Electronic Supplementary Material (ESI) for New Journal of Chemistry

Supporting Information

Reversible iodine absorption of nonporous coordination polymer Cu(TCNQ)

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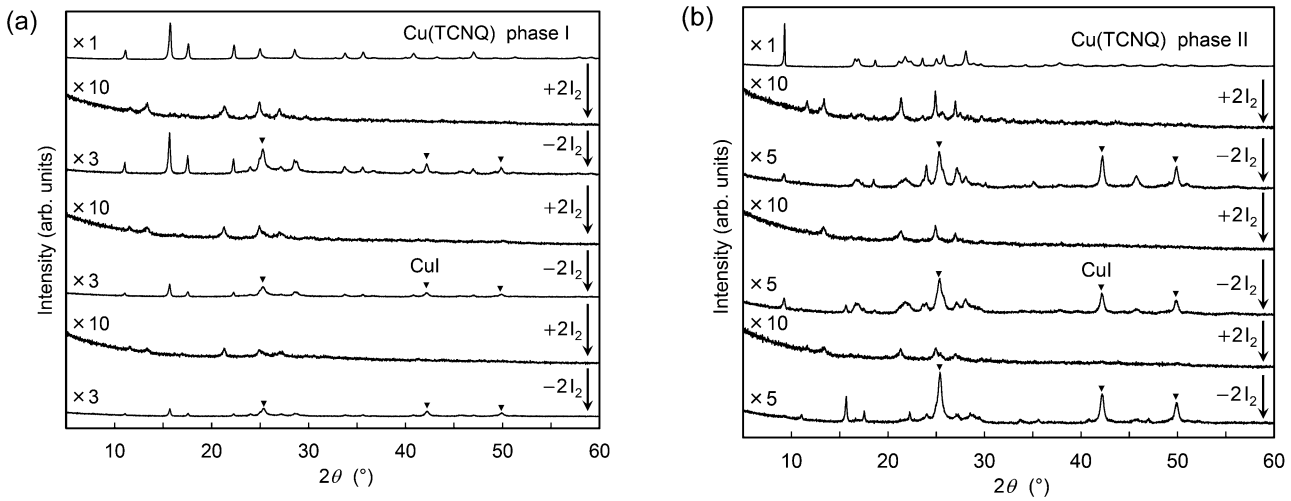

Fig. S1 Powder XRD patterns of (a) phase I and (b) phase II Cu(TCNQ) in the repeated absorption–desorption experiments. Triangles indicate peaks for CuI.

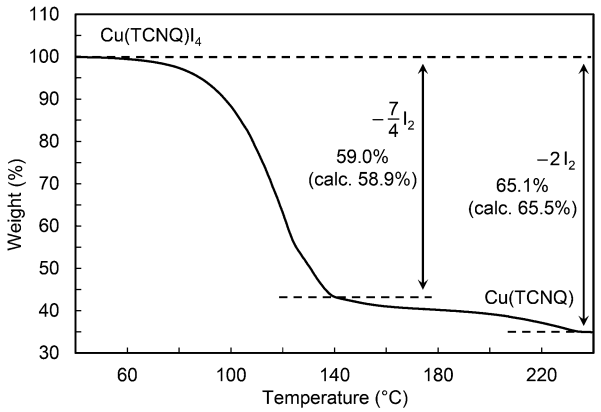

Fig. S2 Thermogravimetric traces of Cu(TCNQ)I_4 prepared by solid-state reactions of phase II Cu(TCNQ).

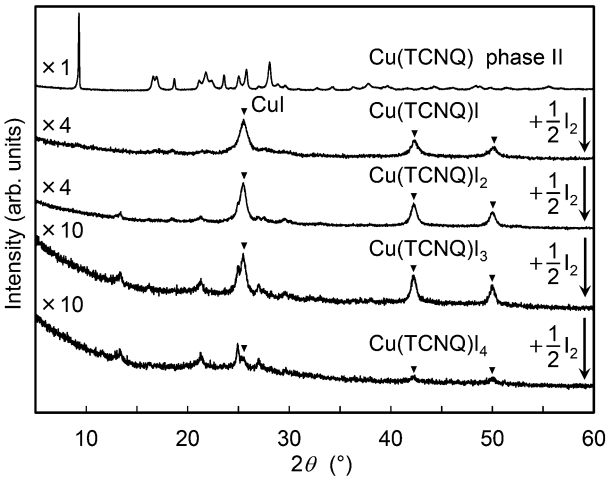

Fig. S3 Powder XRD patterns of Cu(TCNQ)I_n (n = 1, 2, 3, and 4) obtained from stepwise solid-state reactions of phase II Cu(TCNQ). Triangles indicate peaks for CuI.

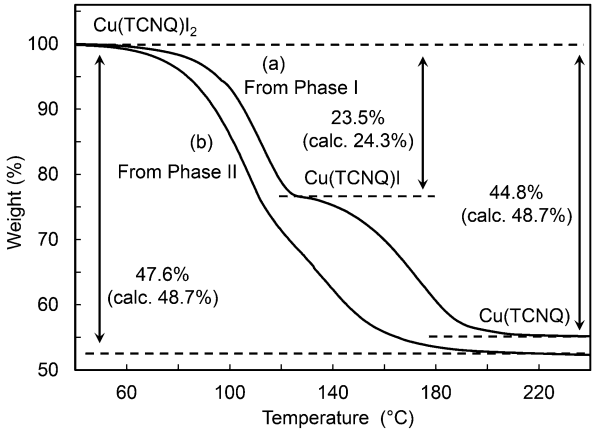

Fig. S4 Thermogravimetric traces of Cu(TCNQ)I_n (n = 2) formed by solid-state reactions of phase I Cu(TCNQ) and phase II Cu(TCNQ).

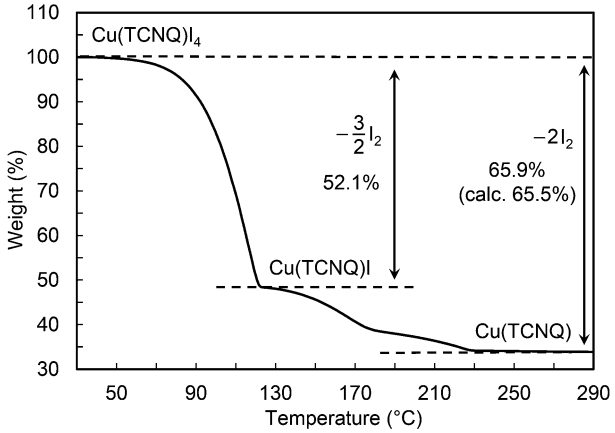

Fig. S5 Thermogravimetric trace of Cu(TCNQ)I_4 prepared by liquid-phase reaction of phase I Cu(TCNQ).

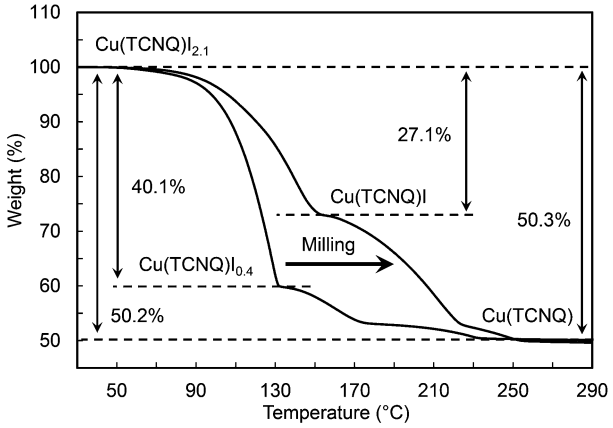

Fig. S6 Thermogravimetric trace of Cu(TCNQ)I_n (n = 2.1) obtained from liquid-phase reaction of phase I Cu(TCNQ). TG trace after grinding phase II Cu(TCNQ) sample (n = 2.1) is also shown.

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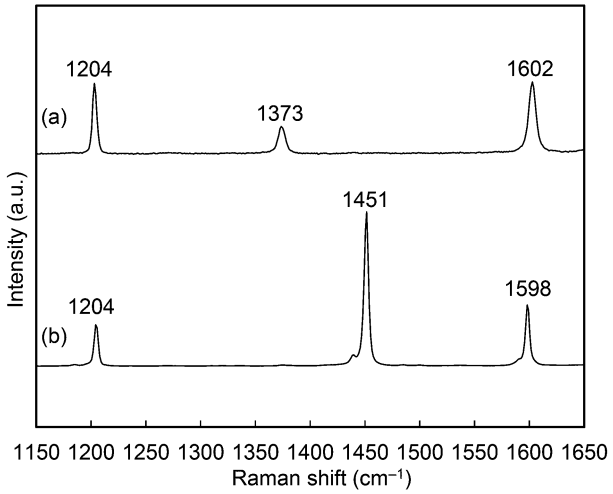

Fig. S7 Raman spectra of (a) phase I Cu(TCNQ) and (b) Cu(TCNQ)I_4 prepared by liquid-phase reaction of phase I.

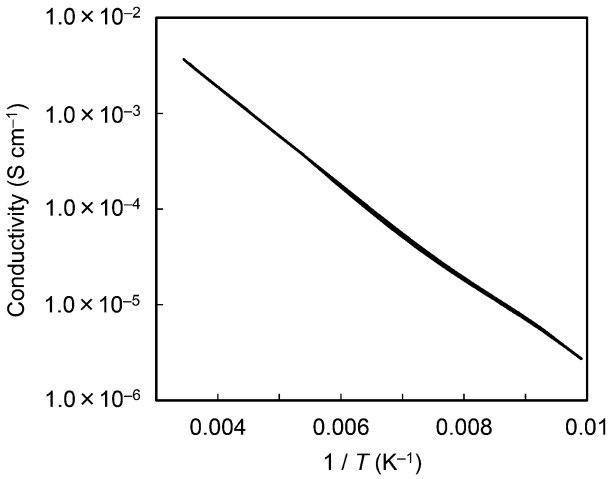

Fig. S8 Temperature dependence of electrical conductivity of compaction pellet of Cu(TCNQ)I_n (n = 3.7) prepared from phase I Cu(TCNQ) using liquid-phase reaction.