

Supplementary Information

Preparation of Graphene-TiO₂ Composites with Enhanced Photocatalytic Activity

Kangfu Zhou, Yihua Zhu, Xiaoling Yang, Xin Jiang, Chungzhong Li*

Key Laboratory for Ultrafine Materials of Ministry of Education, School of Materials
Science and Engineering, East China University of Science and Technology,
Shanghai 200237, China

*Corresponding author: Fax: +86-21-64250624;

E-mail address: yhzhu@ecust.edu.cn (Y. Zhu)

Figure S1. (a) XRD pattern of graphite oxide (GO), (b) AFM image of graphene oxide. Graphene oxide was obtained by sonicating a sufficiently dilute colloidal suspension of GO for a desired time. The uniform thickness of about 1.1 nm, which is the typical thickness of a hydrated individual graphene oxide layer, can be observed from the AFM image.

Figure S2. Photocurrents of TiO_2 and G- TiO_2 composites under intermittent irradiation of a 150 W high-pressure Xenon lamp at a bias potential of 0.2 V.

Figure S3. XRD pattern of the reduced exfoliated GO obtained by solvothermal reduction ($180\text{ }^\circ\text{C}$) for 4 h.

Figure S4. Photographic images of the dispersions of G- TiO_2 composites with reaction time of 4 h (left) and 8 h (right). The left sample exhibits the separation of TiO_2 (white layer) and graphene (black layer) while the right sample is a gray dispersion with homogeneous appearance.

Figure S5. C1s XPS spectra of G- TiO_2 composites with reaction time of (a) 4 h and (b) 8 h.

Figure S6. UV-vis absorption spectra showing the degradation of methylene blue (MB) by using different photocatalysts, (a) pure TiO_2 , (b) P25, (c) G- TiO_2 composites with reaction time of 4 h and (d) G- TiO_2 composites with reaction time of 8 h.

Figure S7. (a) TEM image and (b) FTIR spectrum of used G-TiO₂ composites.

Figure S1.

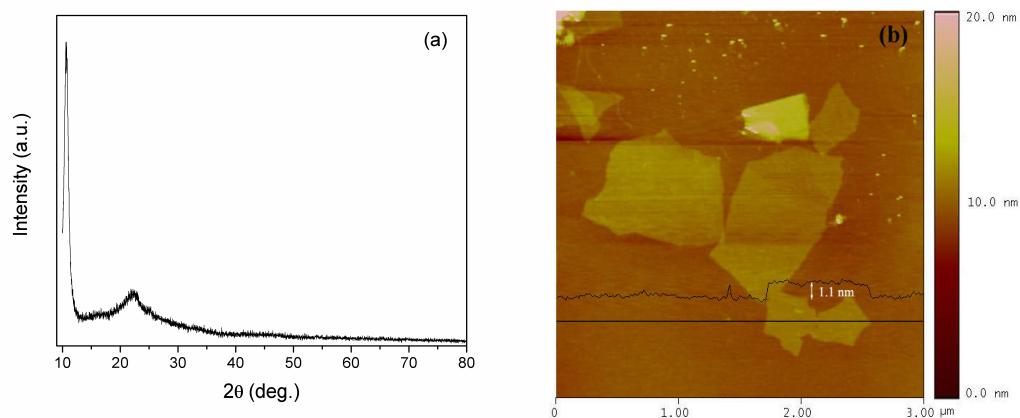


Figure S2.

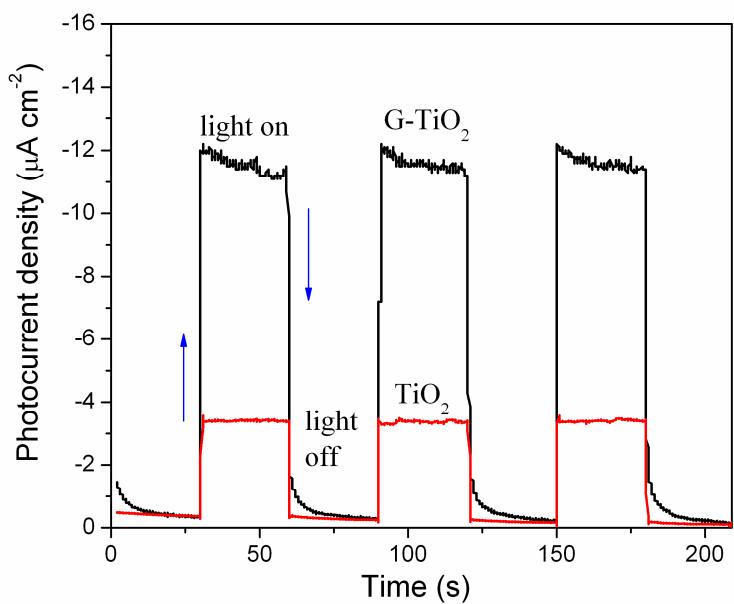


Figure S3.

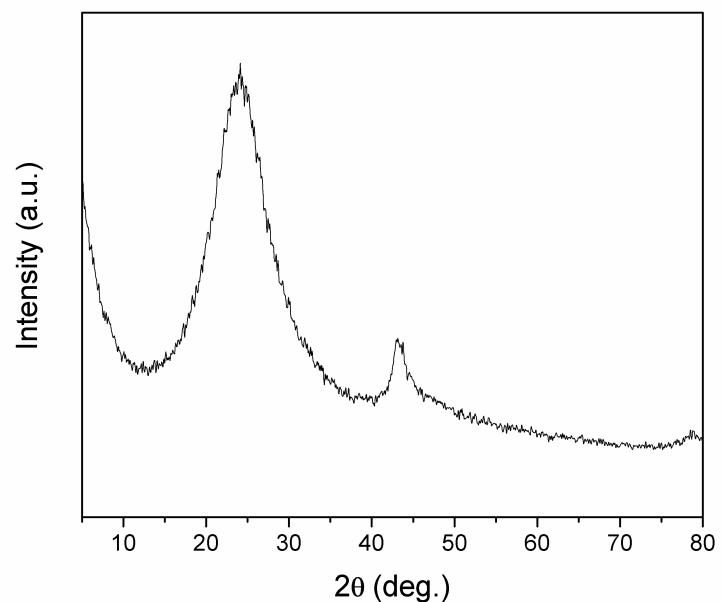


Figure S4.



Figure S5.

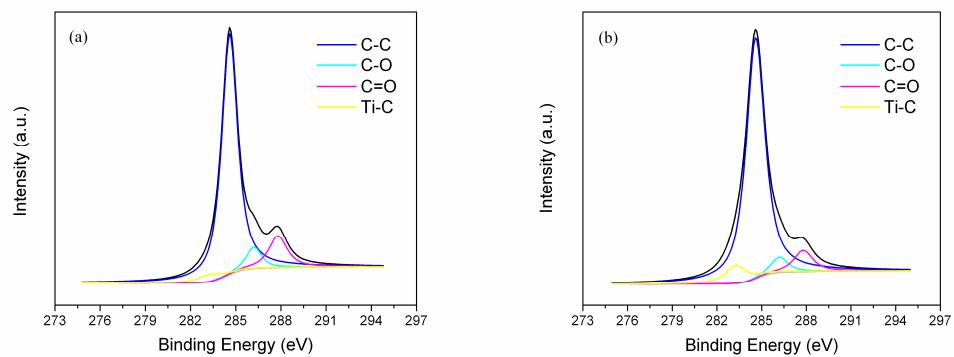


Figure S6.

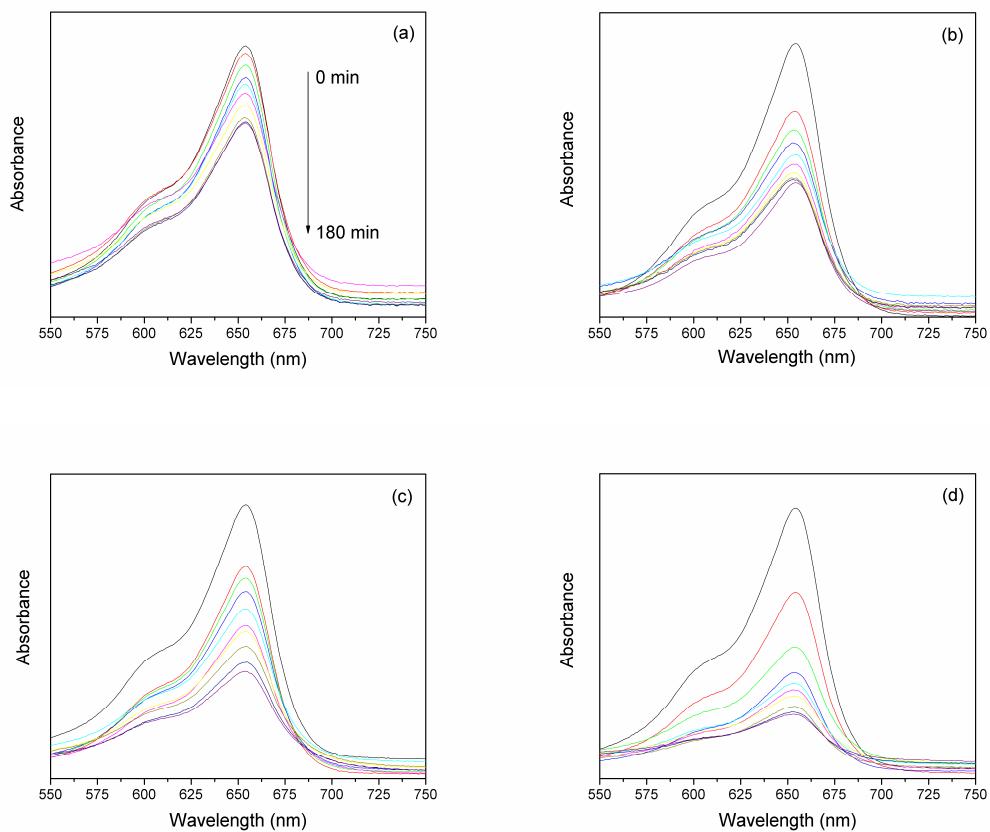


Figure S7.

