

Supplementary Information

Enhancing Photocatalytic Activity of One-Dimensional KNbO₃ Nanowires by Au Nanoparticles Under Ultraviolet and Visible-light

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Figure Captions

Figure S1. SEM image of commercial bulk-like KNbO₃ from Alfa Aesar.

Figure S2. Nitrogen adsorption and desorption isotherms measured at 77 K and BJH pore size distribution for NWs (a-b), NRs (c-d) and commercial KNbO₃ (e-f). The BET surface area for NWs, NRs, and commercial KNbO₃ is measured to be 4.8, 3.25 and 1.0 m²/g, respectively.

Figure S3. UV-vis spectral changes of RhB aqueous solutions as a function of UV irradiation time. (a) NWs. (b) NRs. (c) Commercial. Reaction conditions: RhB concentration 1×10⁻⁵ M; catalyst concentration 1 g/L; 300 W Xe-lamp (UV bandpass at 365 nm, light intensity is 158 mW/cm²).

Figure S4. UV-vis spectral changes of RhB aqueous solutions as a function of irradiation time. (a) Au 4.2 wt% on NWs under UV. (b) Au 8.0 wt% on NWs under UV. (c) Au 4.2 wt% on NWs under visible-light. (d) Au 8.0 wt% on NWs under visible-light. Reaction conditions: RhB concentration 1×10⁻⁵ M; catalyst concentration 1 g/L; 300 W Xe-lamp (visible-light λ>420 nm, light intensity is 230 mW/cm²; UV bandpass at 365 nm, light intensity is 158 mW/cm²).

Figure S5. Valence band structures for KNbO₃ nanostructures before and after Au deposition.

Table S1. Comparison of photocatalytic activity of photocatalyst (NWs+Au 8 wt%) with P25 (Degussa). The reaction rate constant (k) is normalized to the specific surface area, denoted as k_s [*J. Phys. Chem. C* **2010**, *114*, 17051-17061].

$k_s = k / (\text{catalyst concentration} \times \text{SSA})$, herein the catalyst concentration is 1 g/L.

	k (min^{-1})	SSA (m^2/g)	k_s ($\text{min}^{-1}\text{Lm}^{-2}$)
P25	0.151	50	3.02×10^{-3}
NWs+Au 8 wt%	0.0162	5.0	3.22×10^{-3}

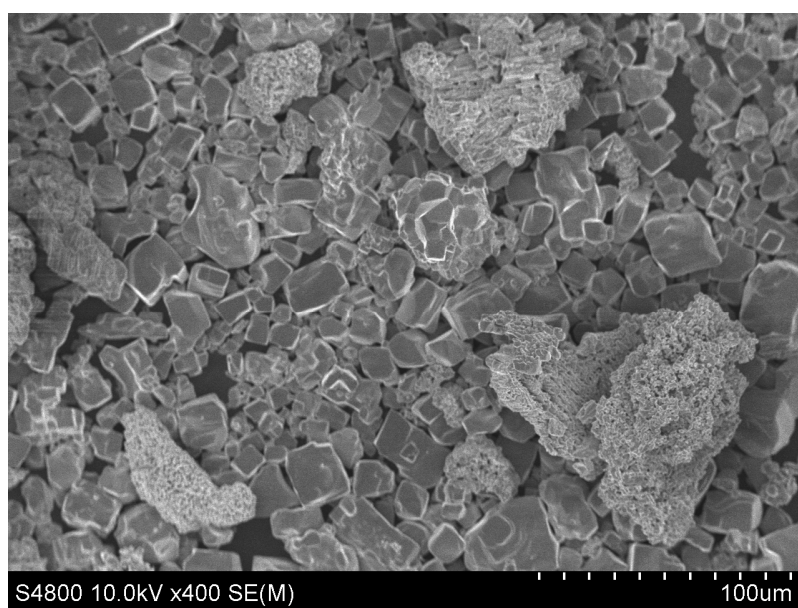


Figure S1.

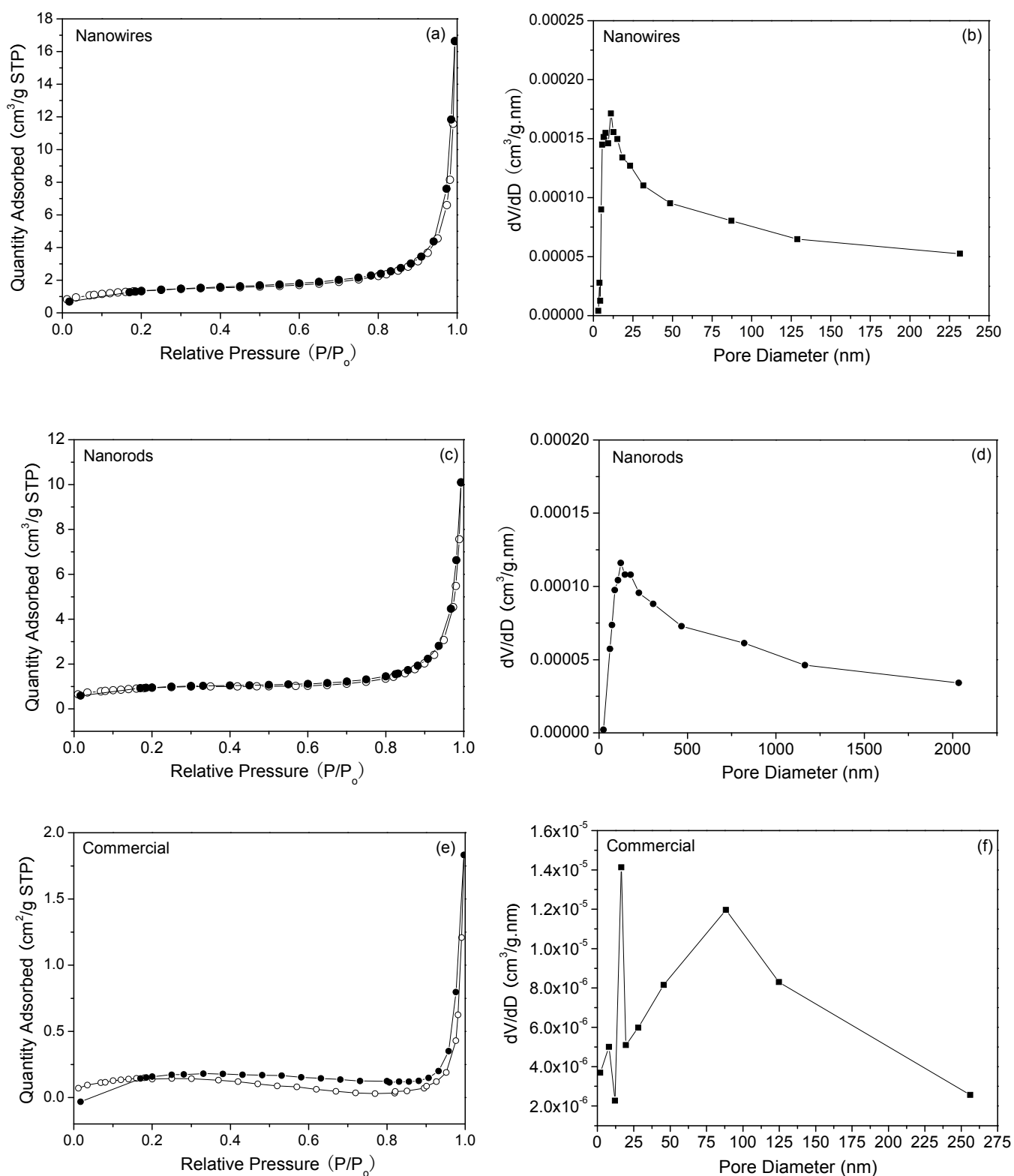


Figure S2 (a-f).

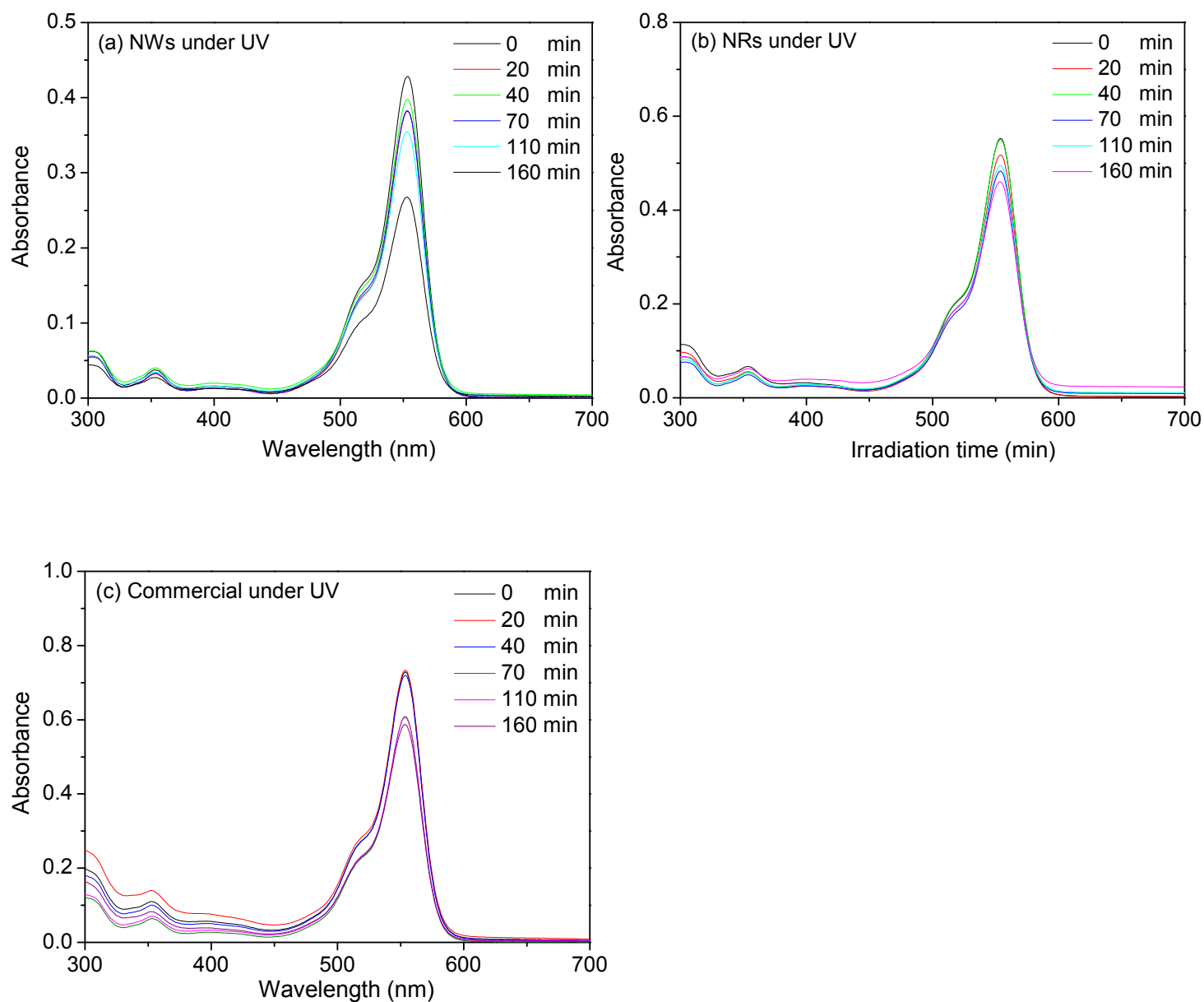


Figure S3 (a-c).

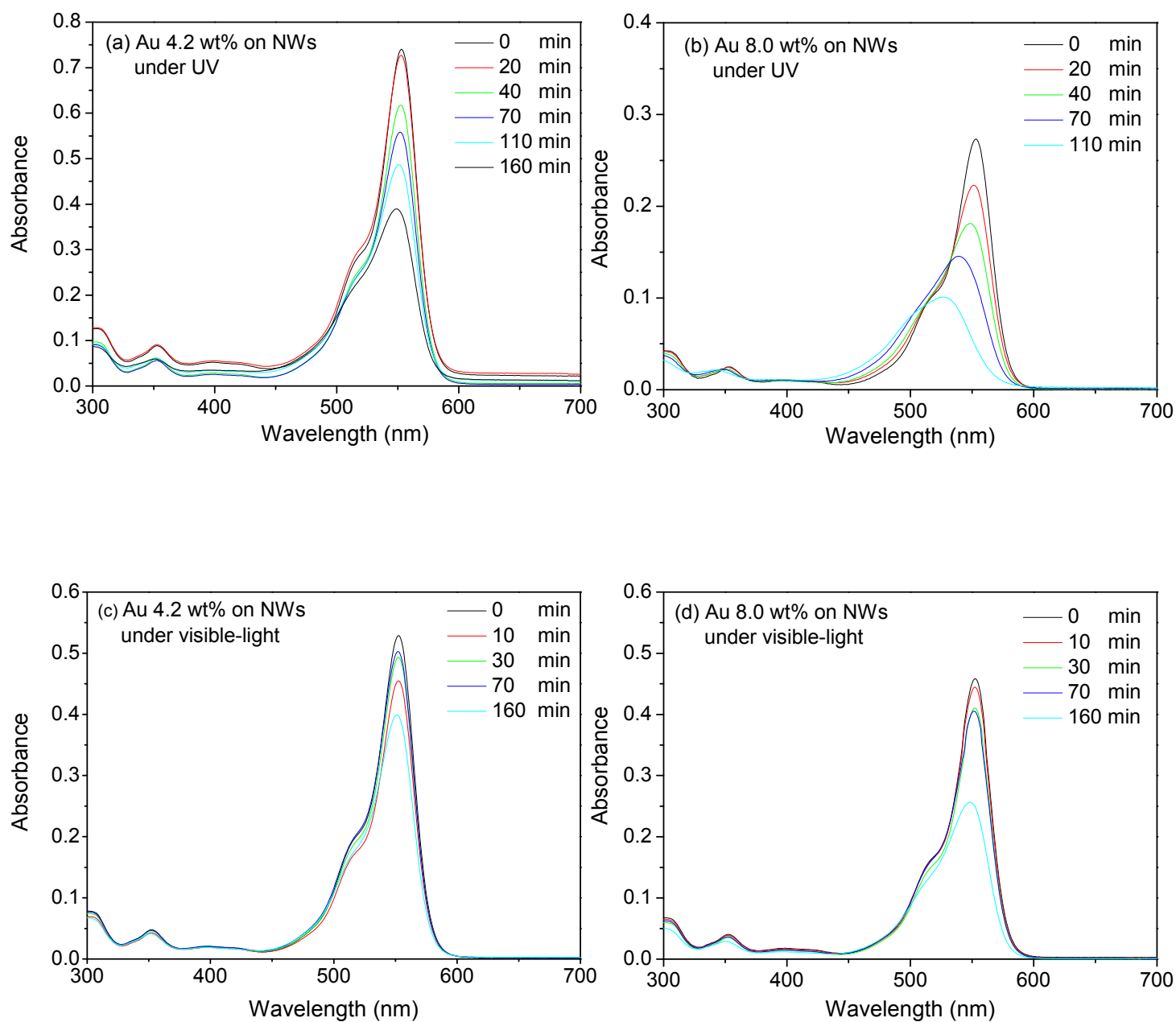


Figure S4 (a-d).

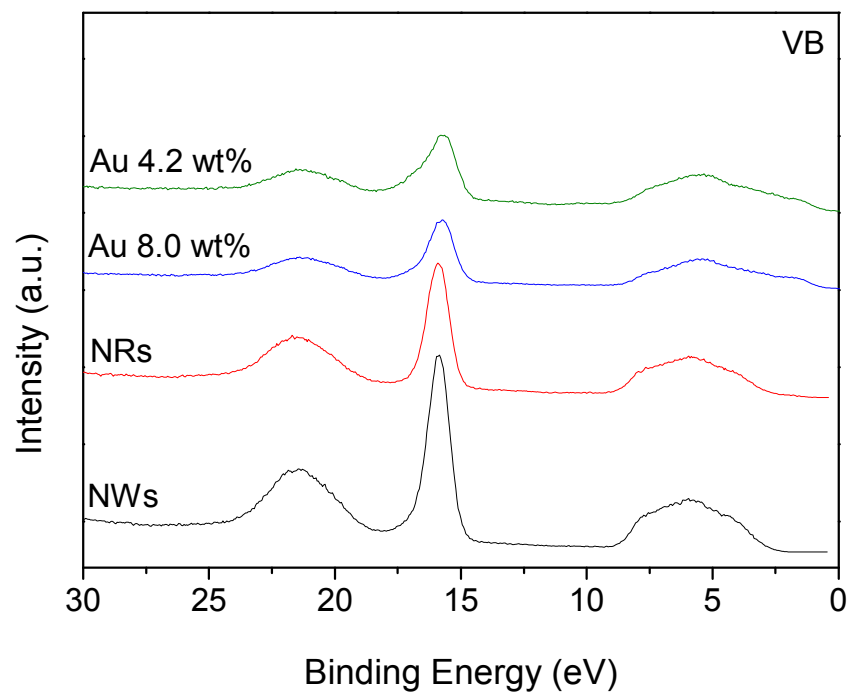


Figure S5.