

Supplementary Material (ESI) for New Journal of Chemistry

Towards sensory Langmuir monolayers consisting of macrocyclic pentaaminoanthraquinone

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1. Hydrogen bond distances and angles

Table S1. Hydrogen bonds for [(PENTAQ)H]CF₃SO₃•H₂O

D	H	A	<i>d</i> (D–H) (Å)	<i>d</i> (H⋯A) (Å)	<i>d</i> (D⋯A) (Å)	∠(D–H⋯A) (°)
O10	H10A	O6B ^a	1.00(2)	1.80(3)	2.77(2)	163(5)
O10	H10A	O6A ^a	1.00(2)	1.99(9)	2.94(9)	157(5)
O10	H10B	O6B ^b	0.997(19)	1.96(4)	2.91(1)	160(7)
O10	H10B	O4A ^b	0.997(19)	1.91(9)	2.75(7)	141(6)
N1	H1	O1	0.77(5)	2.02(5)	2.631(4)	136(4)
N5	H5	O1	0.80(5)	2.07(5)	2.630(4)	127(4)
N5	H5	N4	0.80(5)	2.35(4)	2.765(6)	113(4)
N3	H3A	O10	0.99	1.85	2.829(7)	169.0
N3	H3B	O2 ^c	0.99	2.34	3.054(5)	128.4

Symmetry transformations used to generate equivalent atoms: (a) $-x, 1-y, 1-z$; (b) $x, y, 1+z$. (c) $1-x, 1-y, 1-z$.

Table S2. Hydrogen bonds for [(PENTAQ)H₂](CF₃SO₃)₂

D	H	A	<i>d</i> (D–H) (Å)	<i>d</i> (H⋯A) (Å)	<i>d</i> (D⋯A) (Å)	∠(D–H⋯A) (°)
N1	H1	O1	0.82(2)	1.92(2)	2.582(2)	137(2)
N2	H2C	O3 ^a	0.87(3)	1.99(3)	2.853(3)	174(2)
N2	H2D	O7 ^b	0.91(2)	2.00(2)	2.844(3)	152(2)
N3	H3	O8 ^b	0.86(2)	2.32(2)	3.085(3)	147(2)
N4	H4C	O4	0.92(2)	2.36(2)	2.886(2)	115.8(16)
N4	H4C	O5 ^a	0.92(2)	2.03(2)	2.802(2)	139.9(18)
N4	H4D	O6 ^b	0.90(2)	2.11(2)	2.877(2)	143(2)
N5	H5	O1	0.81(2)	1.94(2)	2.590(2)	137(2)
N5	H5	O5 ^a	0.81(2)	2.59(2)	3.216(2)	135(2)

Symmetry transformations used to generate equivalent atoms: (a) $2-x, 1-y, 1-z$; (b) $1-x, 1-y, 1-z$.

2. Spectrophotometric protonation studies

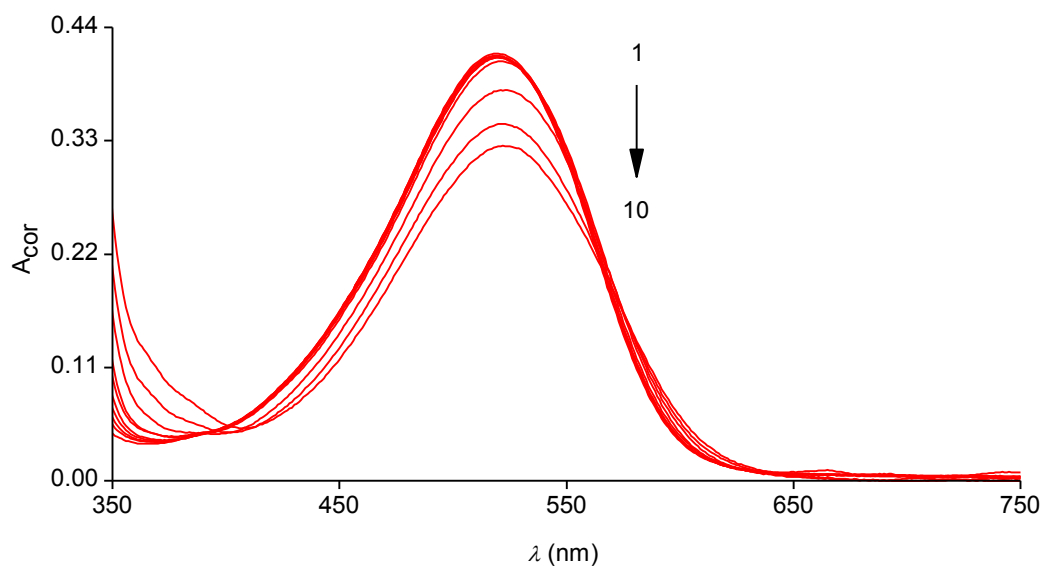


Fig. S1 Spectrophotometric batch titration of PENTAQ (L) at high acidities in CH₃OH/H₂O 80:20 v/v. [L]_{tot} = 47 μ M, [N(CH₃)₄NO₃]₀ = 0.1 M (the total ionic strength increased with each added aliquot of HNO₃), V₀ = 2 mL; T = 298.2(2) K, l = 1 cm, spectra 1–10 (corrected for dilution effects): p[H] = $-\log [\text{HNO}_3]_{\text{tot}}$ = 1.67, 1.17, 0.94, 0.72, 0.51, 0.30, 0.07, -0.2, -0.4, -0.6.

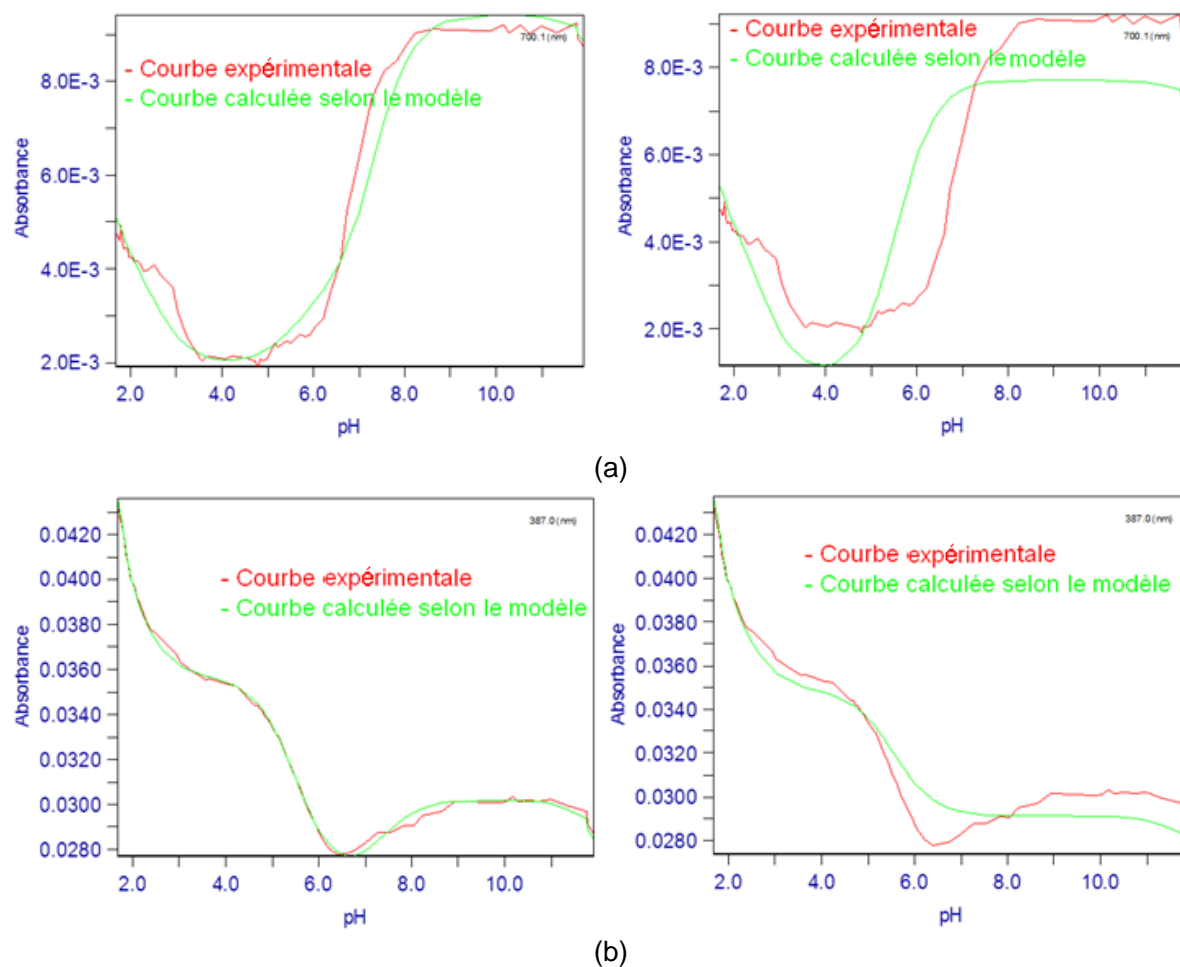


Fig. S2 Absorbance vs. p[H] titration profiles at (a) 700 and (b) 387 nm of PENTAQ (L) in CH₃OH/H₂O 80:20 v/v. The red curves correspond to the experimental data and the green ones to the best fit obtained with a chemical model comprising either 3 (L, LH⁺, LH₂²⁺; right graph) or 4 absorbing species (L, LH⁺, LH₂²⁺, LH₃³⁺; left graph). [L]_{tot} = 47 μM, [N(CH₃)₄OH] = 0.1040 M, V₀ = 30 mL, I = 0.1 M N(CH₃)₄NO₃, T = 298.2(2) K, l = 1 cm.

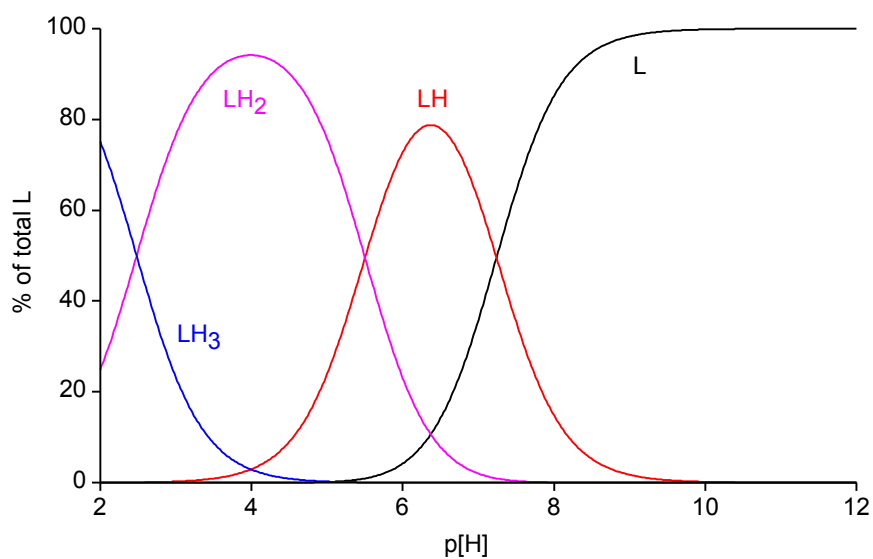


Fig. S3 Species distribution diagram for PENTAQ (L) in CH₃OH/H₂O 80:20 v/v. $I = 0.1 \text{ M N(CH}_3)_4\text{NO}_3$, $T = 298.2(2) \text{ K}$.

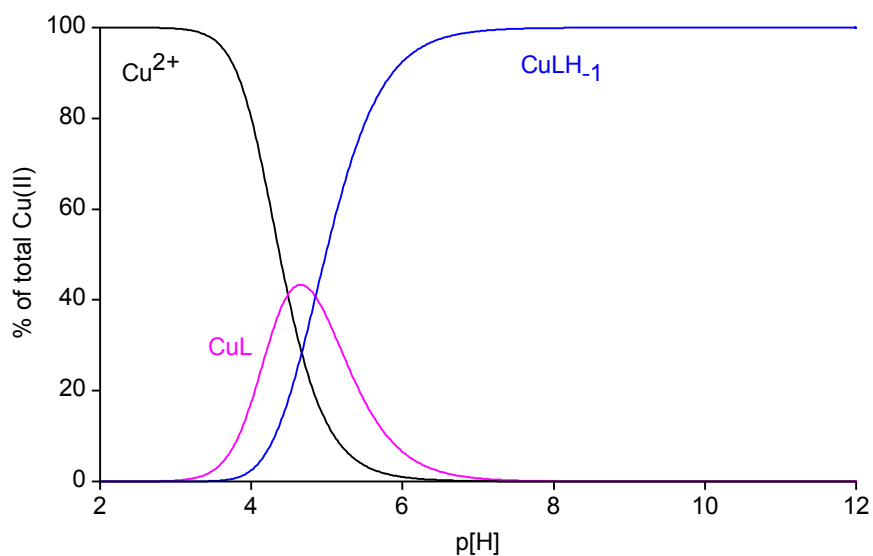
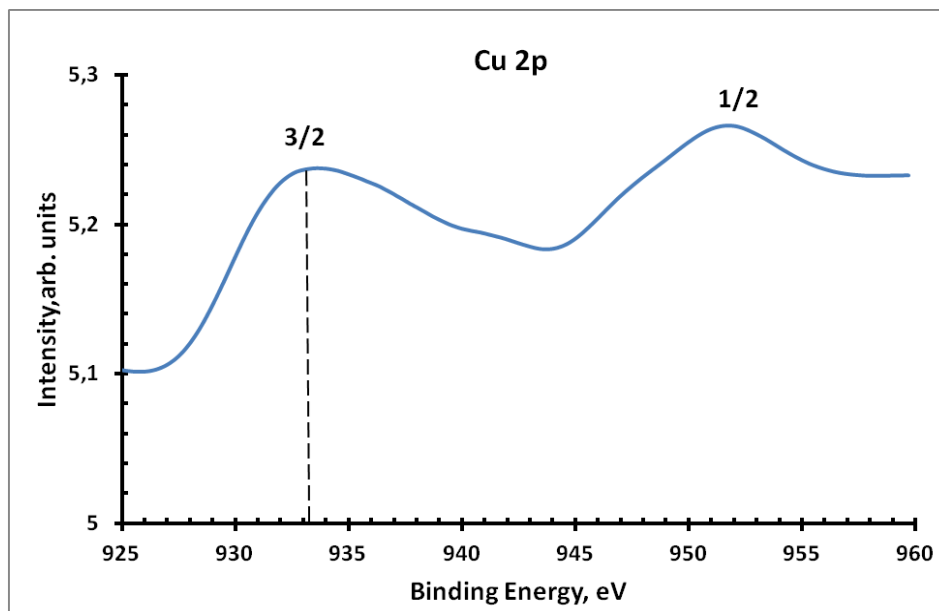
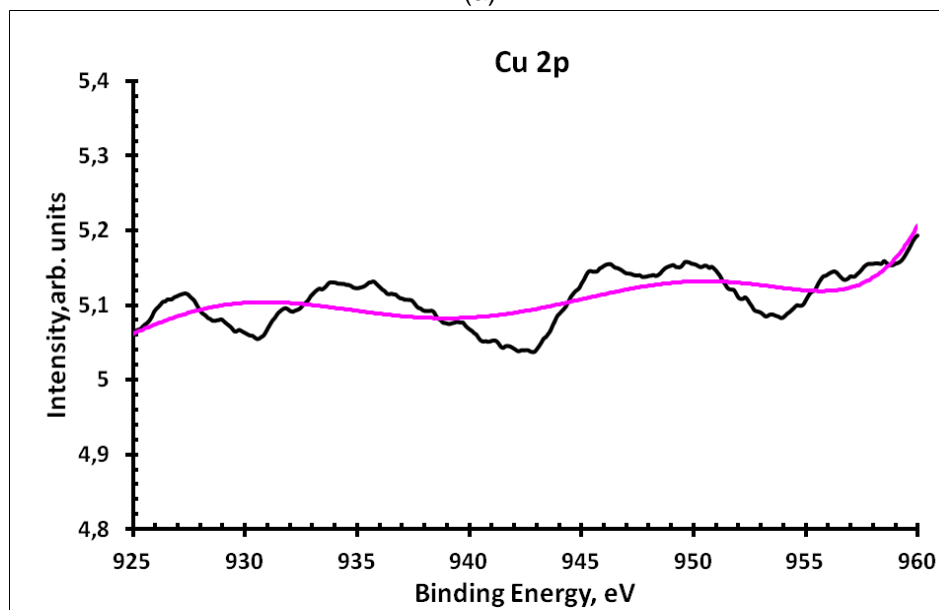


Fig. S4 Species distribution diagram for the Cu²⁺/PENTAQ (L) system in CH₃OH/H₂O 80:20 v/v. $[\text{L}]_{\text{tot}} = [\text{Cu}]_{\text{tot}} = 42.9 \text{ } \mu\text{M}$, $I = 0.1 \text{ M N(CH}_3)_4\text{NO}_3$, $T = 298.2(2) \text{ K}$.

3. XPS studies of Cu^{2+} uptake by Langmuir-Blodgett monolayers of PENTAQ



(a)



(b)

Fig. S5 XPS spectra in the Cu 2p region of PENTAQ Langmuir-Blodgett films transferred from a water subphase (pH = 9.8) to a quartz slide. (a) Sample A was air-dried directly after deposition and analyzed. (b) Sample B was rinsed with deionized water for 30 s after deposition, then air-dried and analyzed. The magenta line corresponds to the smoothed spectrum.