

Supporting Information

The approximate calculation of N_{bulk} and N_{ads}

N_{bulk} is determined base on the 10^{-4} mol/L R6G solution and the illuminated volume of our Raman system. When determining N_{ads} in the illuminated volume of our laser, we assume that R6G molecules may be absorbed as a monolayer on the surface of ZnO/Ag and the surface of ZnO/Ag has been covered with R6G molecules after dropping R6G ethanol solution. Ag NPs are also assumed to closely arrange in the ZnO. This hypothesis represents a theoretical maximum number of R6G molecules absorbed on the surface of ZnO/Ag and is surely an overestimate, thus the EF reported here is likely an underestimate rather than an overestimate of the actual EF value.

The details of the calculation:

1. Ag NPs (29 nm)

$$V_{\text{NP}} = \frac{3}{4} \pi r^3 = 7.18 \times 10^{-18} \text{ cm}^3$$

$$m_{\text{NP}} = \rho V_{\text{NP}} = 7.56 \times 10^{-17} \text{ g}$$

$$S_{\text{NP}} = 4\pi r^2 = 2590 \text{ nm}^2$$

$$n_{\text{NP}} = m/M = 7 \times 10^{-19} \text{ mol}$$

$$N = n_{\text{NP}} \cdot N_{\text{A}} = 4.2 \times 10^5 \text{ ind}$$

2. Concentrations of Ag atoms and Ag NPs

$$C_{\text{atoms}} = n_{\text{atoms}}/V_{\text{total}} = (0.01/170) \text{ mol} / 40 \text{ ml} = 1.47 \times 10^{-3} \text{ mol/L}$$

$$C_{\text{NPs}} = C_{\text{atoms}}/N = 1.47 \times 10^{-3} \text{ mol/L} / 4.2 \times 10^5 = 3.5 \times 10^{-9} \text{ mol/L}$$

3. Mole of Ag NPs

$$n_{\text{all}} = C_{\text{NPs}} \cdot V_{\text{total}} = 3.5 \times 10^{-9} \text{ mol/L} \cdot 40 \text{ ml} = 1.4 \times 10^{-10} \text{ mol}$$

4. Area of the substrate

$$S = 2 \times 1 \text{ cm}^2 = 2 \times 10^{-4} \text{ m}^2$$

5. Mole of Ag NPs in per unit substrate

$$c_{\text{Ag}} = n_{\text{all}}/S = 1.4 \times 10^{-10} \text{ mol} / 2 \times 10^{-4} \text{ m}^2 = 7 \times 10^{-7} \text{ mol/m}^2$$

6. R6G numbers absorbed in single Ag NP

$$N_{\text{ads}} = S_{\text{NP}} \times 0.25 \text{ nm}^2 = 647.5$$

7. R6G numbers absorbed in per unit substrate and concentration of R6G in ethanol

$$c_{\text{ads}} = c_{\text{Ag}} \times N_{\text{ads}} = 7 \times 10^{-7} \text{ mol/m}^2 \times 647.5 = 4.5 \times 10^{-4} \text{ mol/m}^2$$

$$C_{\text{bulk}} = 2 \text{ mol/m}^3$$

8. EF (L is the coke column length of laser penetrating R6G solution; D1 and D2 are naphthalene column diameter of the liquid sample and the solid sample, respectively. L = 1.9 × 10⁻² m;

$$D_1/D_2 = 10)$$

$$EF_{1360} = EF = (I_{\text{SERS}}/I_{\text{Raman}}) \times (N_{\text{bulk}}/N_{\text{ads}}) = (I_{\text{SERS}}/I_{\text{Raman}}) \times (C_{\text{bulk}} \times L/c_{\text{ads}}) \times (D_1/D_2)^2 = (3169/1\text{s}) / (520/60\text{s}) \times (2/4.5 \times 10^{-4}) \times 1.9 = 3 \times 10^6.$$

$$EF_{1575} = EF = (I_{\text{SERS}}/I_{\text{Raman}}) \times (N_{\text{bulk}}/N_{\text{ads}}) = (I_{\text{SERS}}/I_{\text{Raman}}) \times (C_{\text{bulk}} \times L/c_{\text{ads}}) \times (D_1/D_2)^2 = (1474/1\text{s}) / (251/60\text{s}) \times (2/4.5 \times 10^{-4}) \times 1.9 = 3 \times 10^6 = 2.98 \times 10^6.$$