% yield (3.46 g, 26.18 mmol). Spectroscopic data of **S2**: δ_{H} (400 MHz, CDCl_3) 7.03 (d, 1H, $J = 8.5$ Hz), 6.84 (d, 1H, $J = 2.5$ Hz), 6.76 (dd, 1H, $J = 2.5, 8.0$ Hz), 3.68 (br, s, 2H), 2.39 (s, 3H); δ_{C} NMR (100 MHz, CDCl_3) 144.4, 131.3, 130.9, 119.6, 118.3, 117.7, 112.9, 19.2; m/z (EI) 132.0687 (M^+ , $\text{C}_8\text{H}_8\text{N}_2$ requires 132.0687).

5-Bromo-2-methylbenzonitrile (S3)

A mixture of compound **S2** (3 g, 22.65 mmol), water (36 mL) and 48% HBr (10.8 mL) was stirred for 2 h. To it was then added an aqueous solution (20.3 mL) of NaNO_2 (17.1 g, 2.7 mmol), followed by CuBr (9.75 g, 67.65 mmol) after 30 mins. To the mixture was again added a combined solution of 48% HBr (22.5 mL) and water (54 mL) at 70 °C in a period of 30 mins. Then the mixture was heated to reflux for 2 h. After being cooled, the reaction was quenched by adding $\text{KOH}_{(\text{aq})}$, then was extracted with ethyl acetate. The organic layer was dried over anhydrous MgSO_4 and evaporated under vacuum. The products were purified by silica gel column chromatograph eluted with dichloromethane/hexane (1/1). Colorless crystalline solids of **S3** were obtained in 56% yield (2.47 g, 12.68 mmol). Spectroscopic data of **S3**: δ_{H} (CDCl_3) 7.7 (d, 1H, $J = 2.0$ Hz), 7.59 (dd, 1H, $J = 2, 8.4$ Hz), 7.19 (d, 1H, $J = 8.0$ Hz), 2.49 (s, 3H); δ_{C} (CDCl_3) 140.8, 135.7, 134.7, 131.7, 119.2, 116.6, 114.5, 19.9; m/z (EI)

194.9682 (M^+ , C_8H_6BrN requires 194.9684).

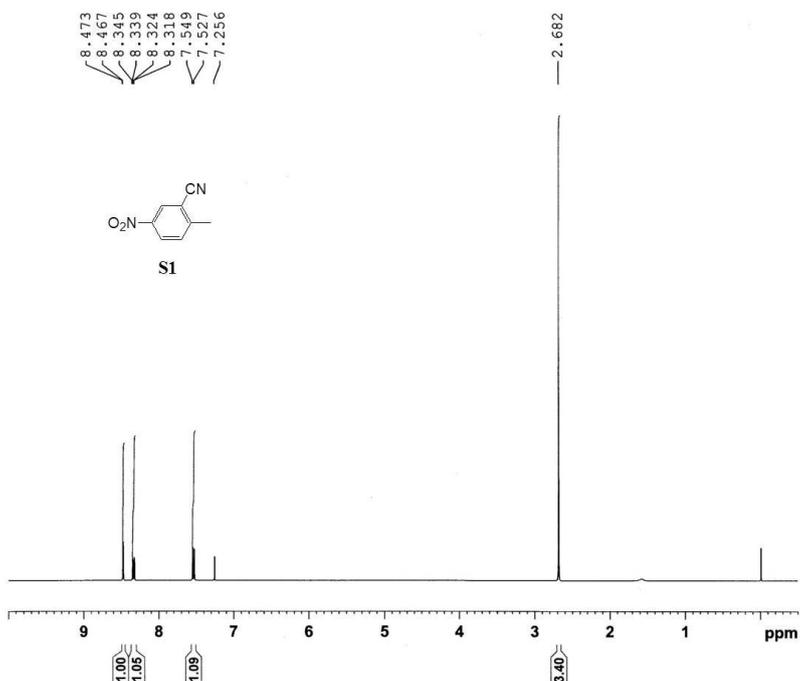
5-Bromo-2-(bromomethyl)benzonitrile (S4)

A mixture of compound **S3** (1 g, 5.13 mmol), NBS (1.0 g, 5.64 mmol), and BPO (40 mg, 0.165 mmol) was dissolved in CCl_4 (20 mL), then the solution was heated to 90 °C with stirring for 9 h. After cooling, the reaction was quenched by adding water, then was extracted with dichloromethane. The organic layer was dried over anhydrous $MgSO_4$ and evaporated under vacuum. The products were purified by silica gel column chromatograph eluted with ethyl acetate/hexane (1/9). White solids of **S4** were obtained in 90% yield (1.26 g, 4.62 mmol). Spectroscopic data of **S4**: δ_H (400 MHz, $CDCl_3$) 7.78 (d, 1H, $J = 2.0$ Hz), 7.71 (dd, 1H, $J = 2, 8.4$ Hz), 7.42 (d, 1H, $J = 8.4$ Hz), 4.57 (s, 2H); δ_C (100 MHz, $CDCl_3$) 140.0, 136.4, 135.5, 131.8, 122.4, 115.2, 114.0, 28.3; m/z (EI) 272.8793 (M^+ , $C_8H_5Br_2N$ requires 272.8789).

5-Bromo-2-formylbenzonitrile (1)

The solution of trimethylamine N-oxide (1.1 g, 14.64 mmol) in DMSO (8 mL) was kept in an ice bath, and to it was added dropwise compound **S4** (1.0 g, 3.66 mmol) in dry CH_2Cl_2 , and the resulted solution was warm to room temperature and stirred with a magnetic bar for 8 h. The reaction was quenched by adding saturated $NaCl_{(aq)}$, then was extracted with ethyl acetate. The organic layer was dried over anhydrous $MgSO_4$ and evaporated under vacuum. The products were purified by silica gel column chromatograph eluted with dichloromethane. White solids of **1** were obtained in 56% yield (428 mg, 2.05 mmol). Spectroscopic data of **1**: δ_H (400 MHz, $CDCl_3$) 10.30 (s, 1H), 7.97 (s, 1H), 7.91 (s, 2H); δ_C (100 MHz, $CDCl_3$) 187.4, 136.7, 136.6, 135.4, 130.56, 129.3, 115.3, 114.5; m/z (EI) 208.9477 (M^+ , C_8H_4BrNO requires 208.9476).

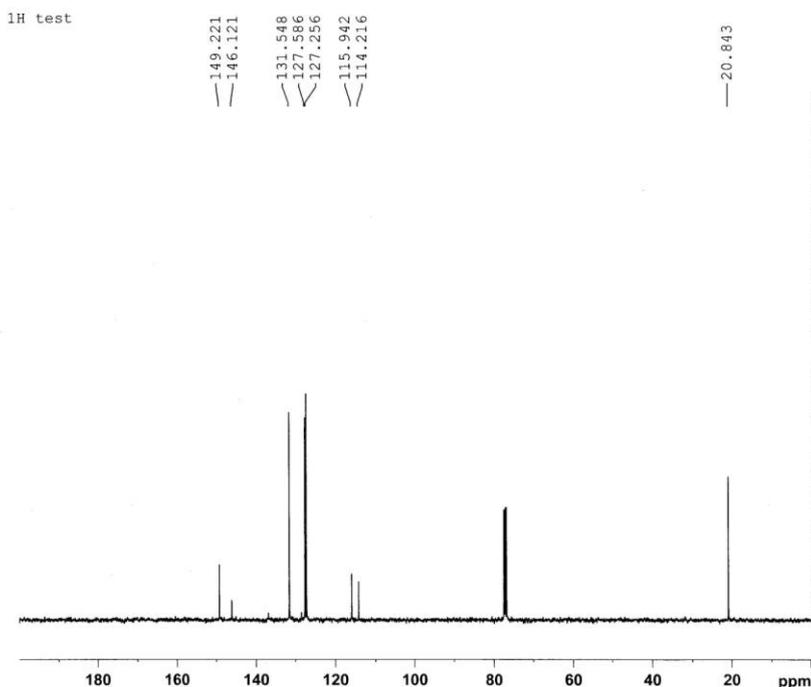
2. ¹H and ¹³C NMR spectra



```

NAME      Jay-20100423-375
EXPNO     1
PROCNO    1
Date_     20100423
Time      17.05
INSTRUM   spect
PROBHD    5 mm Dual 13C/
PULPROG   zg30
TD         16384
SOLVENT   CDCl3
NS         16
DS         2
SWH       4789.272 Hz
FIDRES    0.292314 Hz
AQ         1.7105396 sec
RG         512
DW         104.400 usec
DE         6.50 usec
TE         298.0 K
D1         1.0000000 sec
TDO        1

===== CHANNEL f1 =====
NUC1      1H
P1        10.30 usec
PL1       -2.00 dB
PL1W      23.88643074 W
SFO1      400.1320424 MHz
SI         16384
SF         400.1300108 MHz
WDW        no
SSB        0
LB         0.00 Hz
GB         0
PC         1.00
    
```



```

NAME      Jay-20100423-375
EXPNO     2
PROCNO    1
Date_     20100924
Time      12.03
INSTRUM   spect
PROBHD    5 mm Dual 13C/
PULPROG   zg0dc
TD         32768
SOLVENT   CDCl3
NS         104
DS         8
SWH       23980.814 Hz
FIDRES    0.731836 Hz
AQ         0.6832628 sec
RG         20642.5
DW         20.850 usec
DE         6.50 usec
TE         298.3 K
D1         3.0000000 sec
D11        0.03000000 sec
TDO        1

===== CHANNEL f1 =====
NUC1      13C
P1         6.00 usec
PL1       -1.00 dB
PL1W      47.43416595 W
SFO1      100.6228293 MHz

===== CHANNEL f2 =====
CPDPRG2   waltz16
NUC2      1H
PCPD2     90.00 usec
PL2       -2.00 dB
PL12      15.90 dB
PL2W      23.88643074 W
SFO2      0.38739258 W
SFO2      400.1318764 MHz
SI         32768
SF         100.6127690 MHz
WDW        EM
SSB        0
LB         3.00 Hz
GB         0
PC         1.40
    
```

Figure S1: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **S1** in CDCl_3 .

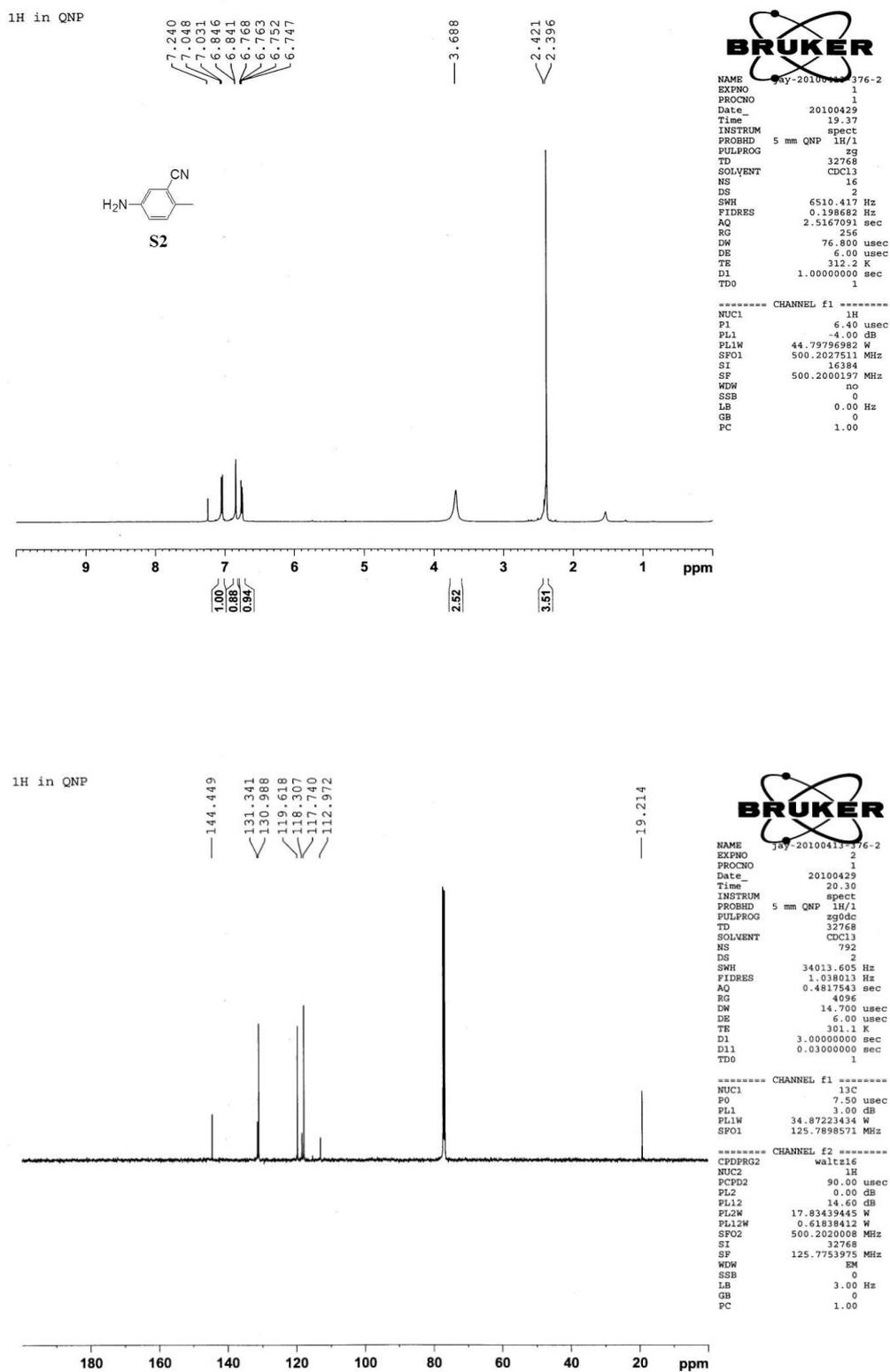


Figure S2: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **S2** in CDCl_3 .

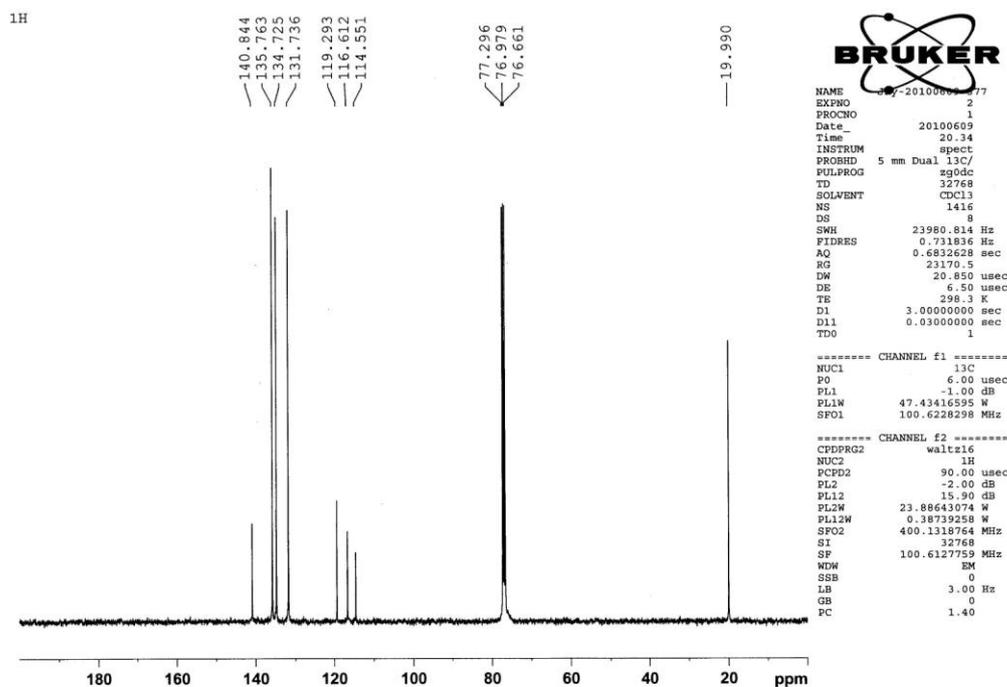
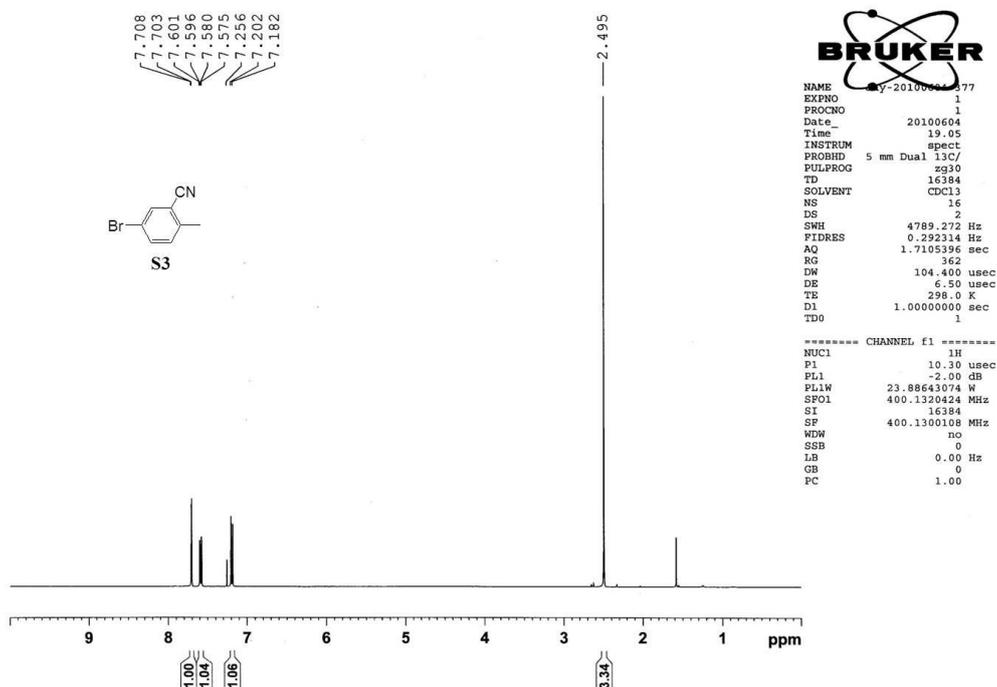


Figure S3: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **S3** in CDCl_3 .

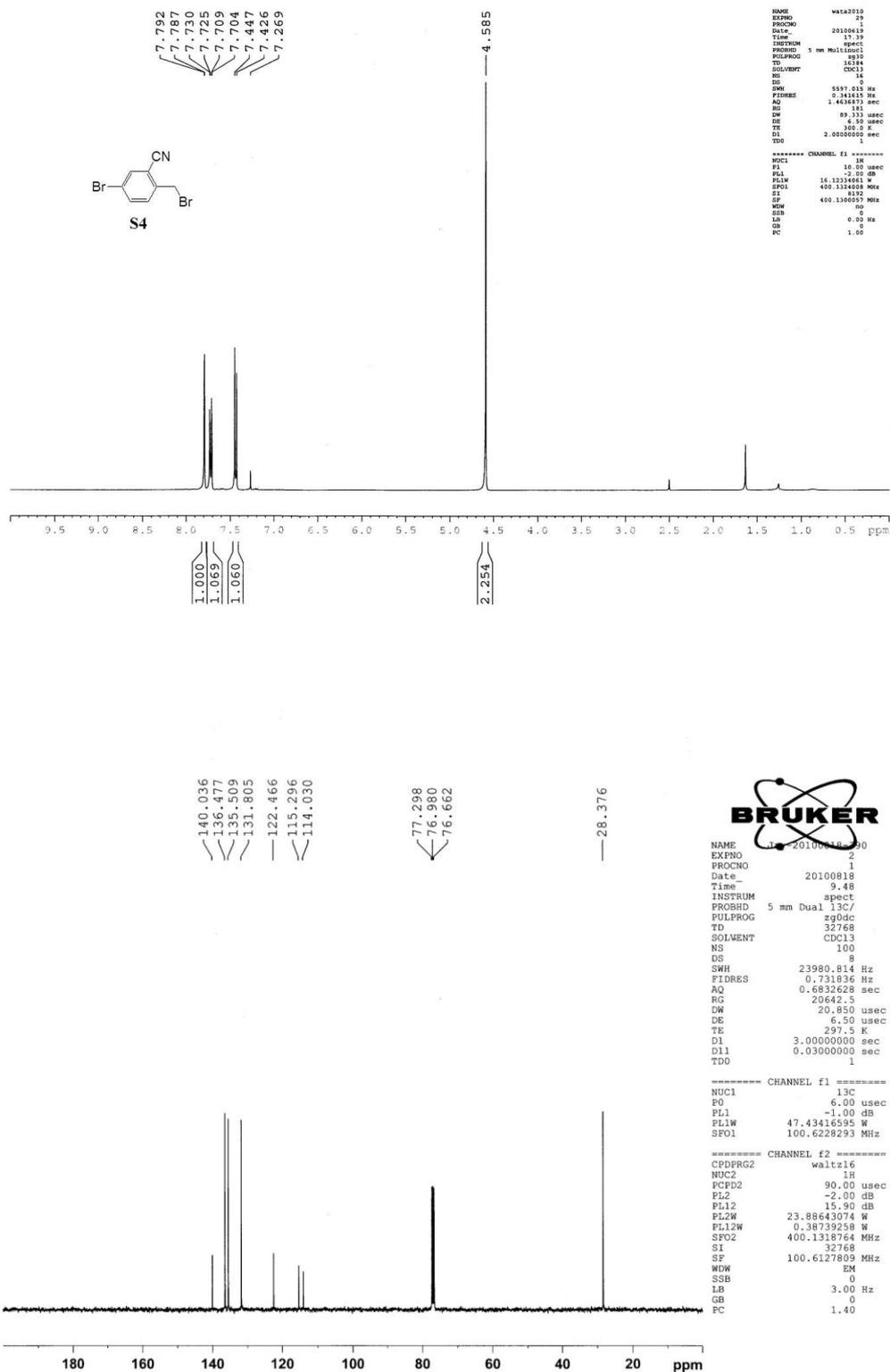


Figure S4: ¹H NMR (upper) and ¹³C NMR (lower) spectra of S4 in CDCl₃.

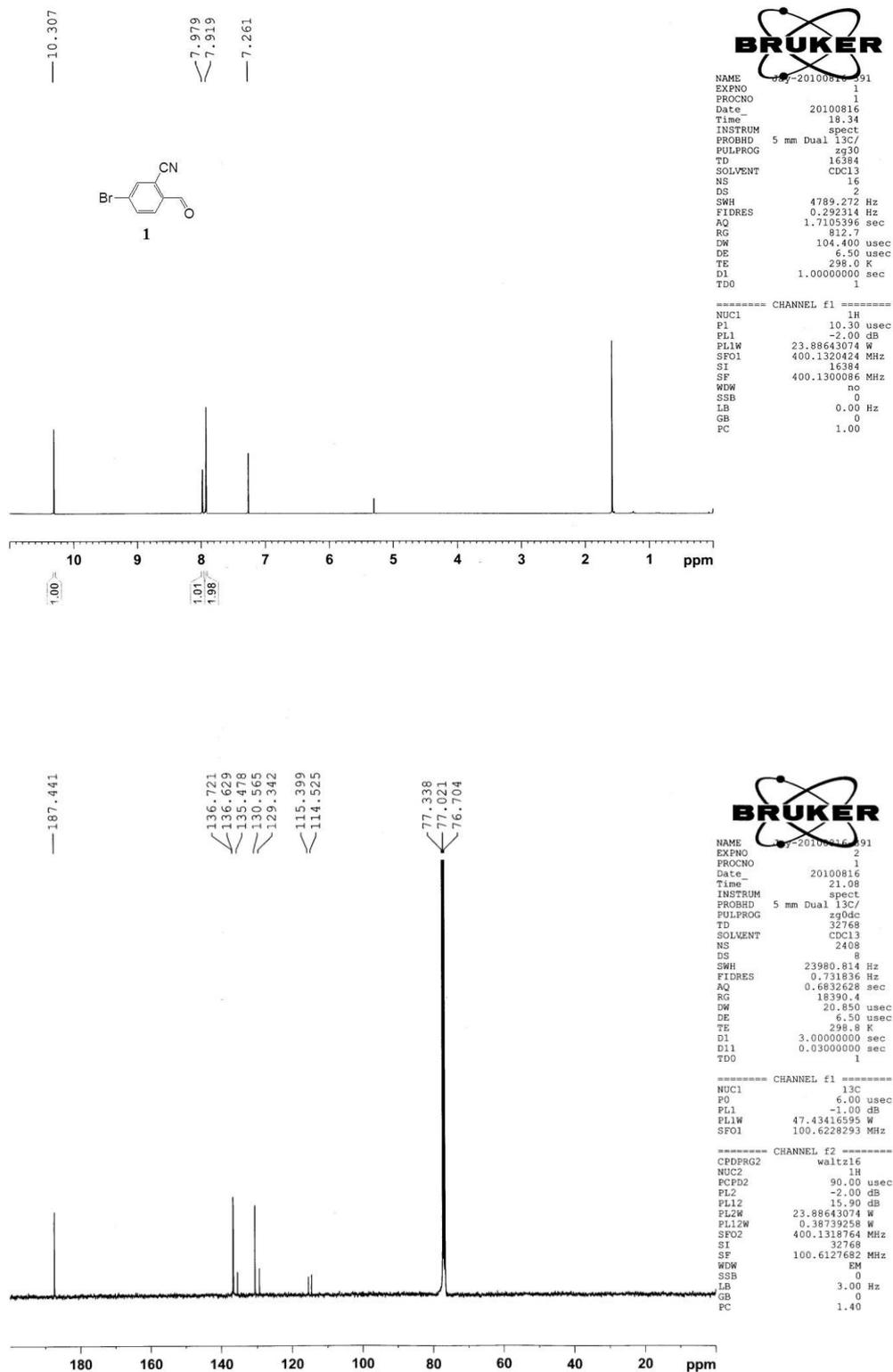


Figure S5: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **1** in CDCl_3 .

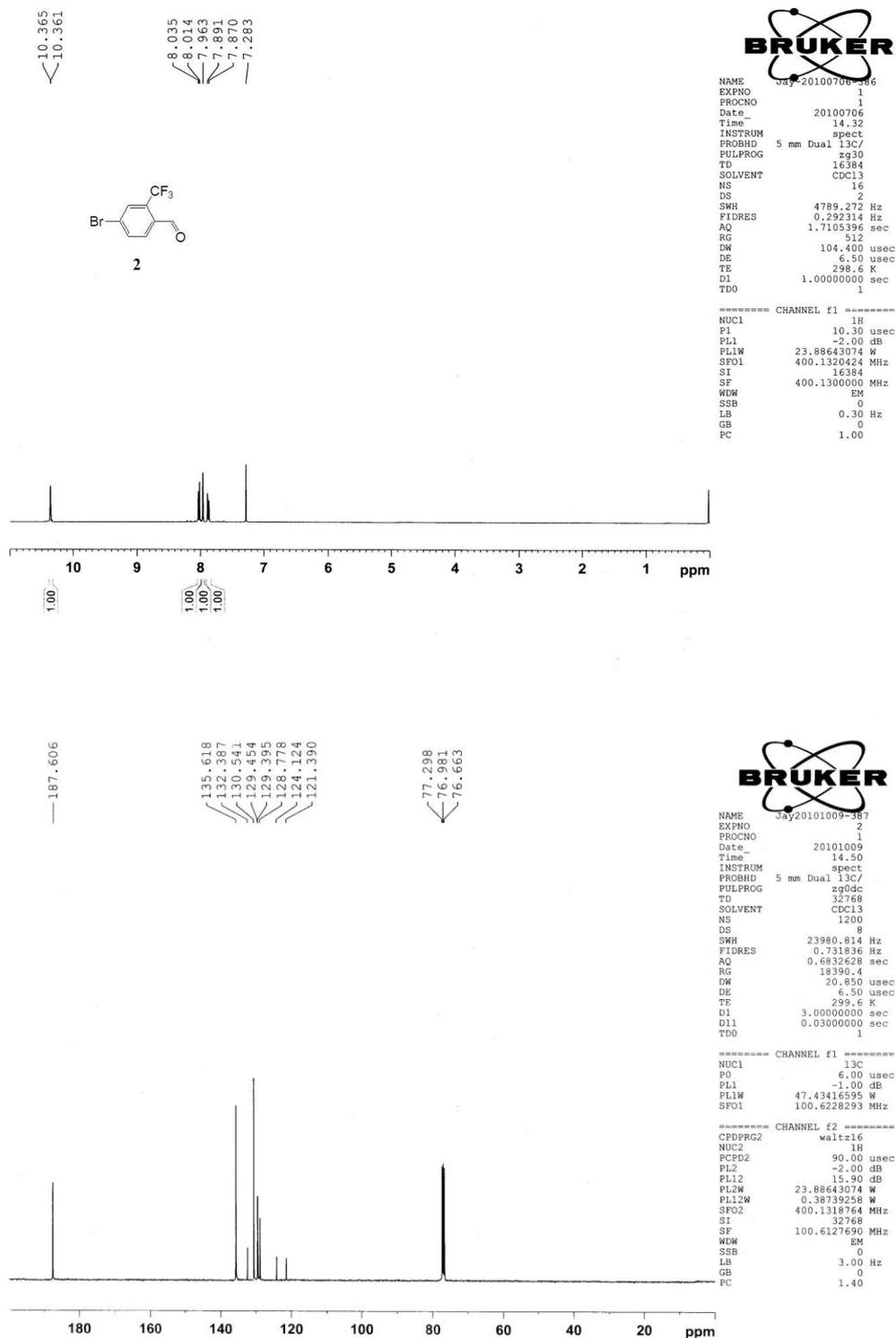


Figure S6: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **2** in CDCl_3 .

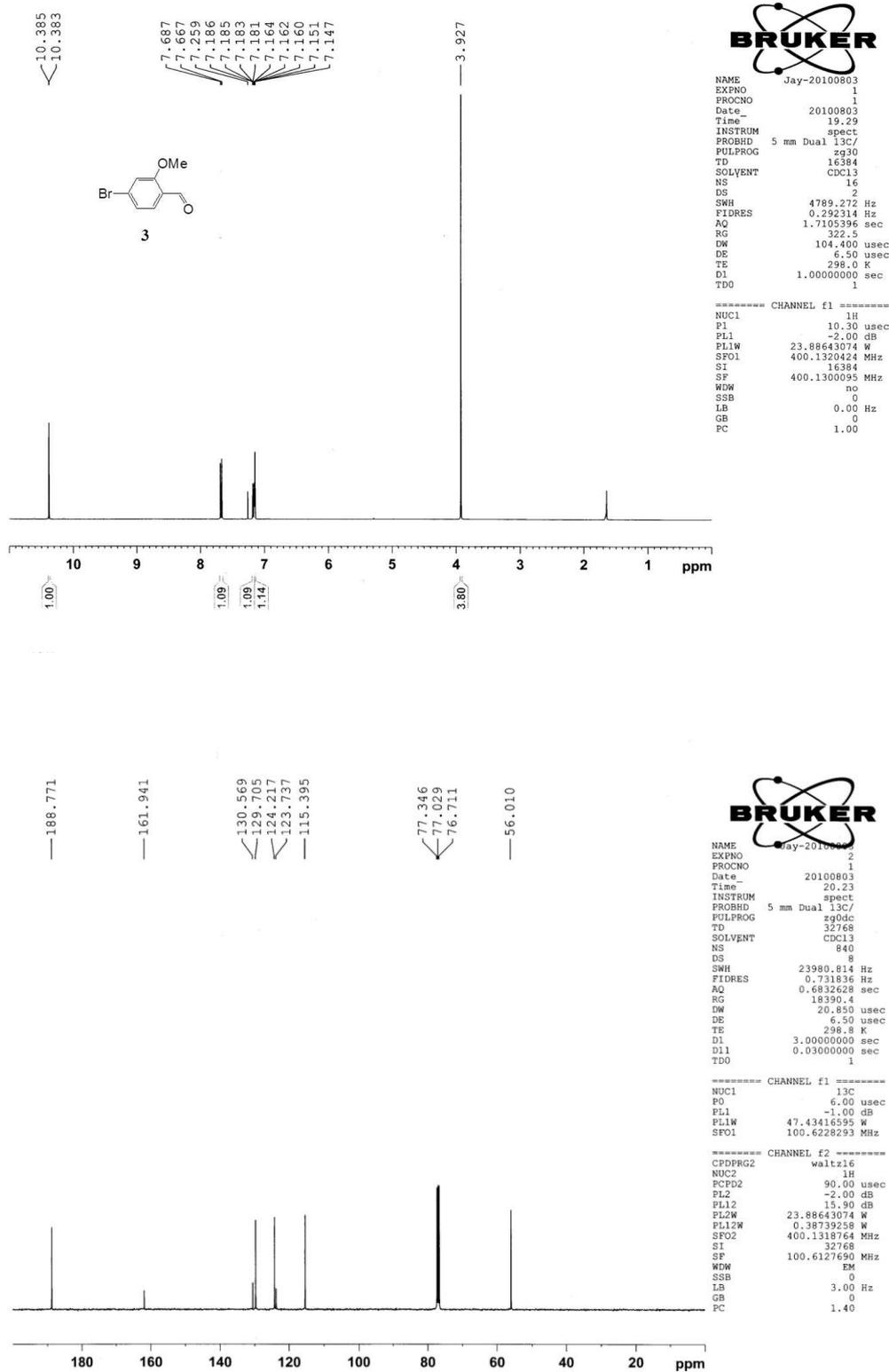


Figure S7: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **3** in CDCl_3 .

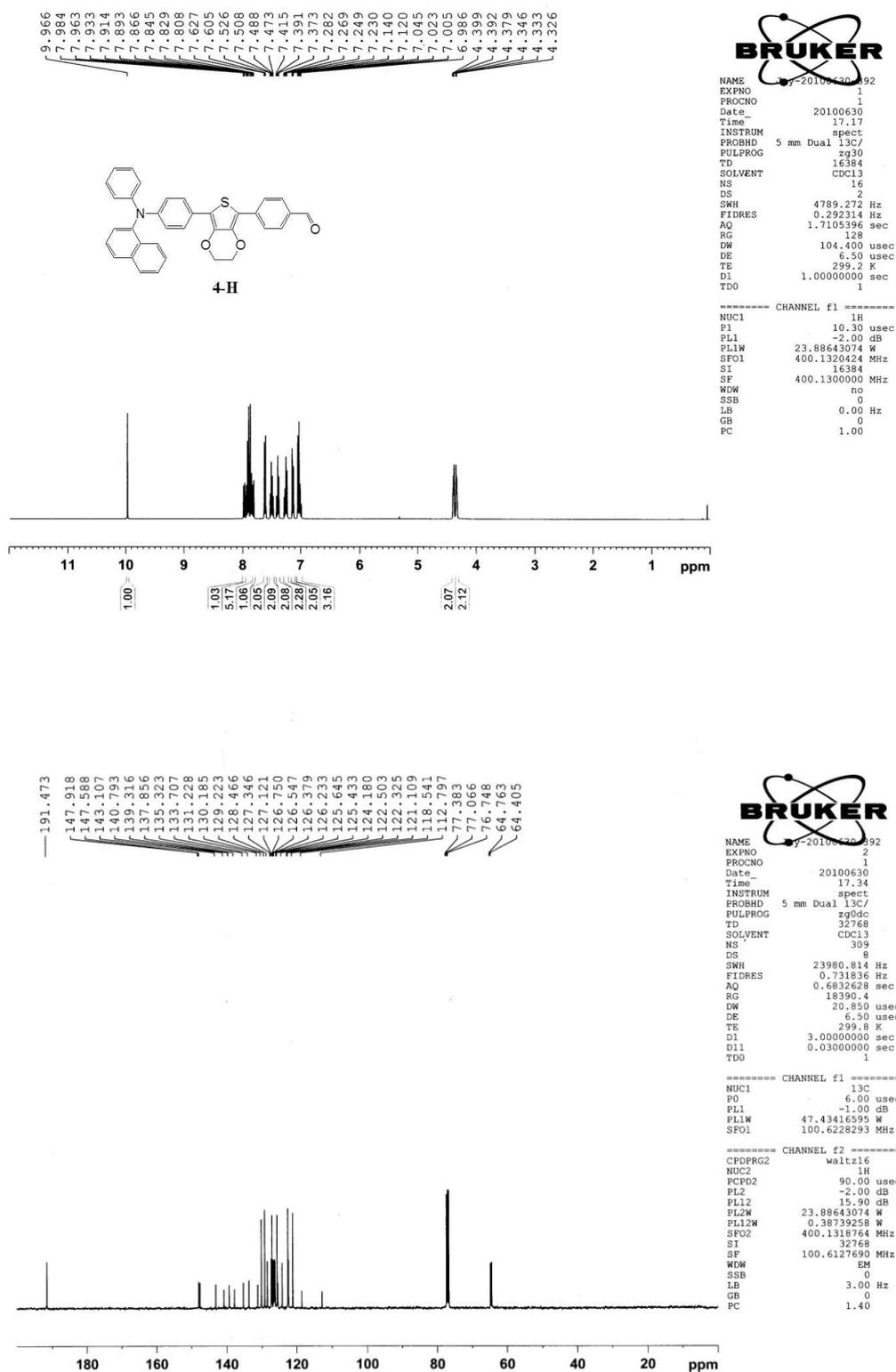


Figure S8: ¹H NMR (upper) and ¹³C NMR (lower) spectra of **4-H** in CDCl₃.

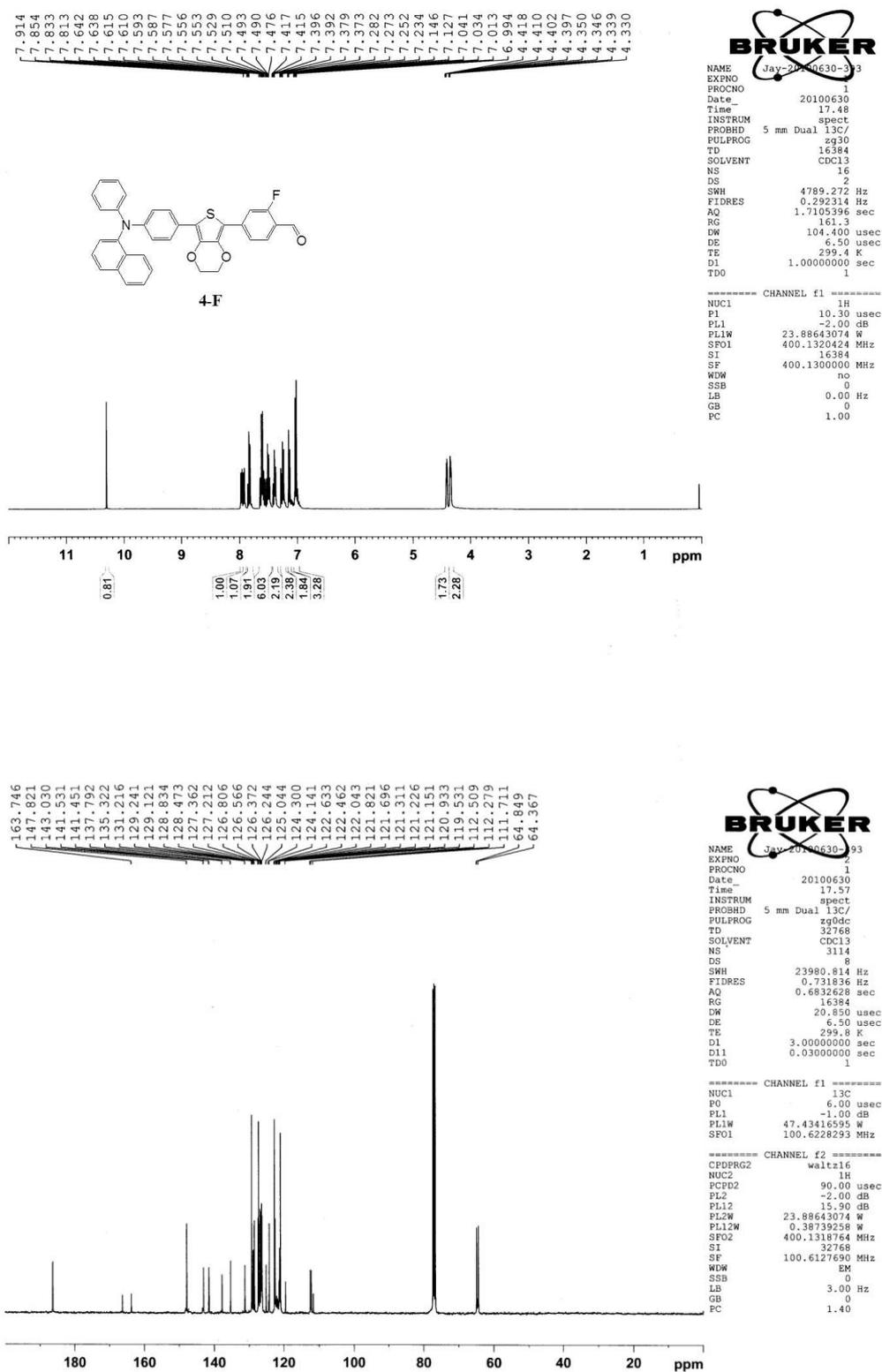


Figure S9: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **4-F** in CDCl_3 .

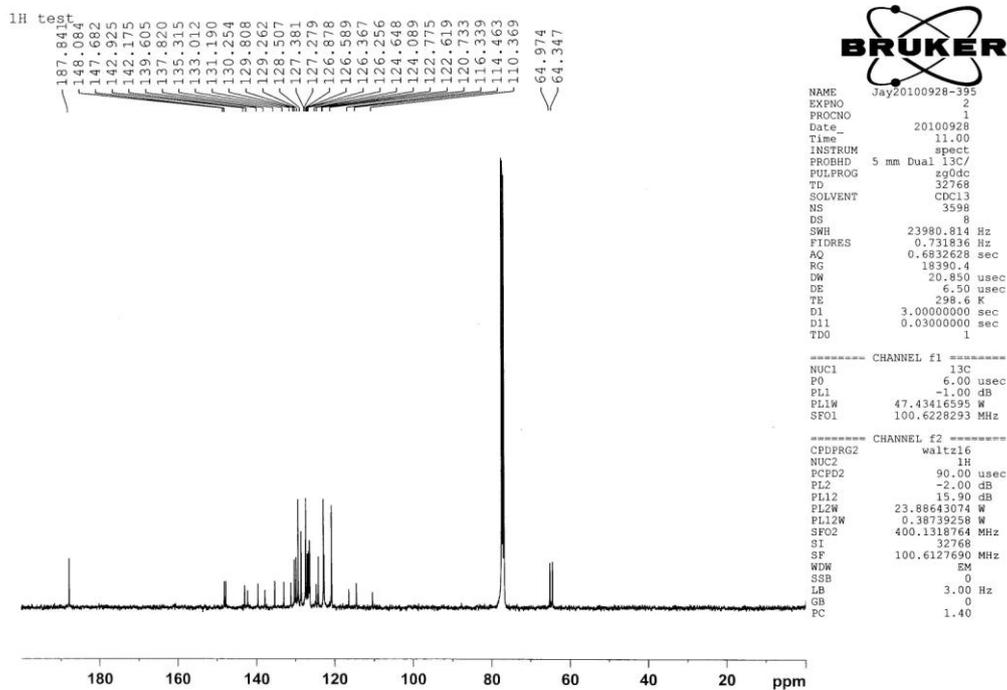
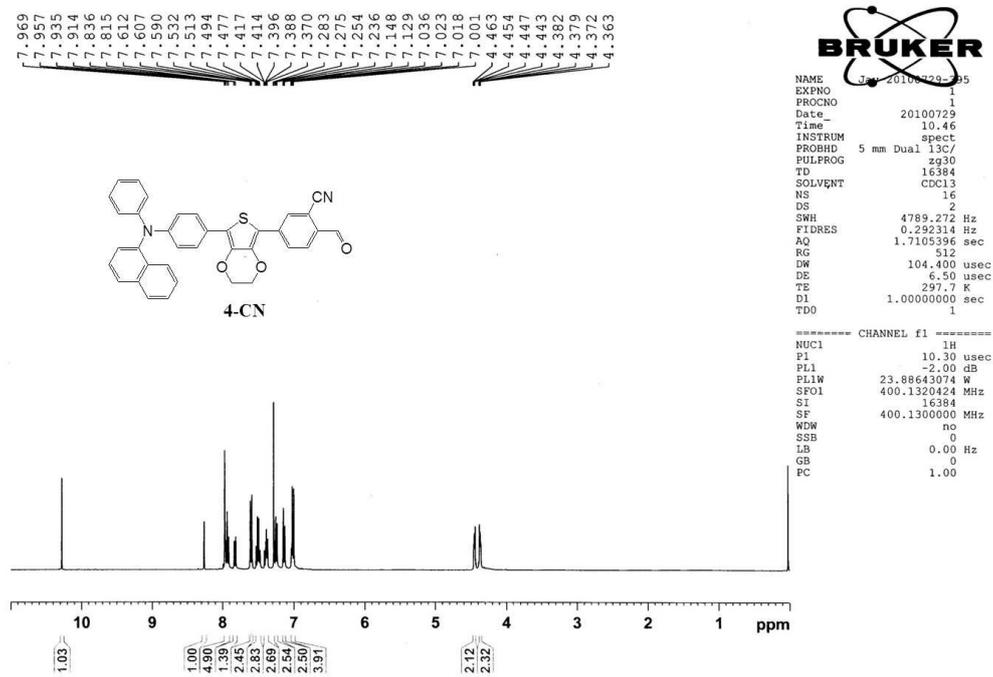


Figure S10: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **4-CN** in CDCl_3 .

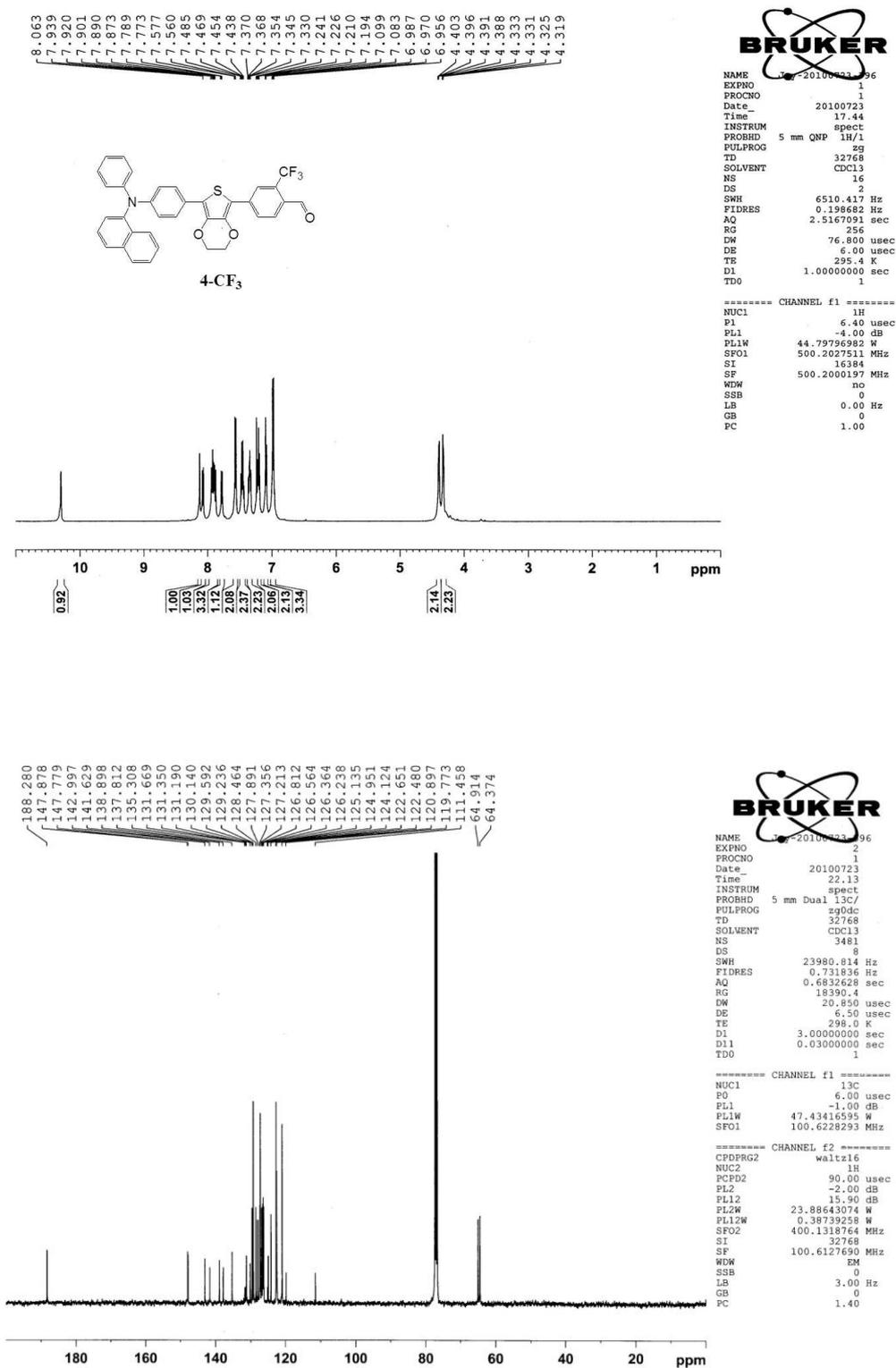


Figure S11: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **4-CF₃** in CDCl_3 .

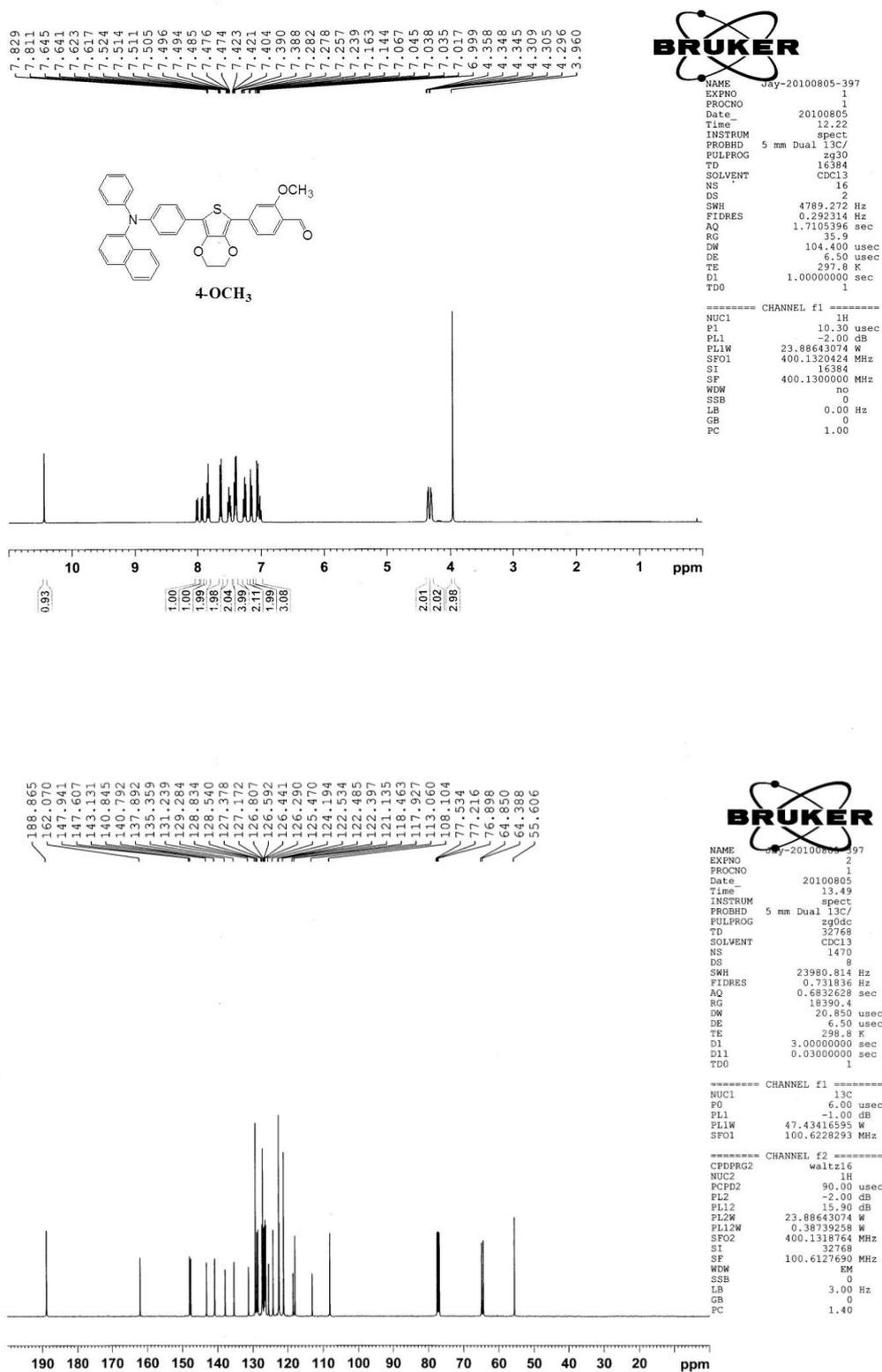


Figure S12: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **4-OCH₃** in CDCl_3 .

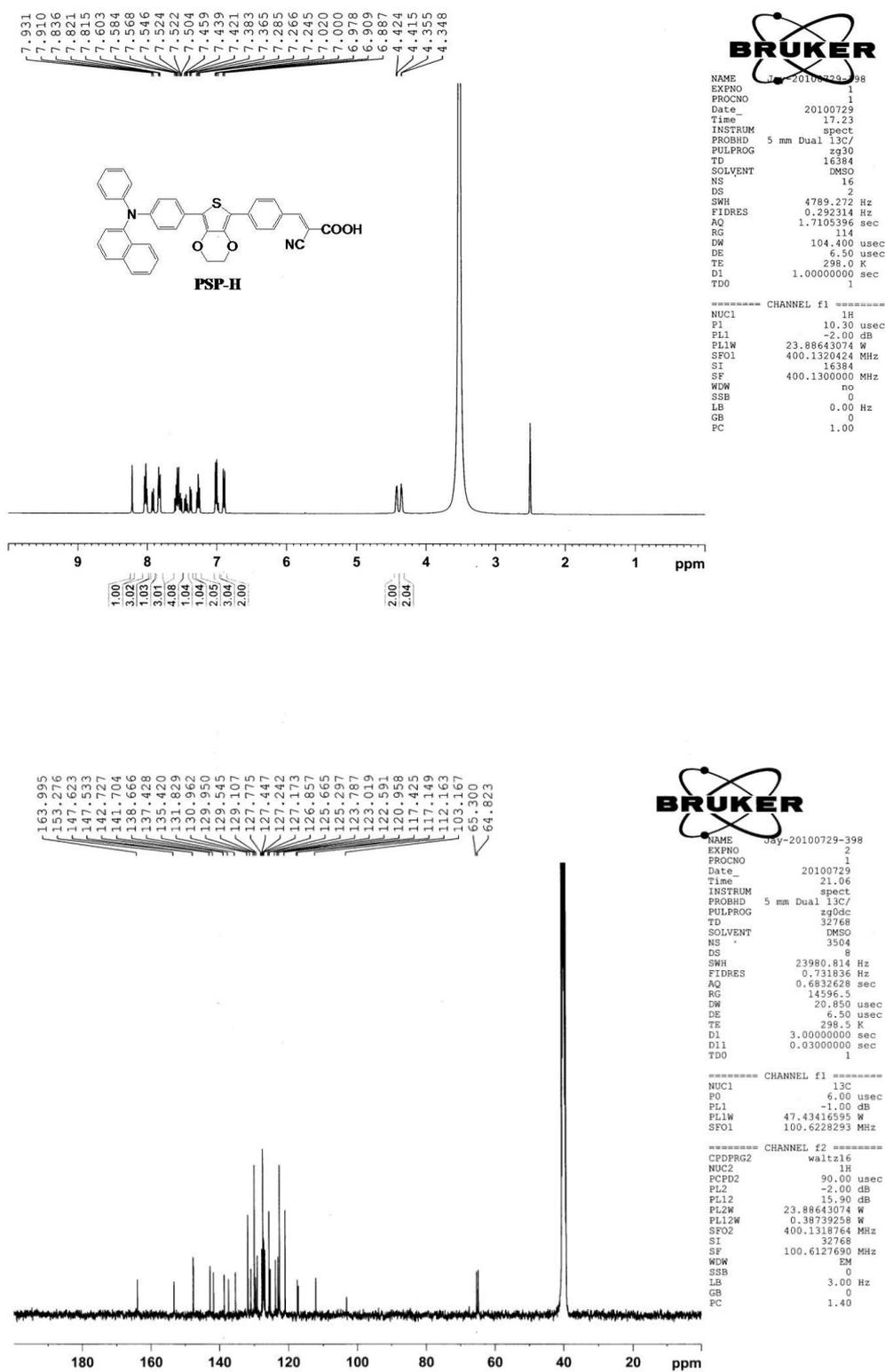


Figure S13: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **PSP-H** in $\text{DMSO-}d_6$.

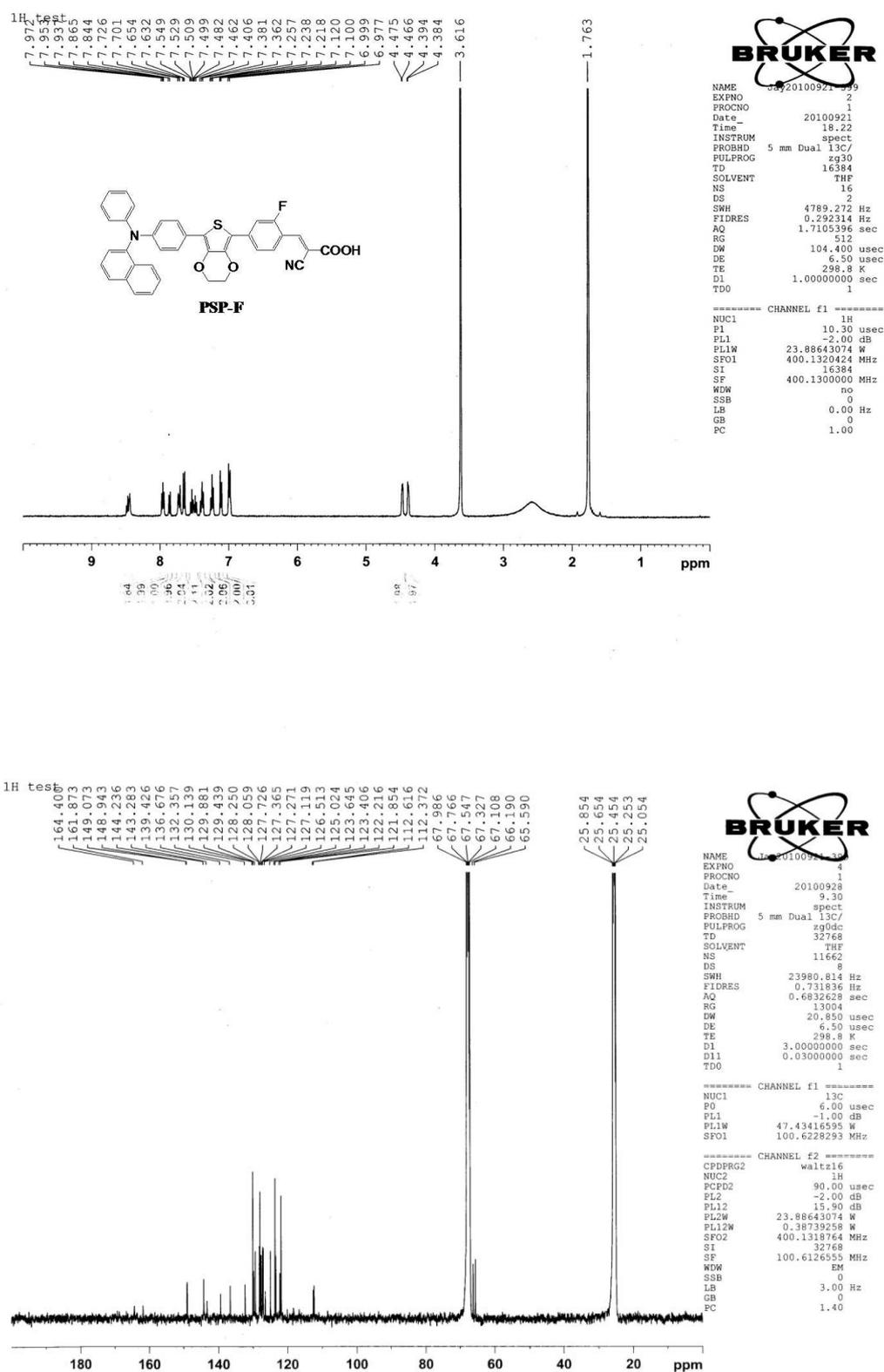


Figure S14: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **PSP-F** in $\text{THF-}d_8$.

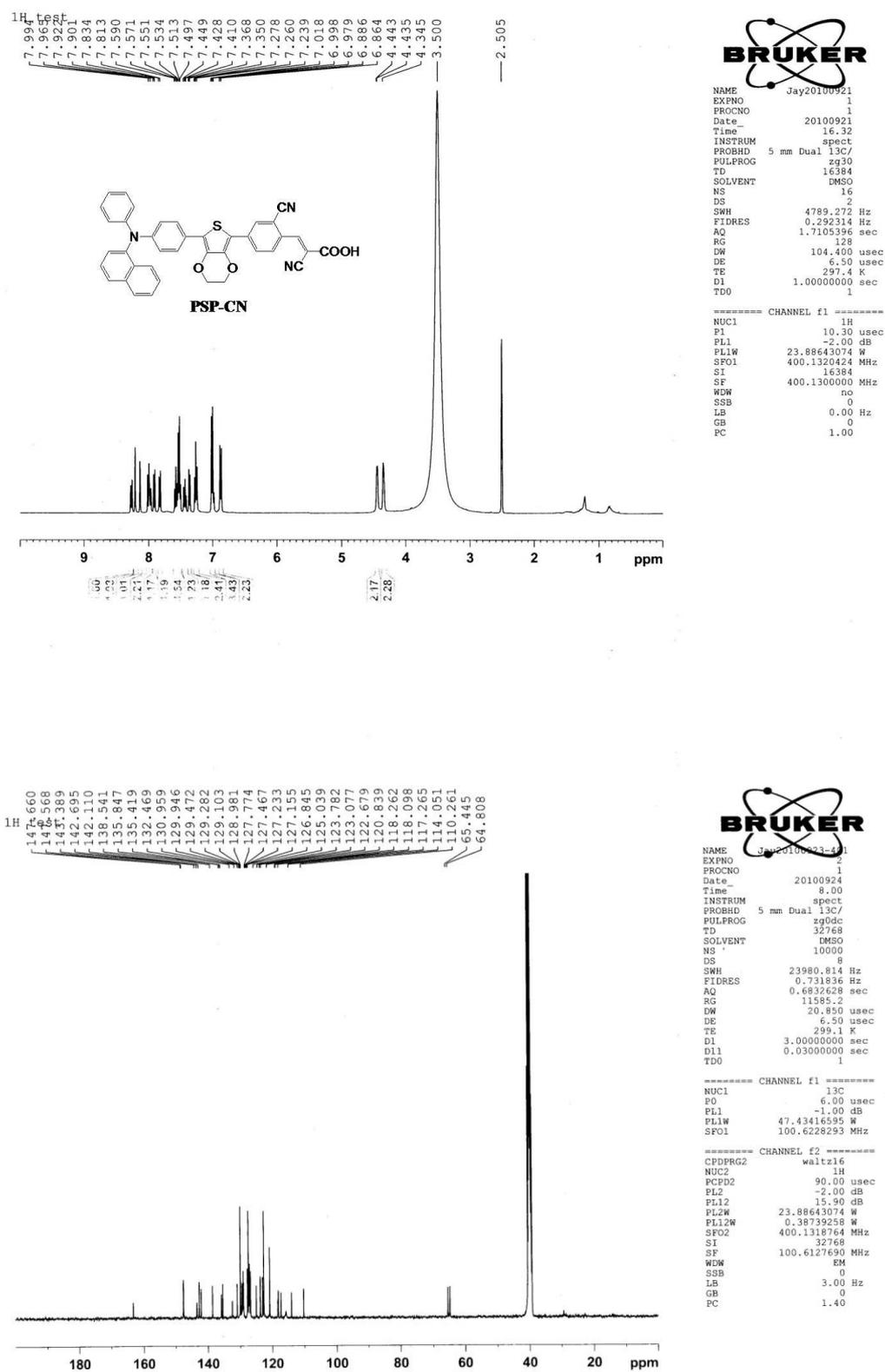


Figure S15: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of PSP-CN in DMSO- d_6 .

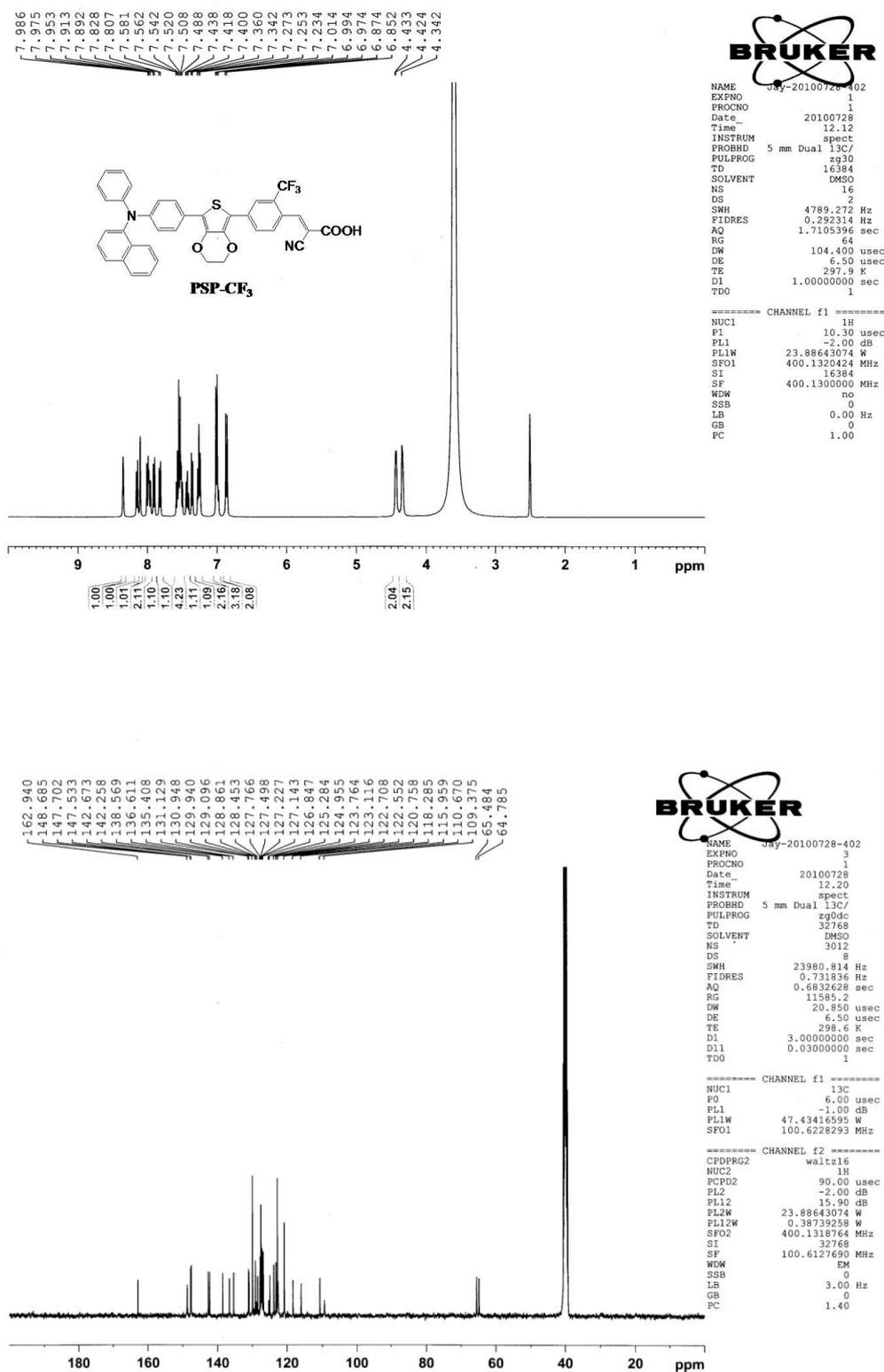


Figure S16: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **PSP-CF₃** in $\text{DMSO-}d_6$.

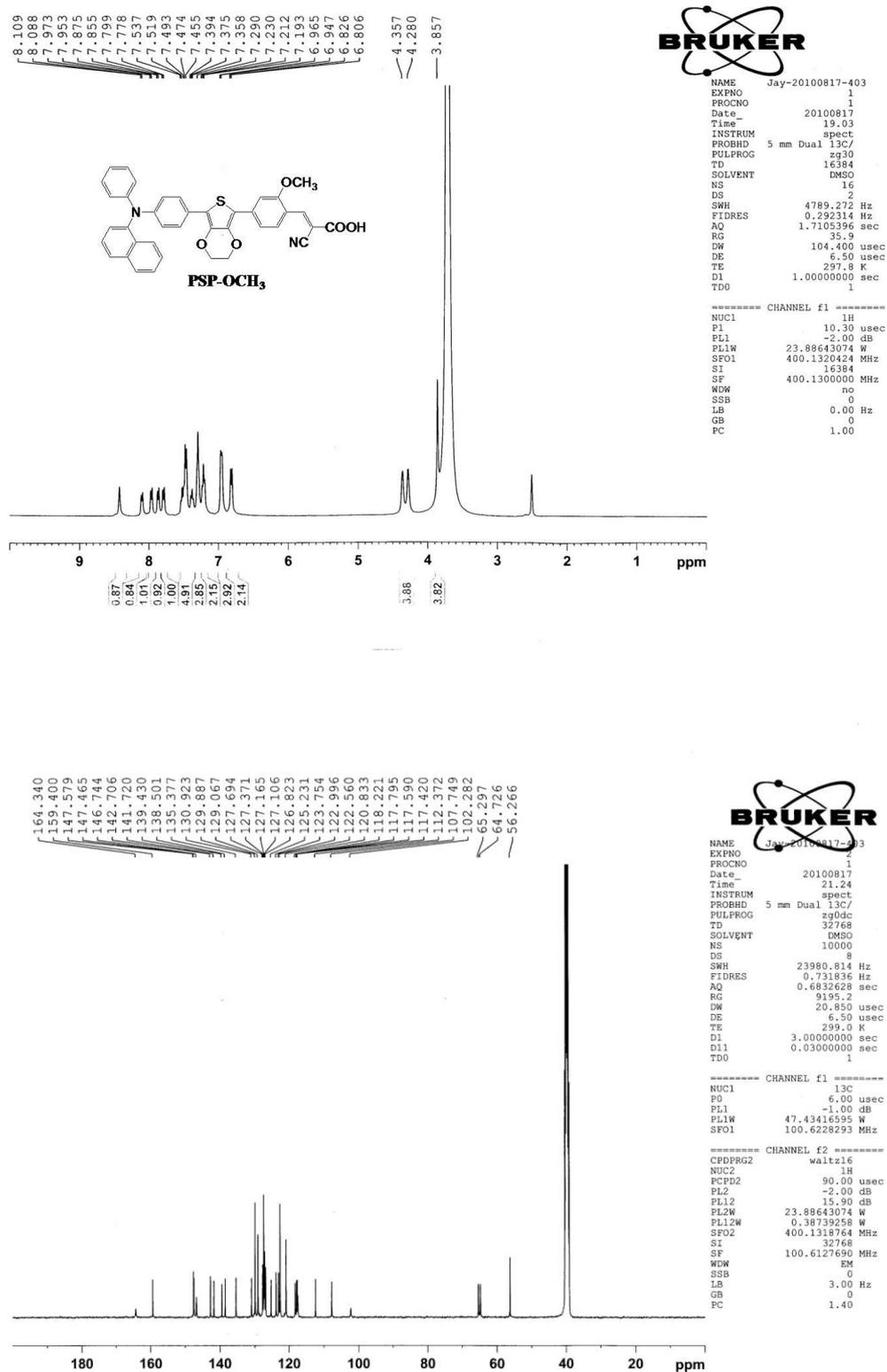


Figure S17: ^1H NMR (upper) and ^{13}C NMR (lower) spectra of **PSP-OCH₃** in $\text{DMSO-}d_6$.

3. Infrared spectra

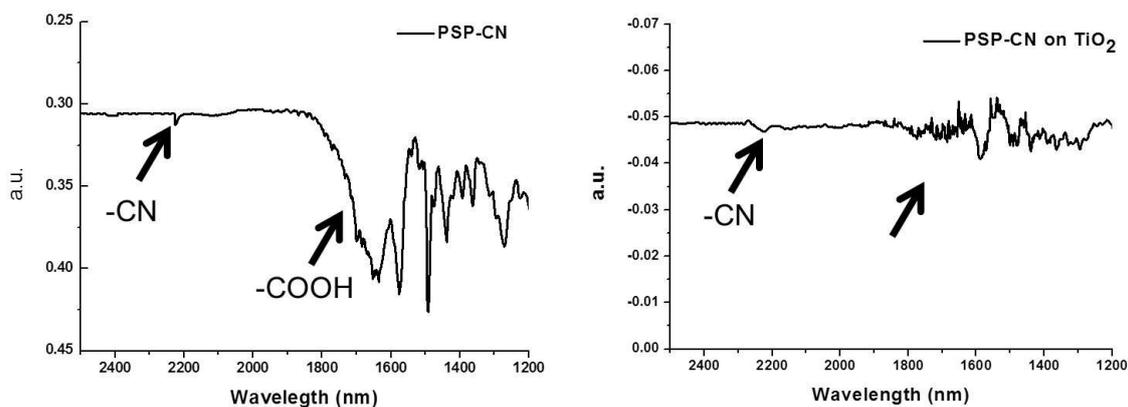


Figure S18. The IR spectra of **PSP-CN** (Left: neat; Right: dye on TiO₂ film).

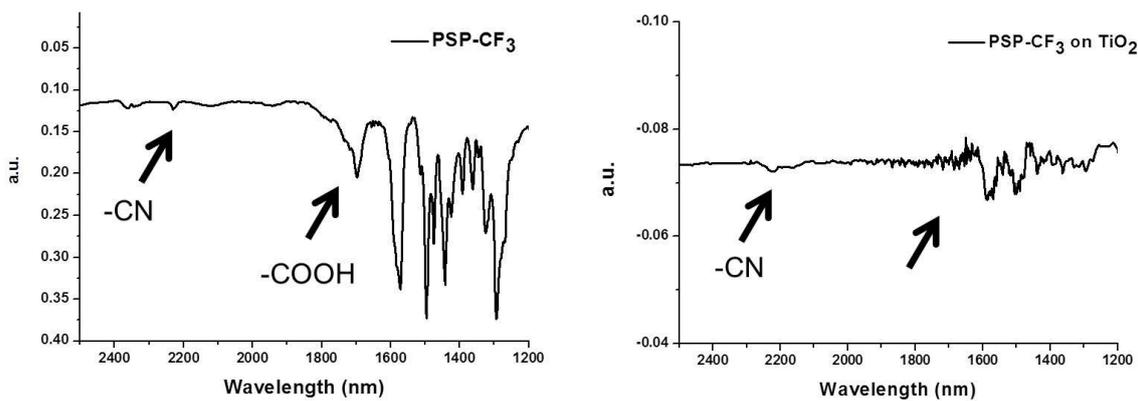


Figure S19. The IR spectra of **PSP-CF₃** (Left: neat; Right: dye on TiO₂ film).

4. Theoretical calculation

Table S1. Calculated Low-Lying Transition for dyes.

dye	state	excitation ^a	λ_{cal} (eV, nm)	f^b B3LYP/631G*	HOMO/LUMO
N-PSP	S1	99.40% H→L	2.26(550)	0.7740	-5.12/ -2.63
	S2	71.30% H-1→L	3.08(402)	0.7722	
	S3	54.90% H-1→L	3.20(388)	0.0913	
PSP-H	S1	99.18% H→L	2.28(554)	0.8956	-4.96/ -2.47
	S2	70.86% H→L+1	3.04(404)	0.6274	
	S3	62.41% H-1→L	3.14(395)	0.1592	
PSP-F	S1	99.26% H→L	2.25(550)	0.8884	-5.02/ -2.55
	S2	77.40% H-1→L	3.04(408)	0.6846	
	S3	64.74% H→L+1	3.16(392)	0.1256	
PSP-CN	S1	99.52% H→L	2.06(600)	0.7257	-5.10/ -2.84
	S2	92.46% H-1→L	2.90(427)	0.6711	
	S3	51.34% H→L+1	3.13(396)	0.1800	
PSP-CF ₃	S1	99.48% H→L	2.13(582)	0.8416	-5.02/ -2.69
	S2	90.38% H-1→L	2.94(422)	0.6589	
	S3	65.35% H→L+1	3.15(393)	0.1248	
PSP-OCH ₃	S1	98.88% H→L	2.37(523)	0.9994	-4.96/ -2.36
	S2	49.85% H→L+1	3.07(494)	0.4682	
	S3	46.06% H-1→L	3.14 (395)	0.3173	

^aH=HOMO, L=LUMO, H+1=HOMO+1, L+1=LUMO+1, and L+2=LUMO+2. ^bOscillator strengths.

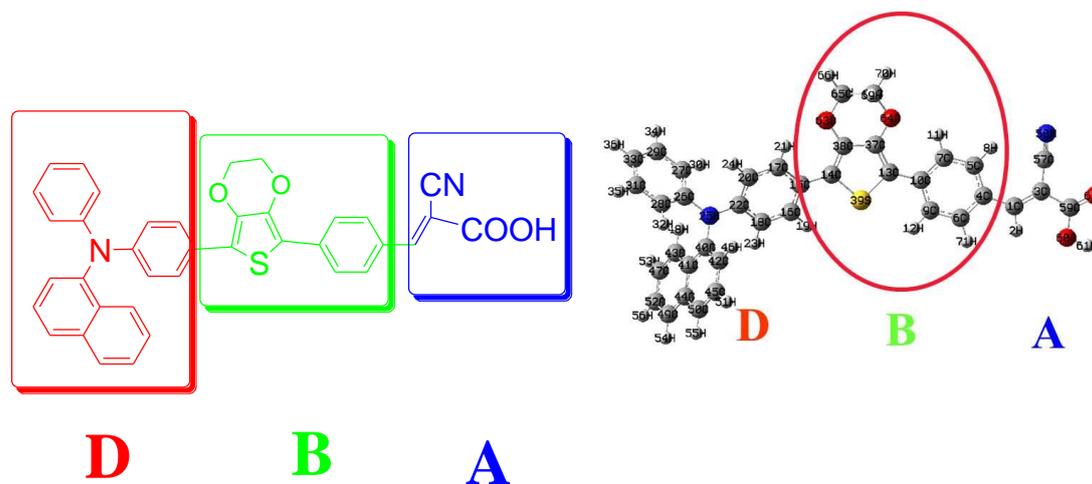
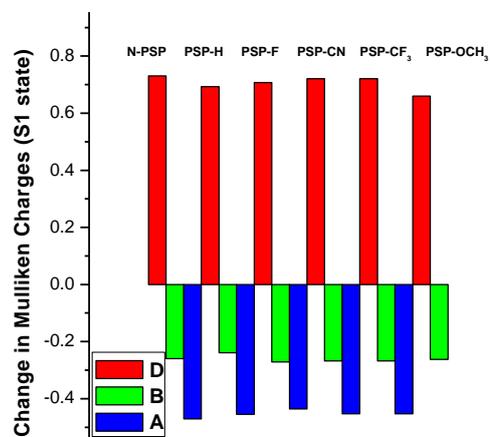


Table S2. Difference of Mulliken charges between ground state (S_0) and excited state (S_1), estimated by time dependent DFT/B3LYP model.

dye	D	B	A
N-PSP	0.73091	-0.26014	-0.47078
PSP-H	0.69312	-0.23877	-0.45435
PSP-F	0.70760	-0.27141	-0.43619
PSP-CN	0.75501	-0.30760	-0.44741
PSP-CF ₃	0.72077	-0.26813	-0.45263
PSP-OMe	0.66031	-0.26239	-0.39792

Difference of Mulliken charge between ground state and excited state.



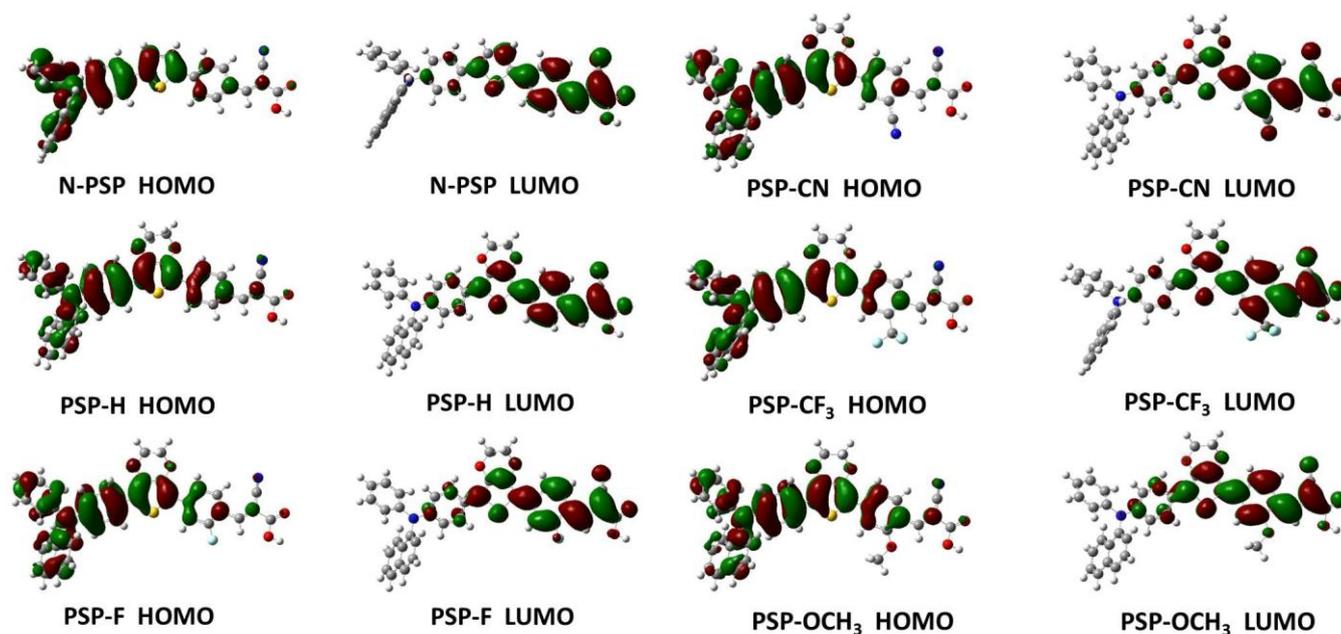


Figure S20. Bar-chart plots for the difference of Mulliken charge listed in Table S2.

Figure S21. Computed HOMO and LUMO orbitals of PSP-series compounds.

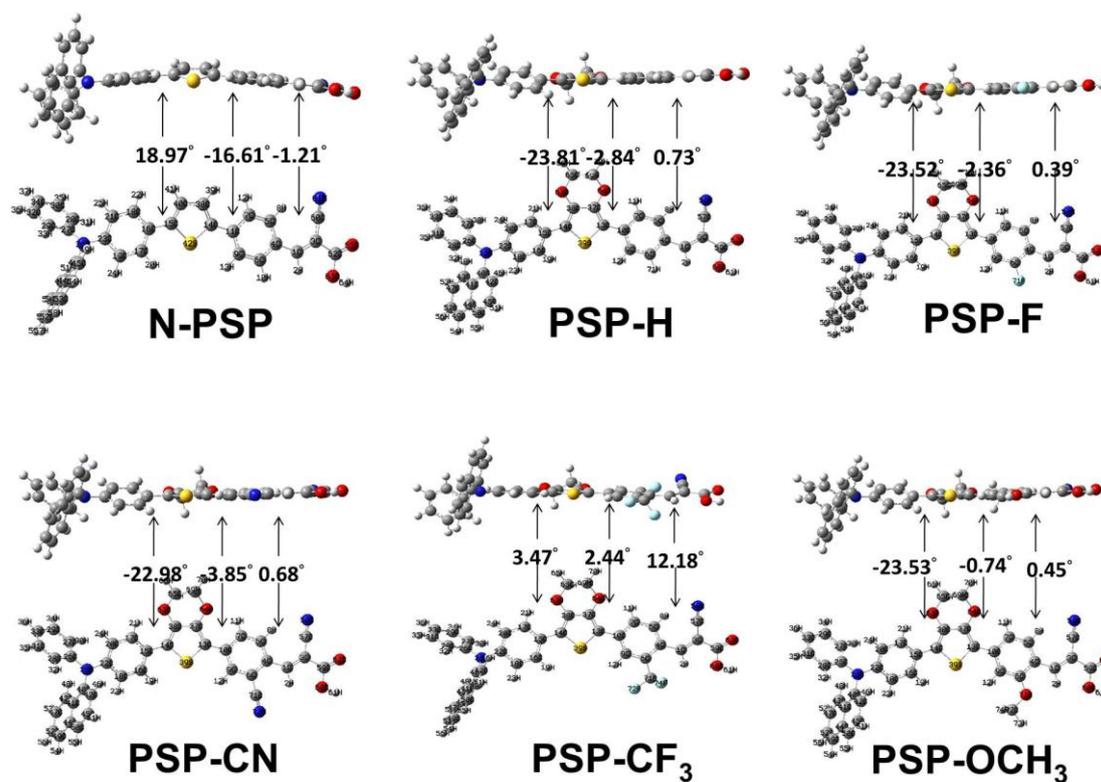


Figure S22. Computed dihedral angles of PSP-series.

5. Loading amount

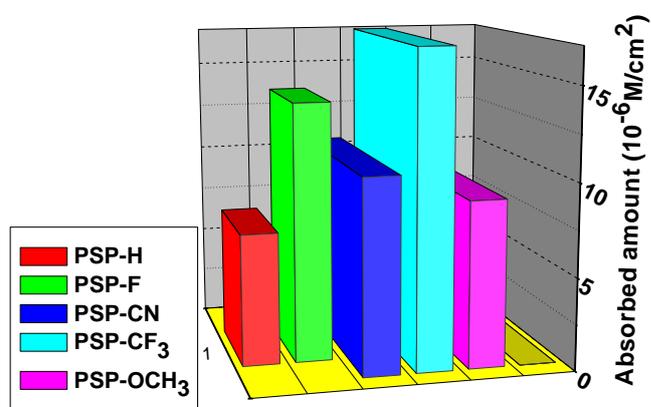


Figure S23. Loading amount of organic dyes on TiO₂ film.

6. HOMO-LUMO level

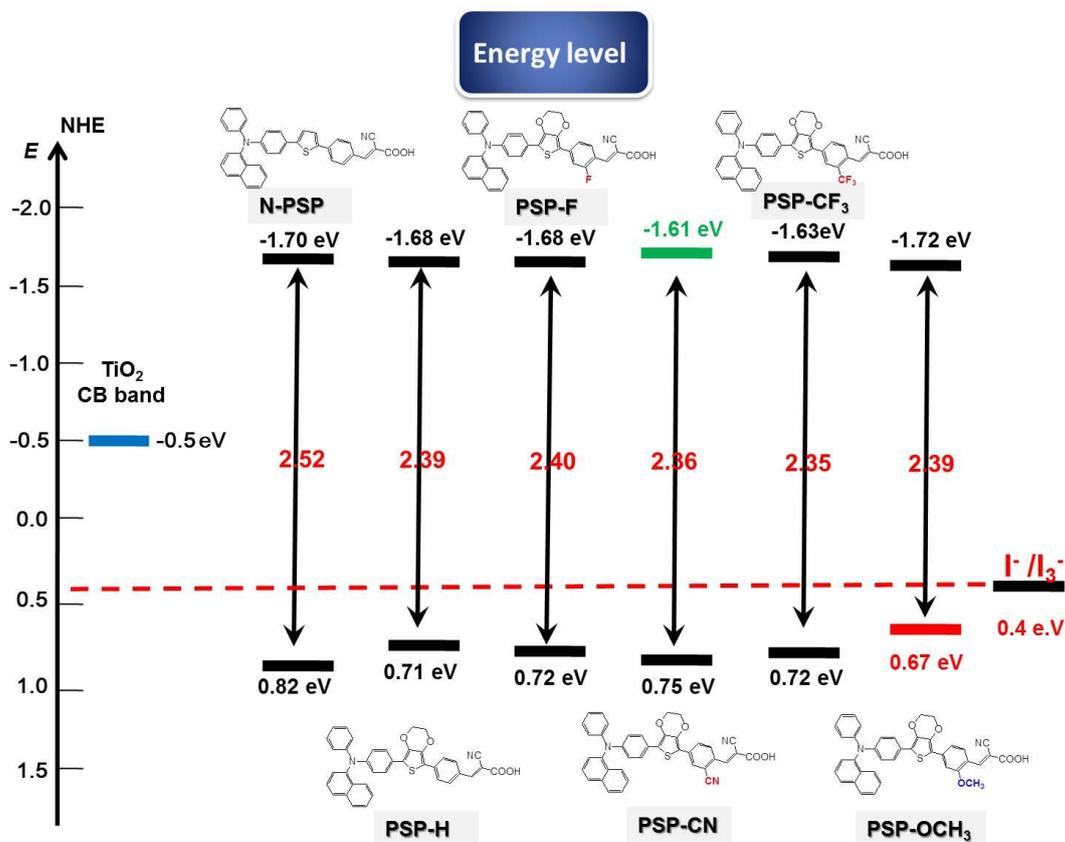


Figure S24. HOMO - LUMO energy levels of PSP-series.

7. EIS spectra

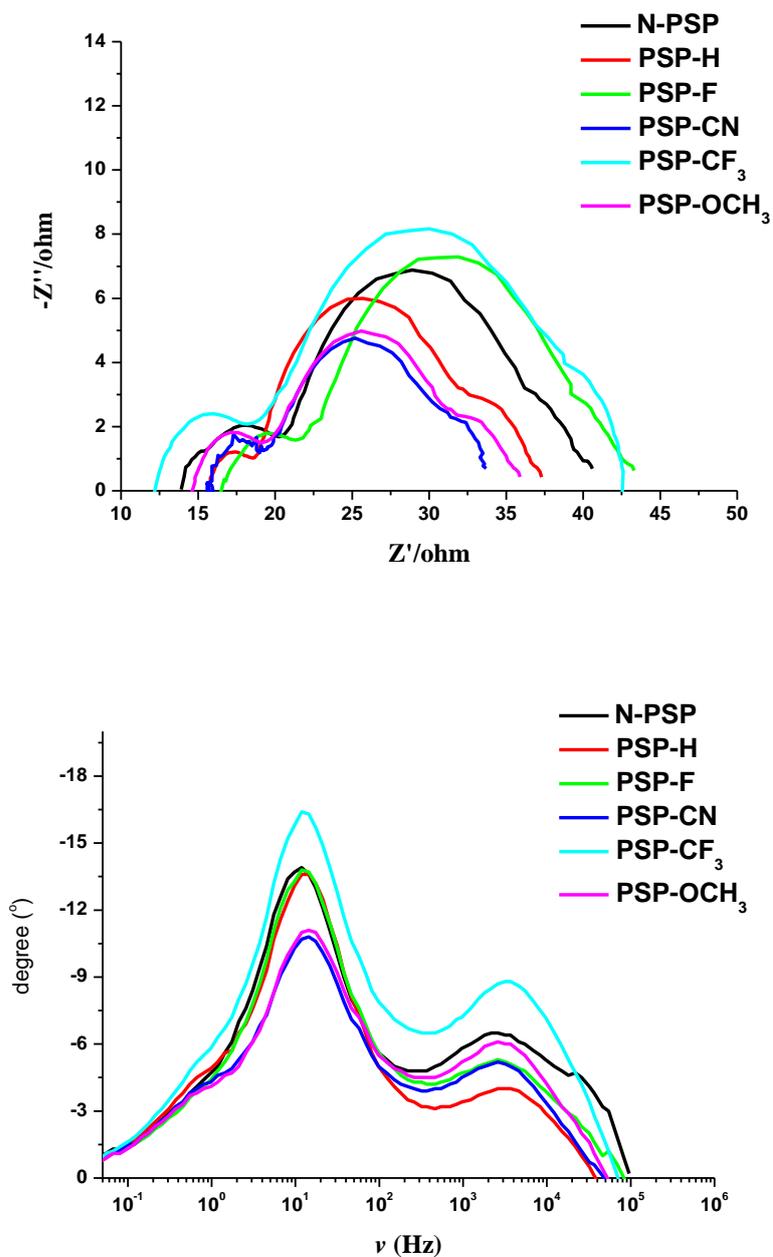


Figure S25. Impedance spectra of **PSP**-series organic dyes at -0.73 V bias in the dark. (upper) Nyquist plots; (lower) Bode phase plots.