Electronic Supplementary Material (ESI) for Journal of Materials Chemistry This journal is  $\ensuremath{\mathbb{C}}$  The Royal Society of Chemistry 2011

### Supplementary information

### Fused Alq<sub>3</sub> derivatives: syntheses and photophysical characteristics

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<sup>1</sup>H NMR spectrum of 8-(benzyloxy)-4-chloroquinoline (2)



<sup>13</sup>C NMR spectrum of 8-(benzyloxy)-4-chloroquinoline (2)



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### HRMS of 8-(benzyloxy)-4-chloroquinoline (2)

Elemental Composition Report

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

116 formula(e) evaluated with 4 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
270.0679	270.0681	-0.1	-0.5	15.0	C19 H10 O2
	270.0686	-0.6	-2.4	10.5	C16 H13 N O Cl
	270.0662	1.8	6.6	7.5	C14 H14 N O Na Cl
	270.0657	2.3	8.4	12.0	C17 H11 O2 Na





<sup>1</sup>H NMR spectrum of 4-benzyloxy-11*H*-indol[3,2-*c*]quinoline (**3**)



<sup>13</sup>C NMR spectrum of 4-benzyloxy-11*H*-indol[3,2-*c*]quinoline (**3**)



# Electronic Supplementary Material (ESI) for Journal of Materials Chemistry This journal is ${\ensuremath{\mathbb O}}$ The Royal Society of Chemistry 2011

### HRMS of 4-benzyloxy-11*H*-indol[3,2-*c*]quinoline (**3**)

Elemental Composition Report

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

### Monoisotopic Mass, Odd and Even Electron Ions

74 formula(e) evaluated with 2 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
325.1347	325.1341	0.6	1.9	15.5	C22 H17 N2 O
	325.1317	3.0	9.3	12.5	C20 H18 N2 O Na

J222



<sup>1</sup>H NMR spectrum of 11*H*-indolo[3,2-*c*]quinolin-4-ol (4)



<sup>13</sup>C NMR spectrum of 11*H*-indolo[3,2-*c*]quinolin-4-ol (4)



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### HRMS of 11*H*-indolo[3,2-*c*]quinolin-4-ol (4)

Elemental Composition Report

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Even Electron Ions

53 formula(e) evaluated with 1 results within limits (up to 50 closest results for each mass)

235.0889	235.0871	1.7	7.4	11.5	C15 H11 N2 O
Mass	Calc. Mass	mDa	PPM	DBE	Formula
Maximum:		5.0	10.0	100.0	
Minimum:				-0.5	

J226



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<sup>1</sup>H NMR spectrum of 4-((2-bromobenzyl)(methyl)amino)quinolin-8-yl 4-methylbenzenesulfonate (5)



<sup>13</sup>C NMR spectrum of 4-((2-bromobenzyl)(methyl)amino)quinolin-8-yl 4-methylbenzenesulfonate (5)

# Electronic Supplementary Material (ESI) for Journal of Materials Chemistry This journal is ${\ensuremath{\mathbb O}}$ The Royal Society of Chemistry 2011

HRMS of 4-((2-bromobenzyl)(methyl)amino)quinolin-8-yl 4-methylbenzenesulfonate (5)

**Elemental Composition Report** 

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Even Electron Ions

822 formula(e) evaluated with 11 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
497.0531	497.0535	-0.3	-0.6	14.5	C24 H22 N2 O3 S Br
	497.0528	0.3	0.7	5.5	C16 H26 N4 O5 S2 Br
	497.0541	-1.0	-2.0	23.5	C32 H18 O Br
	497.0560	-2.8	-5.7	10.5	C20 H22 N2 O8 Br
	497.0503	2.8	5.7	9.5	C20 H26 N4 S3 Br
	497.0562	-3.0	-6.1	0.5	C13 H30 N4 O5 S3 Br
	497.0501	3.1	6.1	19.5	C27 H18 N2 O3 Br
	497.0568	-3.7	-7.4	9.5	C21 H26 N2 O3 S2 Br
	497.0494	3.7	7.4	10.5	C19 H22 N4 O5 S Br
	497.0490	4.2	8.4	4.5	C19 H30 O4 S3 Br
	497.0575	-4.3	-8.7	18.5	C29 H22 O S Br

#### J264

lct-24913 361 (7.226) AM (Cen,2, 80.00, Ar,5000.0,609.28); Sm (SG, 1x2.00); Sb (4,40.00); Si 100<sub>Π</sub> 744





### <sup>1</sup>H NMR spectrum of 5-methyl-5,6-dihydrodibenzo[c,h][1,6]naphthyridin-1-yl 4-methylbenzenesulfonate (**6**)



### $^{13}$ C NMR spectrum of 5-methyl-5,6-dihydrodibenzo[*c*,*h*][1,6]naphthyridin-1-yl 4-methylbenzenesulfonate (**6**)

# Electronic Supplementary Material (ESI) for Journal of Materials Chemistry This journal is O The Royal Society of Chemistry 2011

### HRMS of 5-methyl-5,6-dihydrodibenzo[c,h][1,6]naphthyridin-1-yl 4-methylbenzenesulfonate (**6**)

**Elemental Composition Report** 

Single Mass Analysis (displaying only valid results) Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

#### 338 formula(e) evaluated with 11 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
439.1104	439.1109	-0.5	-1.1	28.0	C32 H13 N3
	439.1099	0.6	1.3	24.5	C32 H16 O Na
	439.1116	-1.2	-2.7	18.5	C26 H19 N2 O3 S
	439.1092	1.2	2.8	15.5	C24 H20 N2 O3 Na S
	439.1119	-1.5	-3.4	20.0	C27 H18 N3 Na S
	439.1123	-1.8	-4.2	27.5	C34 H15 O
	439.1085	1.9	4.3	25.0	C30 H14 N3 Na
	439.1083	2.2	5.0	23.5	C29 H15 N2 O3
	439.1133	-2.8	-6.4	19.5	C29 H20 O Na S
	439.1143	-3.9	-8.8	23.0	C29 H17 N3 S
	439.1059	4.6	10.4	20.5	C27 H16 N2 O3 Na





<sup>1</sup>H NMR spectrum of 5-methyl-5,6-dihydrodibenzo[c,h][1,6]naphthyridin-1-ol (7)



4

115.38

5806.0

0.0785

8

128.06

6444.3

0.0769

12

130.70

J287.006.esp VerticalScaleFactor = 1 DMSO-d6 0.14 0.13 0.12-Normalized Intensity \_\_\_\_\_\_\_132.30 -130.70 0.11-155.14 43.49 <u>~128.0</u>8128.14 \_\_\_\_\_\_124.85 120.72 -152.02 0.10 140.34 115.38 112.29 54.89 145.48 0.09-0.08 0.07 0.06 TMS 0.05-وبالإغلار واركروا والأليس Chemical Shift (ppm) No. (Hz) (Hz) No. (Hz) No. (Hz) (ppm) Height No. (ppm) Height (ppm) Height No. (ppm) (Hz) Height (ppm) Height 1 43.49 2188.3 0.0968 5 120.72 6074.9 0.0889 9 128.14 6448.3 0.0891 13 132.30 6657.7 0.0961 17 155.14 7807.3 0.0947 2 54.89 2762.0 0.0792 6 123.54 6216.9 0.0767 10 129.28 6505.9 0.0788 14 140.34 7062.3 0.0821 3 112.29 5650.9 0.0797 7 124.85 6282.9 0.0875 11 129.52 6517.6 0.0786 15 145.48 7320.9 0.0748

6577.4

0.0894

16

152.02

7649.9

0.0862

### $^{13}$ C NMR spectrum of 5-methyl-5,6-dihydrodibenzo[*c*,*h*][1,6]naphthyridin-1-ol (7)

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HRMS of 5-methyl-5,6-dihydrodibenzo[*c*,*h*][1,6]naphthyridin-1-ol (7)

Elemental Composition Report

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

64 formula(e) evaluated with 2 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
263.1181	263.1184	-0.3	-1.3	11.5	C17 H15 N2 O
	263.1160	2.1	7.9	8.5	C15 H16 N2 O Na



LCT-24998 14 (0.281) AM (Cen.2, 90.00, Ar,5000.0,609.28); Sm (SG, 1x2.00); Sb (4,40.00) 263.1181 243





### <sup>1</sup>H NMR spectrum of 4-(2-bromobenzyloxy)quinolin-8-yl 4-methylbenzenesulfonate (9)



### <sup>13</sup>C NMR spectrum of 4-(2-bromobenzyloxy)quinolin-8-yl 4-methylbenzenesulfonate (9)

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HRMS of 4-(2-bromobenzyloxy)quinolin-8-yl 4-methylbenzenesulfonate (9)

**Elemental Composition Report** 

Single Mass Analysis (displaying only valid results) - displaying only valid results

Tolerance = 3.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Even Electron Ions

387 formula(e) evaluated with 4 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		3.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
484.0208	484.0216	-0.8	-1.6	11.5	C15 H15 N7 O7 Br
	484.0218	-1.0	-2.0	14.5	C23 H19 N O4 S Br
	484.0191	1.7	3.5	15.5	C19 H15 N7 O2 S Br
	484.0184	2.4	5.0	19.5	C26 H15 N O4 Br

J254

LCT-24762 24 (0.482) AM (Cen,2, 70.00, Ar,5000.0,556.28); Sm (SG, 1x2.00); Sb (4,40.00 ) 484.0208 2.23e3





<sup>1</sup>H NMR spectrum of 6*H*-isochromeno[4,3-*c*]quinolin-1-yl 4-methylbenzenesulfonate (10)



### $^{13}$ C NMR spectrum of 6*H*-isochromeno[4,3-*c*]quinolin-1-yl 4-methylbenzenesulfonate (10)

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HRMS of 6*H*-isochromeno[4,3-*c*]quinolin-1-yl 4-methylbenzenesulfonate (10)

**Elemental Composition Report** 

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Even Electron Ions

393 formula(e) evaluated with 7 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
404.0950	404.0955	-0.4	-1.1	12.5	C15 H14 N7 O7
	404.0957	-0.6	-1.5	15.5	C23 H18 N O4 S
	404.0930	2.1	5.1	16.5	C19 H14 N7 O2 S
	404.0923	2.7	6.8	20.5	C26 H14 N O4
	404.0916	3.4	8.4	11.5	C18 H18 N3 O6 S
	404.0988	-3.8	-9.4	7.5	C12 H18 N7 O7 S
	404.0995	-4.5	-11.0	16.5	C20 H14 N5 O5



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<sup>1</sup>H NMR spectrum of 6*H*-isochromeno[4,3-*c*]quinolin-1-ol (**11**)



<sup>13</sup>C NMR spectrum of 6*H*-isochromeno[4,3-*c*]quinolin-1-ol (**11**)



# Electronic Supplementary Material (ESI) for Journal of Materials Chemistry This journal is ${\ensuremath{\mathbb O}}$ The Royal Society of Chemistry 2011

### HRMS of 6*H*-isochromeno[4,3-*c*]quinolin-1-ol (11)

Elemental Composition Report

Single Mass Analysis (displaying only valid results)

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

#### Monoisotopic Mass, Even Electron Ions

106 formula(e) evaluated with 2 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
250.0873	250.0868	0.5	2.1	11.5	C16 H12 N O2
	250.0828	4.5	18.2	7.5	C11 H12 N3 O4

J259

LCT-24905 280 (5.604) AM (Cen,2, 70.00, Ar,5000.0,556.28); Sm (SG, 1x2.00); Sb (4,40.00 ) 100- 1.14e3



<sup>1</sup>H NMR of tris-(11*H*-indolo[3,2-*c*]quinolin-4-ol)aluminum (**12**)



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HRMS of tris-(11*H*-indolo[3,2-*c*]quinolin-4-ol)aluminum (12)

Elemental Composition Report

Single Mass Analysis (displaying only valid results) Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

616 formula(e) evaluated with 8 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
749.1842	749.1850	-0.8	-1.1	46.5	C58 H26 A1
	749.1858	-1.6	-2.1	35.5	C45 H27 N6 O3 Na Al
	749.1826	1.6	2.2	43.5	C56 H27 Na Al
	749.1871	-2.9	-3.9	35.0	C47 H29 N3 O4 Na Al
	749.1810	3.2	4.3	42.5	C53 H26 N2 O2 A1
	749.1882	-4.0	-5.3	38.5	C47 H26 N6 O3 A1
	749.1885	-4.3	-5.7	40.0	C48 H25 N7 Na Al





### <sup>1</sup>H NMR of tris-(5-methyl-5,6-dihydrodibenzo[c,h][1,6]naphthyridin-1-ol)aluminum (13)

Due to overlapping peaks and poor solubility of the studied compound, <sup>1</sup>H NMR spectrum was measured both in CDCl<sub>3</sub> and DMSO-*d*<sub>6</sub>.



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HRMS of tris-(5-methyl-5,6-dihydrodibenzo[*c*,*h*][1,6]naphthyridin-1-ol)aluminum (**13**)

**Elemental Composition Report** 

Single Mass Analysis (displaying only valid results) - displaying only valid results

Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

232 formula(e) evaluated with 4 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
811.2981	811.2980	0.1	0.1	37.0	C52 H39 N7 Na Al
	811.2977	0.4	0.5	35.5	C51 H40 N6 O3 Al
	811.3004	-2.3	-2.9	40.0	C54 H38 N7 Al
	811.2953	2.8	3.4	32.5	C49 H41 N6 O3 Na Al

J281

lct-25009 347 (6.944) AM (Cen,2, 80.00, Ar,5000.0,609.28); Sm (SG, 1x2.00); Sb (4,40.00); Si 100<sub>Π</sub> 2.23e3



<sup>1</sup>H NMR of tris-(6*H*-isochromeno[4,3-*c*]quinolin-1-ol)aluminum (**14**)



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HRMS of tris-(6*H*-isochromeno[4,3-*c*]quinolin-1-ol)aluminum (14)

Elemental Composition Report

Single Mass Analysis (displaying only valid results) Tolerance = 5.0 mDa / DBE: min = -0.5, max = 100.0

Monoisotopic Mass, Odd and Even Electron Ions

604 formula(e) evaluated with 10 results within limits (up to 50 closest results for each mass)

Minimum:				-0.5	
Maximum:		5.0	10.0	100.0	
Mass	Calc. Mass	mDa	PPM	DBE	Formula
794.1868	794.1872	-0.4	-0.5	38.5	C50 H29 N3 O6 Al
	794.1861	0.7	0.8	35.0	C50 H32 O7 Na Al
	794.1875	-0.7	-0.9	40.0	C51 H28 N4 O3 Na Al
	794.1885	-1.8	-2.2	38.0	C52 H31 O7 Al
	794.1848	2.0	2.5	35.5	C48 H30 N3 O6 Na Al
	794.1888	-2.0	-2.6	39.5	C53 H30 N O4 Na Al
	794.1899	-3.1	-3.9	43.0	C53 H27 N4 O3 Al
	794.1826	4.1	5.2	47.0	C59 H27 O2 Al
	794.1912	-4.4	-5.6	42.5	C55 H29 N O4 Al
	794.1915	-4.7	-5.9	44.0	C56 H28 N2 O Na Al

#### J279

LCT-25010 90 (1.801) AM (Cen,2, 70.00, Ar,5000.0,995.31); Sm (SG, 2x2.00); Sb (4,40.00); § 100 794.1868 1.29e3





Fluorescence quantum yield determination in steady state

**Figure S1.** Absorption (a and c) and emission (b and d) spectra in absolute scale of the parent  $Alq_3$ , derivative **12** (in ethanol) and derivatives **13** and **14** (in chloroform) solutions. The parent  $Alq_3$  was used as a reference both in chloroform and ethanol. Excitation wavelength was the crossing point in absorption (365 nm), in order to cancel out the effect of transmittance. Therefore, relative fluorescence quantum yield is the ratio of integrated intensity of the sample to the reference throughout the measured range.

Note: In Figure S1c, the transmittances of samples 13, 14 and the parent  $Alq_3$  at the excitation wavelength (365 nm) are not exactly equal. This was taken into account in the calculation and it turned out, that the contribution rising from this difference is negligible. Experiments were repeated by using solutions of lower concentration, and by using different excitation wavelengths. In each case, the results corresponded very well with the ones presented here.

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Fluorescence quantum yield ( $\Phi$ ) in steady state was calculated by using the equation 1:

$$\frac{\Phi}{\Phi_{ref}} = \frac{\int_{\lambda_1}^{\lambda_2} I(\lambda) d\lambda}{\int_{\lambda_1}^{\lambda_2} I_{ref}(\lambda) d\lambda} \times \frac{(1 - T_{ex, ref})}{(1 - T_{ex})} \times \frac{n^2}{n_{ref}^2}$$
(1)

 $\lambda_1$  and  $\lambda_2$  are determined from the measured wavelength range. Integrals can be estimated as sums of the measured intensities throughout the spectrum (corrected according to the instruments wavelength sensitivity). By selecting the excitation wavelength in a way, that transmittance of the sample equals with the reference ( $T_{ex} = T_{ex,ref}$ ), using the same solvent ( $n = n_{ref}$ ) and normalizing  $\Phi_{ref}$  to 1 (the parent Alq<sub>3</sub>), relative fluorescence quantum yield can be obtained by comparing intensities only (equation 2).

$$\Phi = \frac{\sum_{\lambda_1}^{\lambda_2} I(\lambda)}{\sum_{\lambda_1}^{\lambda_2} I_{ref}(\lambda)}$$

(2)

#### Fluorescence decays obtained with TCSPC method and calculation of parameters



Figure S2. Fluorescence decays of the parent Alq<sub>3</sub> ( $\lambda_{mon} = 520 \text{ nm}$ ) and compounds **13** and **14** ( $\lambda_{mon} = 490 \text{ nm}$ ) in CHCl<sub>3</sub> (a), the parent Alq<sub>3</sub> ( $\lambda_{mon} = 520 \text{ nm}$ ) and compound **12** ( $\lambda_{mon} = 440 \text{ nm}$ ) in ethanol (b) and thin films of the parent Alq<sub>3</sub> and compound **12** (c).



Figure S3. Decay associated spectra of the parent Alq<sub>3</sub> ( $\lambda_{mon} = 520 \text{ nm}$ ) and compounds **13** and **14** ( $\lambda_{mon} = 490 \text{ nm}$ ) in CHCl<sub>3</sub> (a), two components of the compound **13** in CHCl<sub>3</sub> (b), and the parent Alq<sub>3</sub> ( $\lambda_{mon} = 520 \text{ nm}$ ) and compound **12** ( $\lambda_{mon} = 440 \text{ nm}$ ) in ethanol (c).

Relative fluorescence quantum yield can be obtained from decay associated spectrum by comparing the sum intensities ( $I_{sum}$ , equations 4a and b) of the sample and the reference, and it should correlate with the quantum yield in steady-state ( $\Phi$ ).

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Fluorescence decays were fitted to the well known exponential model (*A* = Amplitude, *I* = Intensity):

$$I_{1-\exp} = A_0 + A_1 e^{-\frac{t}{\tau_1}}$$
(3a)
$$I_{n-\exp} = \sum_{i=0}^n (A_i e^{-\frac{t}{\tau_i}})$$
(3b)

Integrating time from zero to infinity, summing the intensities throughout the measured spectrum and by taking transmittance (T) of the excitation pulse (340 nm) into account leads to:

$$I_{sum,1-\exp} = \frac{1}{(1-T)} \sum_{\lambda_{1}}^{\lambda_{2}} A_{1}(\lambda)\tau_{1}$$
(4a)  
$$I_{sum,n-\exp} = \frac{1}{(1-T)} \sum_{\lambda_{1}}^{\lambda_{2}} (\sum_{i=0}^{n} A_{i}(\lambda)\tau_{i})$$
(4b)

Constant  $A_0$  was negligible and was therefore excluded from the analysis. In solution, derivative 13 revealed two-exponential decay (equation 4b, n = 2), while the decays of the parent Alq<sub>3</sub> and derivatives 12 and 14 were single exponential (equation 4a).

Radiative fluorescence rate constants  $(k_r)$  were calculated by using equation:

$$k_r = \frac{\Phi}{\tau_F} \tag{5}$$

where  $\Phi$  is an absolute fluorescence quantum yield and  $\tau_F$  fluorescence lifetime.

Non-radiative rate constants  $(k_{nr})$  were calculated by using equation:

$$k_{nr} = \frac{1}{\tau_F} - k_r \tag{6}$$