

Supplementary Information for

Ultralong $\text{H}_2\text{V}_3\text{O}_8$ Nanowire Bundles as a Promising

Cathode for Lithium Batteries†

**Qinyou An,^a Jinzhi Sheng,^a Xu Xu,^a Qiulong Wei,^a Yaqin Zhu,^b Chunhua Han,^{*a}
Chaojiang Niu^a and Liqiang Mai^{*a}**

^a WUT-Harvard Joint Nano Key Laboratory, State Key Laboratory of Advanced Technology for Materials Synthesis and Processing, Wuhan University of Technology, Wuhan, 430070, P. R. China. Fax: +86-027-87644867; Tel: +86-027-87467595;
E-mail: mlq518@whut.edu.cn, hch5927@whut.edu.cn

^b Shanghai Institute of Ceramics, Chinese Academy of Sciences, Shanghai, 2000 50, P. R. China.

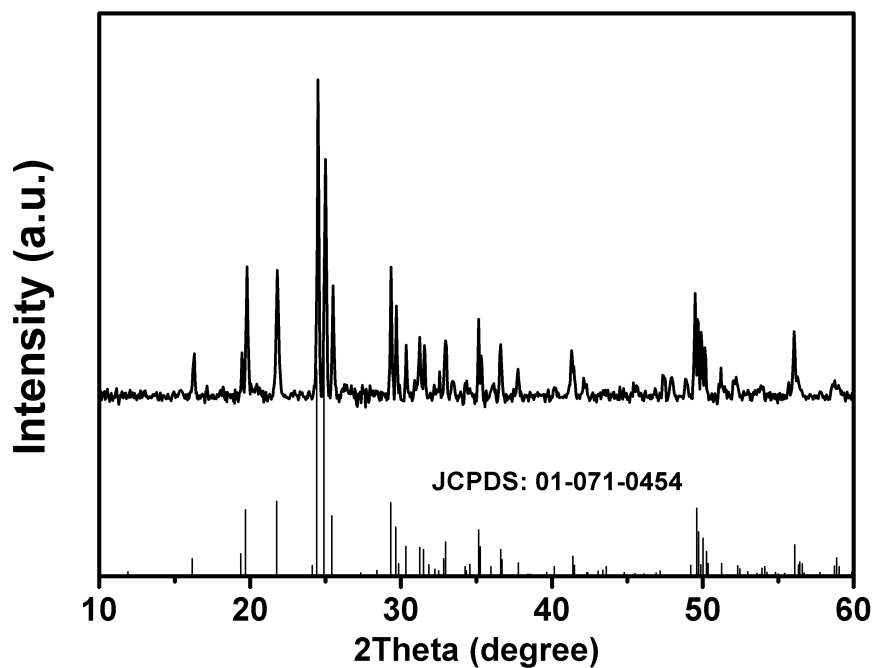


Figure S1. XRD pattern of V_3O_7 nanowires

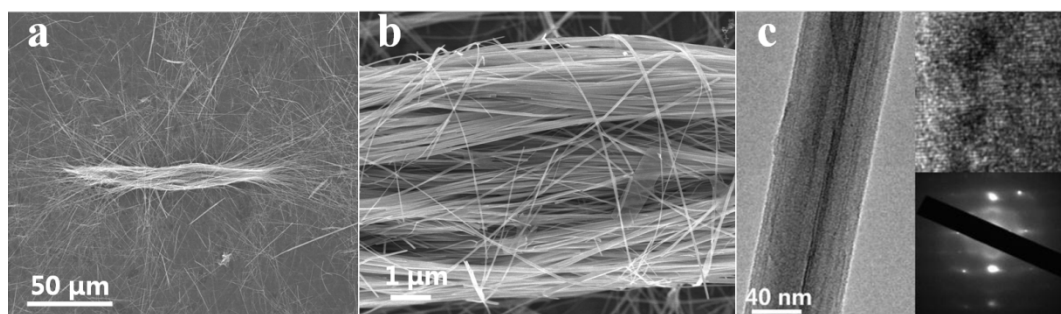


Figure S2. (a, b) SEM images of the bunched V_3O_7 nanowires. (c) TEM images and SAED pattern (inset) of V_3O_7 nanowires.

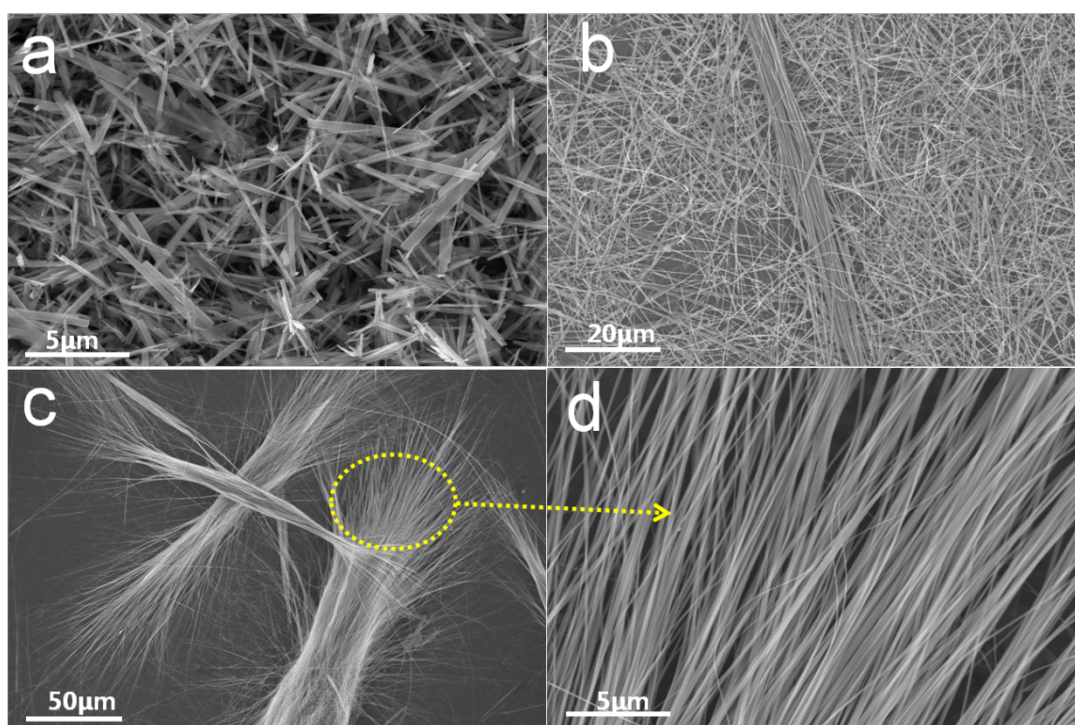


Figure S3. SEM images of vanadium oxides nanowires with different ratio of V_2O_5 sol to PEG: (a) 4:0, (b) 4:1 and (c, d) 4:2.

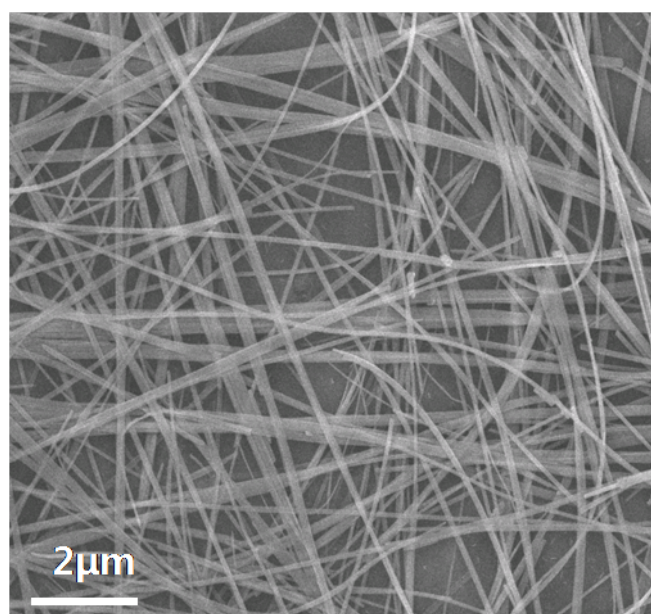


Figure S4. SEM of vanadium oxide nanowires when replace V_2O_5 sol with V_2O_5 powder (V_2O_5 to PEG is 4:2).

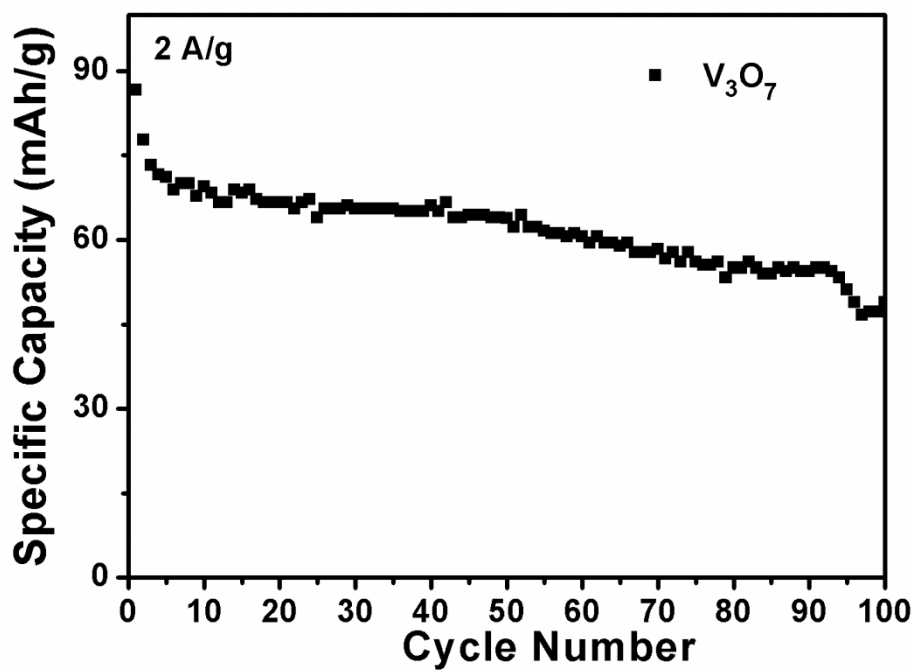


Figure S5. The cycle stability of V_3O_7 at the current density of 2000 mA/g.

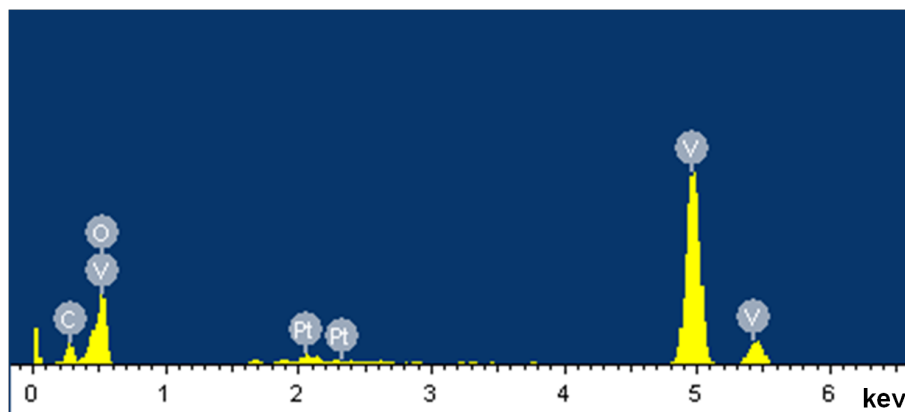


Figure S6. EDS spectra of the $H_2V_3O_8$ nanowires.

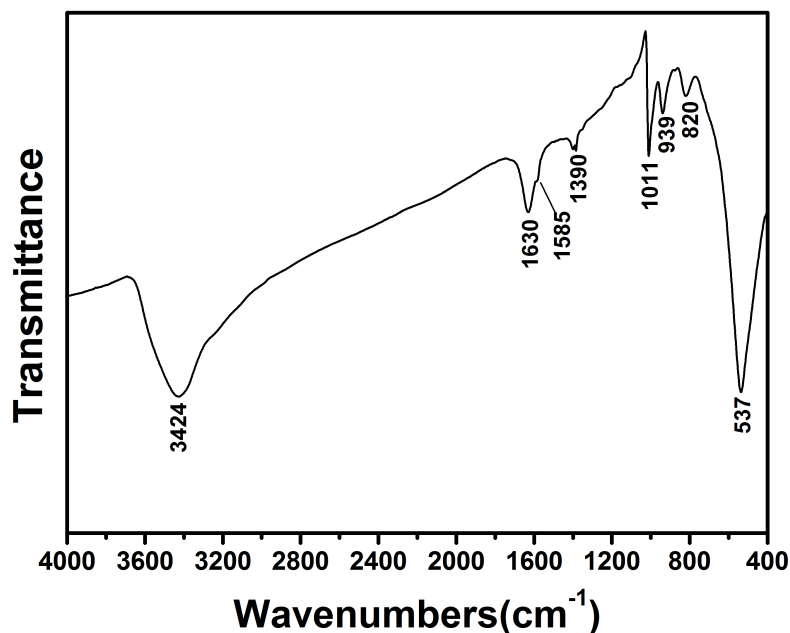


Figure S7. FT-IR spectra of the H₂V₃O₈ nanowires.

The signals at 537 cm⁻¹ and 820 cm⁻¹ can be attributed to the symmetric and asymmetric V–O–V stretching vibrations^{1,6}, respectively. The absorptions at 939 cm⁻¹ and 1011 cm⁻¹ are due to V=O stretching vibrations of trigonal bipyramids and distorted octahedras^{2,5}, respectively. The peaks at 3424 cm⁻¹ and 1630 cm⁻¹ are attributed to O–H stretching vibration and H–O–H bending vibration³, respectively. In addition, signals at 1585 cm⁻¹ and 1390 cm⁻¹ can be attributed to the symmetric and asymmetric –COOH stretching vibrations⁴, respectively. With H–O–H bending vibration at 1630 cm⁻¹ effecting, the peak derived from the symmetric –COOH stretching vibration is not obvious.

1. Y. F. Zhang, X. H. Liu, D. Z. Chen, L. Yu, J. R. Nie, S. P. Yi, H. B. Li and C. Huang, *Journal of Alloys and Compounds*, 2011, **509**: L69.
2. N. G. Park, K. S. Ryu, Y. J. Park, M. G. Kang, D. Kim, S. Kang, K. M. Kim and S. Chang, *Journal of power sources*, 2002, **103**: 273.
3. S. F. Shi, M. H. Cao, X. Y. He and H. M. Xie, *Crystal Growth & Design*, 2007, **7**: 1893.
4. A. Synytsya, J. Copikova, P. Matejka, V. Machovic, *Carbohydrate Polymers*, 2003, **54**: 97.
5. A. Doble, K. Ngala, S. F. Yang, P. Y. Zavalij and M. S. Whittingham, *Chemistry of materials*, 2001, **13**: 4382.
6. L. Q. Mai, L. Xu, C. h. Han, X. Xu, Y. Z. Luo, S. Y. Zhao, and Y. L. Zhao, *Nano letters*, 2010, **10**: 4750.

Table S1. The electrochemical performances comparison (cycling performance at relevant current rate or density, and rate capability) of the $\text{H}_2\text{V}_3\text{O}_8$ NWs with the reported ones.

Sample	Voltage range	Capacity (mAh g^{-1}) / Cycle number	Current rate or density	Rate capacity (mA h g^{-1}) at relevant Current rate or density
$\text{H}_2\text{V}_3\text{O}_8$ NWs in this work	1.5 – 3.75 V	~225 / 50	0.1 A/g	~150 at 1 A/g (100 cycles)
$\text{H}_2\text{V}_3\text{O}_8$ NWs ¹	-0.9 – 0.5 V vs. SCE	~150 / 50	0.2 A/g	~170 at 1 A/g (only 1 cycle)
$\text{V}_3\text{O}_7 \cdot \text{H}_2\text{O}$ nanobelts ²	1.5 – 3.75 V	~250 / 20	0.02 A/g	~160 at 1 A/g (20 cycles)
$\text{V}_3\text{O}_7 \cdot \text{H}_2\text{O}$ nanobelts ³	1.7 – 3.8 V	228.6 / 50	0.03 A/g	Not given

1. H. Q. Li, T. Y. Zhai, P. He, Y. G. Wang, E. Hosono and H. S. Zhou, *J. Mater. Chem.*, 2011, **21**, 1780.
2. S. K. Gao, Z. J. Chen, M. D. Wei, K. M. Wei and H. S. Zhou, *Electrochim. Acta*, 2009, **54**, 1115.
3. H. Qiao, X. J. Zhu, Z. Zheng, L. Liu and L. Zhang, *Electrochem. Commun.*, 2006, **8**, 21.