

## Supporting Information

### **Influence of Nitrogen Donor Ligands on the Supramolecular Architecture of Copper (II) 2-Nitrobenzoate Complexes: Structures, DFT Calculations and Magnetic Properties**

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### *Thermogravimetric Analysis.*

Thermogravimetric analysis of complexes **1-3** was examined from 33 to 1000° in air to study the thermal stability (Fig. S1-S3). In the TG curve of **1** the first weight loss step (calc = 23.99% , expt = 22.75%) from 74-158 °C corresponds to loss of one molecule of  $\gamma$ -picoline and one nitro molecule from one nitrobenzoate moiety.-In **2**, the first weight loss step (calc = 5.82% , expt = 6.00%) from 70-108 °C corresponds to the loss of two coordinated water molecule whereas the second and the third steps correspond to loss of one  $\square$ -pic molecule (calc = 15.06%, expt = 14.99% from 108-175 °C) and two nitrobenzoate moieties-(calc = 54.00% , expt = 51.99% from 179-235 °C) respectively. In **3**, the first and second weight loss corresponds to the loss of two lattice water (calc = 4.55%, expt = 4.74% from 30-104 °C) and one 2-nitrobenzoate moiety (calc = 20.99% , expt = 20.52% from 104-198 °C) respectively.

### *IR and electronic spectroscopy*

Infrared spectrum of the newly synthesized complex salts have been recorded in the region 4000-400  $\text{cm}^{-1}$  and tentative assignments have been made on the basis of earlier reports in literature [K. Nakamoto, Infrared and Raman Spectra of Inorganic and Coordination Compounds, 1997, John Wiley & Sons, New York , 5th Ed. , J. Bellamy, The Infrared Spectra of Complex Molecules, 1980, Chapman & Hall, London/ New York, Ed.]. The strong band in the region 3371  $\text{cm}^{-1}$  in **2** and 3418- 3225  $\text{cm}^{-1}$  in **3** may be attributed to OH( $\text{H}_2\text{O}$ ) stretching and OH of triethanolamine and  $\text{H}_2\text{O}$  stretching, respectively. The bands varying from weak to medium intensity in the region 2877–3167  $\text{cm}^{-1}$  and 1607-1601  $\text{cm}^{-1}$  in complexes **1,2** and **3** may be respectively attributed to CH stretching. In **1-3**, the bands between 1593 -1578  $\text{cm}^{-1}$  and 1393- 1385  $\text{cm}^{-1}$  may be ascribed to  $\nu_{\text{as}}(\text{COO})$  and  $\nu_{\text{s}}(\text{COO})$ , respectively. The  $\Delta\nu_{\text{COO}} = \nu_{\text{as}}(\text{COO}) - \nu_{\text{sym}}(\text{COO})$  values are 175  $\text{cm}^{-1}$  in **1**, 200  $\text{cm}^{-1}$  in **2** and 216  $\text{cm}^{-1}$  in **3** which indicates the presence of bidentate bridging in **1** and monodentate in **2** and **3**. The bands between 1340- 1367 $\text{cm}^{-1}$  and 1524-1537  $\text{cm}^{-1}$  in all complexes may be assigned to asymmetric and symmetric stretching mode of nitro group respectively. The bands at 835, 836 and

813  $\text{cm}^{-1}$  may be ascribed to  $\delta(\text{NO}_2)$  in **1**, **2** and **3** respectively. The bands at 4597, 554  $\text{cm}^{-1}$  in **1**, at 463, 538  $\text{cm}^{-1}$  in **2** and 466 and 568  $\text{cm}^{-1}$  in **3** are ascribed to  $\nu(\text{Cu-N})$  and  $\nu(\text{Cu-O})$  vibrations. The bands at 648, 493  $\text{cm}^{-1}$  in **1** and at 649, 513  $\text{cm}^{-1}$  in **2** were ascribed to coordinated in plane and out plane ring deformation of  $\gamma$ -picoline and  $\beta$ -picoline, respectively. IR spectra are available as Figures S4-S6 .

The electronic spectra of complex **1**, **2** and **3** were recorded in methanol. Complex **1** showed absorption at 732 ( $\epsilon_{\text{max}} = 94.07 \text{ mol}^{-1} \text{ L cm}^{-1}$ ) whereas complex **2** and **3** showed maximum absorption 738 nm with  $\epsilon_{\text{max}} = 102.5 \text{ mol}^{-1} \text{ L cm}^{-1}$  and with  $\epsilon_{\text{max}} 63.41 \text{ Lmol}^{-1}\text{cm}^{-1}$  respectively, due to the d-d transitions. These values are agreement with those observed by  $\text{O}_4\text{N}_2$  chromophore [F. A Cotton, G. Wilkinson, C. A Muritto, M. Bochmann, *Advanced Inorganic Chemistry*, sixth ed., 1999, John Wiley and Sons, New York, 869; R. P. Sharma, A. Saini, S. Singh, A. Singh, P. Venugopalan, P. Starynowicz, J. Jezierska, *J. Mol. Struct.* 2011, **994**, 21–28].

Figure S1. The TGA of complex salt **1**

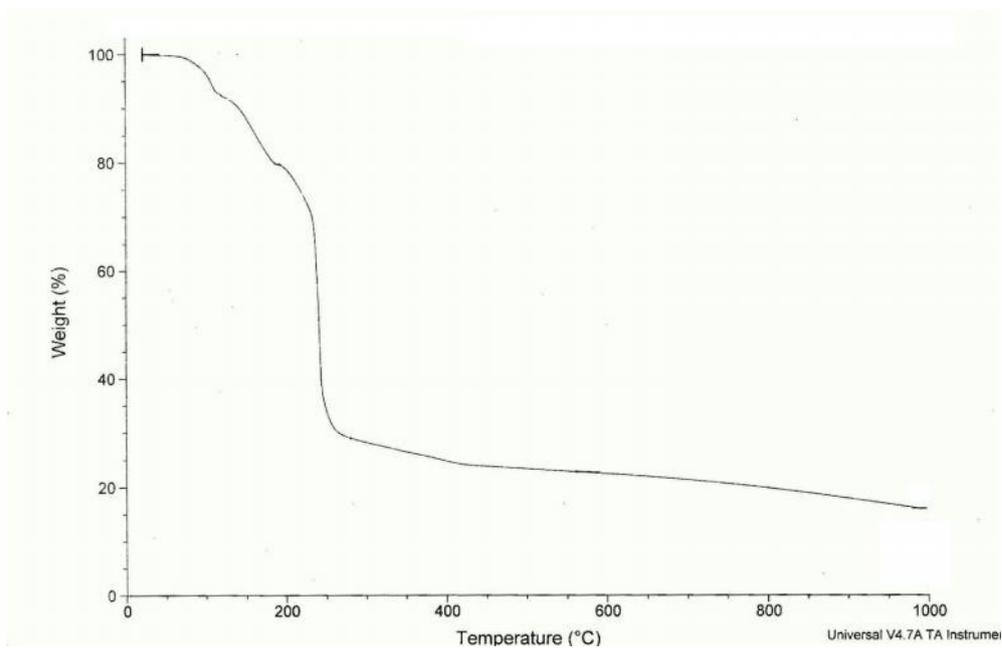


Figure S2. The TGA of complex salt **2**

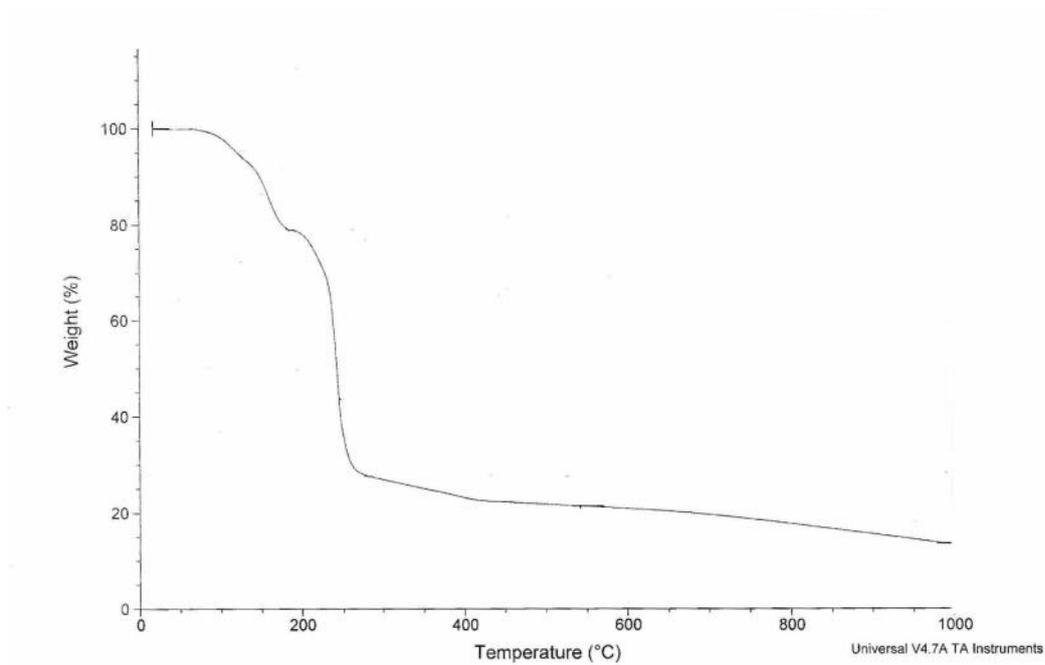


Figure S3. The TGA spectrum of **3**

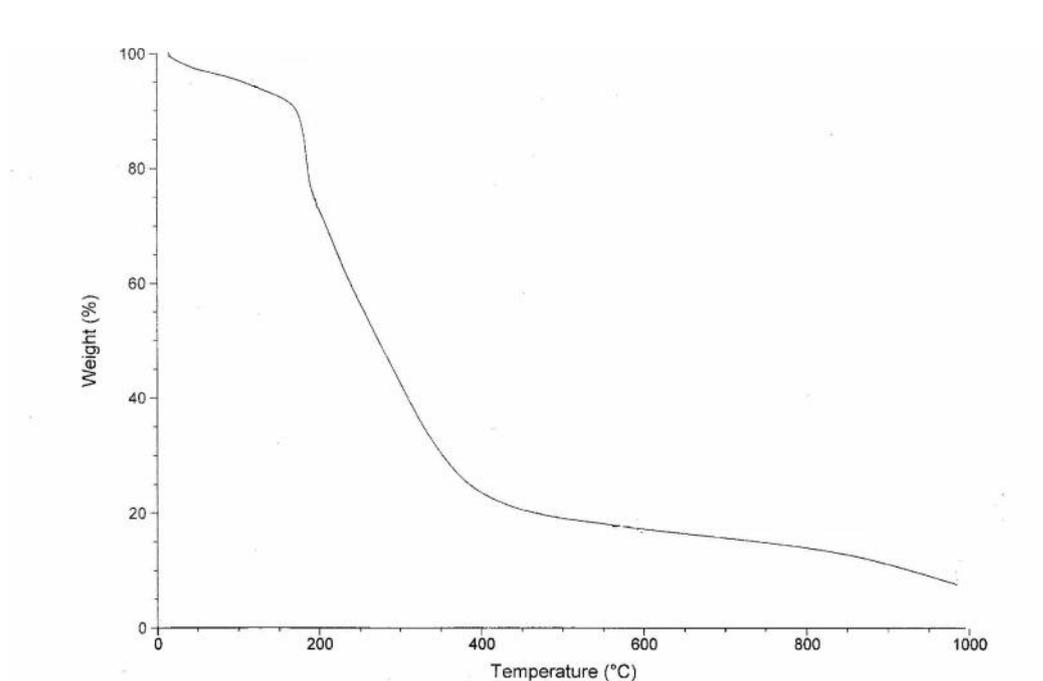


Figure S4. IR spectrum of complex **1**

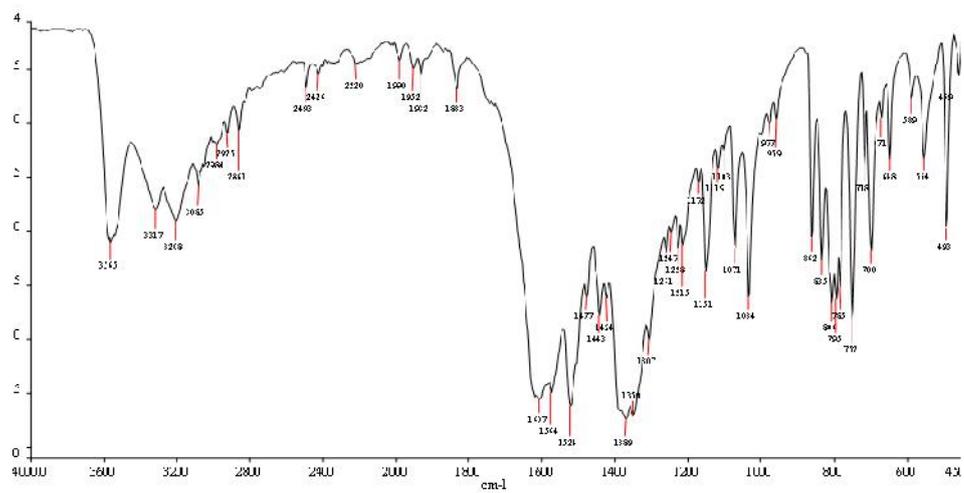


Figure S5. IR spectrum of complex **2**

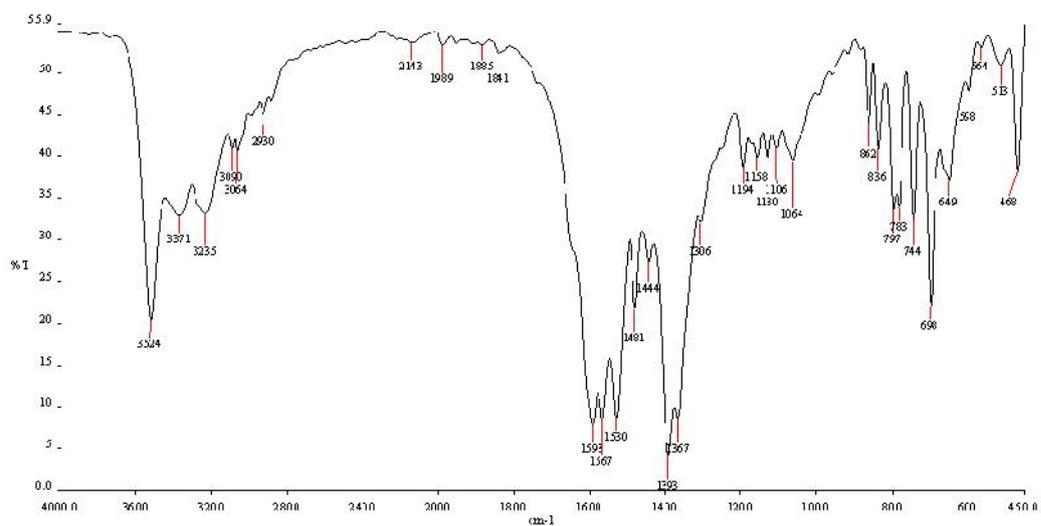


Figure S6. IR spectrum of complex **3**

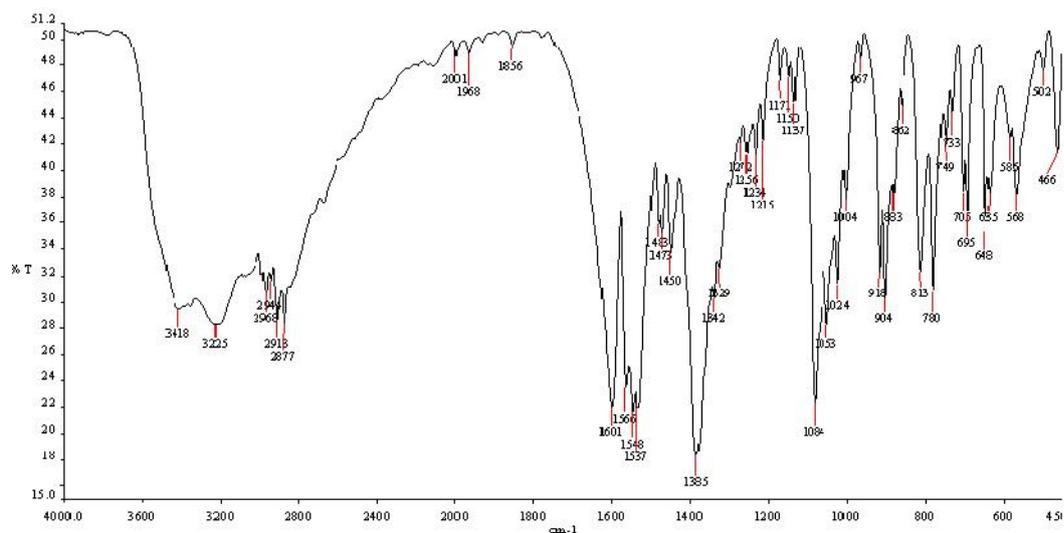


Figure S7. Packing diagram of complex **1** viewed down a axis

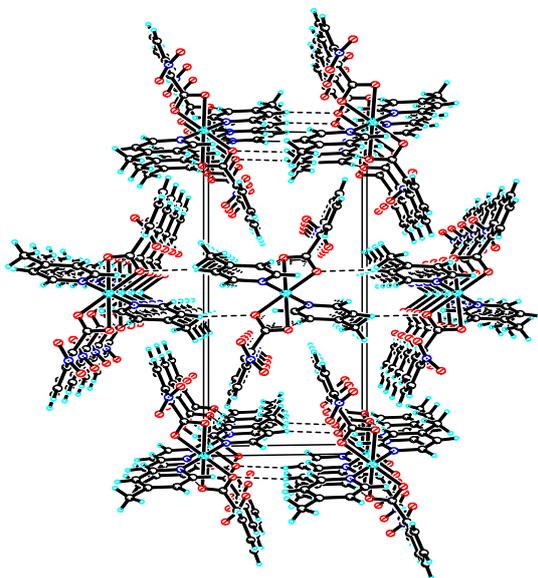


Figure S8. Packing diagram of complex **2** viewed down b axis

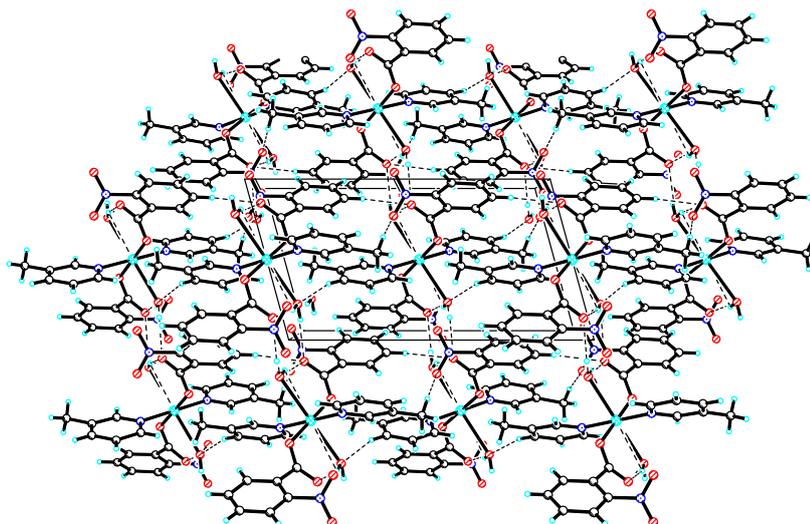


Figure S9. Packing diagram of complex **3** viewed down a axis

