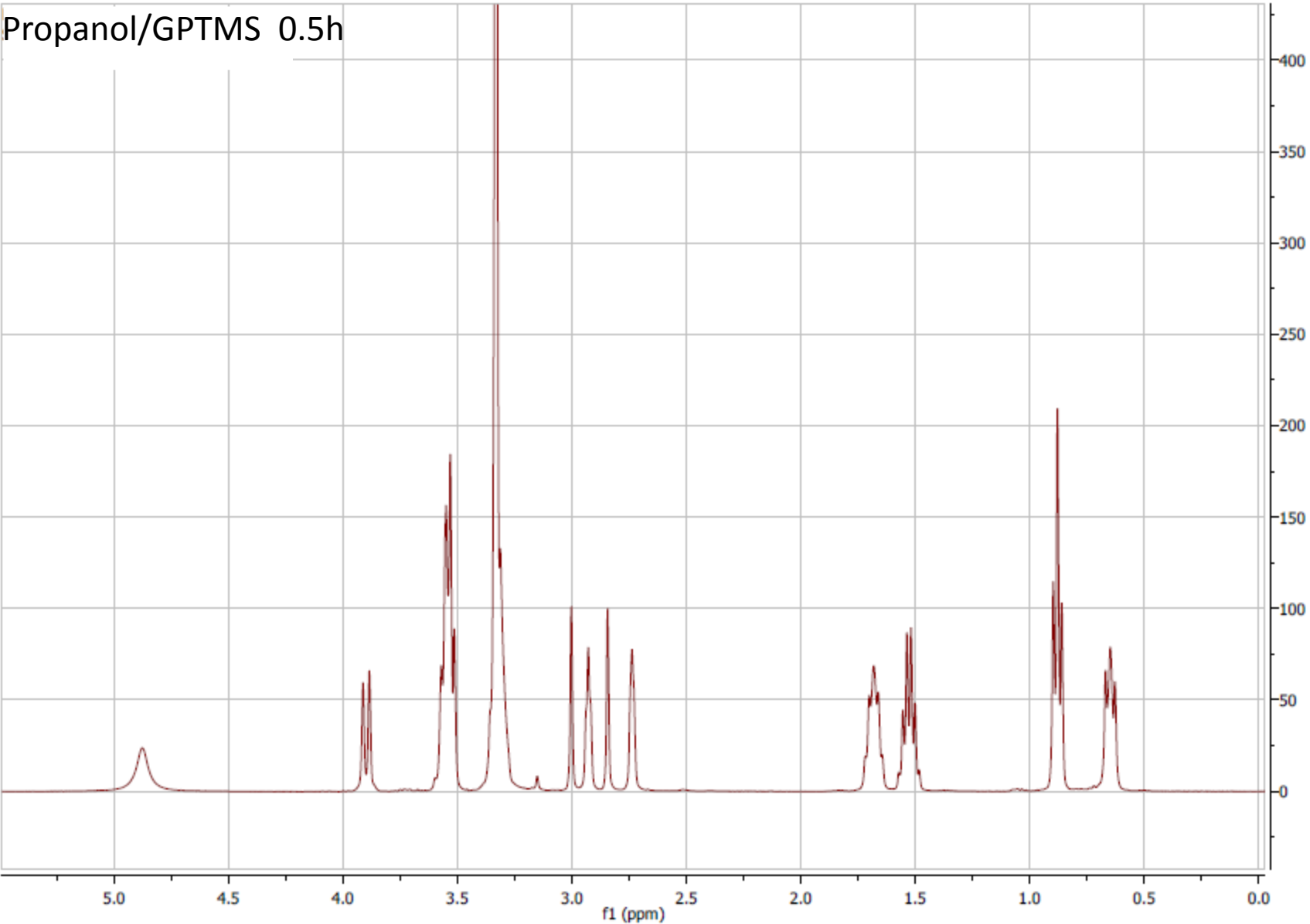
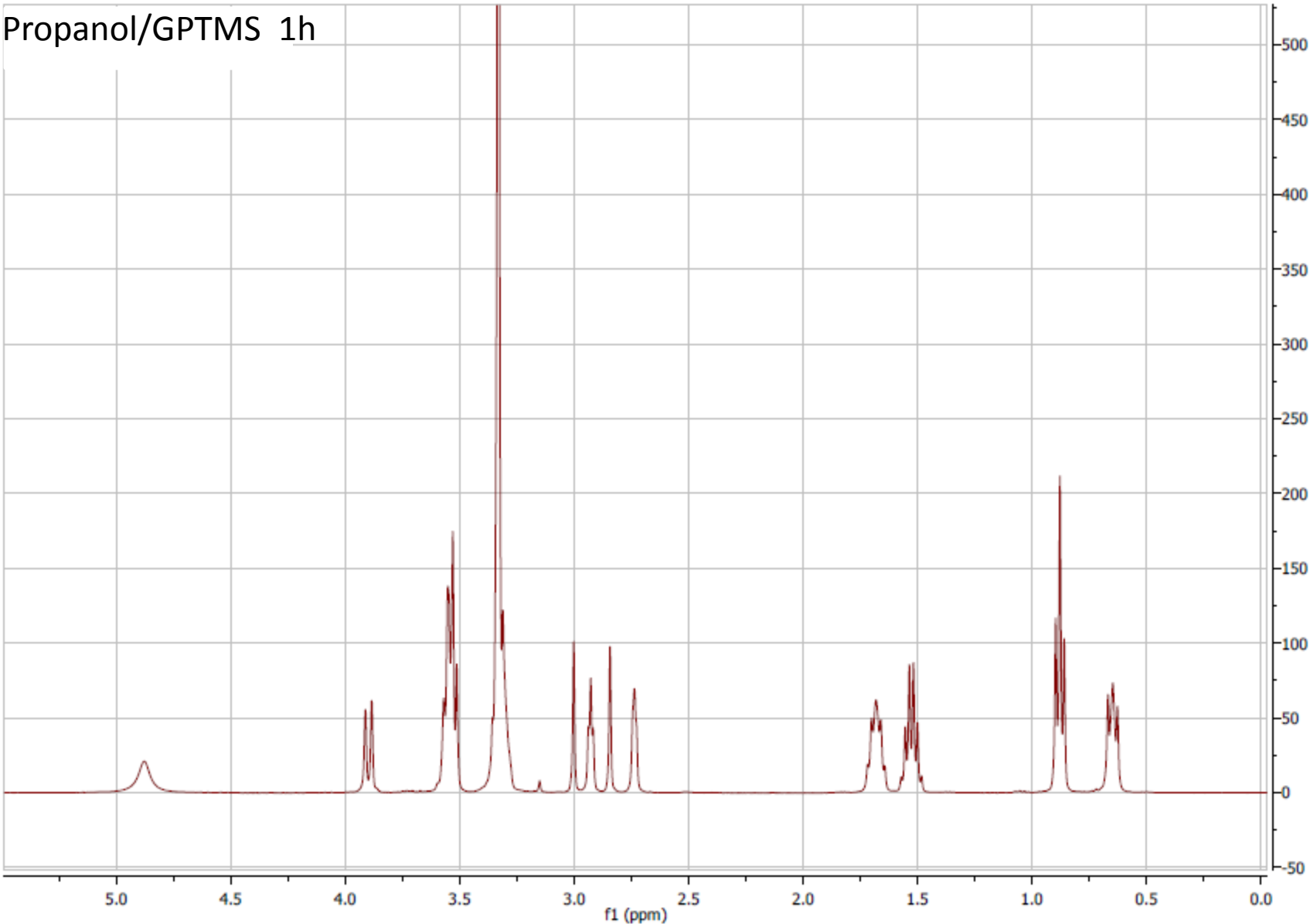
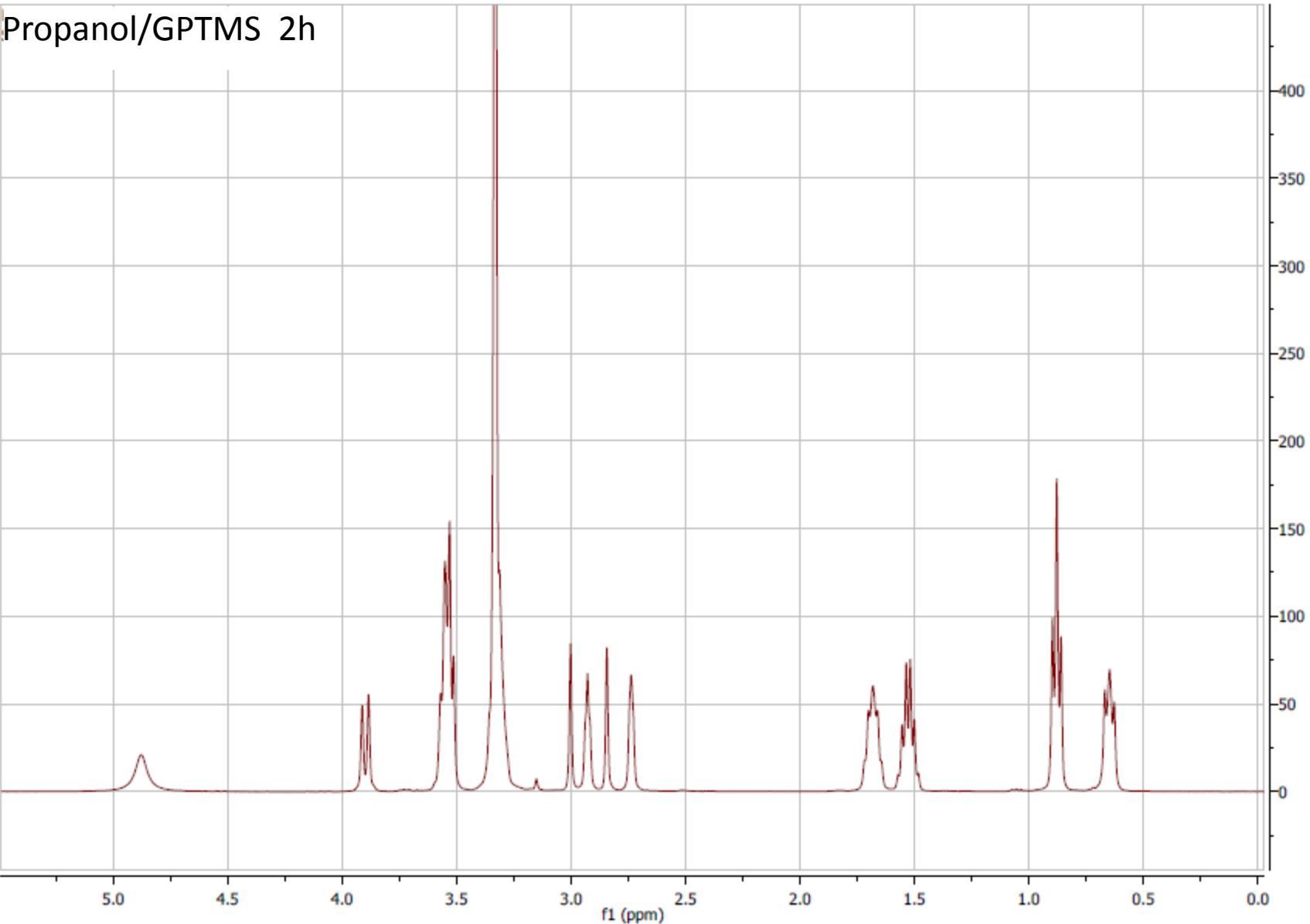
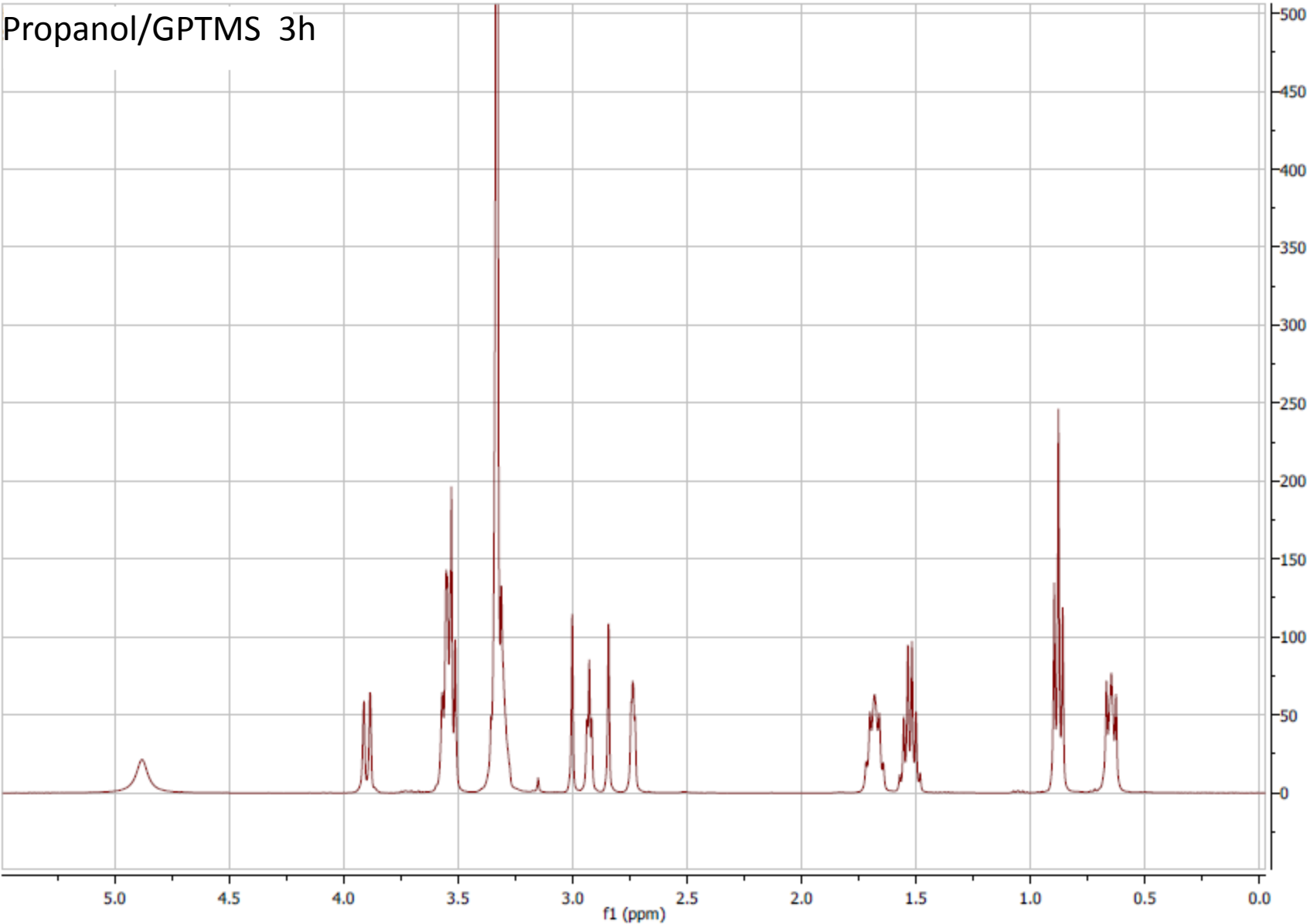


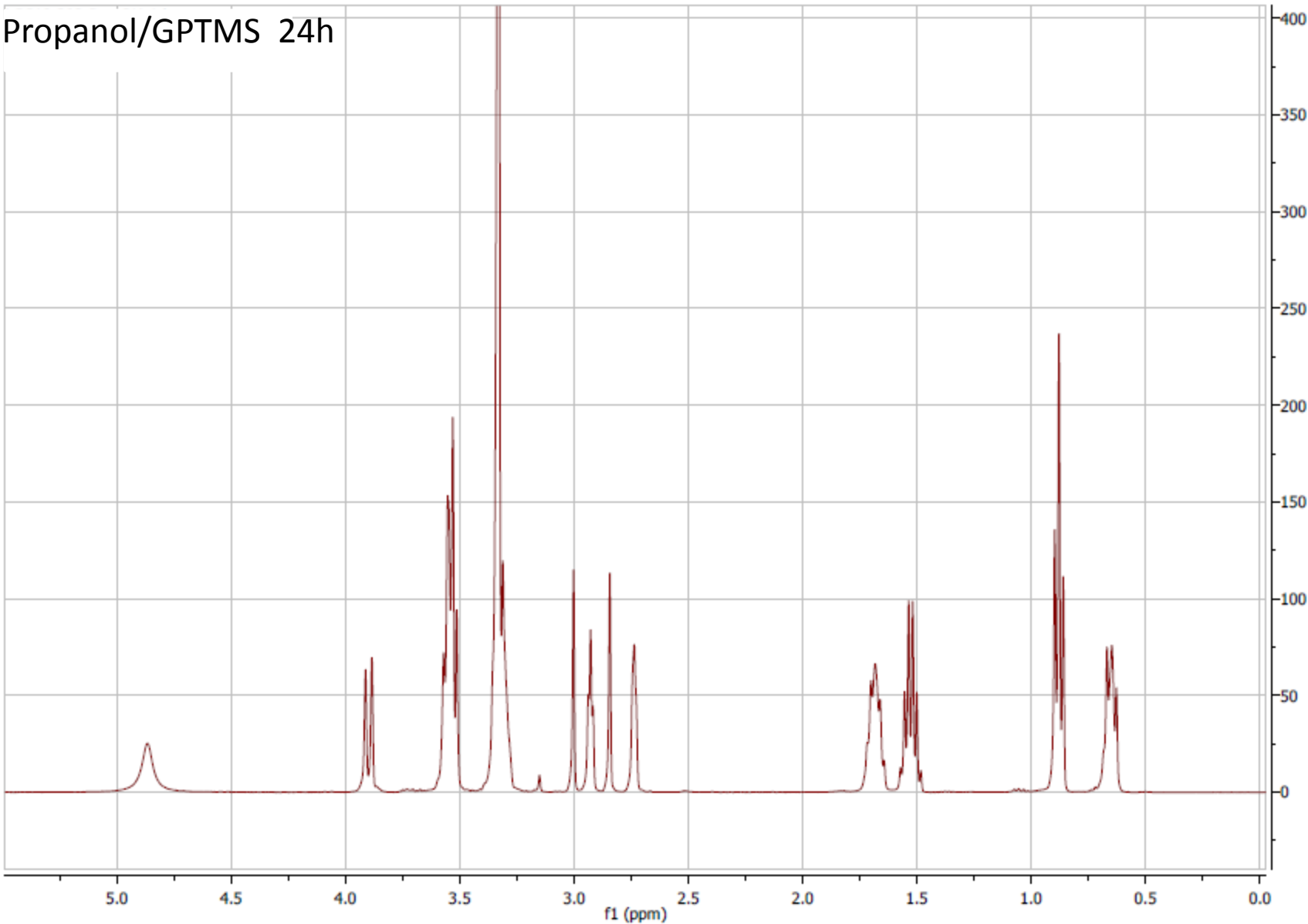
Propanol reaction NMR spectra



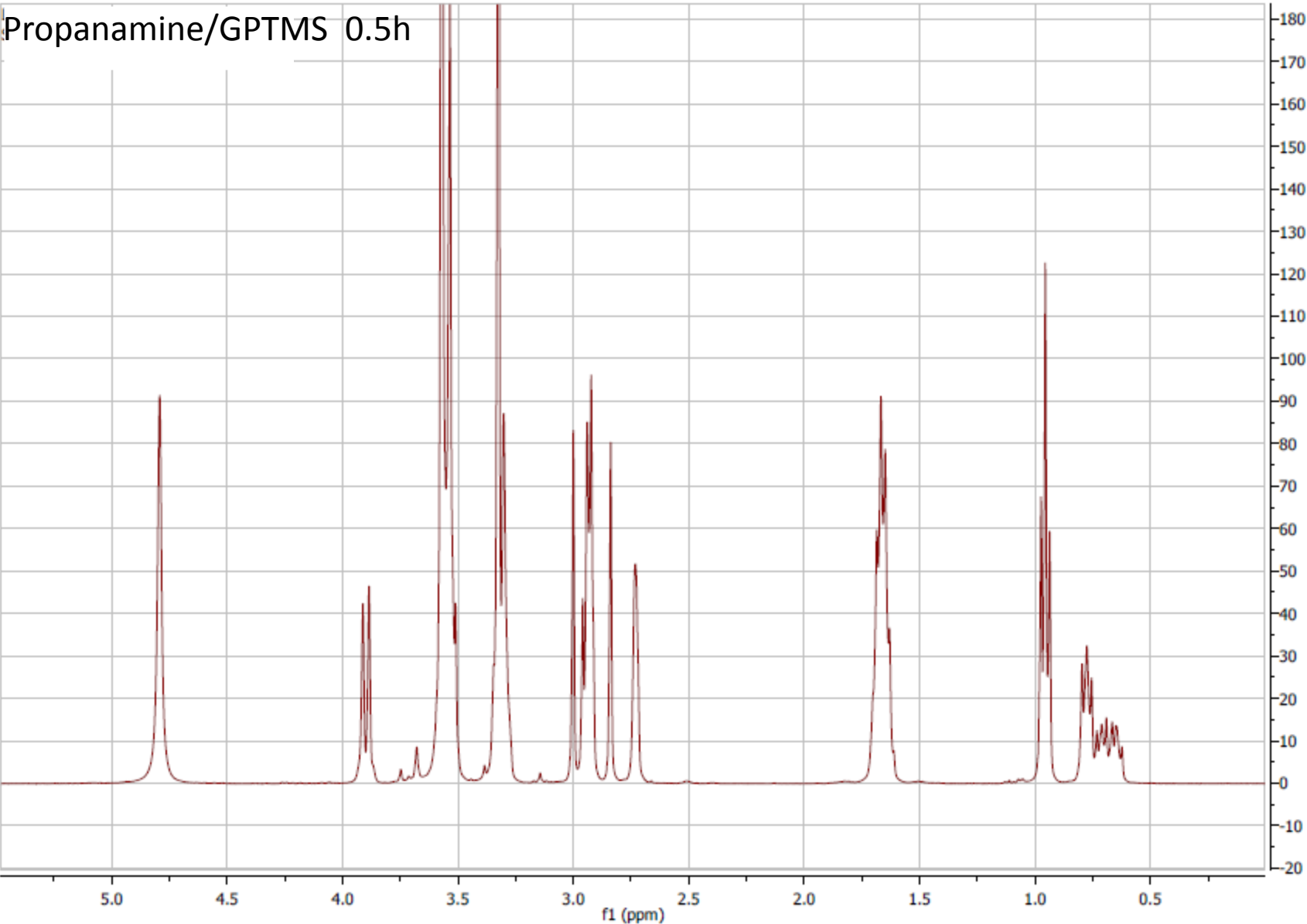


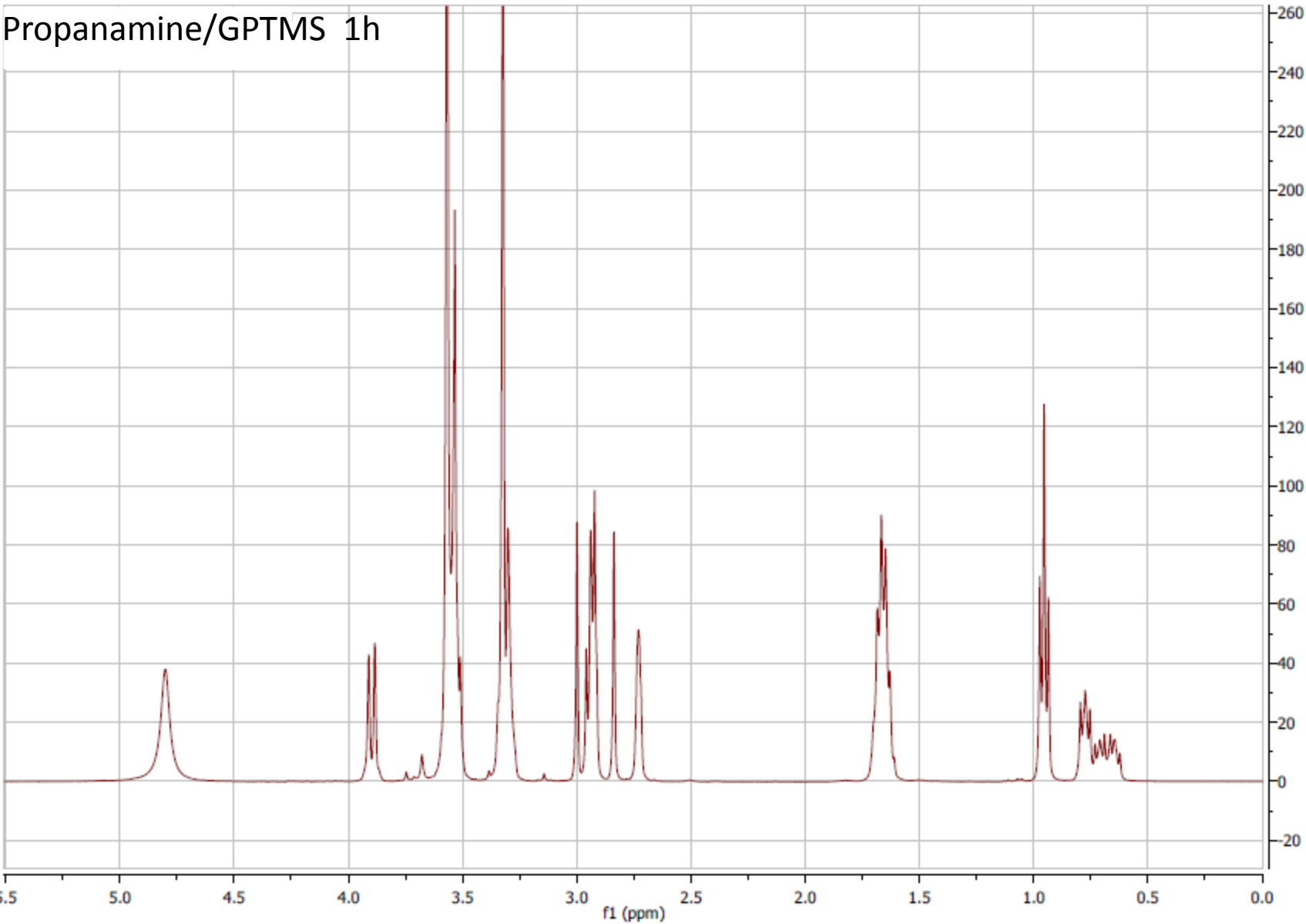


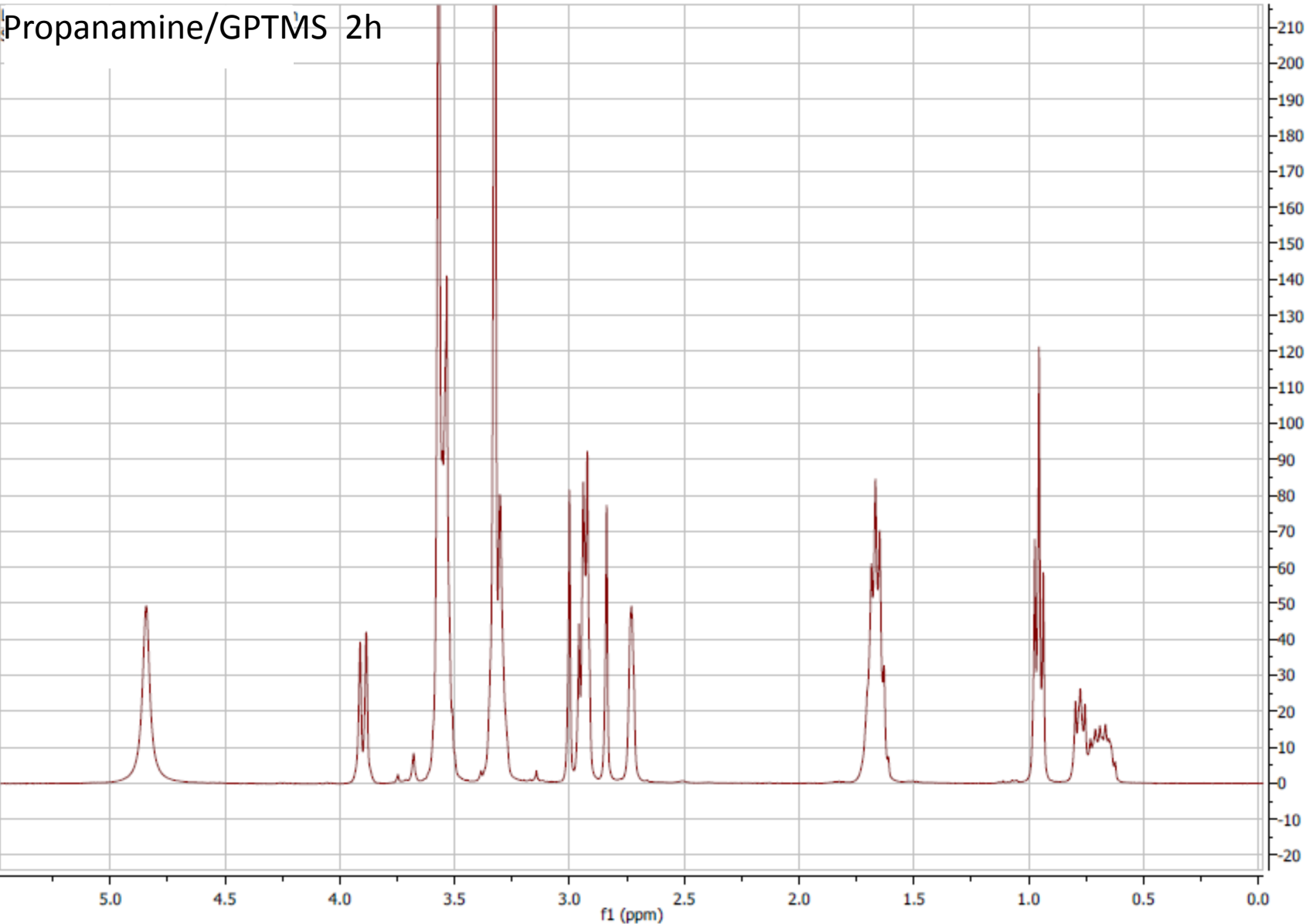


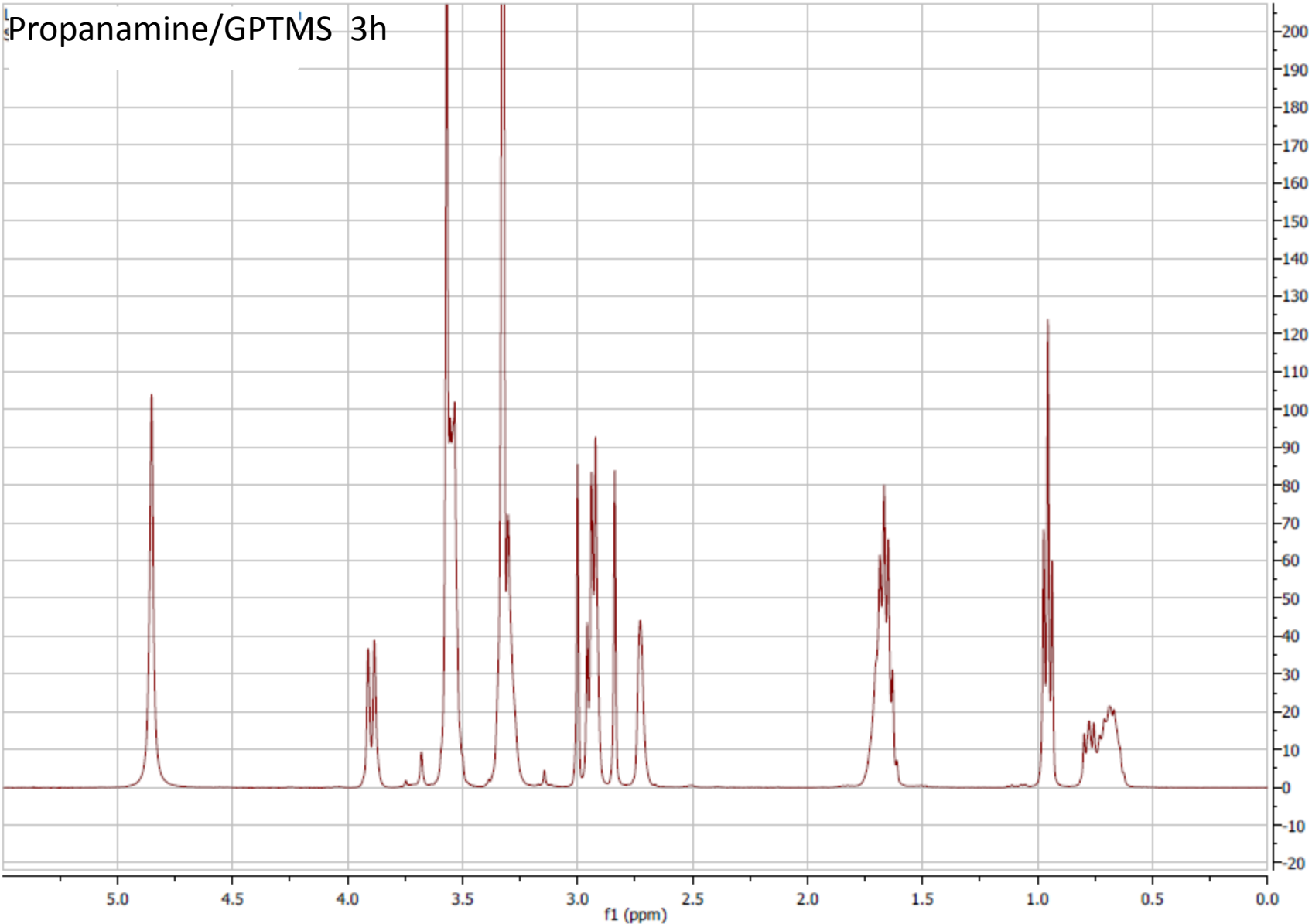


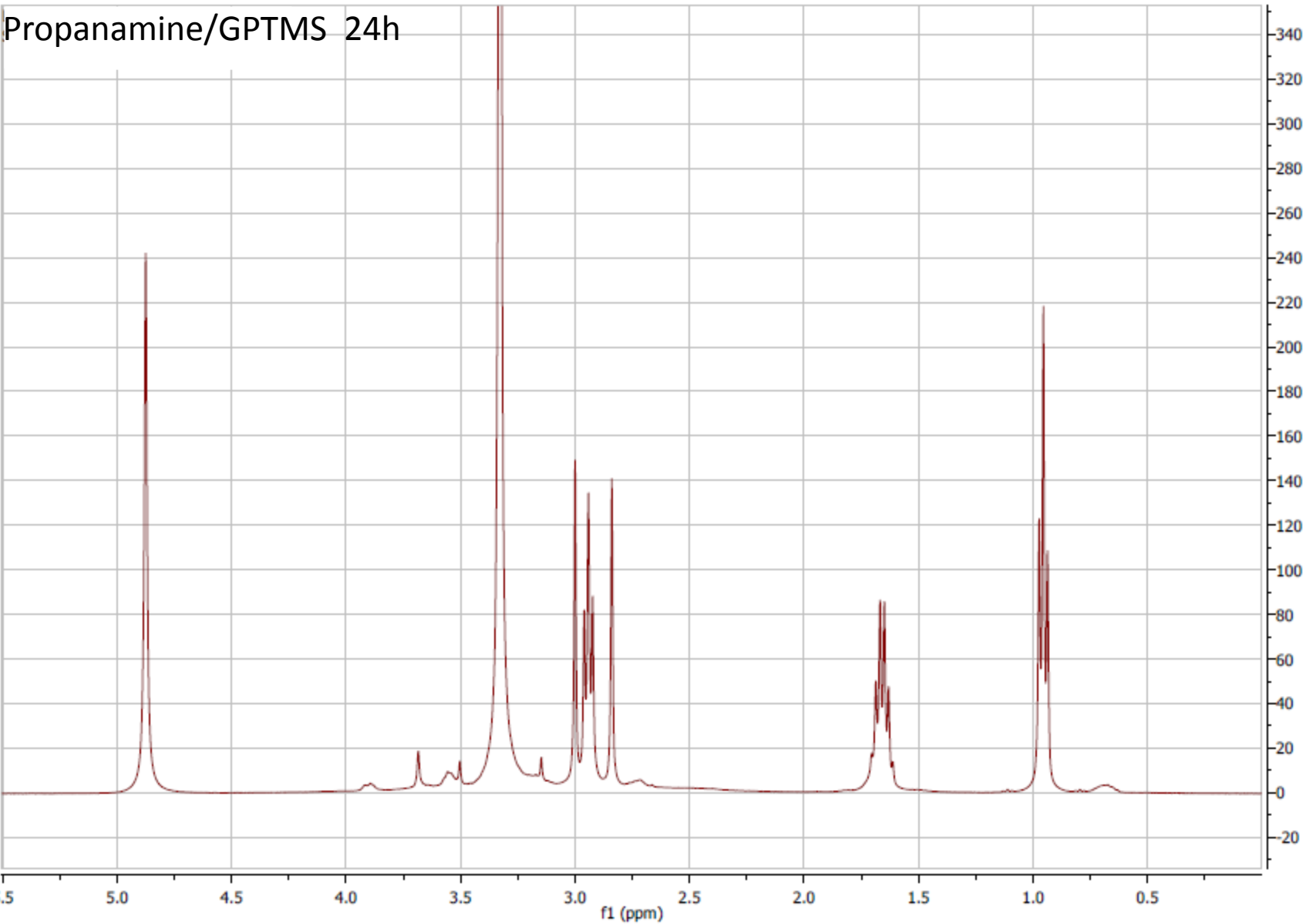
Propanamine reaction NMR spectra



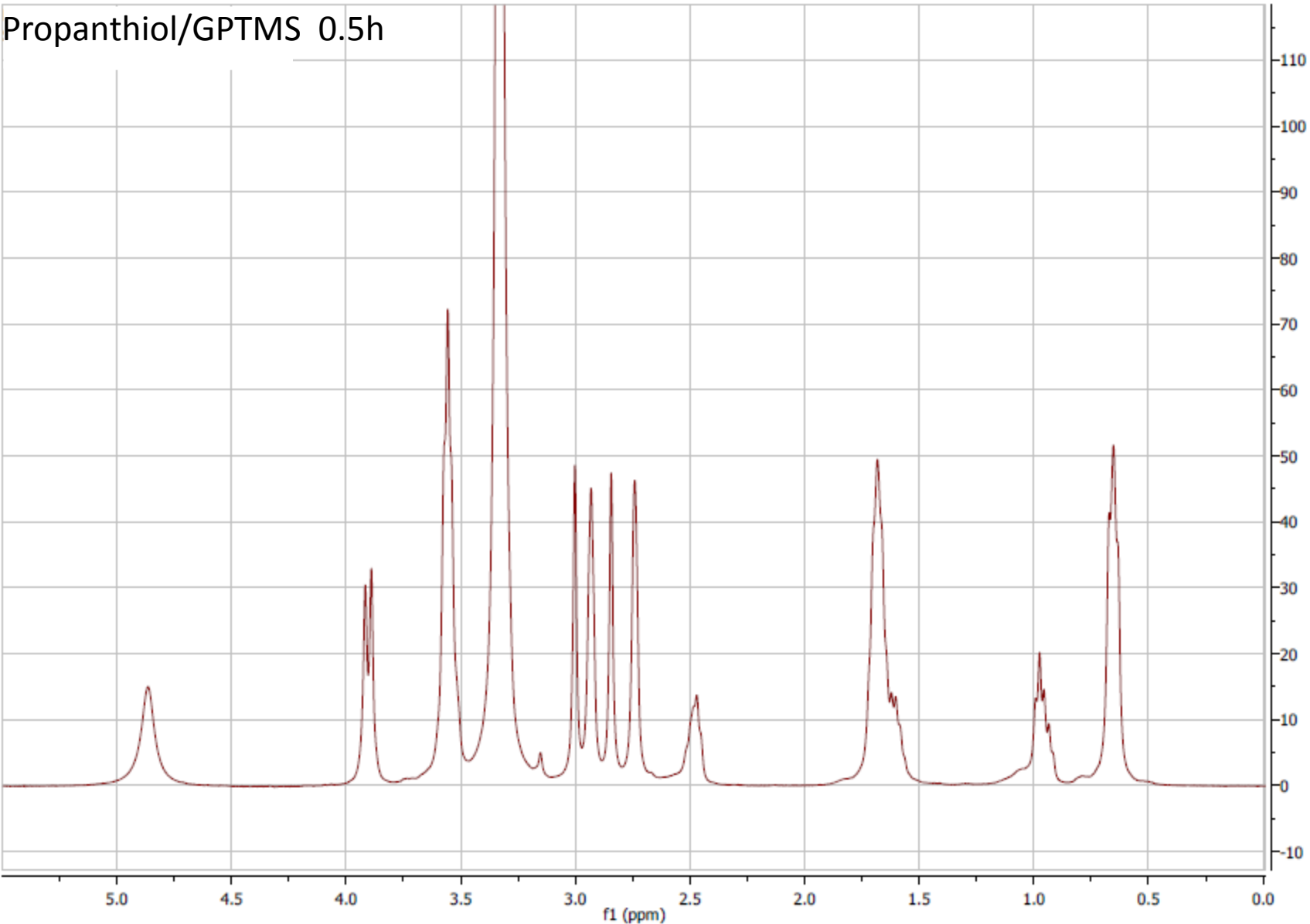


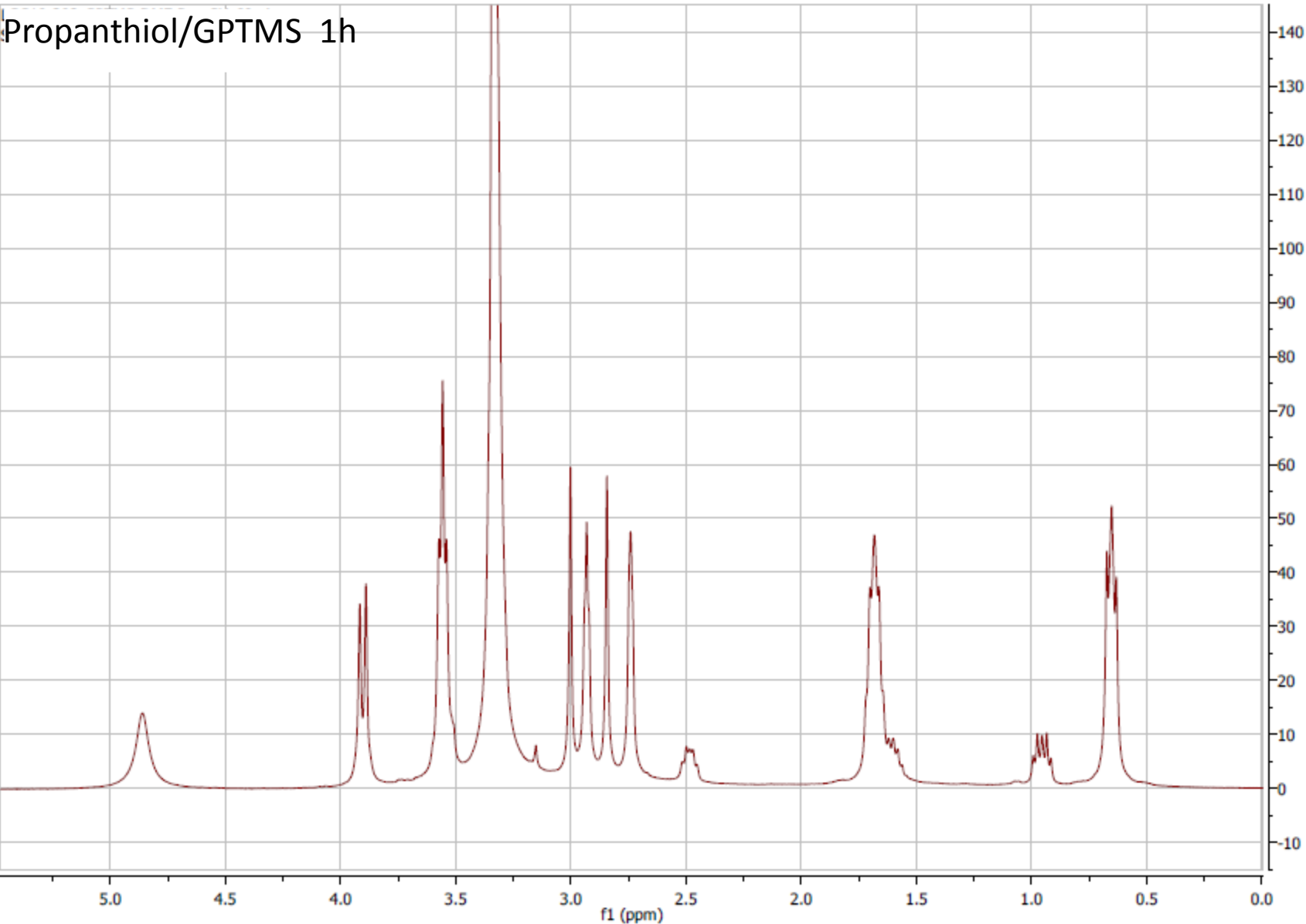


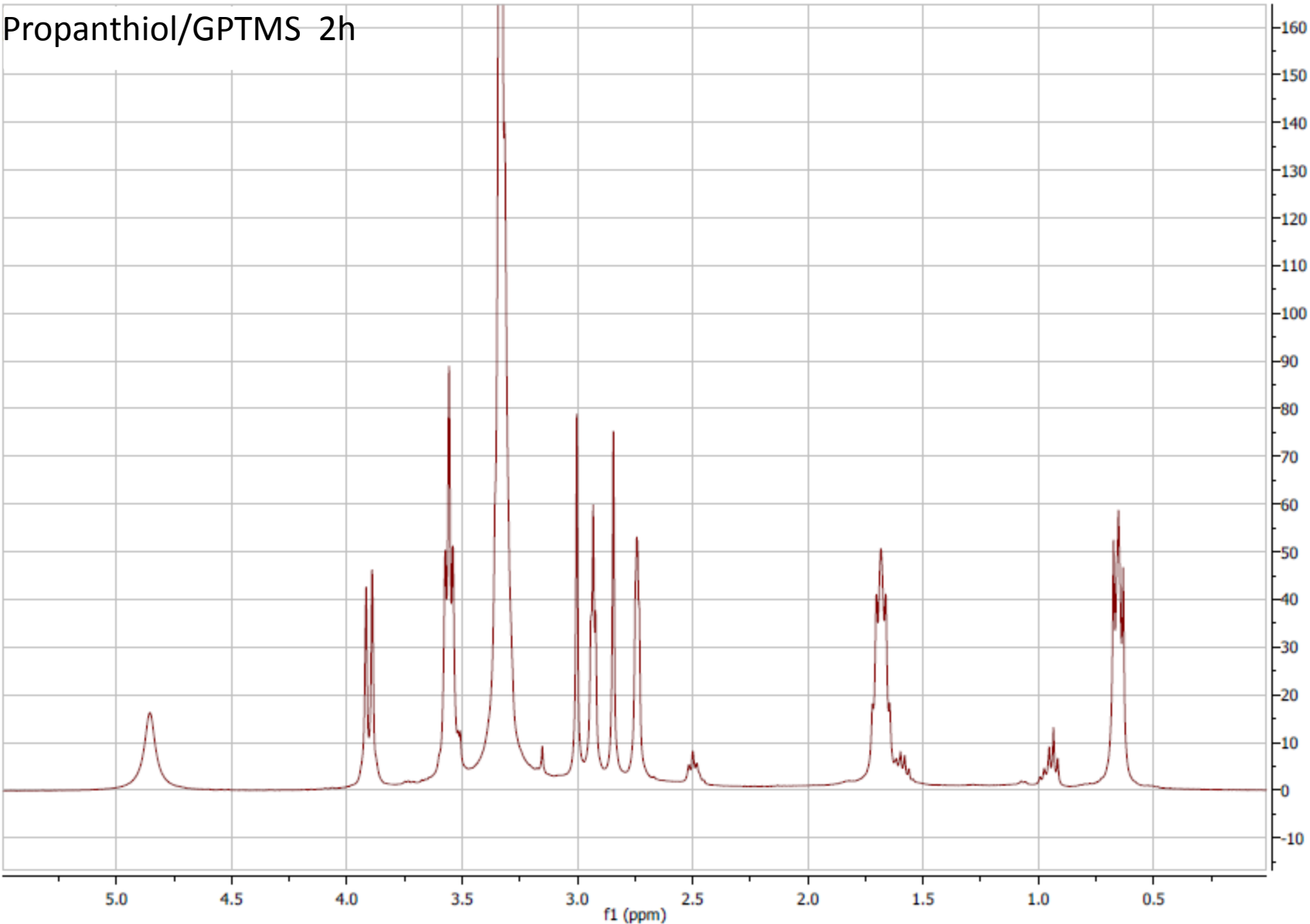


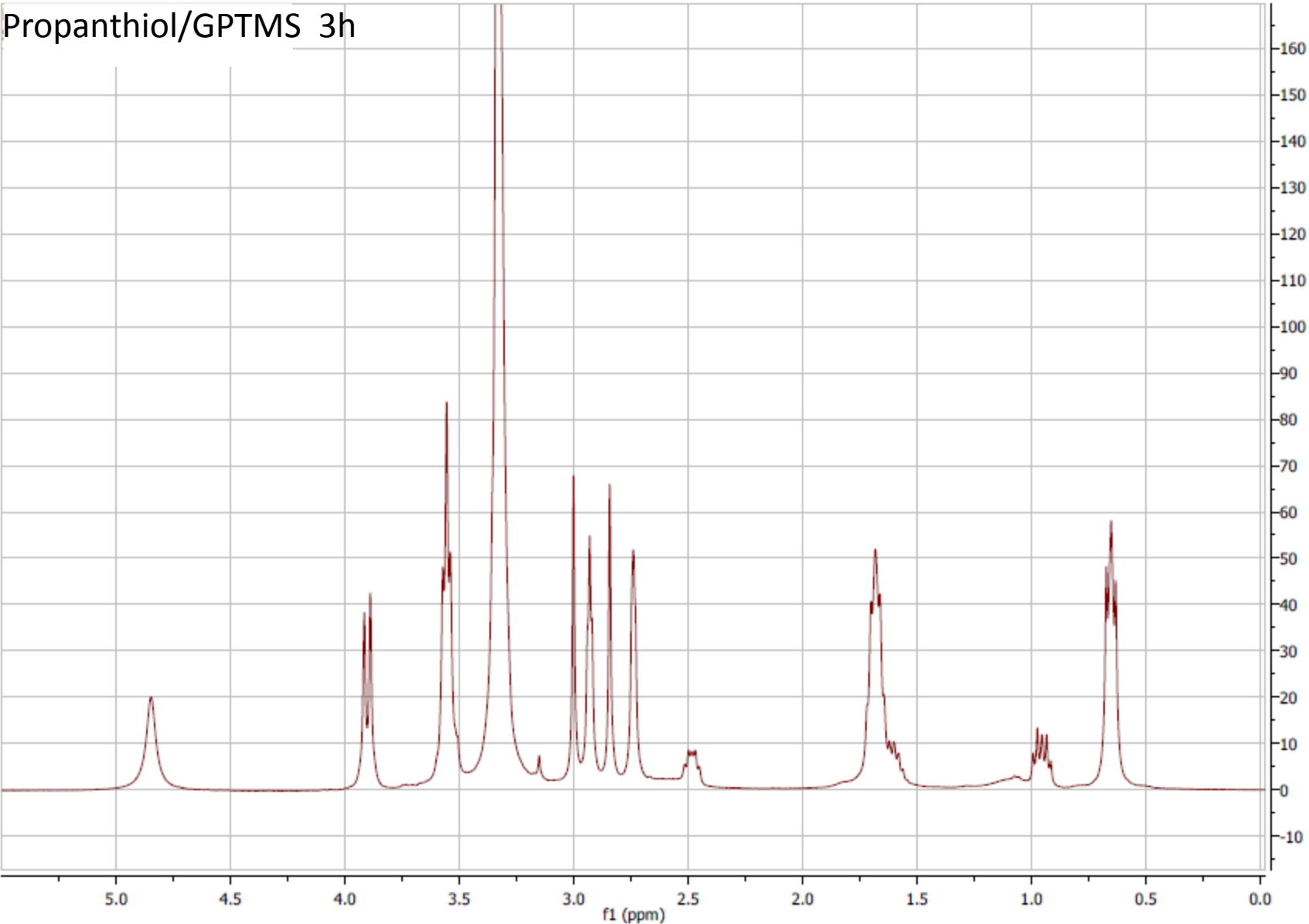


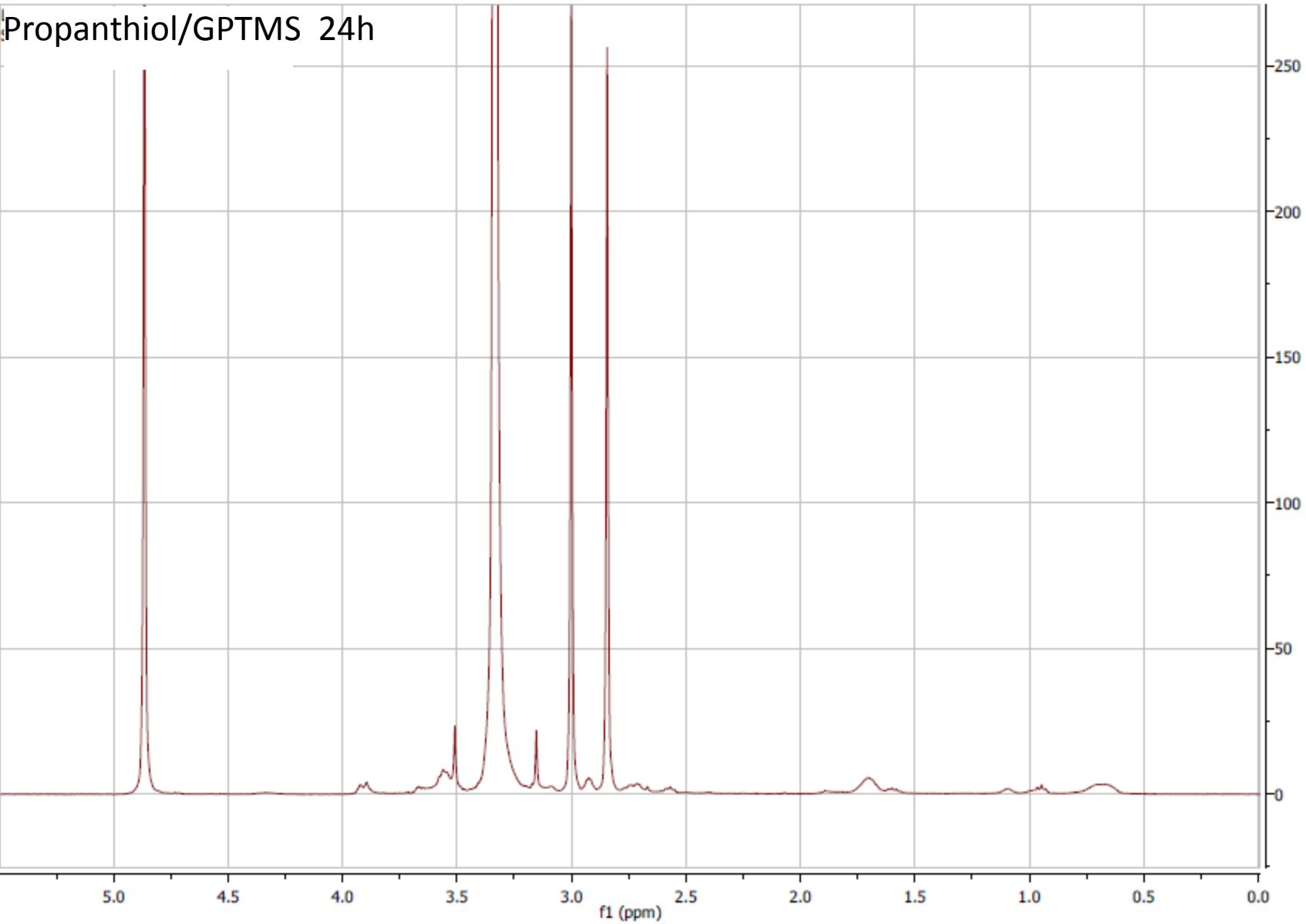
Propanthiol reaction NMR spectra



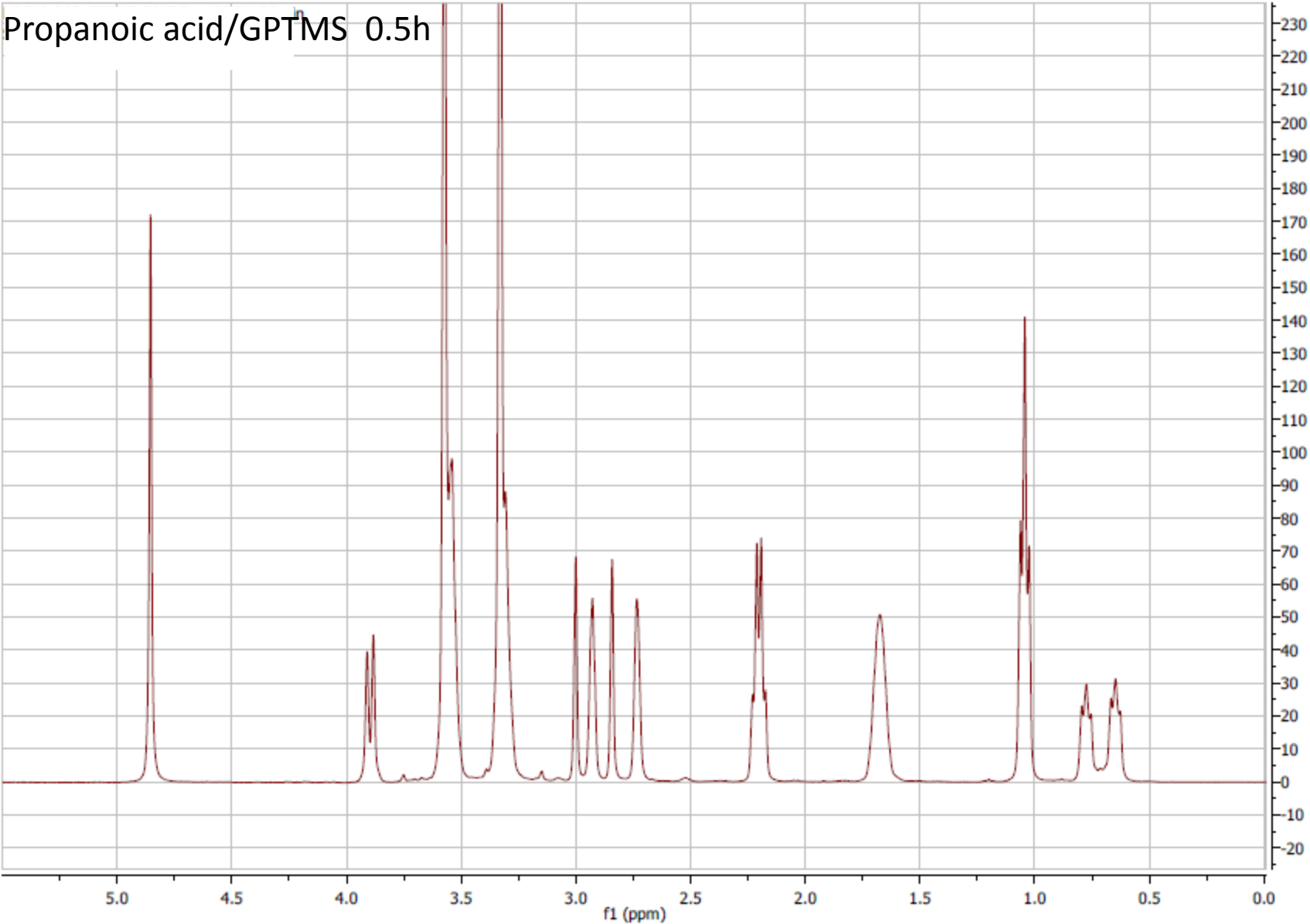


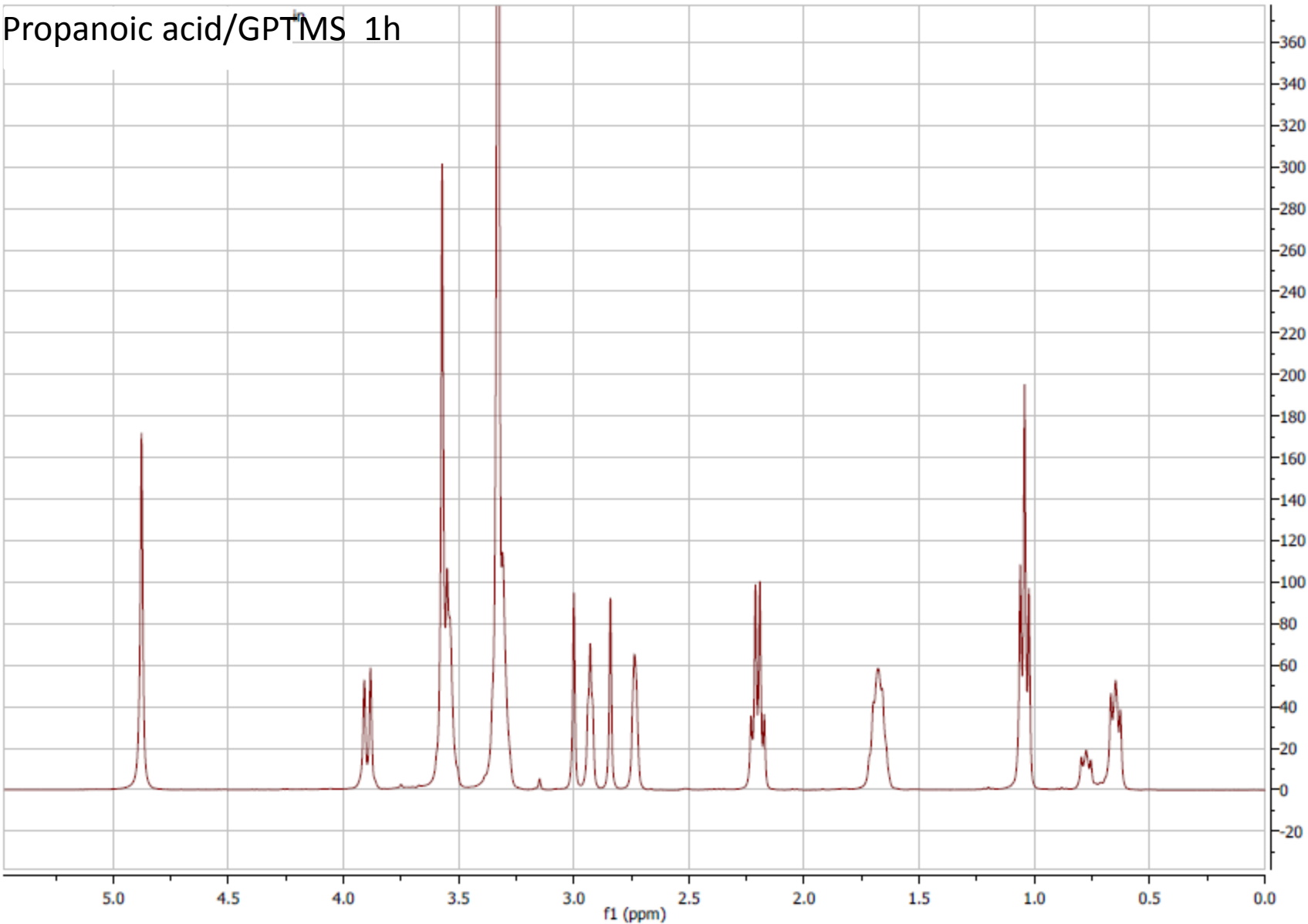


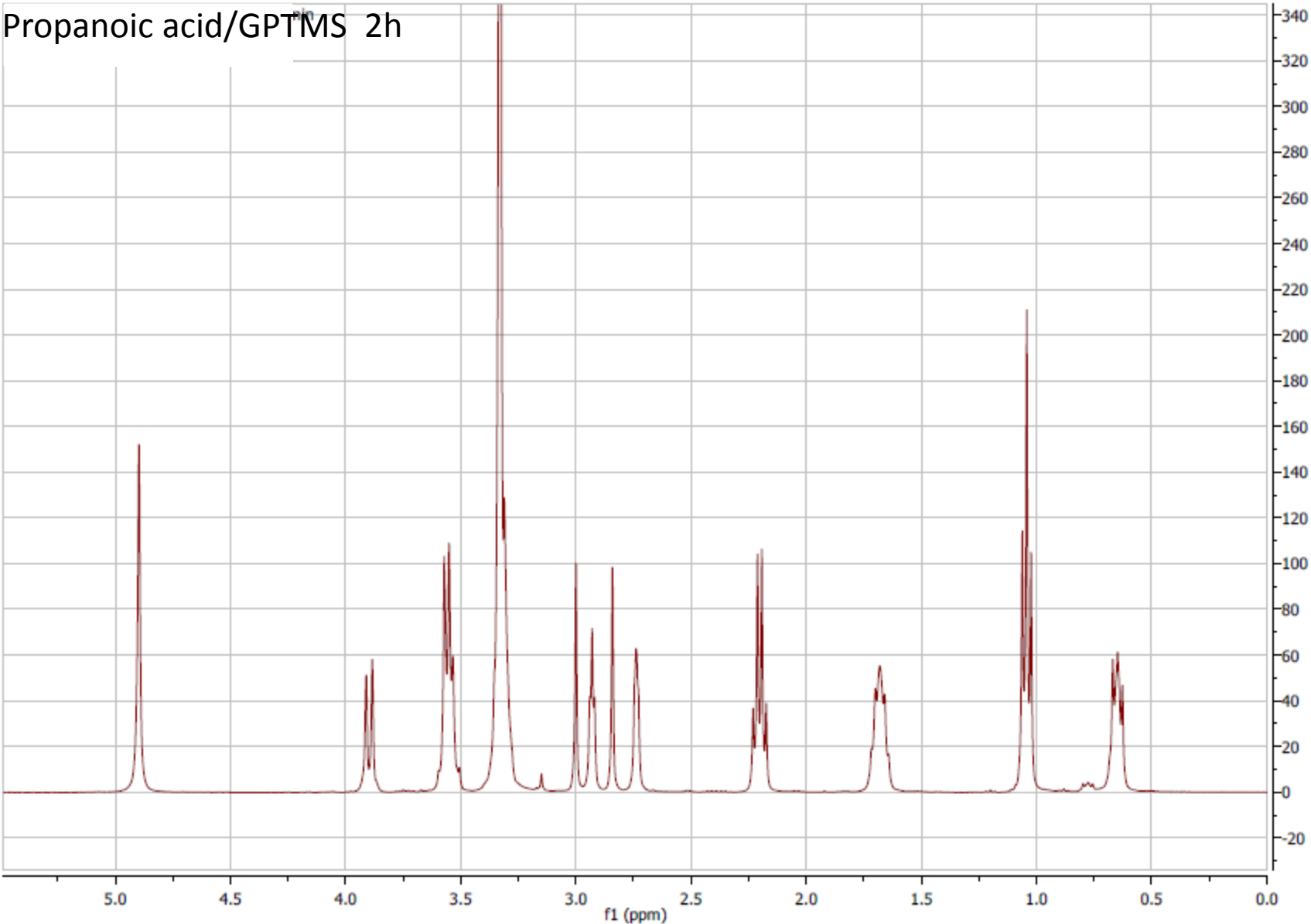




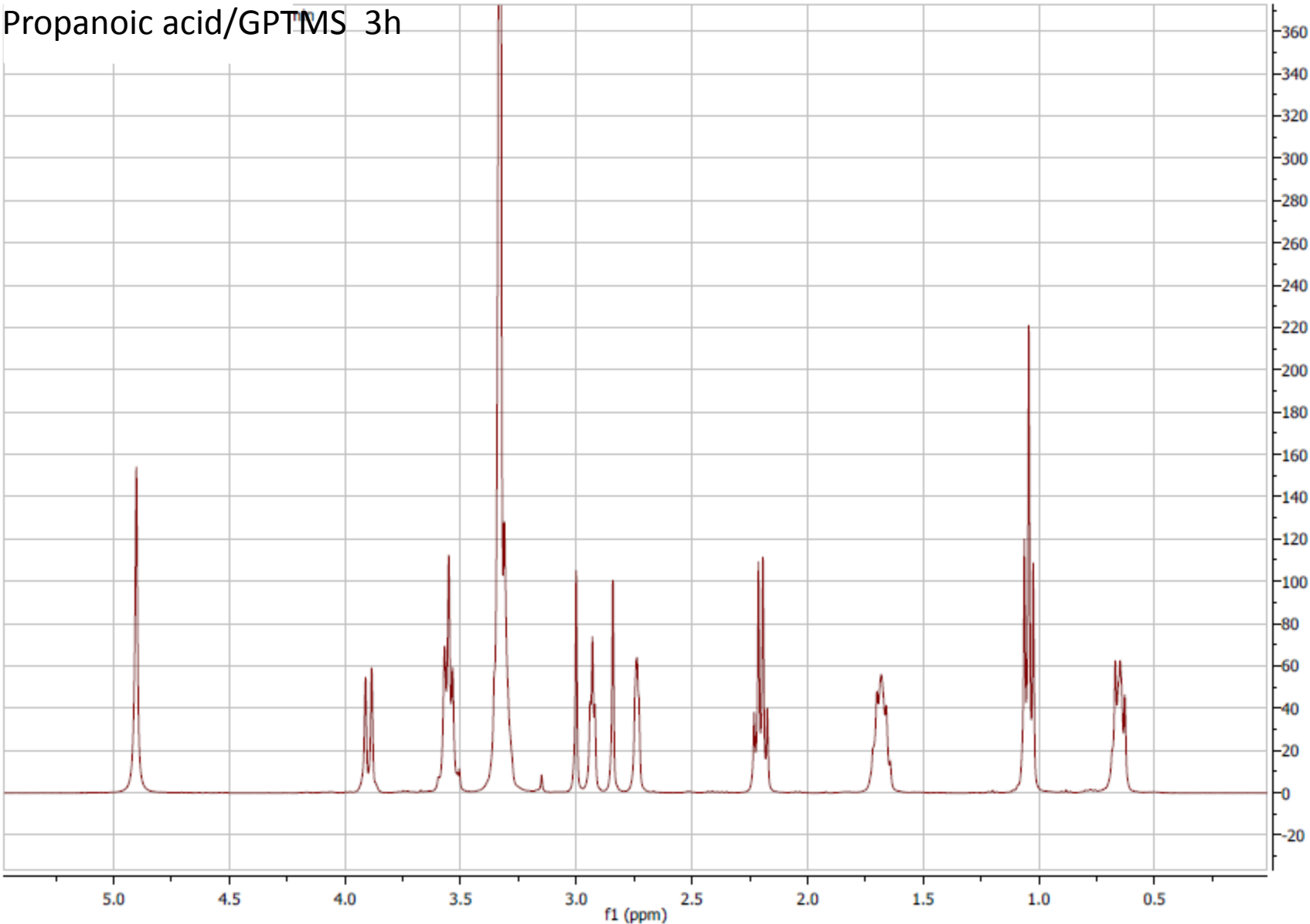
Propanoic acid reaction NMR spectra

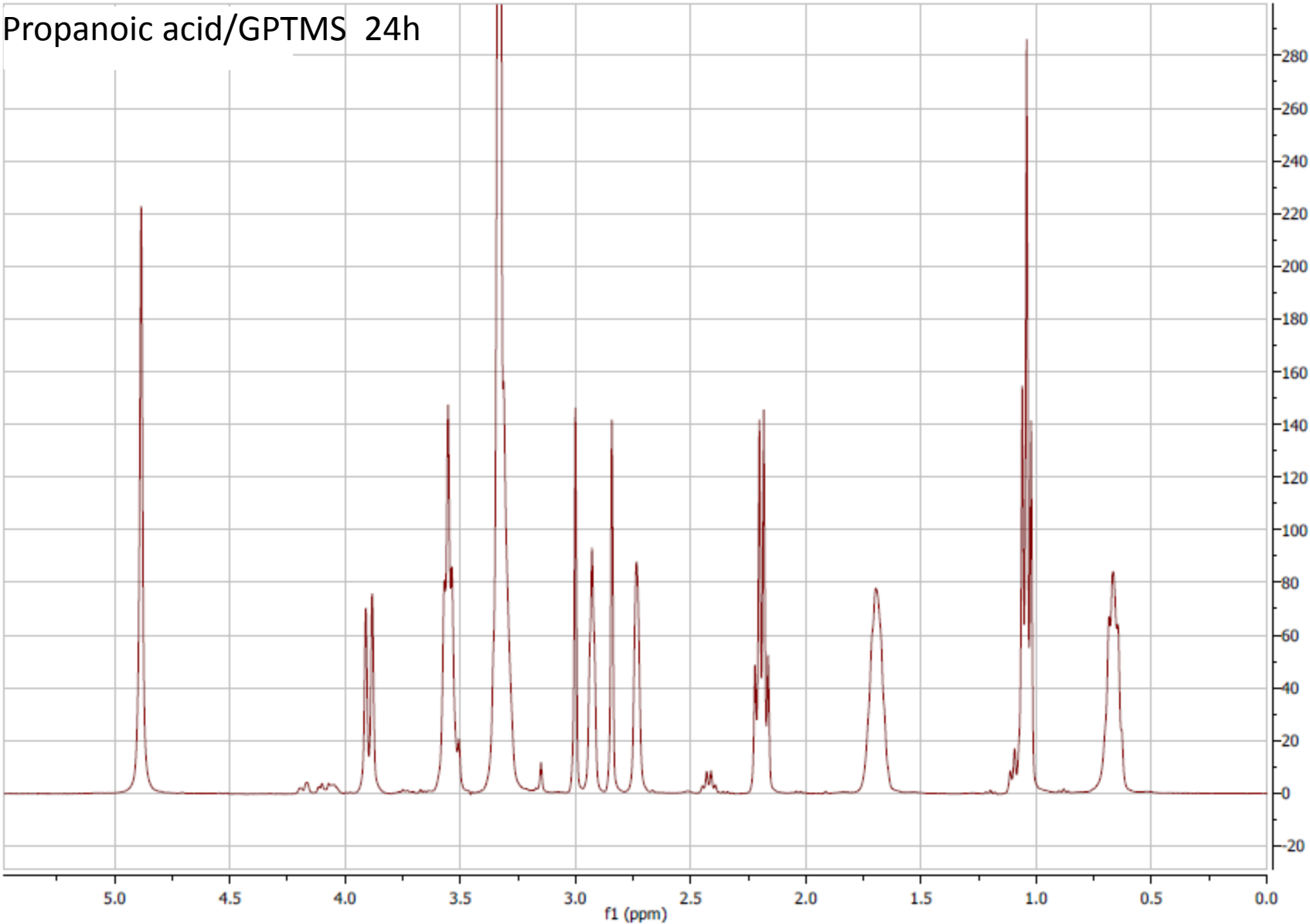


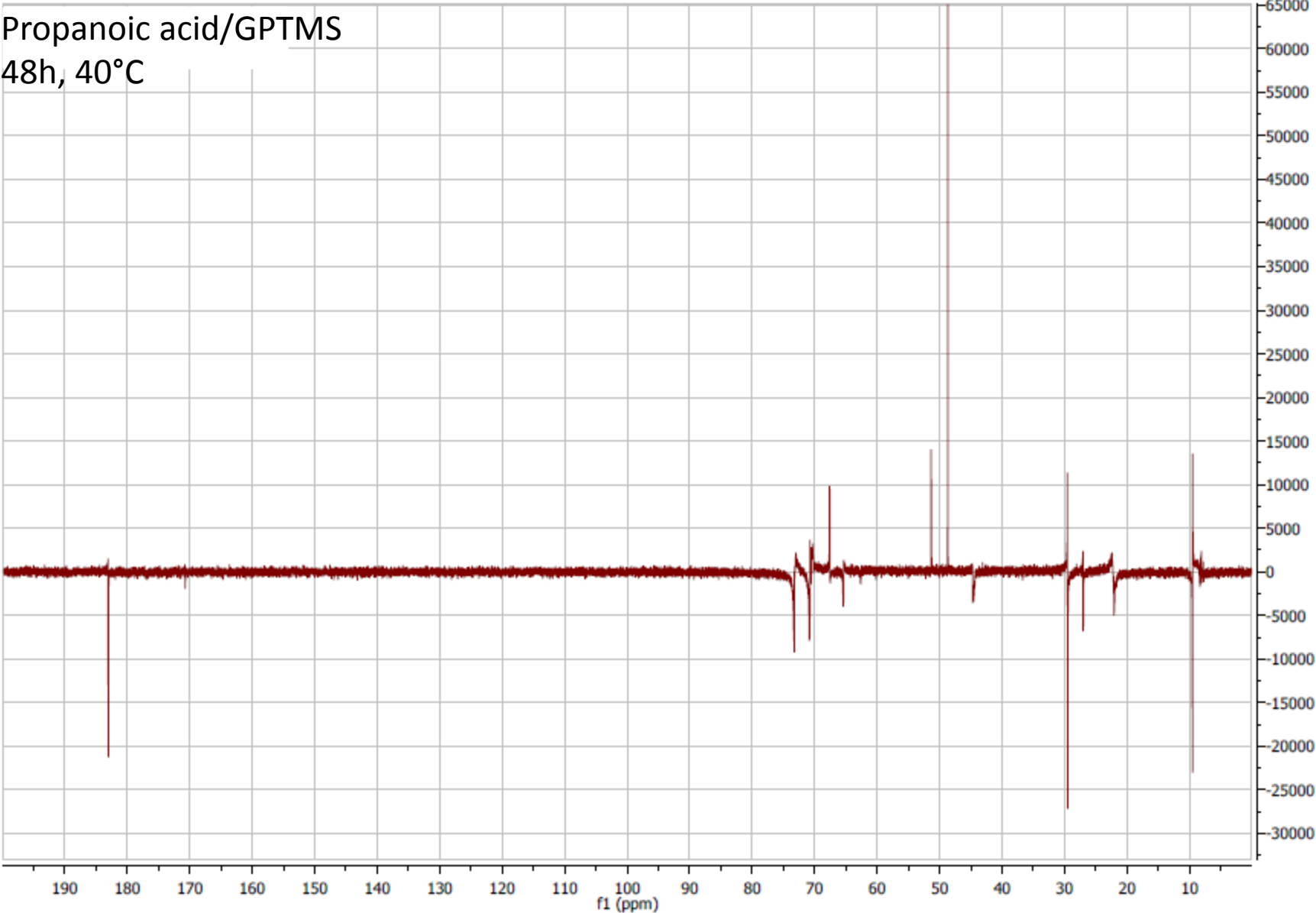




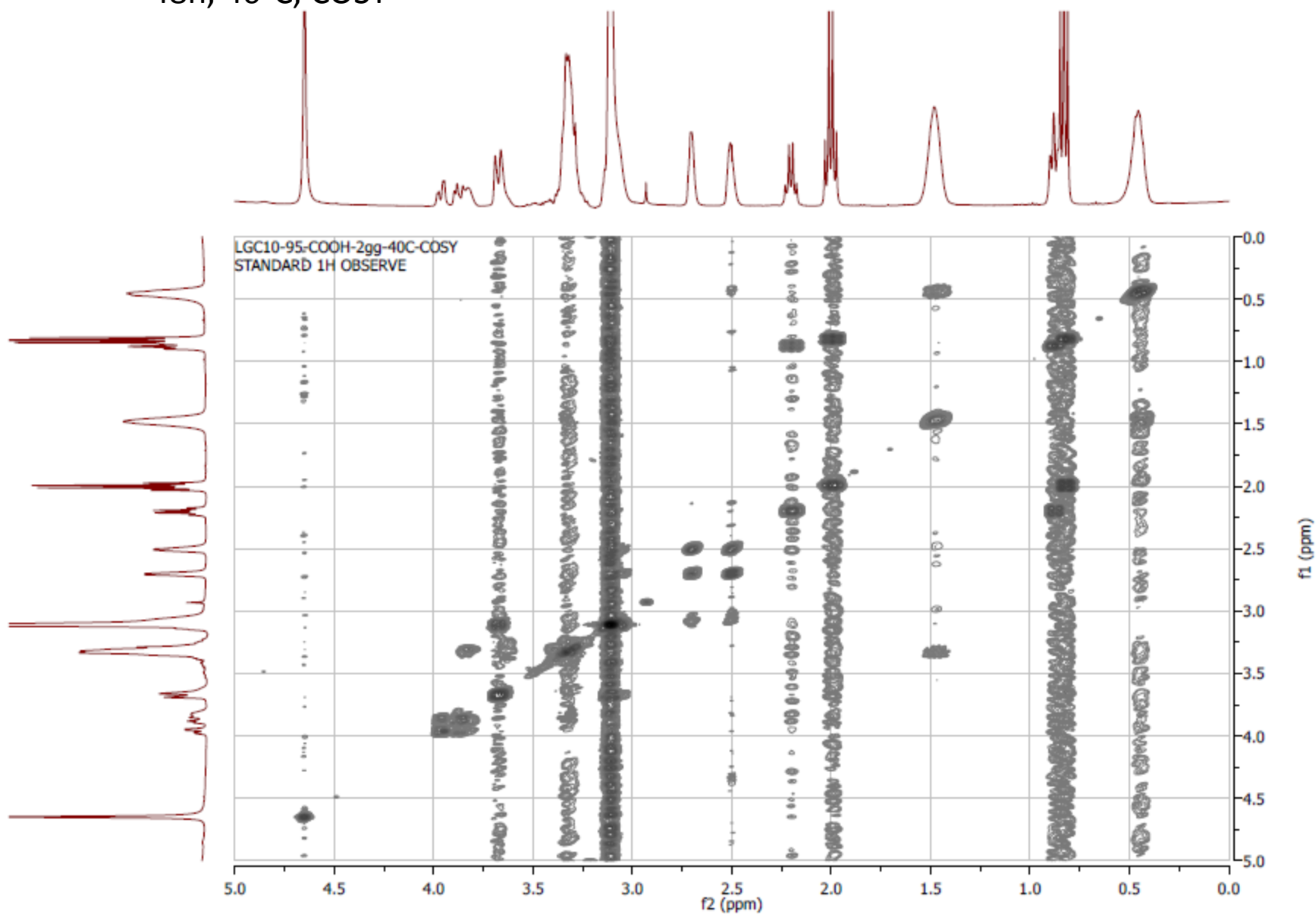
Propanoic acid/GPTMS 3h





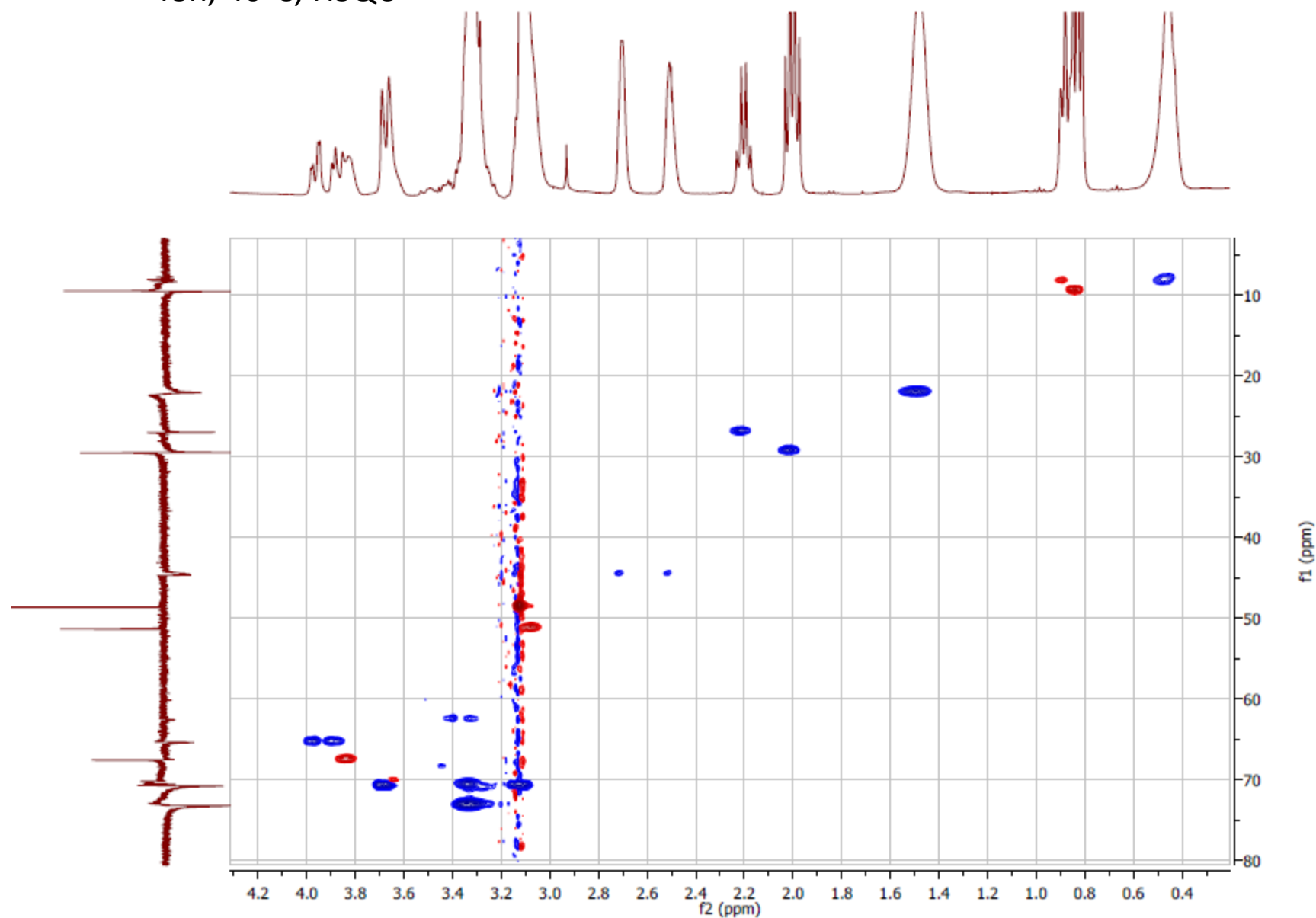


48h, 40°C, COSY



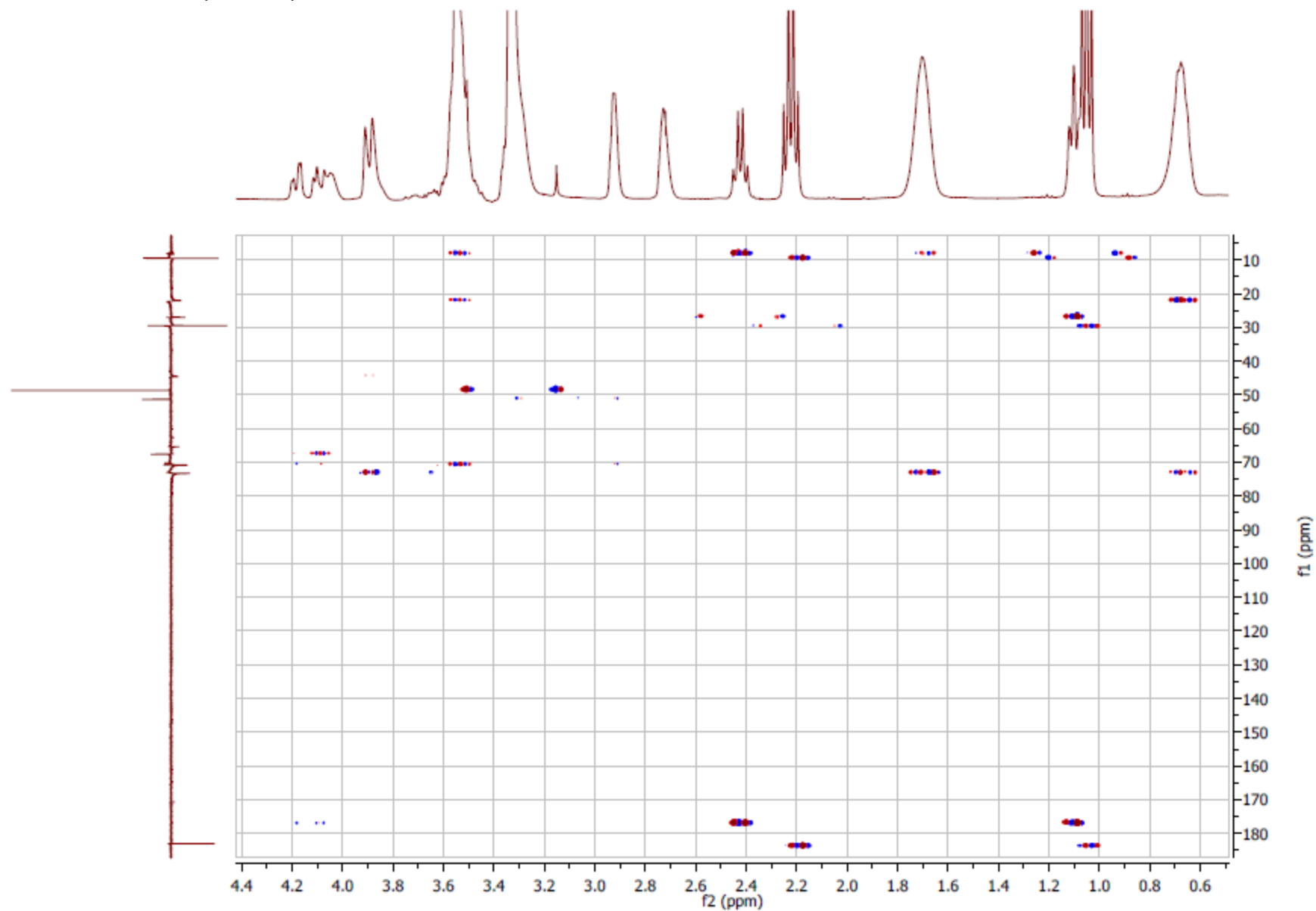
Propanoic acid/GPTMS

48h, 40°C, HSQC



Propanoic acid/GPTMS

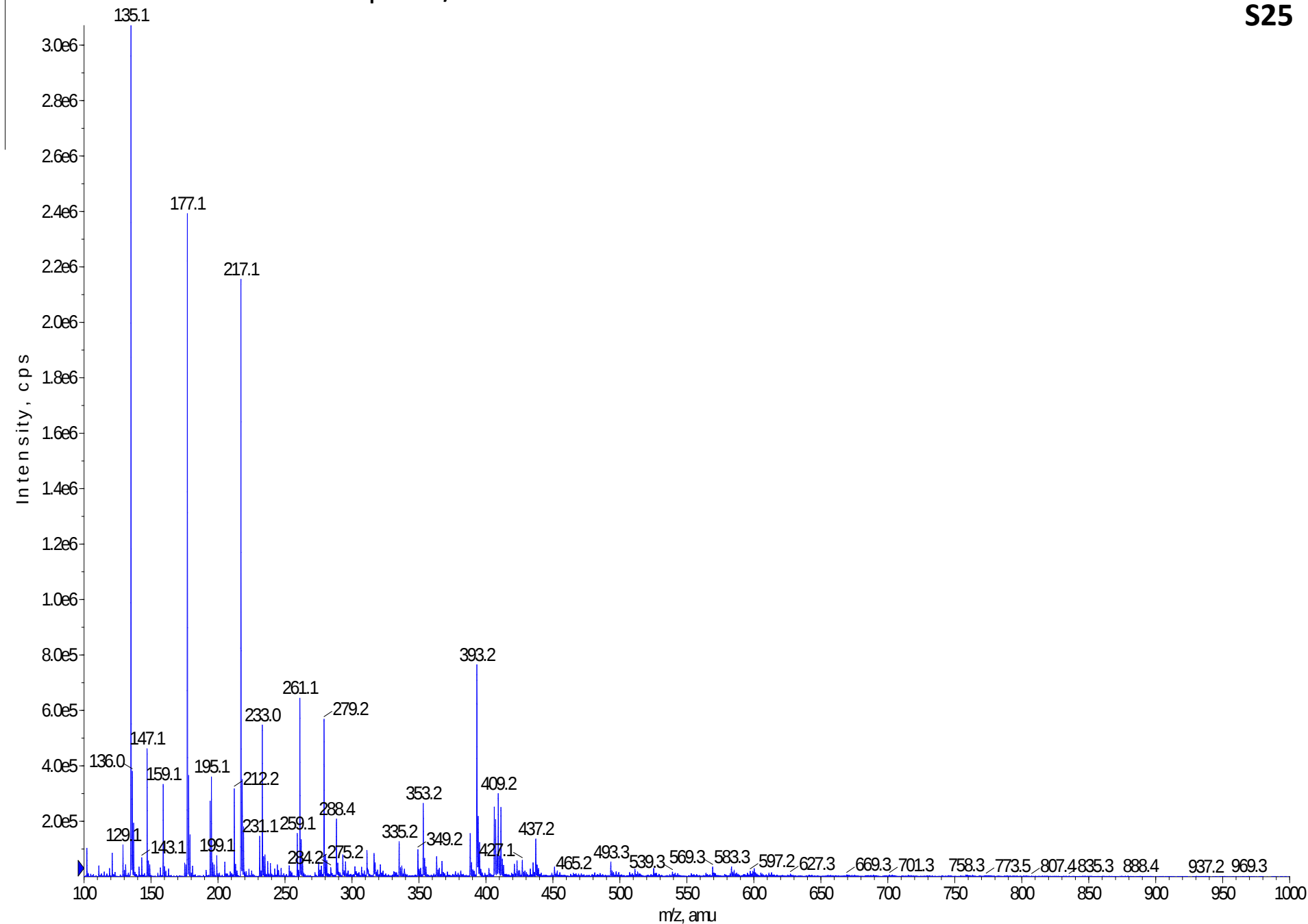
48h, 40°C, HMBC



Propanol reation MS spectra

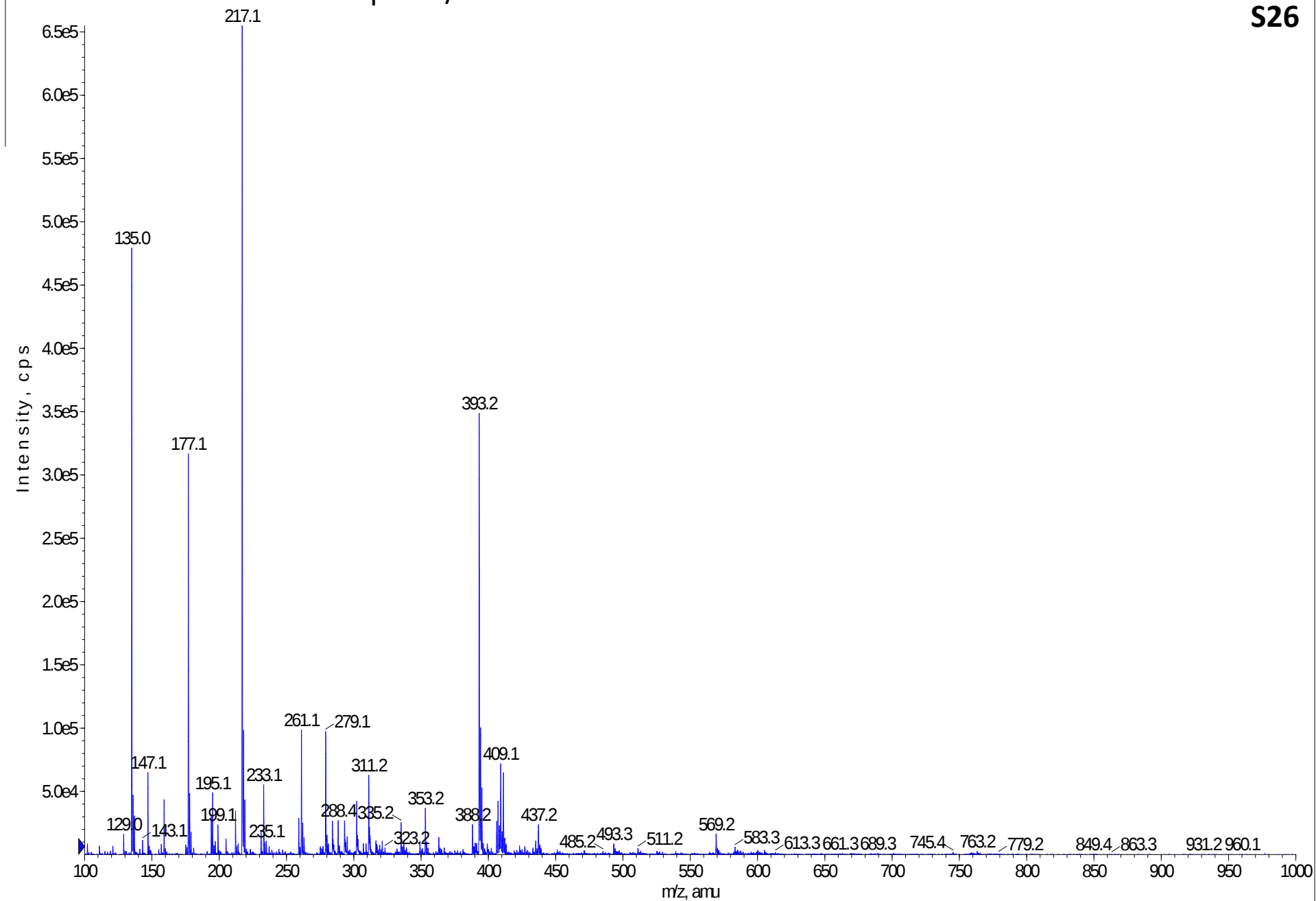
Propanol/GPTMS 0.5h

S25



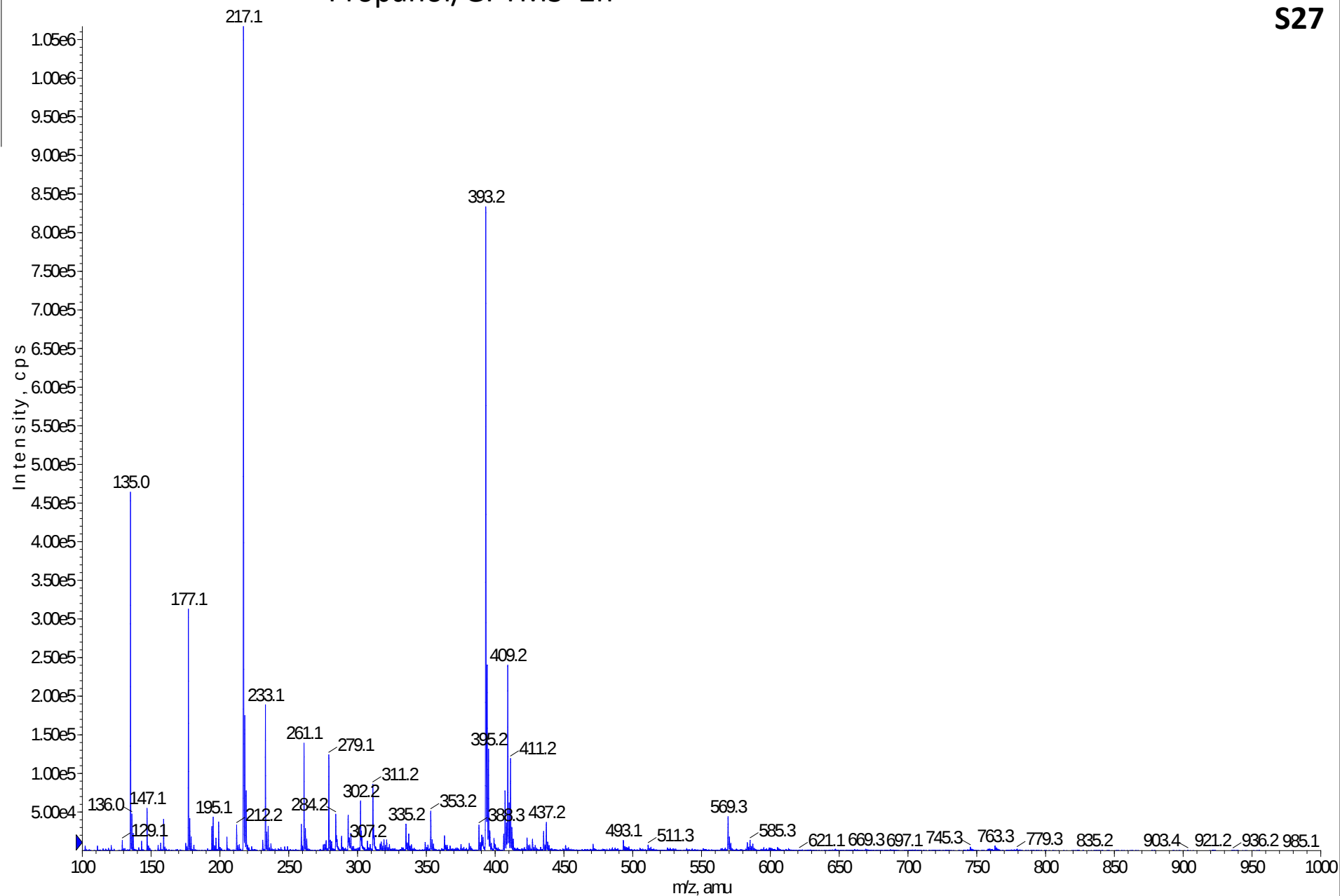
Propanol/GPTMS 1h

S26



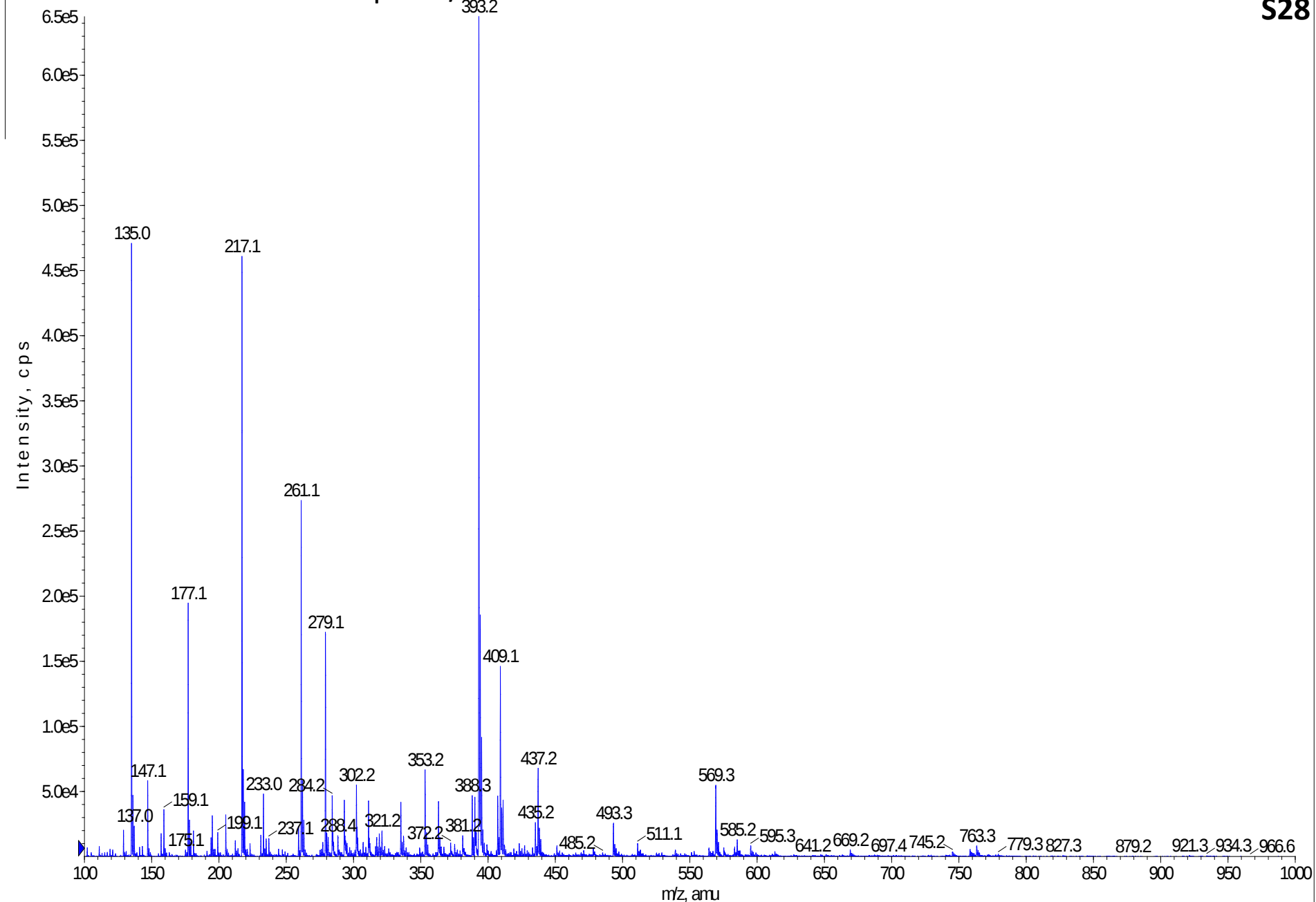
Propanol/GPTMS 2h

S27



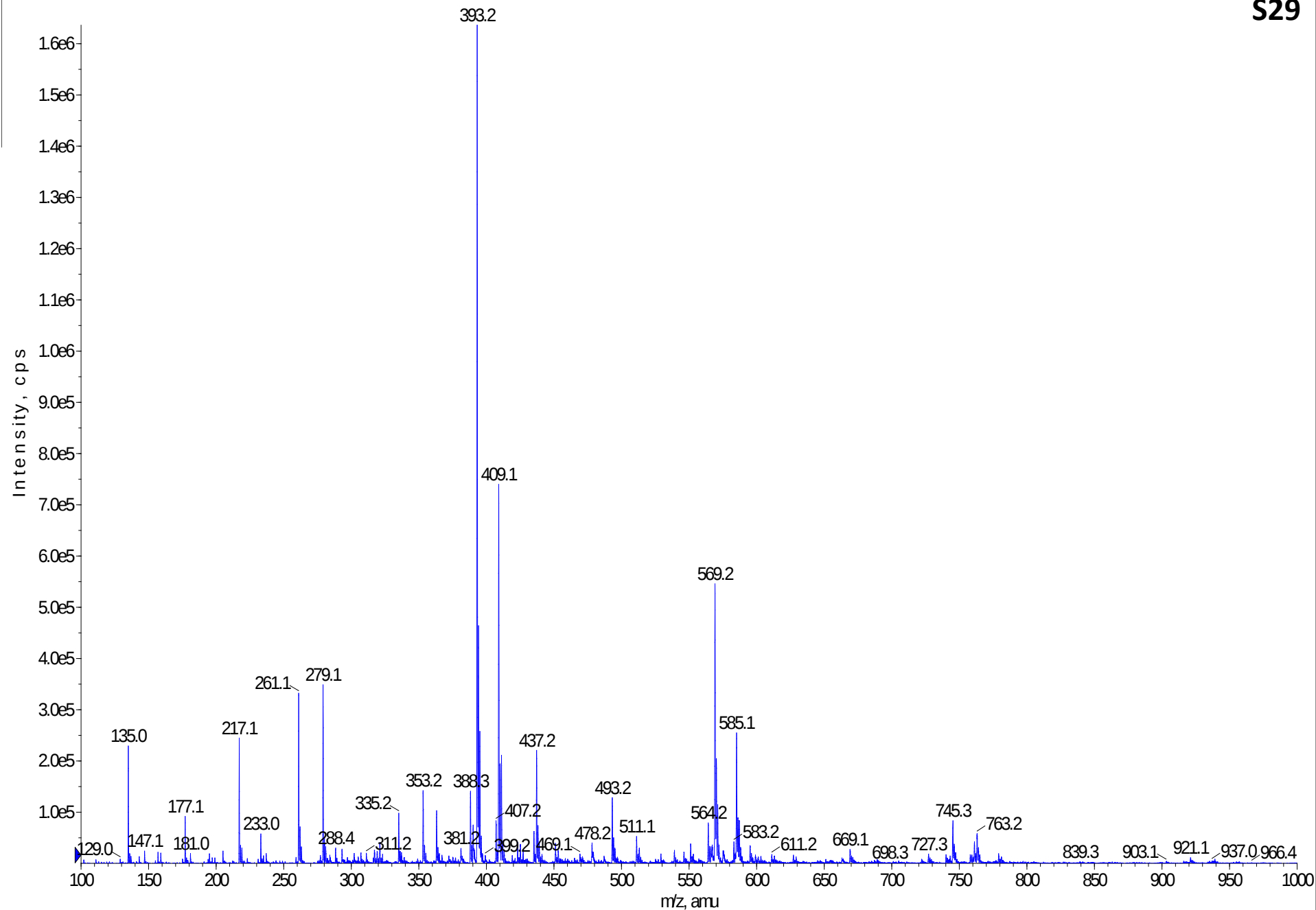
Propanol/GPTMS 3h

S28



Propanol/GPTMS 24h

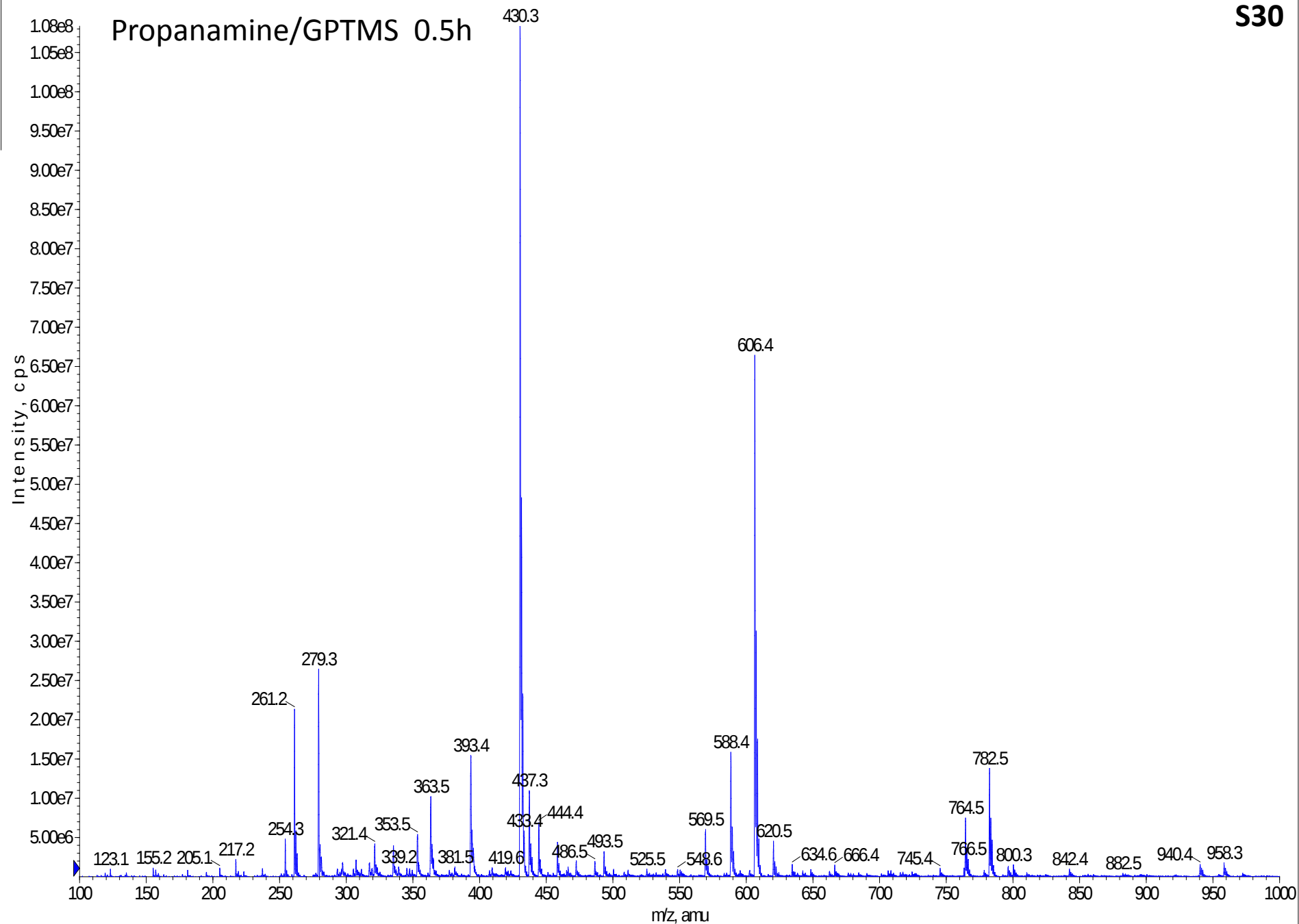
S29



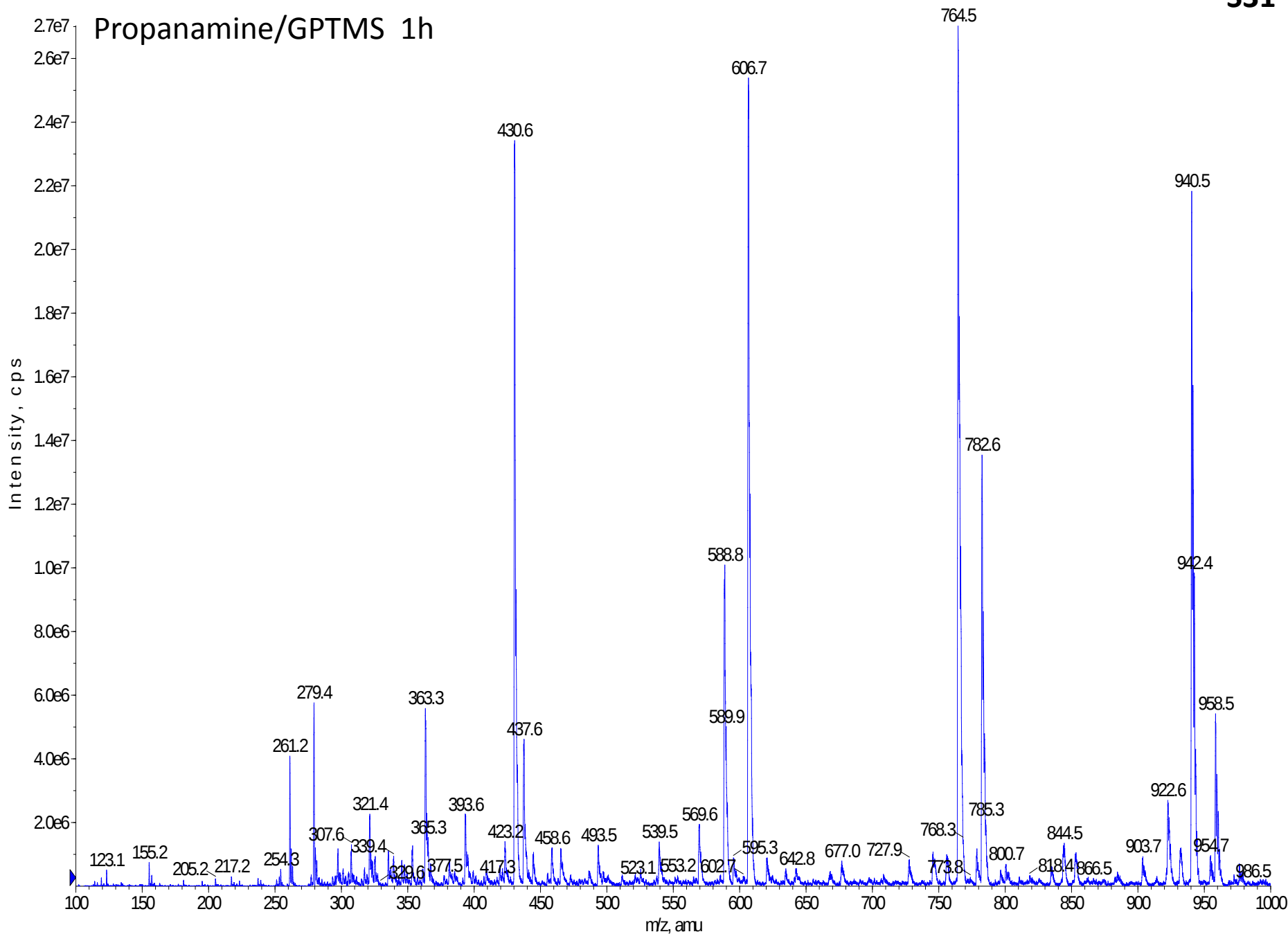
Propanamine reation MS spectra

S30

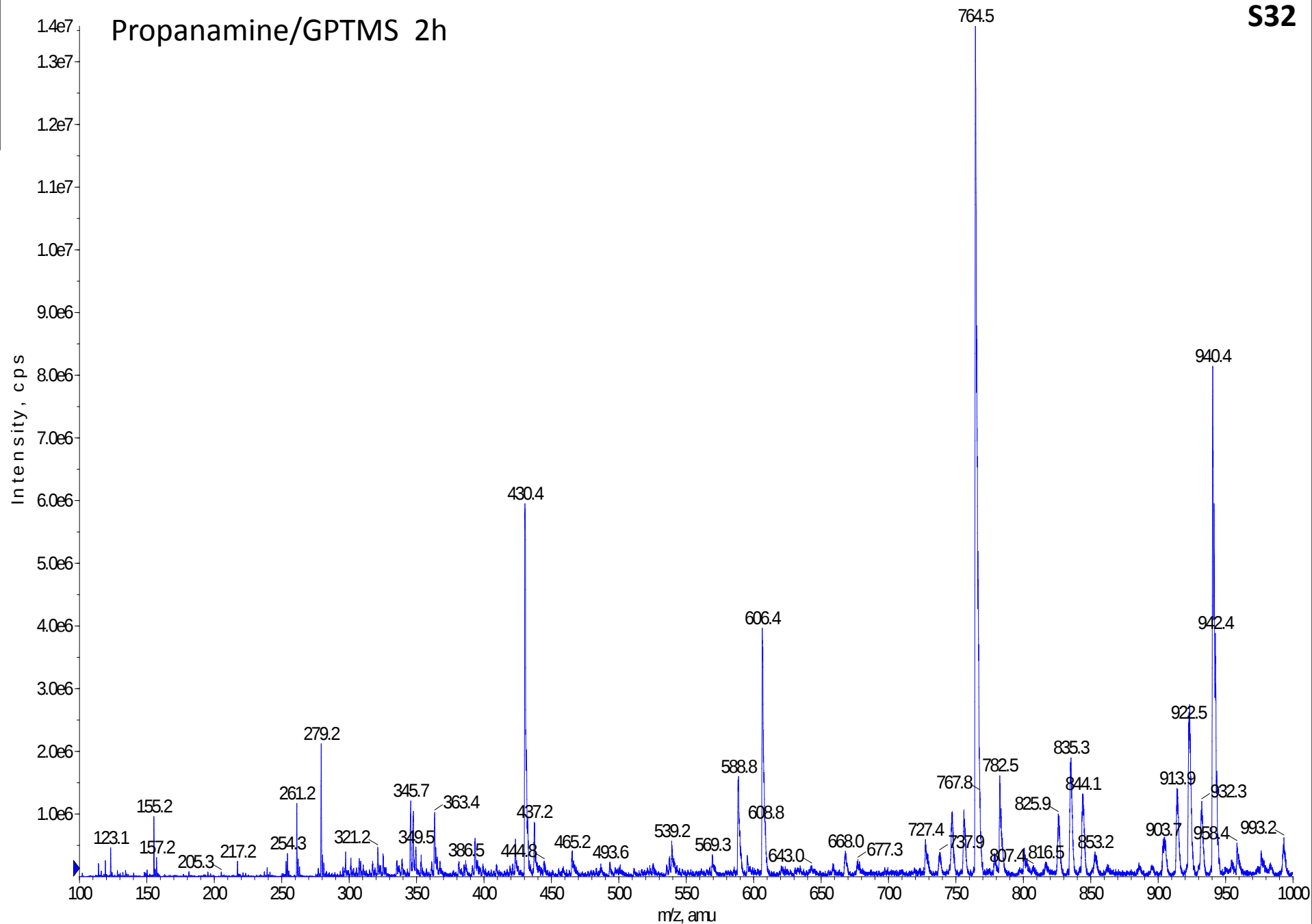
Propanamine/GPTMS 0.5h



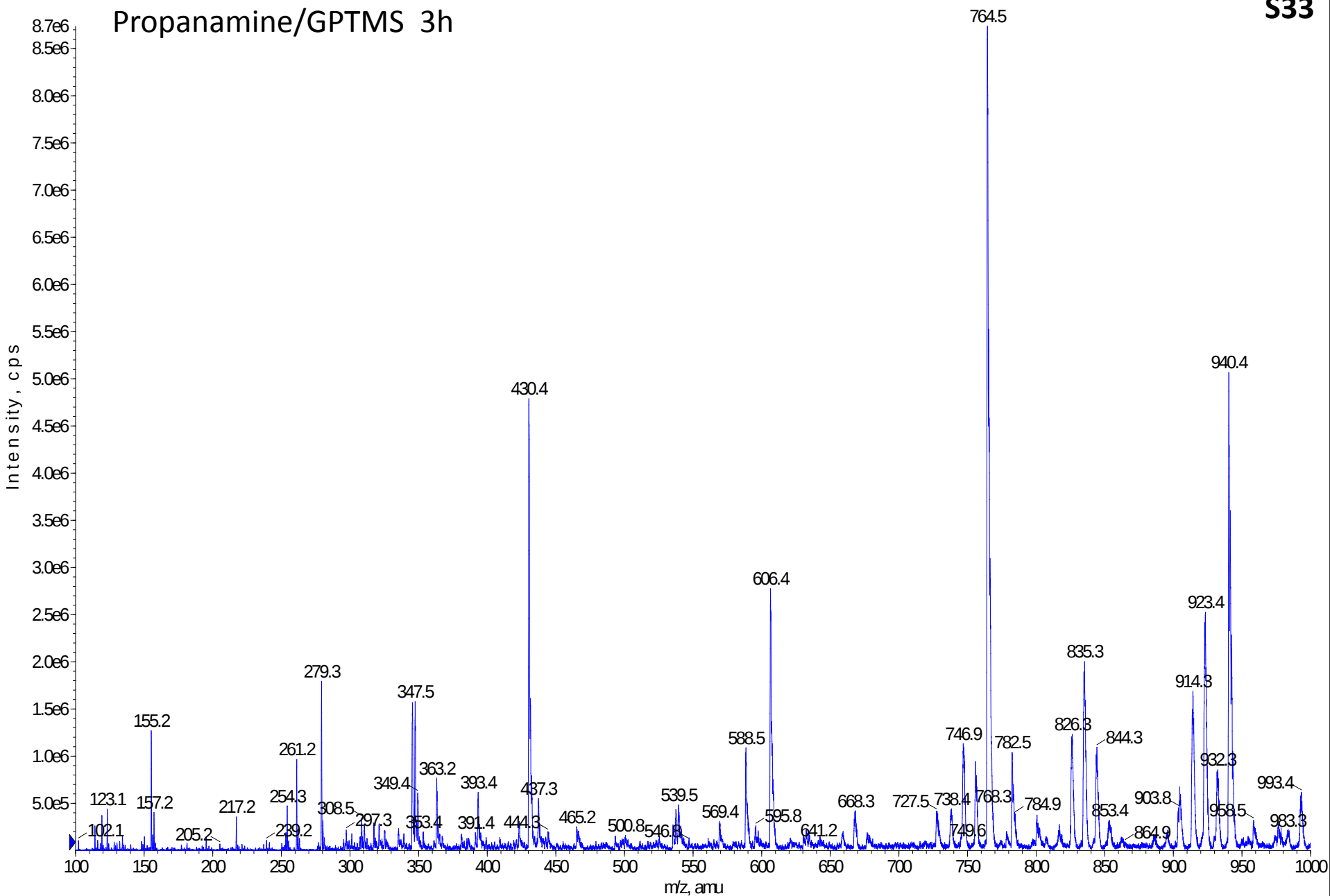
Propanamine/GPTMS 1h



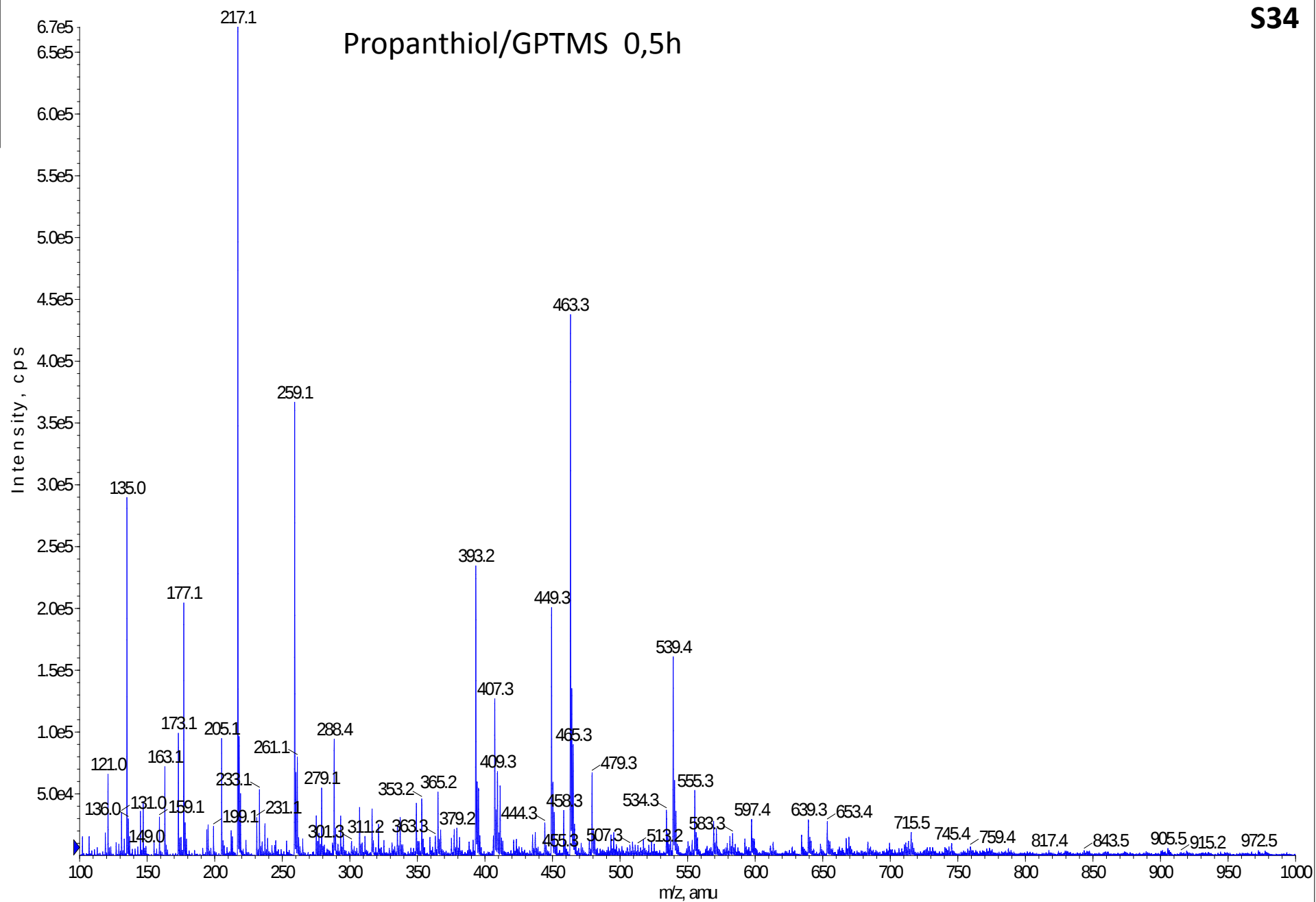
Propanamine/GPTMS 2h



Propanamine/GPTMS 3h

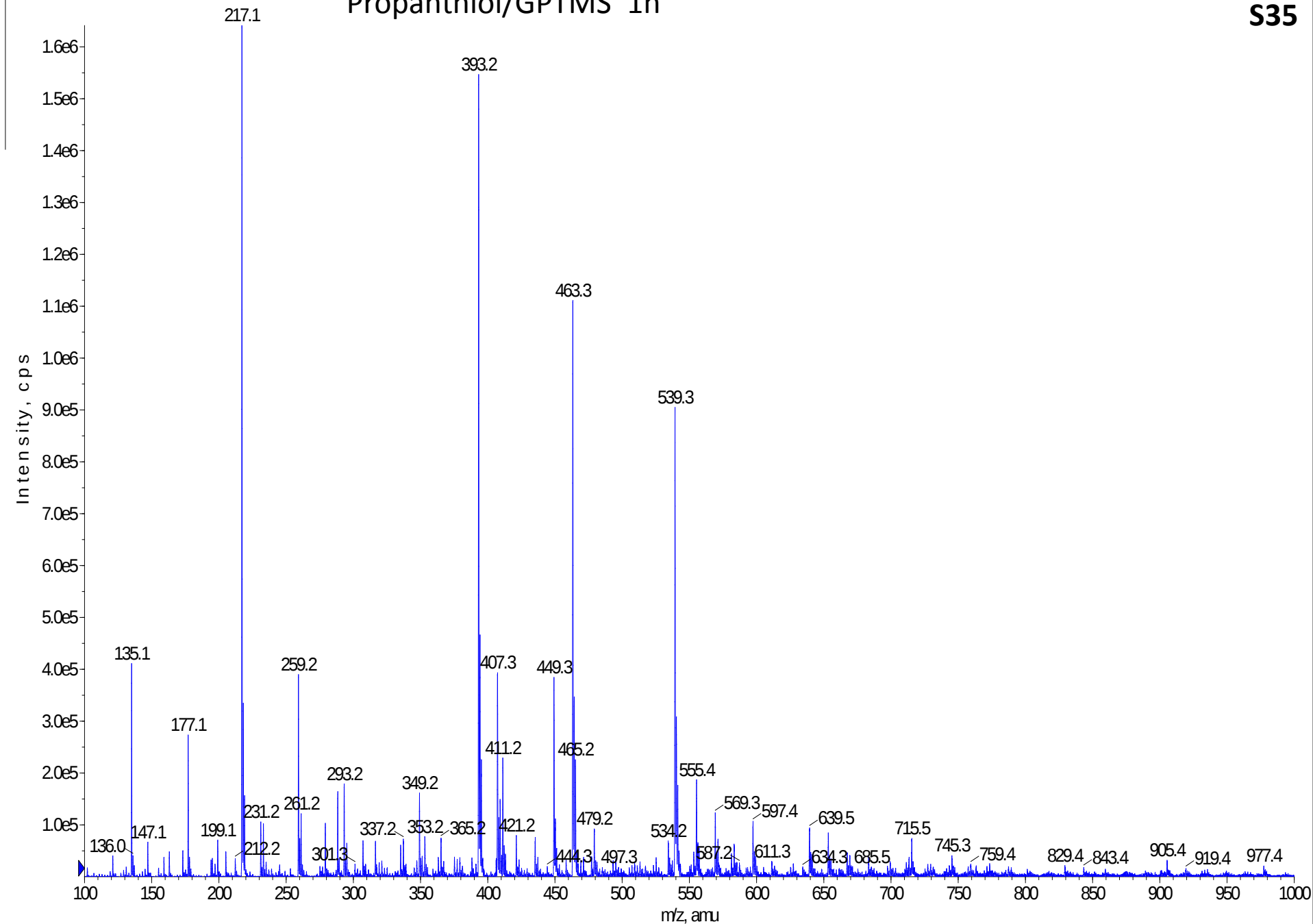


Propanthiol reation MS spectra



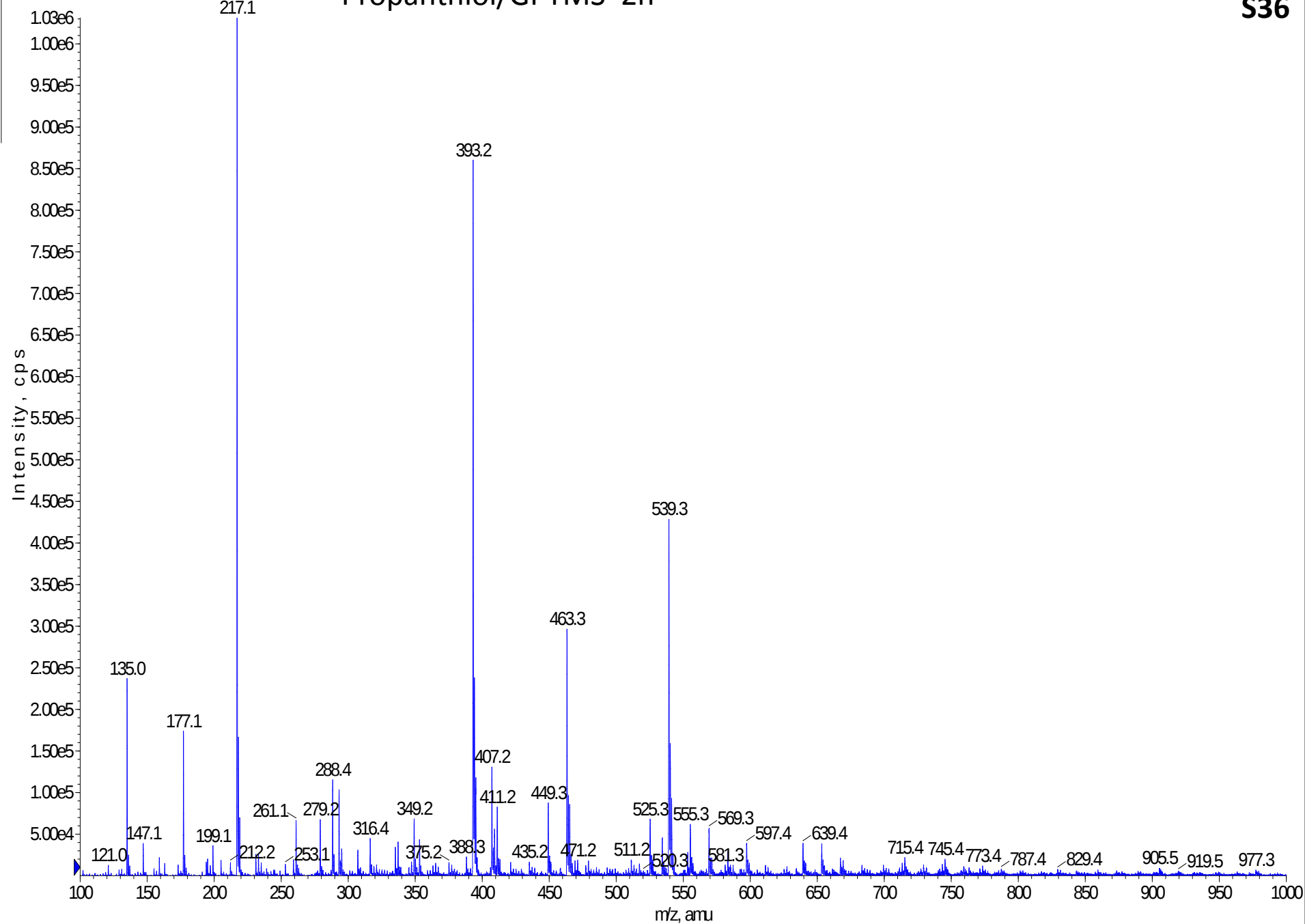
Propanthiol/GPTMS 1h

S35



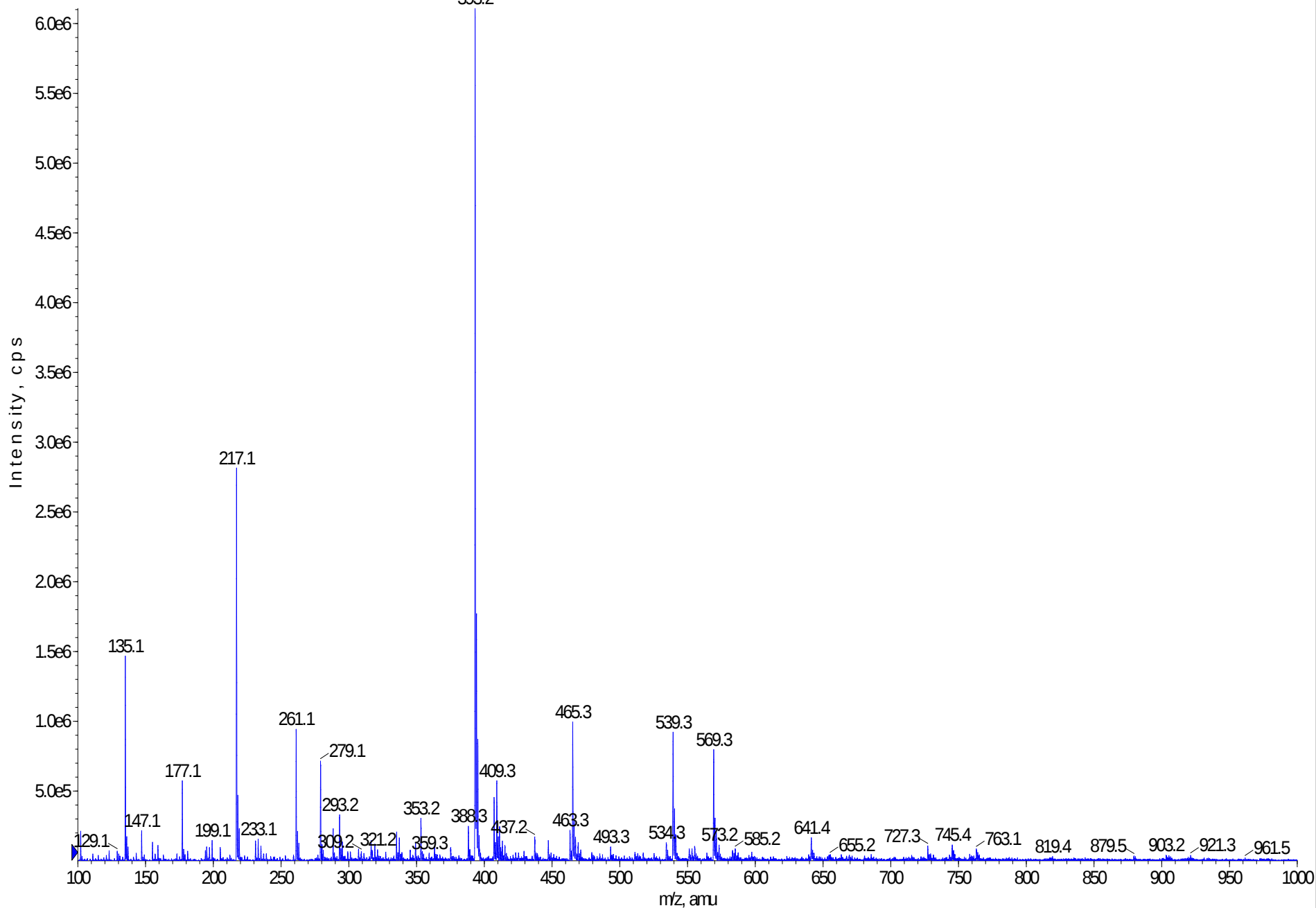
Propanthiol/GPTMS 2h

S36

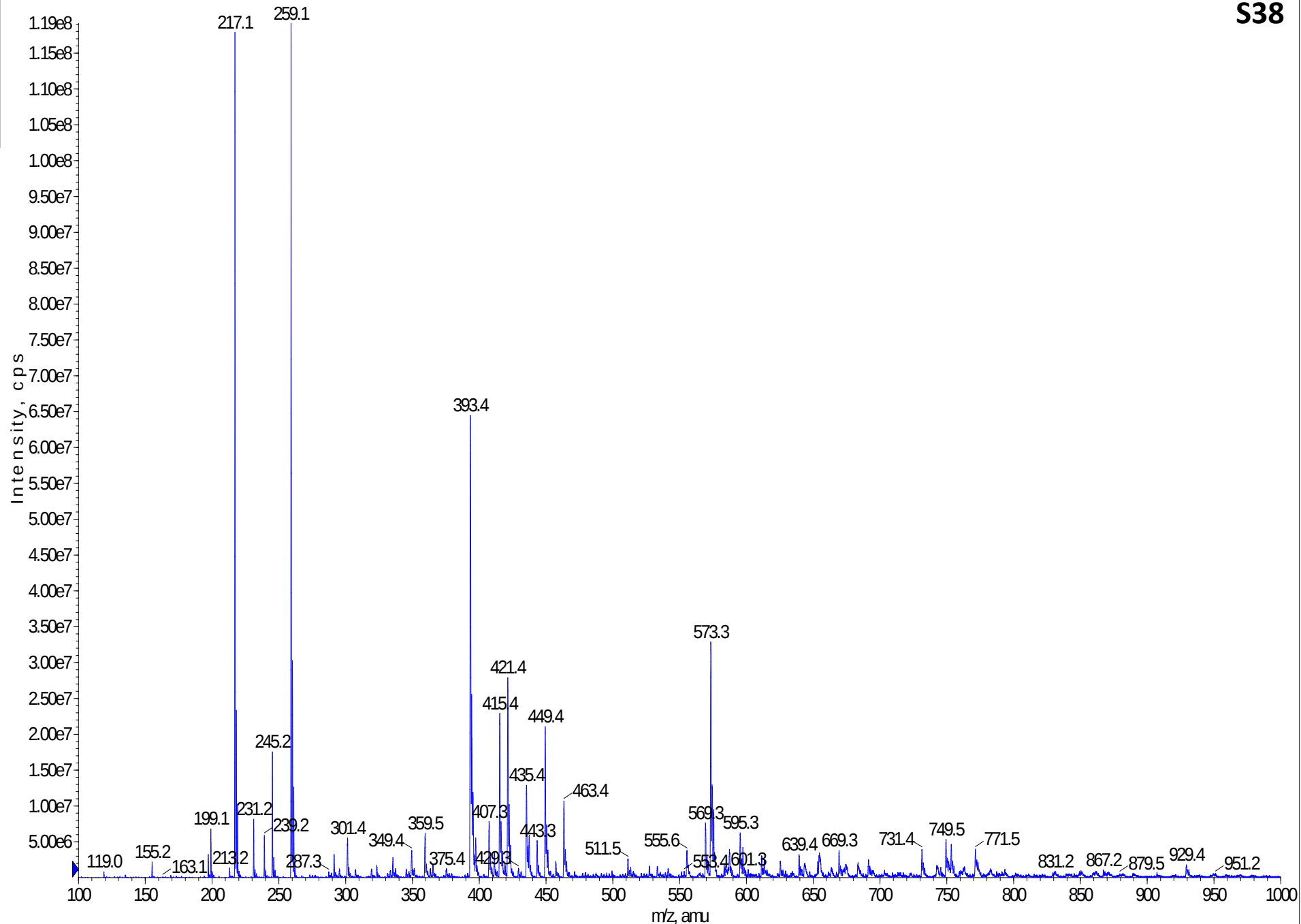


Propanthiol/GPTMS 3h

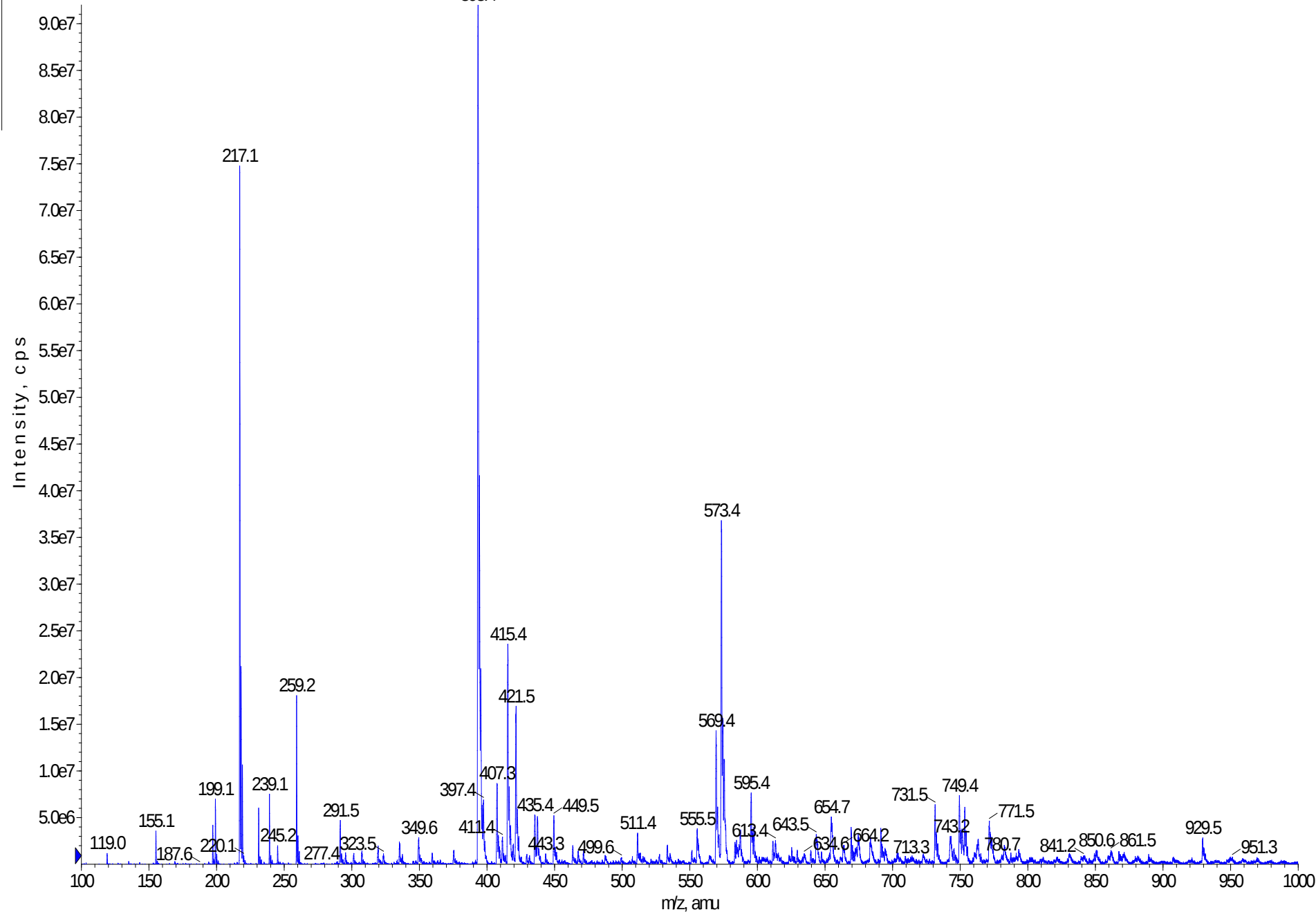
S37



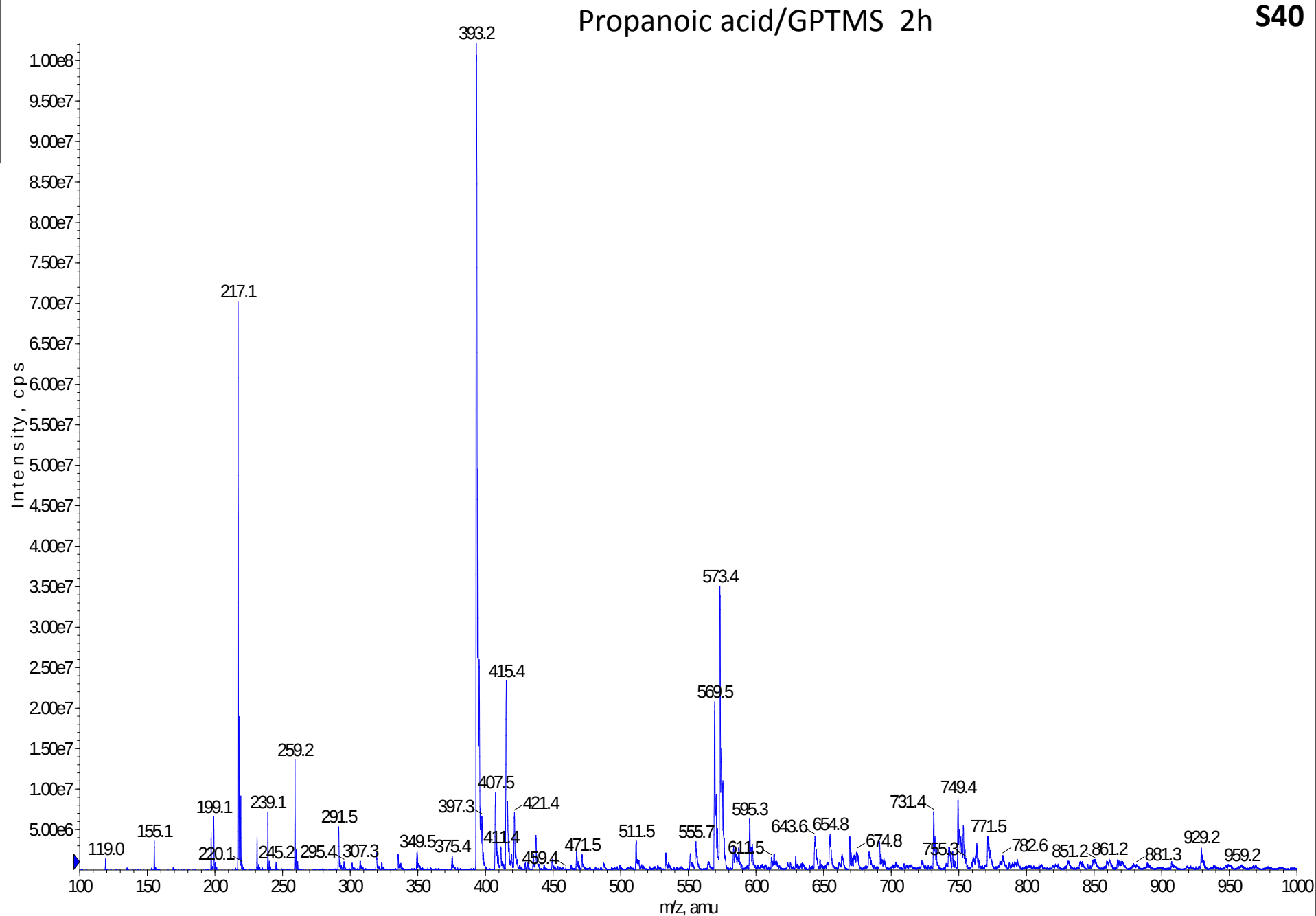
Propanoic acid reaction MS spectra



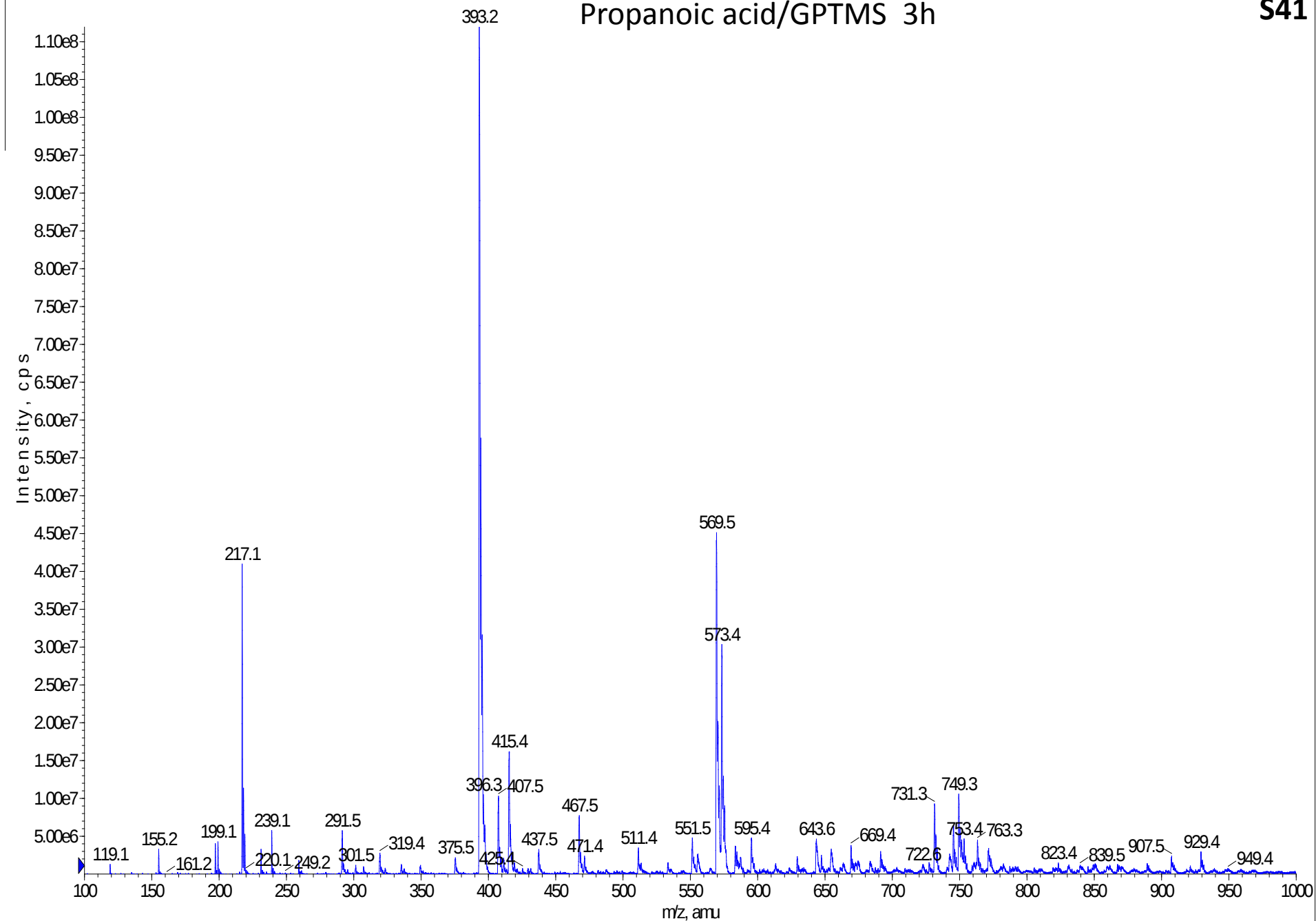
S39

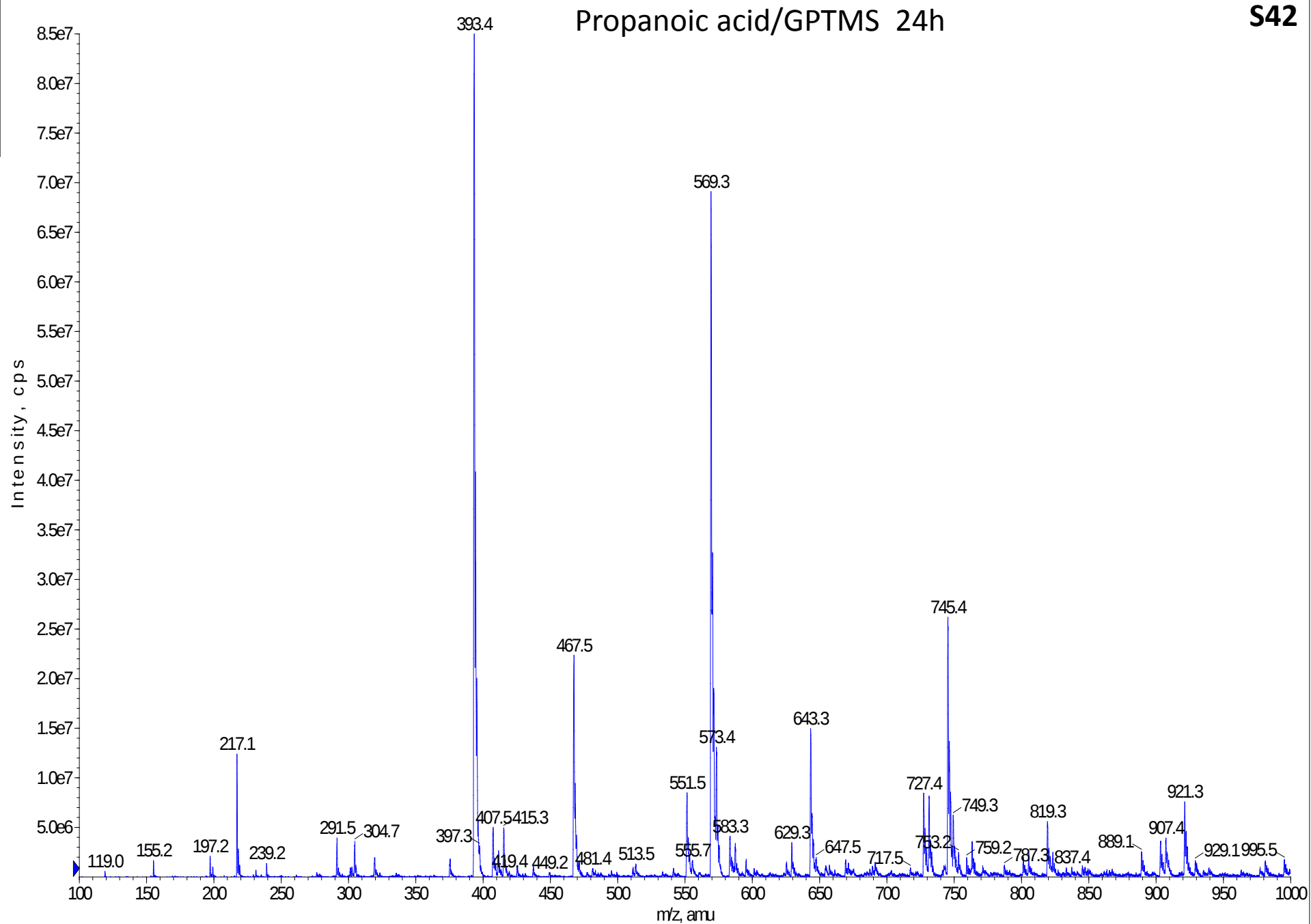


S40



Propanoic acid/GPTMS 3h

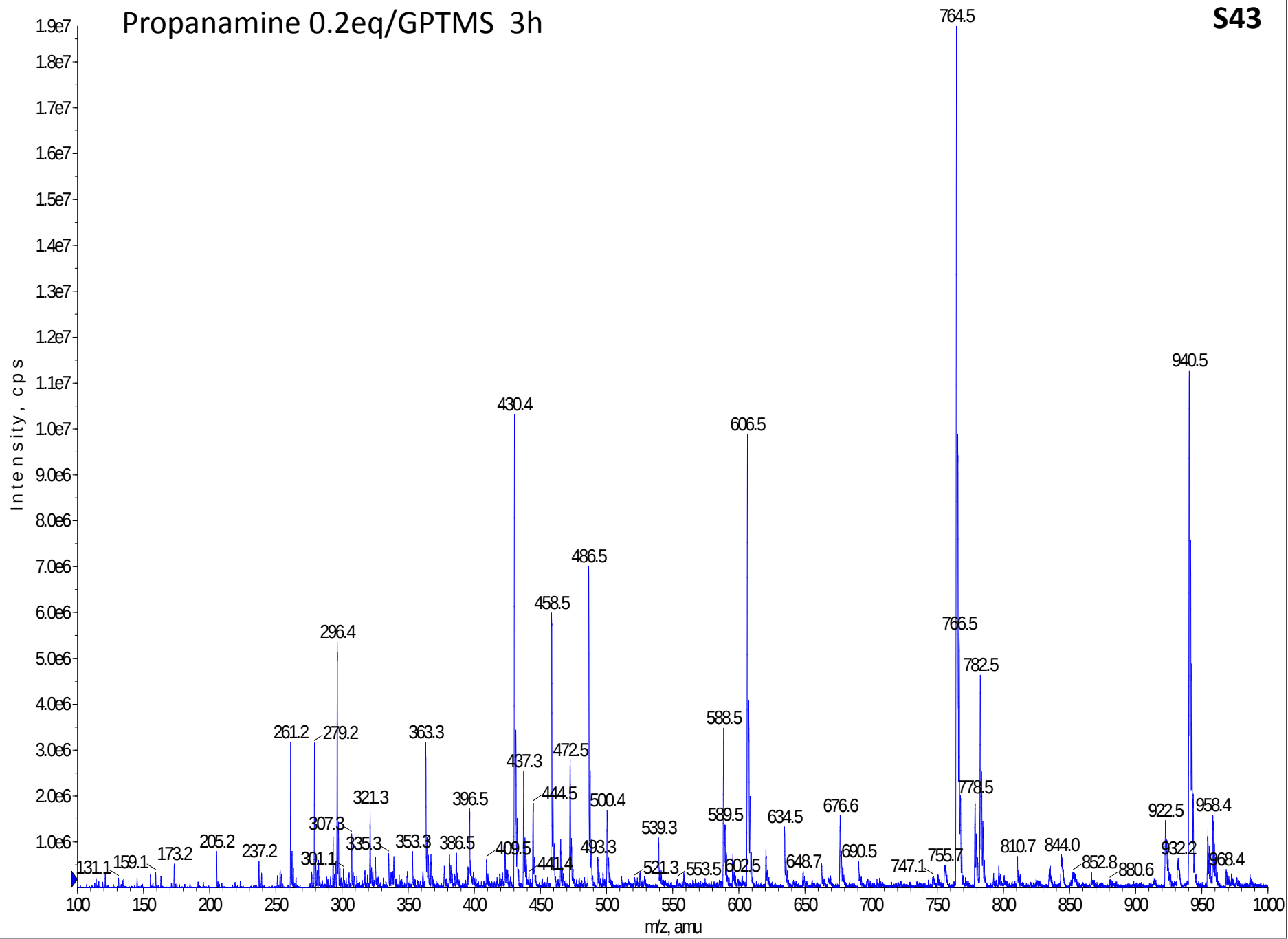




Propanamine 0.2 eq reation MS spectra

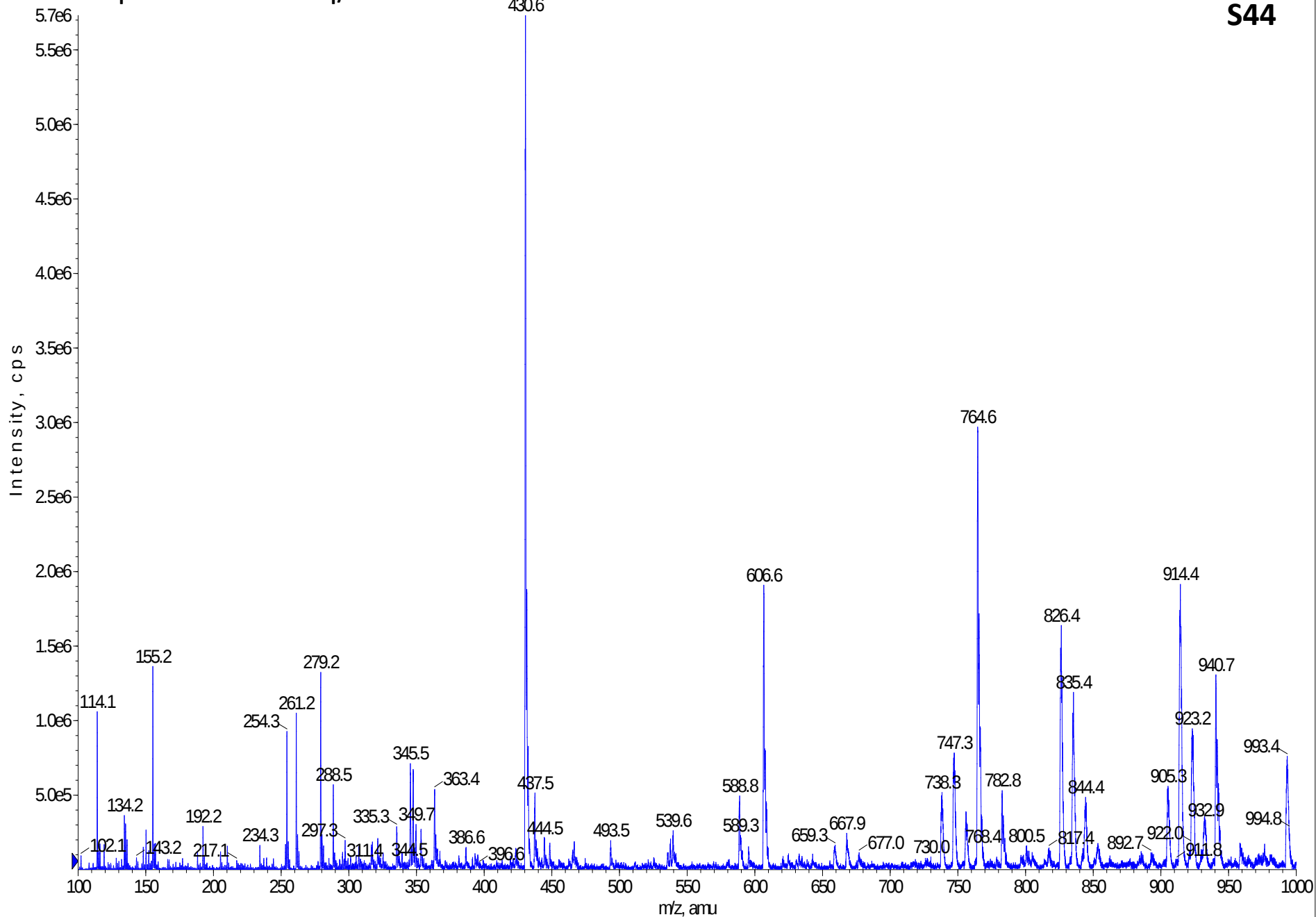
Propanamine 0.2eq/GPTMS 3h

S43



Propanamine 0.2eq/GPTMS 24h

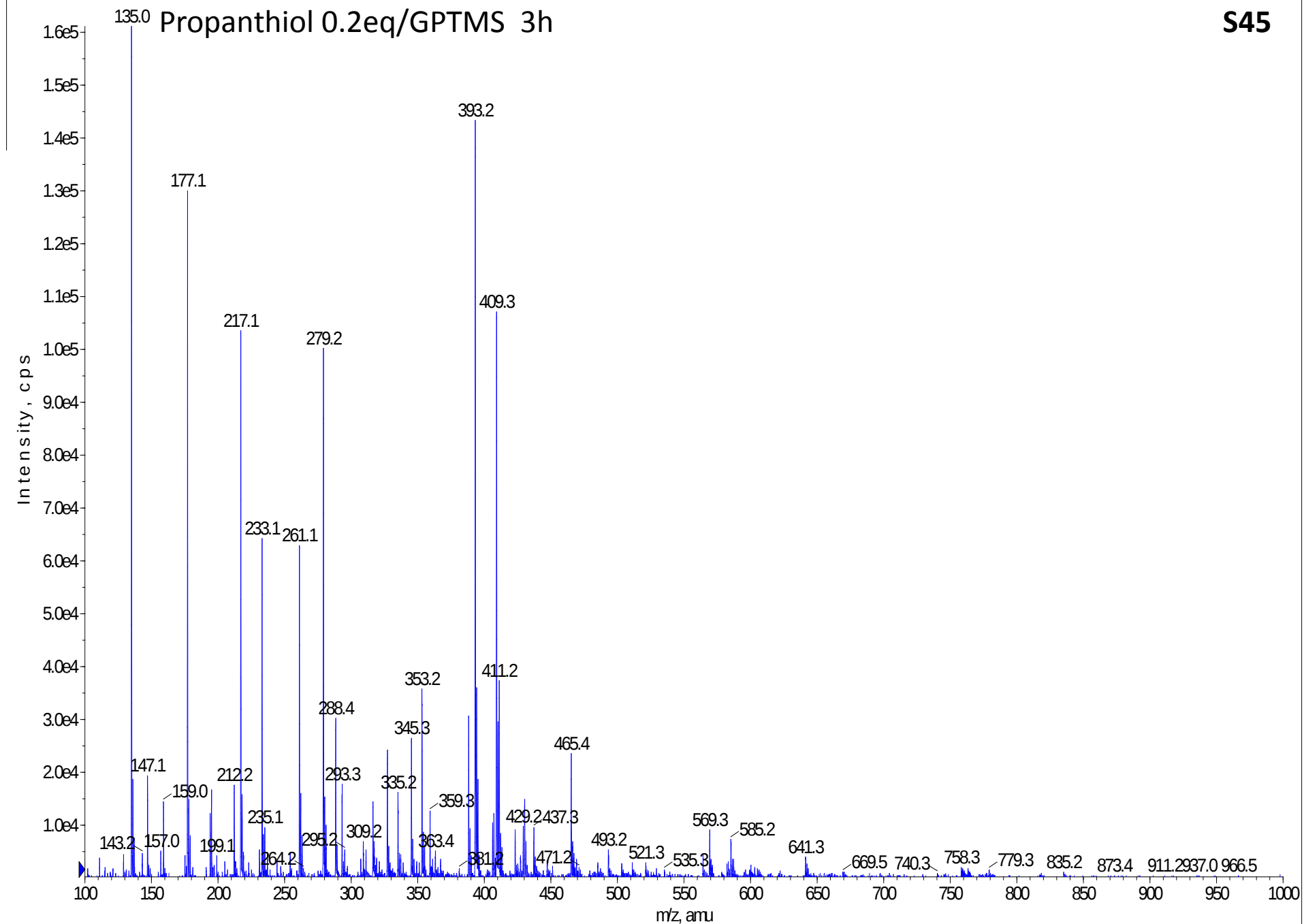
S44



Propanthiol 0.2 eq reation MS spectra

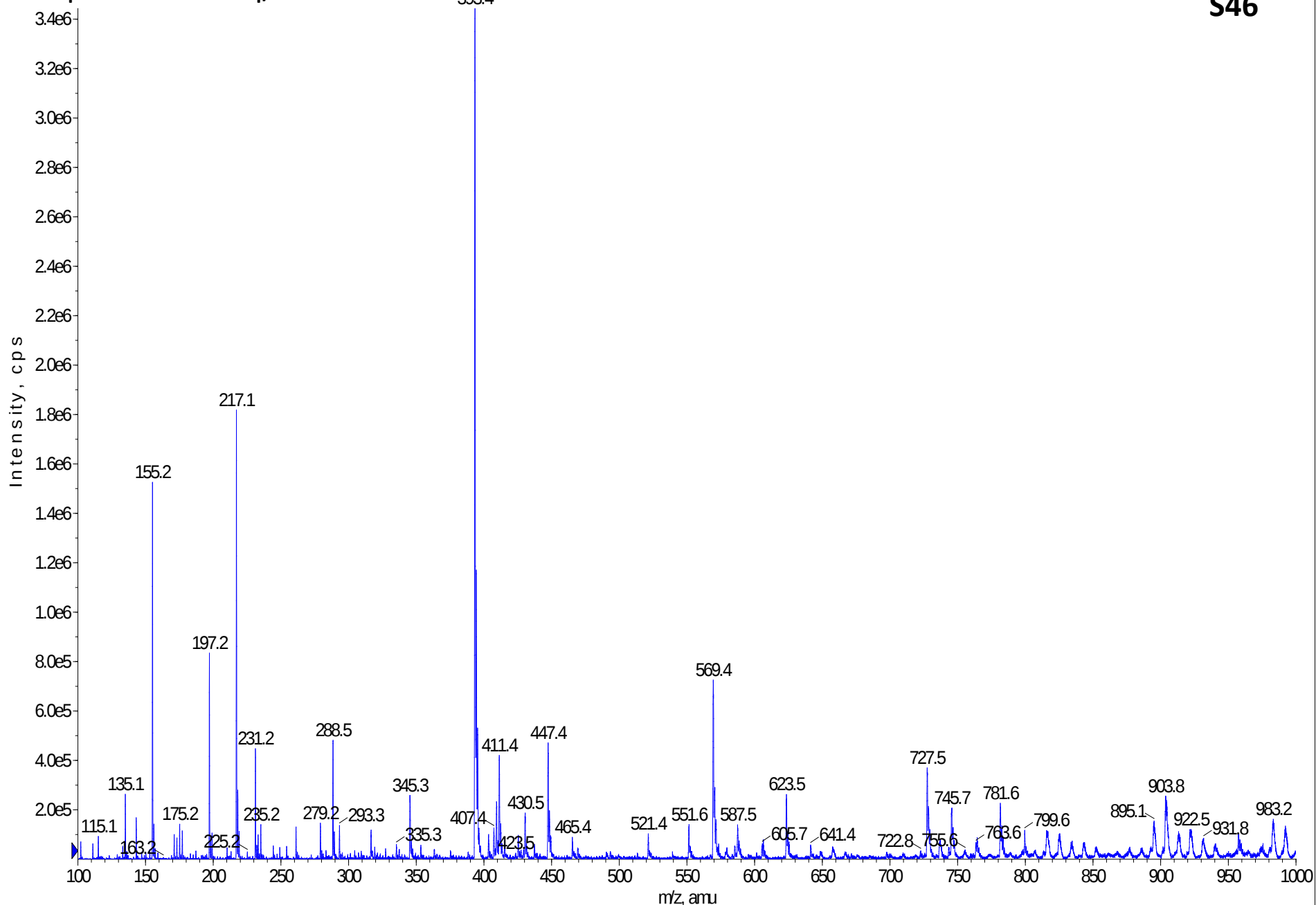
S45

Propanthiol 0.2eq/GPTMS 3h



Propanthiol 0.2eq/GPTMS 24h

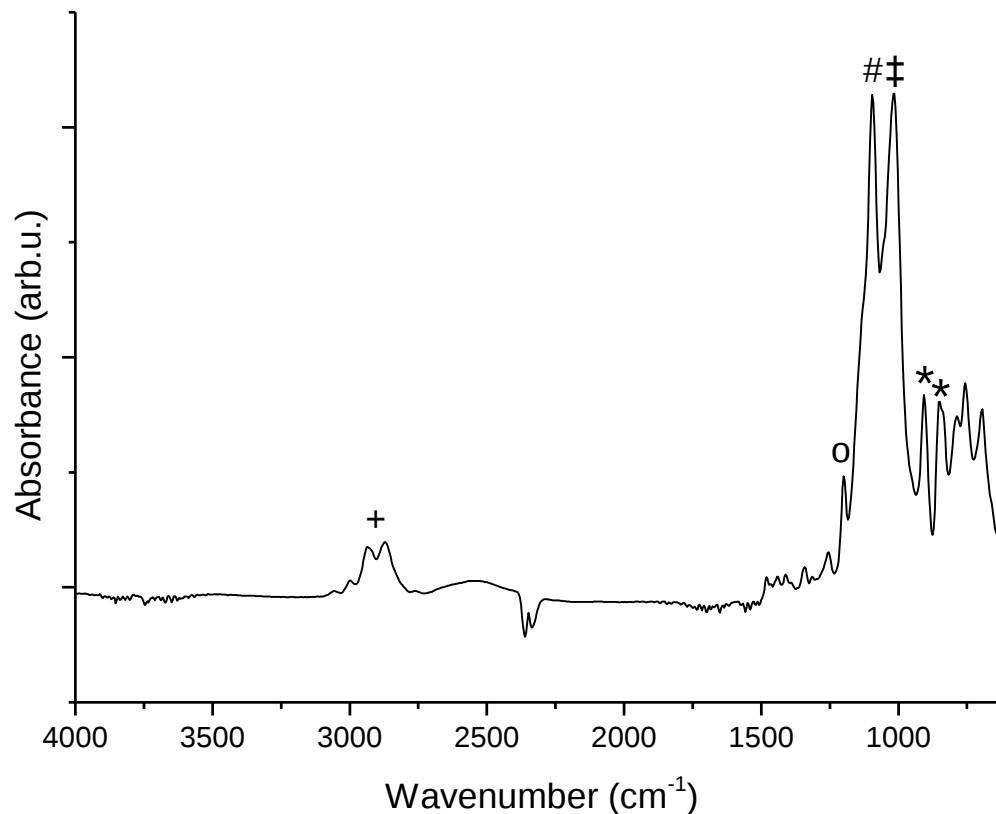
S46



FT-IR spectra of precipitated solids

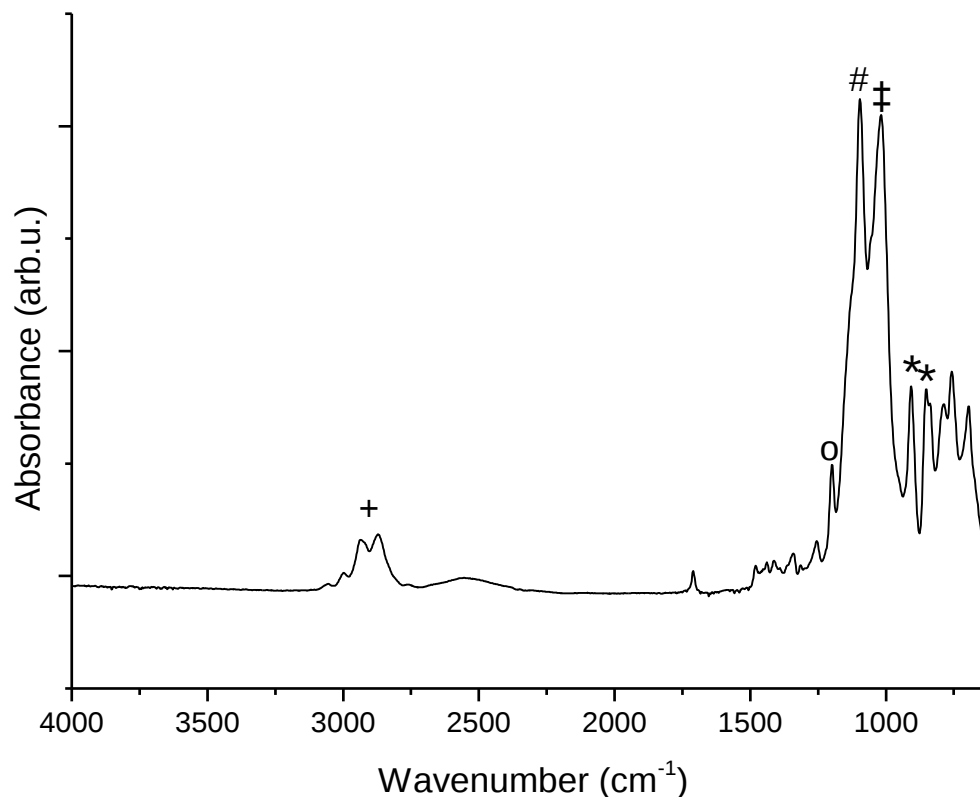
Experimental

Precipitated solids were analyzed by FT-IR using a Thermo Scientific Nicolet iN10 instrument operating in ATR mode and operating with a liquid nitrogen cooled MCT detector. Samples were ground to a fine powder and placed on a gold substrate. Spectra were recorded between wavenumbers of 4000-600 cm^{-1} with a resolution of 4 cm^{-1} .



FT-IR spectra of precipitated solids resulting from GPTMS/propylamine reaction after 24h. +: C-H alkanes, o: -Si-CH₂- #: Si-O-Si ‡: Si-OH, *: epoxide ring.

FT-IR analysis of the gel-like precipitated solids confirmed the presence of GPTMS species. The double band due to the Si-O-Si symmetric stretch of the silicate network (1094 cm^{-1}) and a surprisingly high level of non-bridging oxygens, Si-OH (1016 cm^{-1}) which shows that the silicate network is only partially condensed. Unopened epoxide ring species were observed at 904 cm^{-1} and 846 cm^{-1} providing further evidence that the epoxide ring remains unreacted within this timescale. There was no observable bands corresponding to amine groups (typically expected around 3400-3500 cm^{-1} for N-H stretch, 1550-1650 cm^{-1} for N-H bending of primary amines) confirming that the propylamine primarily remains in solution as observed by ^1H NMR and does not react with the epoxide of GPTMS.



FT-IR spectra of precipitated solids resulting from GPTMS/propanthiol reaction after 24h. +: C-H alkanes, o: -Si-CH₂- #: Si-O-Si ‡: Si-OH, *: epoxide ring.

The FT-IR spectra of the gel-like precipitate obtained by reaction between GPTMS and propanthiol (Figure 8) showed characteristic bands from a partially condensed silicate network. The Si-O-Si symmetric stretch appeared at 1098 cm⁻¹ with the non-bridging Si-OH band appearing at 1020 cm⁻¹. The unopened epoxide ring was observed at 847 and 909 cm⁻¹, indicating that no reaction had taken place between the thiol and the epoxide ring. Further confirmation of this was that no band attributed to the thiol group could be observed in the spectra.