

## Electrosorption of LiCl in different solvents by carbon nanotubes film electrodes

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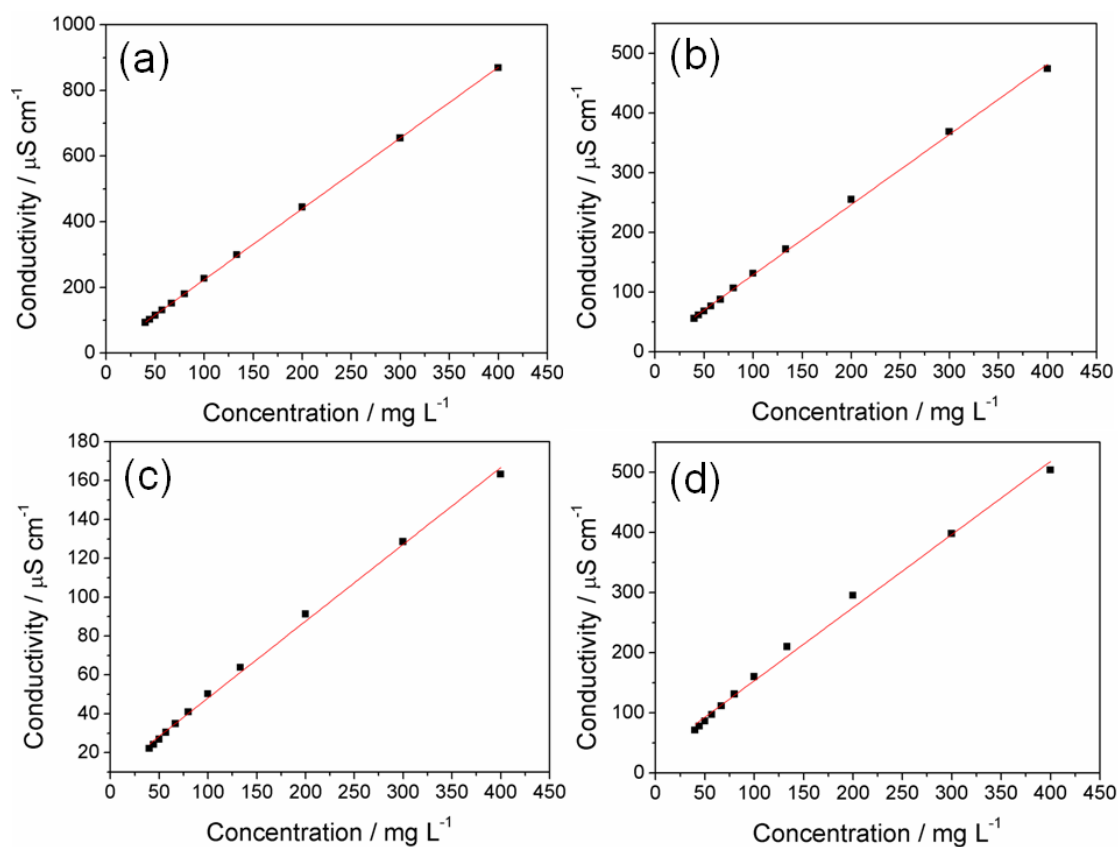


Fig. S1 The relationship between conductivity and LiCl concentration in (a) water (b) methanol (c) ethanol (d) DMF solutions.

Fig. S1(a)-(d) show the relationship between solution conductivity  $z$  ( $\mu\text{S cm}^{-1}$ ) and LiCl concentration  $C$  ( $\text{mg L}^{-1}$ ) in water, methanol, ethanol and DMF solutions, respectively. It can be seen that the relationship follows the linear plot based on the following equations:

Water:  $z = 2.158 \times C + 0.816$  (regression coefficient  $R^2 = 0.999$ )

Methanol:  $z = 1.177 \times C + 11.566$  ( $R^2 = 0.998$ )

Ethanol:  $z = 0.395 \times C + 8.649$  ( $R^2 = 0.997$ )

DMF:  $z = 1.216 \times C + 32.118$  ( $R^2 = 0.994$ )