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Size-Tunable Mesoporous Spherical TiO₂ as a Scattering Overlayer in High-Performance Dye-Sensitized Solar Cells

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Preparation of pore size-controlled MS TiO₂.

Experiments Mesoporous spherical TiO₂ (MS TiO₂) with large surface area and high reflectance was obtained through "dilute mixing"-driven hydrolysis and subsequent solvothermal treatement as shown in **Schematic S1**. For a titanium source, widely used titanium alkoixdes were tested. Titanium tetraethoxide (TTE) was chosen to produce size-tunable amorphous TiO₂ spheres along with [TTE]/[water] concentration. Then, the amorphous TiO₂ spheres underwent sovothermal treatment at 160 °C for 16 hours, resulting in the MS TiO₂ with particle diameter variation (1554 nm for T1, 757 nm for T2, 587 nm for T3). As-prepared three different MS TiO₂ particles interestingly possess the same primary pore sizes of around 10 nm. With the presented synthetic route, not only the particle size but the primary pore size can be adjusted simply by adding NH₃ during the solvothermal process. To investigate the effect of NH₃ on the internal pore sizes, 0-2 mL of NH₃ was added at a fixed [TTE]/[water] ratio (synthetic condition for T3). Therefore, three NH₃-MS TiO₂ samples (T3-0 mL NH₃, T3-1 mL NH₃, T3-2 mL NH₃) were prepared.





Schematic S1. (Color online). Synthetic procedure of size-tunable MS TiO₂.

Nitrogen sorption analysis of NH₃-MS TiO₂ samples Nitrogen isotherms of three NH₃-MS TiO₂ samples are shown in the Figure S1. All the samples had type IV isotherms and H1 hysteresis at high relative pressure region (P/P₀=0.8-0.9), indicating the presence of meso-scale internal pores. Figure S2 shows pore size distribution of each sample. It was observed that pore size was increased as the NH₃ concentration was increased. As summarized in Table S1, T3-0 mL NH₃ had the highest surface area of 110.6 m²/g with the smallest pore size of 8.57 nm and crystallite size of 9.21 nm. After adding 1 and 2 mL of NH₃, their surface areas were reduced with increased pore and crystallite sizes.



Figure S1. (Color online). Nitrogen sorption isotherms of the NH_3 -MS TiO₂ samples (T3-0 mL NH_3 , T3-1 mL NH_3 , T3-2 mL NH_3). The isotherms of T3-1 mL and T3-2 mL are shifted by 100 and 200 cm³ g⁻¹, respectively, for the sake of clear presentation.



Figure S2. (Color online). Pore size distribution of three NH₃-MS TiO₂ samples (pore size: 8.57 nm, 12.85 nm, 15.87 nm)

Smaple	Surface area	Pore size	d _{xrd}
	(m^2/g)	(nm)	(nm) ^[a]
T3-0 mL NH ₃	110.6	8.57	9.21
T3-1 mL NH ₃	80.2	12.85	14.94
T3-2 mL NH ₃	76.0	15.87	15.79

[a] The crystal size was calculated by applying the Scherrer equation to the anantase (101) peak from the XRD spectra.

Table S1. Physical properties of samples T3-0,1,2 mL NH₃.

Diffuse reflection spectra Light scattering effect can be quantified by measuring the diffuse reflection spectra. T3-0mL NH₃ exhibited the highest reflectance among three NH₃- MS TiO₂ samples over whole spectral range (**Figure S3**). T3-1 mL NH₃, T3-2 mL NH₃ ranked behind T3-0 mL NH₃.



Figure S3. (Color online). Diffuse reflectance of three films made of the NH₃-MS TiO₂ samples.

It is reasonable to conclude that MS TiO₂ without adding NH₃ (T3-0 mL NH₃) is the best choice for a scattering overlayer in dye-sensitized solar cells to achieve high photovoltaic performance, due to the largest surface area for dye-loading and the highest reflectance value for light scattering. Therefore, three distinct size-tunable MS TiO₂ particles were prepared without addition of NH₃ in the solvothermal treatment in the present research. Since as-synthesized MS TiO₂ particles possess almost same pore and crystallite sizes, size-dependent scattering efficiency of hierarchical structures on the light harvest could be studied. NH₃-added MS TiO₂ particles are also being tested as a main layer that can facilitate the dual effects of efficient dye loading and optical confinement.