

Supporting information

Novel Glutathione-linked Nitrones as Dual Free Radical Probes

Yuguang Song,^{a,b,†} Yangping Liu,^{a,b,†} Libo Du,^a Frederick A. Villamena,^b Yiqiong Ji,^a Qiu Tian,^a Ke-Jian Liu,^{*c} and Yang Liu^{*,d}

^aState Key Laboratory for Structural Chemistry of Unstable and Stable Species, Center for Molecular Science, Institute of Chemistry, Chinese Academy of Sciences, Zhongguancun, Beijing 100190, P. R. China; E-mail: yliu@iccas.ac.cn

^bCenter for Biomedical EPR Spectroscopy and Imaging, The Davis Heart and Lung Research Institute, the Division of Cardiovascular Medicine, Department of Internal Medicine, and Department of Pharmacology, College of Medicine, The Ohio State University, Columbus, OH, 43210, USA.

^cCollege of Pharmacy, University of New Mexico, 2502 Marble NE, Albuquerque, NM, 87131, USA; E-mail: kliu@salud.unm.edu

[†]Both authors contributed equally to this work.

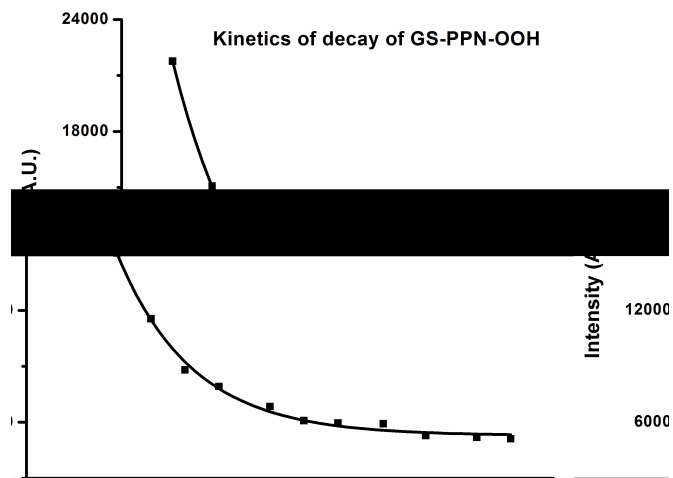


Figure S1. Decay plot of the superoxide spin adduct of GS-PPN. The superoxide spin adduct was generated by adding xanthine oxidase (0.02 U/mL) to an aqueous solution containing hypoxanthine (0.4 mM), DTPA (1 mM) and GS-PPN (50 mM) in PBS (20 mM, pH 7.4). The decay of the adduct was examined 2 min after addition of superoxide dismutase (400 units per mL).

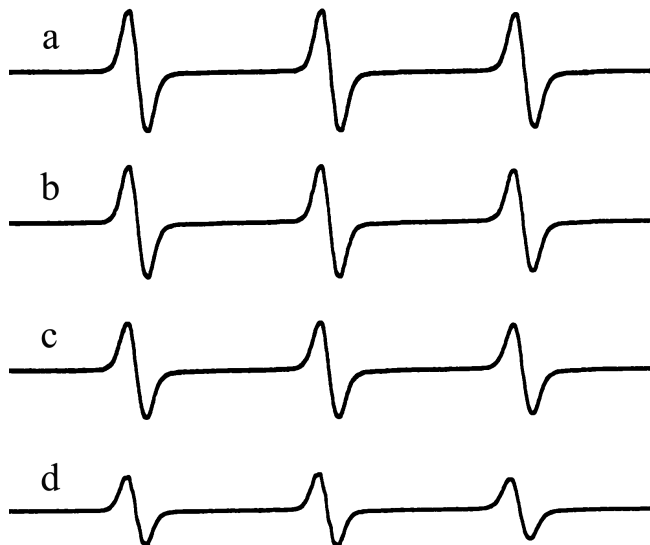


Figure S2. EPR spectra of 4-hydroxy-2,2,6,6-tetramethylpiperidine-1-oxyl (Tempol, 0.5 mM) in the aqueous solutions containing various concentrations of glycerol; a) in the absence of glycerol; b) glycerol (volume percentage, 10%); c) glycerol (30%); d) glycerol (50%).

Table S1. EPR hyperfine splitting constants of GS-PPN spin adducts

Adduct	Source	α_{P} (mT)	α_{N} (mT)	α_{H} (mT)
$\text{O}_2^{\bullet-}$	HX/XOD ^a	4.21	1.35	0.21
	PSII ^b /hv	4.23	1.36	0.22
$\text{CH}_3\text{O}^{\bullet}$	$\text{CH}_3\text{OH}/\text{Pb}(\text{OAc})_4$	3.92	1.32	0.26
EtO^{\bullet}	$\text{C}_2\text{H}_5\text{OH}/\text{Pb}(\text{OAc})_4$	3.94	1.34	0.25
$\text{CH}_2^{\bullet}\text{OH}$	$\text{CH}_3\text{OH}/\text{Fenton}^{\text{c}}$	4.25	1.48	0.34
$\text{CH}_3\text{CH}^{\bullet}\text{OH}$	$\text{C}_2\text{H}_5\text{OH}/\text{Fenton}^{\text{c}}$	4.20	1.48	0.33
$(\text{CH}_3)_2\text{C}^{\bullet}\text{OH}$	iPrOH/ Fenton ^c	3.94	1.46	0.34
$\text{C}_3\text{H}_7\text{CH}^{\bullet}\text{OH}$	$\text{nC}_4\text{H}_9\text{OH}/\text{Fenton}^{\text{c}}$	4.23	1.43	0.34
$\text{CO}_2^{\bullet-}$	$\text{HCO}_2\text{Na}/\text{Fenton}^{\text{c}}$	5.04	1.47	0.45
ClPh [•]	p-CDT ^d /hv, water	4.54	1.48	0.39
	p-CDT ^d /hv, DMSO	4.24	1.40	0.31
$^{\bullet}\text{SO}_3^{2-}$	$\text{Na}_2\text{SO}_3/\text{K}_2\text{Cr}_2\text{O}_7$	4.45	1.37	0.1

a, HX/XOD, hypoxanthine/xanthine oxidase; b, PSII, photosystem II; c, Fenton consists of H_2O_2 , EDTA and Fe^{2+} ; d, p-CDT, para-chlorophenyl diazonium tetrafluoroborate.