

Supplemental Materials

Geometrical Effect of Stilbene on the Performance of Organic Dyes-Sensitized Solar Cells

Yan-Duo Lin and Tahsin J. Chow*

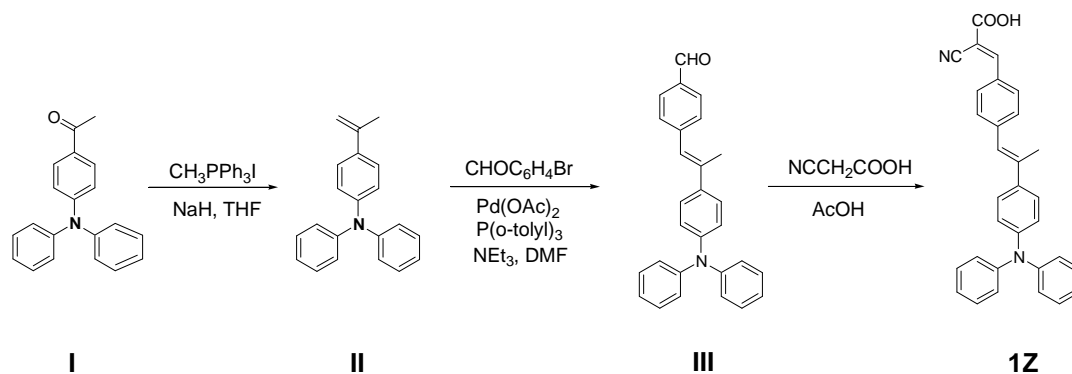
Institute of Chemistry, Academia Sinica, Taipei 115 Taiwan

tjchow@chem.sinica.edu.tw

Table of contents:

Page	Contents
S2~S3	Synthesis of compound 1Z
S4	Fig. S1 Absorption spectra of (a) 1P , (b) 1N , (c) 1M , (d) 1T , and (e) 1Z in different solvents
S5	Fig. S2 Absorption spectra of (a) 2P , (b) 2N , (c) 2M , and (d) 2T in different solvents
S6	Fig. S3 Frontier orbitals of 1N (a), 2N (b), 1M (c), 2M (d), 1T (e), and 2T (f), optimized with DFT at the B3LYP/6-31G(d,p) level
S7	Fig. S4 Optimized geometry of 1Z and its electronic maps of Frontier orbitals
S8	Fig. S5 The electrochemical impedance spectra of (a) Nyquist plots, and (b) Bode phase plots for the DSSCs made with dye 1Z
S9	Table S1. Calculated TDDFT excitation energies (<i>E</i>), Oscillator strengths (<i>f</i>), MO compositions and characters, are compared with experimental absorptions for all of dyes

Synthesis of compound 1Z



(4-Isopropenyl-phenyl)diphenylamine (II). Methyltriphenylphosphonium iodide (6.47 g, 16.0 mmol) was dissolved in THF (53 mL). To it was added NaH (1.92 g, 48.0 mmol) slowly in an ice-water bath, and then the mixture was stirring for another 2 h at ambient temperature. Compound I (2.3 g, 8.0 mmol) in THF (23 mL) was added dropwise to this reaction mixture by using a pressure-equalizing addition funnel. After the completion of reaction, methanol was added to quench excess NaH in an ice-water bath and then THF was removed by rotary evaporation. The reaction mixture was extracted with ethyl acetate. The organic layer was dried over anhydrous MgSO_4 and filtered. The filtrate was concentrated under reduced pressure. Column chromatograph eluted with ethyl acetate/hexane (1/25) afforded the desired product as a white solid (1.43 g, 63% yield). Mp 65–66 °C Spectral data of II: δ_{H} (400 MHz, CDCl_3) 7.38 (d, $J=8.1$ Hz, 1H), 7.29–7.25 (m, 4H), 7.13–7.02 (m, 8H), 5.35 (s, 1H), 5.04 (s, 1H), 2.16 (s, 1H) ppm; δ_{C} (100 MHz, CDCl_3) 147.72, 147.16, 142.48, 135.24, 129.22, 126.24, 124.32, 123.40, 122.83, 111.05, 29.70 ppm. m/z (FAB) 285.1519 ((M^+)). $\text{C}_{21}\text{H}_{19}\text{N}$ requires 285.1517).

4-[2-(4-Diphenylaminophenyl)propenyl]benzaldehyde (III). Compound III was synthesized via a typical Heck reaction procedure as described above. Column chromatography eluted with ethyl acetate/hexane (1/10) afforded the yellow solid of III (0.46 g, 50% yield). Mp 154–155°C. Spectral data of III: δ_{H} (400 MHz, CDCl_3) 10.11 (s, 1H), 7.89 (d, $J=7.9$ Hz, 2H), 7.52 (d, $J=7.9$ Hz, 2H), 7.43 (d, $J=7.3$ Hz, 2H), 7.30–7.26 (m, 4H), 7.17–7.03 (m, 8H), 6.86 (s, 1H), 2.31 (s, 3H) ppm; δ_{C} (100 MHz,

CDCl₃) 191.67, 147.52, 144.95, 139.70, 136.86, 134.25, 132.39, 130.92, 129.64, 129.62, 126.76, 125.29, 124.52, 123.10, 113.81, 17.54. *m/z* (FAB) 389.1785 (M⁺. C₂₈H₂₃NO requires 389.1780).

2-Cyano-3-{4-[2-(4-diphenylaminophenyl)propenyl]phenyl}acrylic acid (1Z). Compound **1Z** was synthesized according to the standard Knoevenagel as described above. Orange solid of **1Z** was afforded in 65%. Mp 242–243°C. Spectral data of **1Z**: δ_{H} (400 MHz, CDCl₃) 8.27 (s, 1H), 8.01 (d, *J* = 8.5 Hz, 2H), 7.47 (d, *J* = 6.8 Hz, 2H), 7.40 (d, *J* = 7.6 Hz, 2H), 7.28–7.24 (m, 4H), 7.12–7.02 (m, 8H), 6.81 (s, 1H), 2.31 (s, 3H) ppm; δ_{C} (100 MHz, CDCl₃) 167.01, 155.87, 147.73, 147.52, 144.63, 140.56, 136.78, 131.55, 129.97, 129.32, 129.03, 126.84, 125.18, 124.62, 123.18, 123.03, 115.61, 17.73. *m/z* (FAB) 456.1840 (M⁺. C₃₁H₂₄N₂O₂ requires 456.1838).

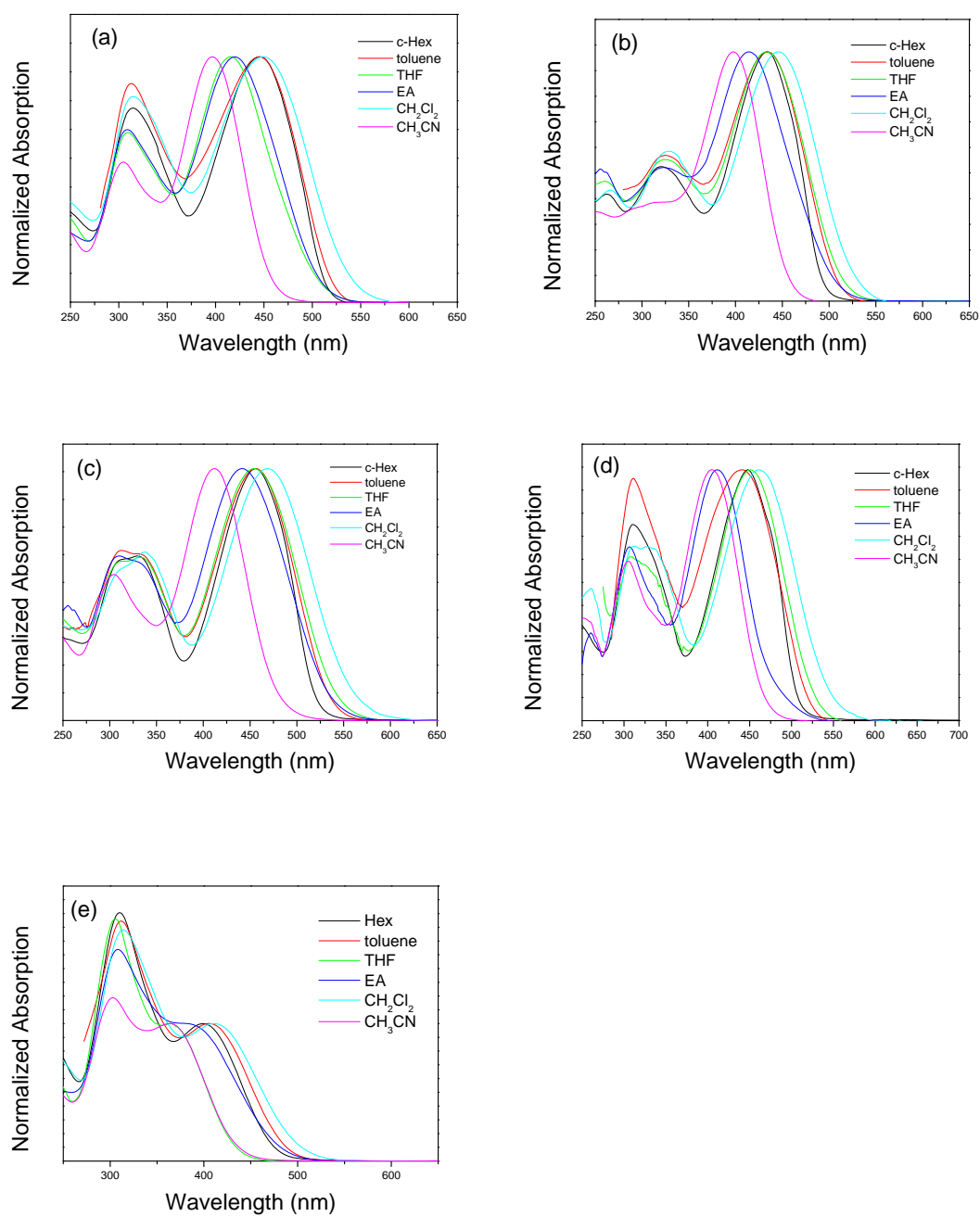


Fig. S1 Absorption spectra of (a) **1P**, (b) **1N**, (c) **1M**, (d) **1T**, and (e) **1Z** in different solvents

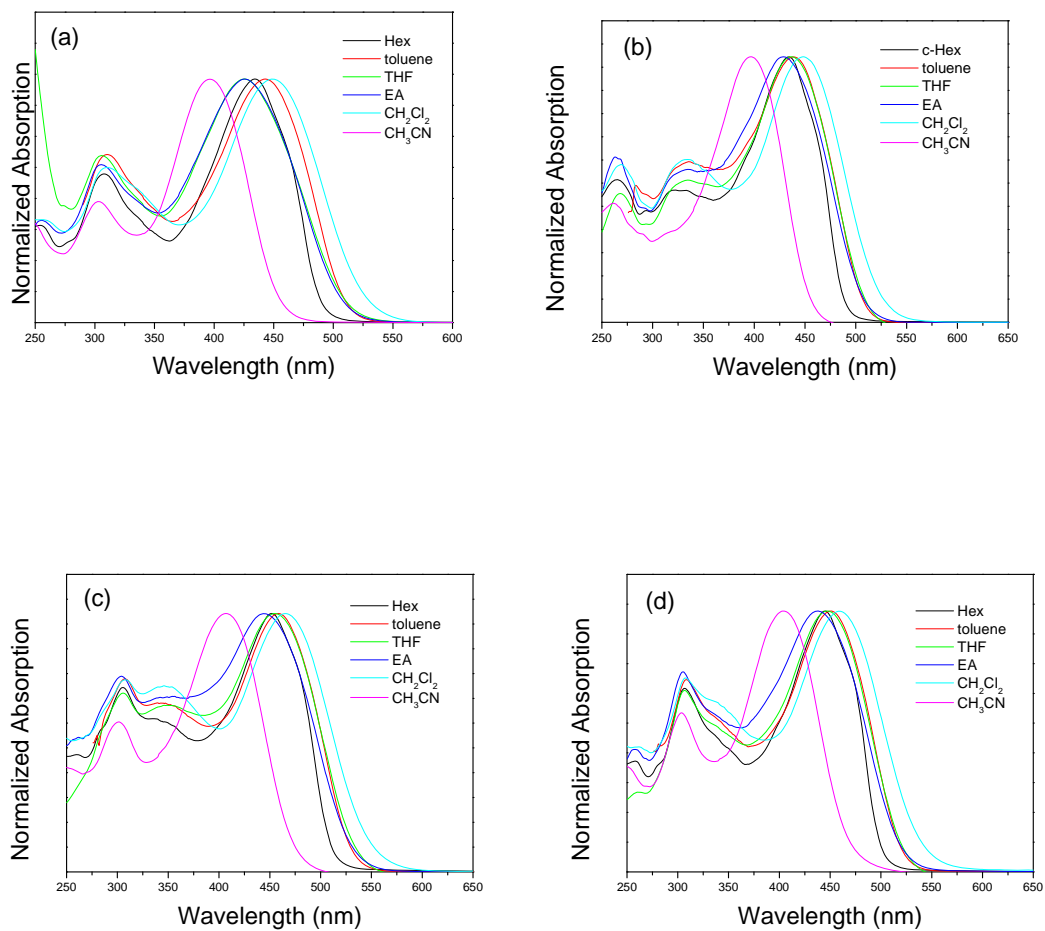


Fig. S2 Absorption spectra of (a) **2P** (b) **2N** (c) **2M**, and (d) **2T** in different solvents

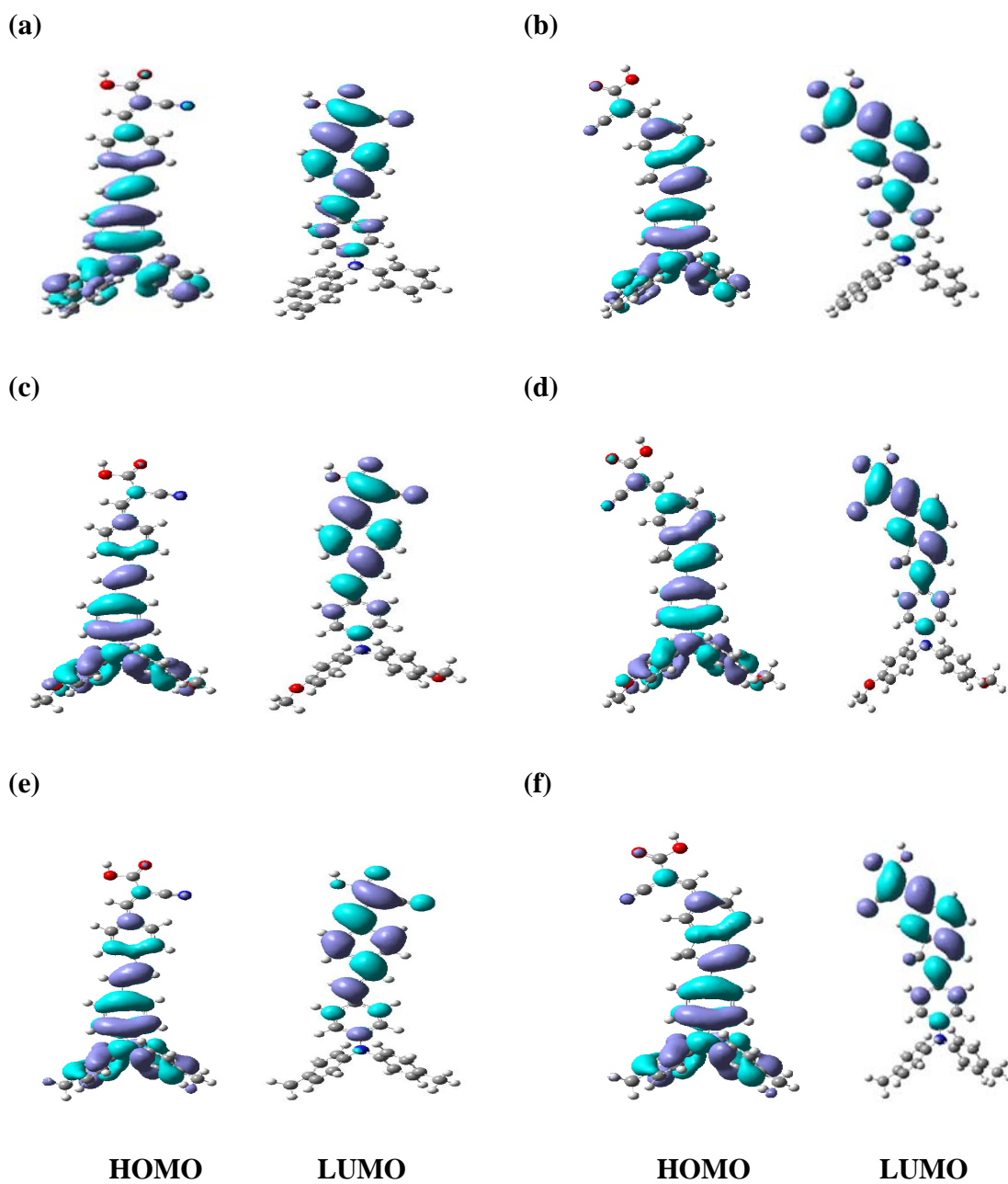


Fig. S3 Frontier orbitals of (a) 1N, (b) 2N, (c) 1M, (d) 2M, (e) 1T, and (f) 2T, optimized with DFT at the B3LYP/6-31G(d,p) level

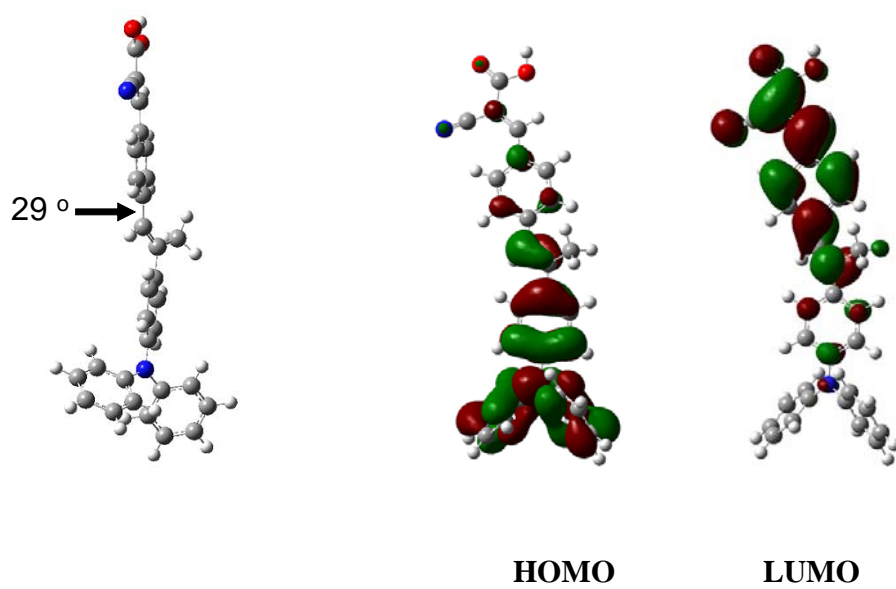


Fig. S4 Optimized geometry of **1Z** and its electronic maps of Frontier orbitals

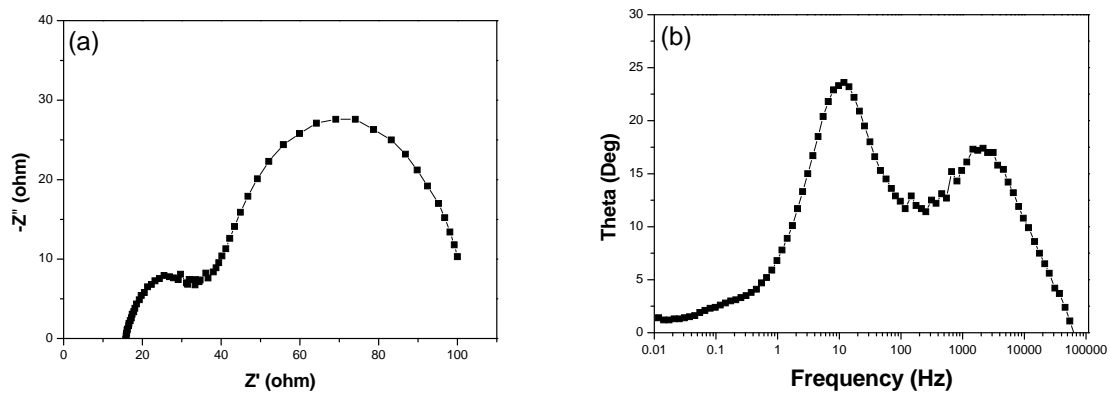


Fig. S5 The electrochemical impedance spectra of (a) Nyquist plots, and (b) Bode phase plots for the DSSCs made with dye **1Z**.

Table S1. Calculated TDDFT excitation energies (E), oscillator strengths (f), MO compositions and characters, are compared with experimental absorptions for all of dyes.

dye	n ^a	E (ev, nm)	f	composition	Character	exptl (ev, nm)
1P	1	2.32 (534)	0.95	98% HOMO→LUMO	CT	2.86 (433)
	2	3.31 (375)	0.98	82% HOMO-1→LUMO	π - π^* (1)	4.00 (310)
	5	4.01 (309)	0.13	88% HOMO→LUMO+3	π - π^* (2)	
2P	1	2.43 (510)	0.88	98% HOMO→LUMO	CT	2.82 (439)
	2	3.33 (372)	0.69	79% HOMO-1→LUMO	π - π^* (1)	4.06 (305)
	6	4.00 (310)	0.14	93% HOMO→LUMO+3	π - π^* (2)	
1N	1	2.34 (530)	1.00	99% HOMO→LUMO	CT	2.86 (433)
	2	3.18 (390)	0.29	80% HOMO→LUMO+1	π - π^* (1)	3.83 (324)
	6	3.29 (377)	0.33	77% HOMO-1→LUMO	π - π^* (2)	
2N	1	2.51 (494)	1.06	96% HOMO→LUMO	CT	2.84 (437)
	2	3.22 (385)	0.17	83% HOMO→LUMO+1	π - π^* (1)	3.30 (335)
	5	3.43 (361)	0.59	71% HOMO-1→LUMO	π - π^* (2)	
1M	1	2.14 (579)	0.87	99% HOMO→LUMO	CT	2.73 (455)
	2	3.16 (392)	1.05	87% HOMO-1→LUMO	π - π^* (1)	3.99 (311)
2M	1	2.30 (539)	0.82	99% HOMO→LUMO	CT	2.73 (455)
	2	3.22 (385)	0.76	86% HOMO-1→LUMO	π - π^* (1)	4.07 (305)
	4	3.55 (349)	0.13	80% HOMO→LUMO+1	π - π^* (2)	
1T	1	2.25 (551)	0.93	99% HOMO→LUMO	CT	2.84 (437)
	2	3.26 (380)	1.03	75% HOMO-1→LUMO	π - π^* (1)	4.04 (307)
2T	1	2.34 (530)	0.82	99% HOMO→LUMO	CT	2.77 (447)
	2	3.27 (379)	0.78	81% HOMO-1→LUMO	π - π^* (1)	4.04 (307)
1Z	1	2.29 (542)	0.53	99% HOMO→LUMO	CT	3.40 (365)
	2	3.33 (373)	1.01	88% HOMO-1→LUMO	π - π^* (1)	4.07 (305)
	3	3.67 (338)	0.18	85% HOMO→LUMO+1	π - π^* (2)	
	6	4.04 (307)	0.15	78% HOMO→LUMO+3	π - π^* (3)	
	10	4.24 (292)	0.11	75% HOMO-4→LUMO	π - π^* (4)	

^a Sequence of calculated transitions in order of energy.