Bronsted acid catalyzed skeletal rearrangements in polycyclic conjugated boracycles: a

thermal route to a ladder diborole

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1. Experimental Procedures

General Details. Toluene, hexane and tetrahydrofuran (THF) solvents were dried and purified using the Grubbs/Dow purification system¹ and stored in 500 mL thick-walled glass vessels under vacuum over sodium/benzophenone ketal. Pentanes were dried over CH₂Cl₂ and stored over sodium/benzophenone ketal in thick-walled glass vessels under vacuum. Dichloromethane were dried over CH₂Cl₂ and stored over molecular sieves (4A) in thick-walled glass vessels under vacuum. Unless otherwise noted, solvents were introduced into the reaction vessels via vacuum distillation with condensation at -78 °C. Silica gel column chromatography was carried out on Silia-P Flash Silica Gel (particle size 40 - 63 µm) from Silicycle. Bis(pentamethylcyclopentadienyl) cobalt(II), silver trifluoromethanesulfonate, 4-methoxyphenol, Bis(trifluoromethane)sulfonamide, 2,6-dimethylphenol, *t*-butanol and bis(triphenylphosphoranylidene)ammonium chloride were purchased from Sigma-Aldrich. Nuclear magnetic resonance spectroscopy (¹H, ¹¹B, ¹³C{¹H}, HSQC, HMBC, and COSY experiments) was performed on Bruker 400 MHz (¹H: 400 MHz; ¹¹B: 128 MHz, ¹³C: 100 MHz) or Bruker 600 MHz (¹H: 600 MHz; ¹¹B: 192 MHz, ¹³C: 150 MHz) spectrometers. All 2D NMR experiments were performed using Bruker 400 MHz or Bruker 600 MHz spectrometers. All ¹H NMR spectra were referenced to SiMe₄ through residual ¹H resonance(s) of the employed solvent: CD₂Cl₂ (5.32 ppm) or THF-d8 (3.58 ppm). ¹¹B NMR spectra were referenced to an external standard of boron trifluoride diethyl etherate (0.0 ppm) in C_6D_6 prior to spectrum acquisition. ¹³C{¹H} NMR spectra were referenced relative to SiMe₄ through the resonance(s) of the employed solvent: CD₂Cl₂ (54.0 ppm) or THF-d8 (25.31 ppm). High-resolution mass spectra were obtained on a Kratos MS-80 spectrometer operating in electron impact (EI) mode or using a Bruker Esquire 3000 spectrometer operating in electrospray ionization (ESI) mode. UV-visible spectra were obtained on a Varian Cary 5000 UV-vis-NIR spectrophotometer operating in single-beam mode. EPR measurements were made on a Bruker EMX10/12, equipped with VT capabilities. X-ray crystallographic analyses were performed on suitable crystals coated in Paratone oil and mounted on a Nonius Kappa CCD diffractometer. Supplementary crystallographic data for CCDC-997324 (2Co), CCDC-997321 (1K₂), CCDC-997322 (2K₂), CCDC-997323 (3), and CCDC-885085 (IV) can be obtained free of charge via www.ccdc.cam.ac.uk/conts/retrieving.html (or from the Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge, UK, CB21EZ; fax: (+44) 1223-336-03; or deposit@ccdc.cam.ac.uk).

Synthesis of 1K₂

A mixture of potassium metal (12.9 mg, 0.33 mmol) and naphthalene (133 mg, 0.36 mmol) in THF (8 mL) was stirred at room temperature under argon for 4 h. The produced solution of potassium naphthalenide was added to a solution of **1** (100 mg, 0.165 mmol) in THF (30 mL) at -78 °C and the solution was stirred for 2 h

at room temperature. After this period of time, the solution turned brown. Solvent and naphthalene were removed in vacuo. The brown solid was dissolved in THF (6 mL), layered with hexanes (6 mL) and allowed to crystallize for 6 days at -35 °C. Compound **1K**₂ (0.081 g) was isolated as a dark brown solid in 61 % yield. ¹H NMR (400 MHz, THF-d8) δ 6.77 (s, 4H), 5.47 (s, 2H), 5.29 (s, 5.7 Hz, 2H), 5.09 (s, 2H), 4.21 (s, 2H), 3.59 (4H buried under the proteo THF signal), 2.76 (s, 2H), 1.19 (s, 36H) all the signals in the ¹H NMR spectrum are broad. ¹³C{¹H} NMR (101 MHz, THF-d8) δ 176.66, 167.45 (small and broad), 152.17, 145.42, 143.42 (small and broad), 129.36 (small and broad), 125.94, 120.72, 119.29, 113.32, 101.04, 35.59, 34.06, 26.55, 24.96. ¹¹B NMR (128 MHz, THF-d8): δ 38.7. HRMS: Calcd for C₄₄H₅₄B₂ [[M]⁺-(K+THF)]: 604.4431, Observed: 604.4423. Anal. Calcd. for C₅₂H₇₀B₂K₂O₂: C, 75.53; H, 8.53. Found: C, 75.89; H, 8.51.

Synthesis of 2K₂ from 1K₂ (acid promoted isomerization)

In a 25 mL round-bottomed flask compound $1K_2$ (70 mg, 0.085 mmol) was dissolved in THF (15 mL). A solution of *t*-butanol (1.1 mL, 0.093 mmol, 87.7 mM) in THF was slowly added to the solution of $1K_2$ while stirring at 0 °C. The cold bath was removed and the solution was stirred for 1 h at room temperature. The solution slowly changed from dark brown to dark red and the solvent was removed under vacuum. The dark red solid was dissolved in 1 mL of THF and layered with 1 mL of hexanes and stored at -35 °C for 3 days. Removal of the mother liquor yielded compound $2K_2$ as red crystals (62 mg, 89 %). ¹H NMR (600 MHz, THF-d8) δ 6.94 (s, 4H), 6.80 – 6.62 (m, 4H), 6.32 – 6.22 (m, 2H), 6.09 (td, J = 7.0, 1.2 Hz, 2H), 3.41 (hept, J = 7.0 Hz, 4H), 2.87 (hept, J = 6.9 Hz, 2H), 1.30 (d, J = 6.9 Hz, 12H), 1.14 (d, J = 7.0 Hz, 12H), 1.12 (d, J = 6.9 Hz, 12H). ¹³C{¹H} NMR (151 MHz, THF) δ 155.13 (small and broad), 154.20, 151.27, 146.77 (small and broad), 145.06, 137.90 (small and broad), 129.86, 123.51, 120.63, 119.42, 118.26, 35.53, 34.57, 26.55, 25.03. ¹¹B NMR (128 MHz, THF-d8): δ 32.0. HRMS: Calcd for C₄₄H₅₄B₂ [[M]⁺-(K+THF)]: 604.4431, Observed: 604.4413. Anal. Calcd. for C₅₂H₇₀B₂K₂O₂: C, 75.53; H, 8.53. Found: C, 75.60; H, 8.63.

Synthesis of 2K₂

A mixture of potassium metal (3.0 mg, 0.051 mmol) and naphthalene (6.4 mg, 0.050 mmol) in THF (3 mL) was stirred at room temperature under argon for 4 h. The produced solution of potassium naphthalenide was added to a solution of **2** (15 mg, 0.0248 mmol) in THF (5 mL) at -78 °C and the solution was stirred for 2 h at room temperature. After this period of time, the solution turned red. Solvent and naphthalene were removed in vacuo. The red solid was dissolved in THF (1 mL), layered with hexanes (1 mL) and allowed to

crystallize for 6 days at -35 °C. Compound $2K_2$ (12.0 mg) was isolated as a red solid in 55 % yield. Spectroscopic data are identical to $2K_2$ prepared by isomerization of $1K_2$.

Synthesis of 3

In a 25 mL round-bottomed flask compound $1K_2$ (55 mg, 0.066 mmol) was dissolved in THF (5 mL). 4methoxyphenol (17 mg, 0.137 mmol) was dissolved in THF (1 mL) and added dropwise to the solution of $1K_2$ while stirring. After the addition was complete the solution turned colorless. Solvent was removed under vacuum. The crude white solid was dissolved in boiling hexanes (6 mL) and filtered through a fine frit to remove the potassium 4-methoxyphenoxide produced. After removal of the solvent, the crude white solid was dissolved in boiling hexanes (3 mL) and stored at -35 °C for 3 days. Removal of the mother liquor yielded compound **3** as a white solid (29 mg, 72 %). ¹H NMR (600 MHz, CD₂Cl₂) δ 7.52 (ddd, *J* = 7.4, 1.3, 0.8 Hz, 2H), 7.38 (ddd, *J* = 7.8, 7.2, 1.4 Hz, 2H), 7.16 – 7.10 (m, 4H), 7.10 (d, *J* = 1.5 Hz, 2H), 6.86 (d, *J* = 1.6 Hz, 1H), 4.16 (s, 2H), 2.93 (hept, *J* = 7.1 Hz, 2H), 2.70 – 2.60 (m, *J* = 6.8 Hz, 2H), 1.41 (d, *J* = 6.7 Hz, 6H), 1.31 (d, *J* = 2.8 Hz, 6H), 1.30 (d, *J* = 2.8 Hz, 6H), 1.28 (hept, *J* = 6.9 Hz, 2H), 1.21 (d, *J* = 6.8 Hz, 6H), 0.77 (d, *J* = 6.7 Hz, 6H), 0.33 (d, *J* = 6.7 Hz, 6H). ¹³C {¹H} NMR (151 MHz, CD₂Cl₂) δ 163.54, 149.64, 148.97, 148.91, 145.05 (small and broad), 136.47 (small and broad), 135.15, 134.89, 125.62, 125.46, 120.17, 120.09, 57.86, 53.84, 36.52, 35.93, 34.72, 25.60, 24.86, 24.68, 24.32, 24.25, 22.37. ¹¹B NMR (193 MHz, CD₂Cl₂) δ 77.0. HRMS: Calcd for C₄₄H₅₆B₂ ([M]⁺): 606.4568, Observed: 606.4573.

Synthesis of 2 from 2K₂

In a 25 mL round-bottomed flask compound $2K_2$ (80 mg, 0.0967 mmol) and AgSO₃CF₃ (49.7 mg, 0.193 mmol) were dissolved in THF (15 mL) at -78 °C. The solution was stirred for 10 minutes at room temperature. After this period of time the solution turned brown-greenish. Solvent was removed in vacuo. In the glove box the greenish solid was dissolved in dichloromethane (4 mL) and filtered through a sintered glass filter to remove the potassium triflate produced in the reaction. Solvent was removed in vacuo and the solid was re-dissolved in dichloromethane (2 mL), layered with hexanes (2 mL) and left in the fridge at -35 °C for 2 days. Compound **2** (0.035 g) was isolated as green crystals in 60 % yield. The spectroscopic data are consistent with previously published data.²

Generation of 1K

In a 20 mL scintillation vial compound 1 (27 mg, 0.0447 mmol) was dissolved in THF (8 mL). KC_8 (6.2 mg, 0.0459 mmol) was added to the solution of 1. After the addition was complete the solution turned violet. The

solution was filtered through a sintered glass filter to remove the graphite and the solvent was removed under vacuum. The crude brown-violet solid was dissolved THF (1 mL) and the solution was layered with hexanes (1.5 mL). The solution was stored at -35 °C for 21 days. Removal of the mother liquor yielded compound **1K** (25 mg) as very sensitive violet crystals. Due to the high sensitivity of the crystals elemental analysis could not be determined.

Generation of 2K

In a 20 mL scintillation vial compound 2 (10 mg, 0.0165 mmol) was dissolved in THF (3 mL). KC₈ (2.2 mg, 0.0163 mmol) was added to the solution of 2. After the addition was complete the solution turned blue. The solution was filtered through a sintered glass filter to remove the graphite and the solvent was removed under vacuum. The crude blue solid was dissolved THF (0.5 mL) and the solution was layered with hexanes (1 mL). The solution was stored at -35 °C for 5 days. After removal of the mother liquor the blue crystals were dried under vacuum, yielding 9.5 mg of compound 1K. Due to the high sensitivity of the crystals elemental analysis could not be determined.

Synthesis of 2Co

In a 3 mL vial compound **2** (4.0 mg, 0.006 mmol) was dissolved in dichloromethane (1 mL) and added into another vial containing decamethylcobaltocene. After the addition was complete the solution turned blue. The solution was layered with hexanes (1 mL) and allowed to crystallize for 6 days at -35 °C. After removal of the mother liquor the blue crystals were dried under vacuum, yielding 2.1 mg of compound **1K** (34% yield). Due to the high sensitivity of the crystals elemental analysis could not be determined.

2. EPR Spectroscopy

The structure of compounds **1K** and **2K** were modeled by DFT using the UB3PW91 functional with the 6-31+G(d) basis set (compounds were modeled as anionic species without the counter cation). The hyperfine coupling constants (hfcc) were calculated using the basis set EPR-II, which is optimized for their computation. The hyperfine coupling constants were used as the initial values for the simulations and they were modified by hand until the correlation between the simulated and the experimental spectrum was higher than 97%. Subsequently, the program algorithm was utilized, which allowed reaching a correlation of 99.6% for **1K** and 99.9% for **2K**



Figure S1. Experimental (black) and simulated (red) EPR spectrum of **1K** (left). Experimental (black) and simulated (red) EPR spectrum of **2K** (right)

Titration Experiment

Compound $1K_2$ was dissolved in THF in an EPR tube and the spectrum was measured (Figure S2, series "1"). Approximately 1 mg of compound 1 was dissolved in 1 mL of THF and a couple of drops were added into the EPR tube and the spectrum was measured again (Figure S2, series "2"). This procedure was repeated two more times (Figure S2, series "3" and "4" respectively). Along with the augmentation of the EPR signal there is also a change in color of the solution from dark brown to violet. Since compound $1K_2$ is capable of reducing compound 1, this experiments generates the radical 1K *in situ* upon subsequent additions of 1 to the EPR tube containing $1K_2$.



Figure S2. Qualitative titration of compound $1K_2$. The increment in the series numbers indicate a qualitative increment in the amount of compound 1 added to the EPR tube containing $1K_2$.

3. UV-Vis spectroscopy

Compound $1K_2$ is dark brown in the solid state and in solution (THF). Contrasting to 1, $1K_2$ is not emissive under UV-light and its absorption spectrum is considerably red-shifted (Figure S3), with the longest absorption maximum at 802 nm ($\mathcal{E} = 9640 \text{ M}^{-1} \text{ cm}^{-1}$). According to TD-DFT calculations (B3LYP/6-311G(d)) this energy transition can be attributed to a transition from HOMO to LUMO (calcd: 805 nm, f = 0.1468) and is mainly $\pi \rightarrow \pi^*$ in character.

 $2K_2$ exhibits a magenta colour in solution (THF) and displays purplish pink emission under UV light ($\lambda_{max} = 620 \text{ nm}$). In contrast, the precursor 2 is non-fluorescent under UV light. Compound $2K_2$ shows major peaks at 531 nm, 352 nm, 337 nm and a shoulder at 507 nm in the absorption spectrum (Figure S3). Based on TD-DFT calculations (B3LYP/6-311G(d)) the longest absorption band is manly comprised of the HOMO-LUMO transition (calcd: 485 nm, f = 0.2530) and is mainly $\pi \rightarrow \pi^*$ in character. Compound $2K_2$ is isoelectronic to dibenzopentalene. However, the former is non-emissive under UV-light and the longest absorption maximum is around 450 nm. This clearly shows that the exchange of carbon for the more electropositive boron atom causes a significant change in the electronic and photophysical properties.



Figure S3. Absorption and emission spectrum of compounds **1**K₂ and **2**K₂ in THF. Blue line: **1**K₂ absorption; red line: **2**K₂ absorption; dotted red line: **2**K₂ emission.

4. NMR Spectroscopy



Figure S5.¹³C $\{^{1}H\}$ spectrum of **1K**₂ in THF-d₈.



Figure S7. ¹³C $\{^{1}H\}$ NMR spectrum of $2K_{2}$ in THF-d₈.



Figure S9. $^{13}C{^{1}H}$ NMR spectrum of **3** in CD₂Cl₂.



Figure S10. ¹¹B NMR spectrum of $1K_2$ in THF-d₈.





Figure S13. ¹H NMR spectrum of KO'Bu in THF-d₈ (top) and ¹H NMR spectrum of KO'Bu after addition of **3** (bottom) showing the immediate formation of $2K_2$.

5. X-ray crystallography

	1K ₂	2K ₂	2C0	3
Empirical formula	$C_{76}H_{118}B_2K_2O_8$	$C_{60}H_{86}B_2K_2O_4$	$C_{66}H_{88}B_2Cl_4Co$	$C_{44}H_{56}B_2$
Crystal system	Orthorhombic	Monoclinic	Triclinic	Triclinic
Fw	1259.52	485.55	1103.71	606.51
F(000)	2744	1052	589	660
<i>T</i> (K)	173(2)	123(2)	173(2)	173(2)
Wavelength (Å)	1.54178	0.71073	1.54178	0.71073
Space group	Pbca	P21/c	P-1	P-1
<i>a</i> (Å)	16.1537(3)	10.3128(2)	11.5296(9)	11.3370(7)
<i>b</i> (Å)	19.9550(4)	14.6667(2)	13.487(2)	11.9746(7)
<i>c</i> (Å)	22.8279(5)	19.3119(4)	14.651(1)	13.9071(6)
a (deg)	90	90	112.337(6)	83.312(3)
β (deg)	90	102.005(1)	111.343(4)	85.920(3)
γ (deg)	90	90	95.399(6)	88.345(2)
Z	4	4	1	2
$V(\text{\AA}^3)$	7358.5(3)	2857.13(9)	1890.0(3)	1870.0(2)
$\rho_{calcd} (g \cdot cm^{-3})$	1.137	1.129	0.97	1.077
$\mu (mm^{-1})$	1.534	0.209	3.303	0.059
θ range (deg)	3.87 - 72.96	2.45 - 27.57	3.62 - 68.19	2.13 - 27.29
Completeness	0.987	0.99	0.964	0.983
Data/Restraints/Param	5969 / 15 / 415	5625 / 0 / 313	5395 / 6 / 356	4895 / 0 / 427
Collected reflections; R_{σ}	77463; 0.0115	23720; 0.0247	24504; 0.0419	15376; 0.1145
Unique reflections; R _{int}	7247; 0.0243	6559; 0.0356	6666; 0.0398	8259; 0.0946
$R_1; wR_2[I > 2\sigma(I)]$	0.0701; 0.2082	0.0682; 0.1642	0.0586; 0.1638	0.1478; 0.2176
R_1 ; w R_2 [all data]	0.0805; 0.2222	0.0791; 0.1742	0.0665; 0.1699	0.2218; 0.2406
GOF	1.001	1.022	1.108	1.495
largest diff peak and hole	0.684 and -0.549	1.137 and - 0.345	0.462 and - 0.346	0.297 and - 0.274

Table S1: Data Collection and Structure Refinement Details for 1K₂, 2K₂, 2Co, and 3

6. Cyclic voltammetry



Figure S15. Cyclic voltammetry traces for **1** (top) and **2** (bottom). Cyclic voltammetry was carried out at a scan rate of 100 mV/s in THF with 1mM substrate and 0.1M [NBu₄][PF₆] as the supporting electrolyte. Ferrocene was added as an internal standard and set to 0 V. Cyclic voltammogram on bottom right shows reversibility of this potential when the scan does not exceed -1.75 V. **2**: $E_{pc} = -2.47$ V, -2.89 V; **3**: $E_{pc} = -1.52$ V, 2.44 V, $E_{1/2} = -1.45$ V.

7. Quantum Chemical Calculations

XYZ-coordinates of optimized structures

1K'

В	-1.748856	1.215171	0.112532
С	-0.331427	0.602230	0.073202
С	0.203182	1.979008	0.284465
С	1.399460	2.671985	0.419843
Η	2.367694	2.178022	0.379032
С	1.306220	4.071330	0.616539
Η	2.221437	4.652529	0.728199
С	0.077777	4.725289	0.671204
Η	0.053388	5.803876	0.824270
С	-1.135854	4.006886	0.530302
Η	-2.088848	4.534713	0.575292
С	-1.061329	2.636344	0.337682
С	-3.242828	0.719241	-0.012579
С	-3.968229	0.305228	1.128470
С	-5.292815	-0.124041	1.001877
Η	-5.833102	-0.441891	1.893790
С	-5.940214	-0.158747	-0.236959
С	-5.219266	0.256517	-1.357985
Η	-5.710721	0.233008	-2.331566
С	-3.893095	0.695112	-1.265528
С	-3.315104	0.338240	2.504267
Η	-2.275253	0.661939	2.356879
С	-3.264824	-1.047531	3.161036
Η	-4.272652	-1.437300	3.357077
Η	-2.734368	-0.999038	4.121290
Η	-2.742659	-1.766706	2.520968
С	-3.986989	1.364178	3.428030
Η	-3.987133	2.361115	2.972996
Η	-3.456091	1.427598	4.387197
Η	-5.029330	1.090024	3.639599
С	-7.378036	-0.632335	-0.372509
Η	-7.639919	-0.557976	-1.438481
С	-8.356439	0.262604	0.402122
Η	-9.394409	-0.057813	0.238609
Η	-8.266068	1.309046	0.088632
Η	-8.160520	0.221867	1.480780
С	-7.542315	-2.102710	0.038238
Η	-7.306736	-2.245216	1.100066
Η	-6.872219	-2.750003	-0.538696
Η	-8.574336	-2.442286	-0.124416
С	-3.157048	1.140699	-2.522289
Η	-2.160735	1.480481	-2.206448
С	-2.945886	-0.019136	-3.505228
Η	-2.398184	-0.840178	-3.030544

Η	-2.369727	0.314971	-4.378455
Η	-3.904238	-0.415063	-3.867421
С	-3.847423	2.327533	-3.208364
Η	-4.839530	2.051784	-3.590039
Η	-3.250149	2.680304	-4.059501
Η	-3.975718	3.164685	-2.512730
В	1.748752	-1.215196	-0.112324
С	0.331327	-0.602251	-0.072912
С	-0.203298	-1.979026	-0.284163
С	-1.399586	-2.672000	-0.419467
Η	-2.367817	-2.178041	-0.378565
С	-1.306359	-4.071340	-0.616215
Η	-2.221585	-4.652535	-0.727822
С	-0.077921	-4.725294	-0.671001
Η	-0.053542	-5.803876	-0.824107
С	1.135720	-4.006894	-0.530169
Η	2.088712	-4.534719	-0.575251
С	1.061208	-2.636359	-0.337490
С	3.242756	-0.719327	0.012651
С	3.968087	-0.305419	-1.128481
С	5.292737	0.123698	-1.002033
Η	5.832964	0.441478	-1.894007
С	5.940273	0.158336	0.236733
С	5.219391	-0.256812	1.357843
Η	5.710958	-0.233358	2.331370
С	3.893149	-0.695226	1.265535
С	3.314808	-0.338351	-2.504205
Η	2.274956	-0.662003	-2.356714
С	3.264524	1.047455	-3.160903
Н	4.272351	1.437183	-3.357027
Н	2.733965	0.999036	-4.121106
Η	2.742460	1.766625	-2.520746
С	3.986533	-1.364279	-3.428092
Н	3.986687	-2.361236	-2.973102
Н	3.455513	-1.427634	-4.387197
Н	5.028862	-1.090164	-3.639781
C	7.378171	0.631745	0.372122
H	7.640162	0.557360	1.438065
C	8.356379	-0.263320	-0.402610
H	9.394408	0.056959	-0.239200
H	8.265902	-1.309752	-0.089114
H	8.160357	-0.222554	-1.481249
C	7.542585	2.102096	-0.038653
H	7.306906	2.244623	-1.100455
H	6.872633	2.749477	0.538352
H	8.5/4666	2.441548	0.123883
	5.15/214	-1.140/85	2.522572
H	2.160538	-1.4/9699	2.206743
C	2.94/253	0.018/29	3.303941

Η	2.400084	0.840448	3.031812
Н	2.371048	-0.315357	4.379146
Η	3.906004	0.413706	3.868112
С	3.847008	-2.328436	3.207621
Н	4.839528	-2.053586	3.588879
Н	3.249898	-2.681102	4.058918
Η	3.974303	-3.165395	2.511572

2K'

С	-0.609339	-0.261099	0.252324
С	-0.329296	-1.354395	1.186597
С	1.085401	-1.556306	1.300017
С	1.560830	-2.557368	2.143926
Η	2.634733	-2.723235	2.241400
С	0.668417	-3.361628	2.874693
Η	1.047577	-4.143529	3.531556
С	-0.707039	-3.155480	2.756154
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Η	-2.287659	-1.999899	1.829688
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Η	-6.037310	-1.413439	-1.342932
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Η	-2.347315	-1.154844	-1.969005
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Η	-4.981425	-1.223135	-3.539215
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Η	-7.921320	0.526362	1.316911
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Η	-7.472355	-1.877609	1.794669
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Η	4.278615	-1.592379	-3.622918
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С	-0.303127	0.448755	1.295504
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Η	-3.680729	-1.109946	-1.123930
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Η	2.295520	-1.977099	-1.883819
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Η	2.523860	0.862362	-2.972603
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Η	-5.981212	-1.692115	1.580776
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Η	-3.166033	3.255551	-0.429404
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Η	-8.269124	0.619737	-1.596435
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С	2.118094	0.381648	-0.373798
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Η	-6.840255	-4.488772	-0.734593
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Η	1.777322	3.662302	4.290232
Η	0.202997	2.874742	4.417906
Η	1.573556	2.042391	3.645199
С	-0.030105	4.961466	2.593058
Н	-0.914293	4.854933	3.230109
H	0.625712	5.697392	3.067406
Н	-0.354015	5.372999	1.634336
В	0.778557	1.731795	0.026666
ĸ	-0.593669	0.113485	2.769944
-			

Ο	0.356237	-1.102786	4.994607
0	-2.709121	1.715833	3.410429
С	1.697966	-1.627010	4.997253
Η	2.078927	-1.577958	3.977971
Η	2.322379	-0.997479	5.644812
С	1.581582	-3.059932	5.548320
Η	2.414497	-3.296746	6.213499
Η	1.592577	-3.785224	4.734030
С	0.215748	-3.069606	6.286660
Η	-0.487933	-3.737257	5.785974
Η	0.297503	-3.394345	7.326079
С	-0.260839	-1.613912	6.180402
Η	0.066086	-1.021854	7.046910
Η	-1.340711	-1.504299	6.070606
С	-3.510110	2.322442	2.373482
Η	-2.891509	2.411881	1.482740
Η	-4.353985	1.660691	2.147012
С	-3.991955	3.665569	2.953550
Η	-3.402461	4.494778	2.560805
Η	-5.034782	3.853828	2.690924
С	-3.790767	3.513663	4.484996
Η	-4.696738	3.728506	5.055570
Η	-3.010027	4.189213	4.841000
С	-3.347505	2.052987	4.645412
Η	-4.210638	1.391685	4.806869
Н	-2 628077	1 889662	5 449647

TD-DFT calculations, td=(nstates=20) B3LYP/6-311G(d)

Excitation energies and oscillator strengths (1K₂):

Excited State	1:	Singlet-A	1.5396 eV	805.32 nm	f=0.1468	<s**2>=0.000</s**2>
303 -> 304		0.70192				
Excited State	2:	Singlet-A	1.6856 eV	735.54 nm	f=0.0017	<s**2>=0.000</s**2>
303 -> 305		0.70436				
Excited State	3:	Singlet-A	1.7014 eV	728.70 nm	f=0.0000	<s**2>=0.000</s**2>
303 -> 306		0.70047				
Excited State	4:	Singlet-A	1.7950 eV	690.73 nm	f=0.0000	<s**2>=0.000</s**2>
303 -> 307		0.65455				
303 -> 309		0.20526				
303 -> 315		0.11832				
Excited State	5:	Singlet-A	2.1039 eV	589.32 nm	f=0.0058	<s**2>=0.000</s**2>
303 -> 308		0.70452				
Excited State	6:	Singlet-A	2.1432 eV	578.50 nm	f=0.0001	<s**2>=0.000</s**2>
303 -> 307		-0.22267				
303 -> 309		0.65681				
Excited State	7:	Singlet-A	2.2229 eV	557.76 nm	f=0.0001	<s**2>=0.000</s**2>

303 -> 309 -0.10418	
303 -> 310 0.67701	
303 -> 312 -0.12764	
Excited State 8: Singlet-A	2.2447 eV 552.35 nm f=0.0047 <s**2>=0.000</s**2>
303 -> 311 0.67336	
$303 \rightarrow 312 -0.12219$	
$303 \rightarrow 313 -0.15560$	
Excited State 9. Singlet-A	2 3248 eV 533 32 nm f=0 0007 <s**2>=0 000</s**2>
$303 \rightarrow 310$ 0 14485	
$303 \rightarrow 312$ 0 51792	
$303 \rightarrow 312 \qquad 0.31792$ $303 \rightarrow 313 \qquad -0.44674$	
Excited State 10: Singlet-A	23270 eV 532.81 nm f=0.0014 <s**2>=0.000</s**2>
303 -> 311 0 19529	
$303 \rightarrow 312 \qquad 0.17527$	
303 > 312 0.45116	
303 > 314 = 0.10448	
$\frac{503 - 514}{503 - 514} = 0.10440$	2.4034 eV 515.87 nm f=0.0110 < S**2>=0.000
202 > 214 0 22212	2.4034 EV 515.87 IIII 1-0.0110 \5*22-0.000
303 - 314 = 0.32212	
303 - 315 = 0.39024	
503 - 2510 - 0.11344	24252 oV = 500.12 mm = f = 0.0042 scs**2 = 0.000
Exclied State 12: Singlet-A	2.4555 eV 509.12 nm 1-0.0945 <5**2>-0.000
303 -> 314 0.59905	
303 -> 315 -0.3180/	
303 -> 318 -0.10917	2 (17() 472 (5 C 0.0001 (0**0) 0.000
Excited State 13: Singlet-A	$2.61/6 \text{ eV} 4/3.65 \text{ nm} = 0.0001 < S^{**}2 >= 0.000$
$303 \rightarrow 315 \qquad 0.13202$	
303 -> 316 0.68/43	
Excited State 14: Singlet-A	2.7137 eV 456.88 nm f=0.0388 <s**2>=0.000</s**2>
303 -> 317 0.69702	
Excited State 15: Singlet-A	$3.0412 \text{ eV} 407.68 \text{ nm} = 0.1661 < S^{**2} = 0.000$
$303 \rightarrow 314 0.10489$	
303 -> 318 0.68496	
Excited State 16: Singlet-A	3.1299 eV 396.13 nm f=0.0001 <s**2>=0.000</s**2>
303 -> 319 -0.22025	
303 -> 320 0.63910	
303 -> 321 0.10613	
303 -> 322 -0.11359	
Excited State 17: Singlet-A	3.2062 eV 386.70 nm f=0.0055 <s**2>=0.000</s**2>
303 -> 319 -0.45360	
303 -> 320 -0.23899	
303 -> 321 0.46068	
Excited State 18: Singlet-A	3.2102 eV 386.22 nm f=0.0045 <s**2>=0.000</s**2>
303 -> 319 0.47981	
303 -> 321 0.49153	
Excited State 19: Singlet-A	3.2390 eV 382.78 nm f=0.0005 <s**2>=0.000</s**2>
302 -> 304 0.69242	
Excited State 20: Singlet-A	3.3039 eV 375.26 nm f=0.0000 <s**2>=0.000</s**2>
303 -> 320 0.10484	
303 -> 322 0.67594	



Excitation energies and oscillator strengths (2K₂):

Excited State 1:	Singlet-A	2.5559 eV	485.10 nm	f=0.2530	<s**2>=0.000</s**2>
262 -> 267	0.16027				
263 -> 264	0.68195				
Excited State 2:	Singlet-A	2.5784 eV	480.86 nm	f=0.0000	<s**2>=0.000</s**2>
263 -> 266	0.70130				
Excited State 3:	Singlet-A	2.6021 eV	476.48 nm	f=0.0007	<s**2>=0.000</s**2>
263 -> 265	0.70494				
Excited State 4:	Singlet-A	2.7211 eV	455.64 nm	f=0.0000	<s**2>=0.000</s**2>
262 -> 264	0.69069				
263 -> 267	-0.12457				
Excited State 5:	Singlet-A	2.8825 eV	430.13 nm	f=0.0000	<s**2>=0.000</s**2>
262 -> 264	0.13523				
263 -> 267	0.65684				
263 -> 271	0.20192				
Excited State 6:	Singlet-A	3.0940 eV	400.72 nm	f=0.0041	<s**2>=0.000</s**2>
263 -> 268	0.70277				
Excited State 7:	Singlet-A	3.1286 eV	396.29 nm	f=0.0048	<s**2>=0.000</s**2>
263 -> 270	0.70075				
Excited State 8:	Singlet-A	3.1509 eV	393.49 nm	f=0.0000	<s**2>=0.000</s**2>
263 -> 269	0.70114				
Excited State 9:	Singlet-A	3.2303 eV	383.82 nm	f=0.0000	<s**2>=0.000</s**2>
263 -> 267	-0.20276				
263 -> 271	0.64716				
263 -> 273	0.18142				
Excited State 10:	Singlet-A	3.2714 eV	379.00 nm	f=0.0020	<s**2>=0.000</s**2>
262 -> 266	0.69572				
Excited State 11:	Singlet-A	3.2941 eV	376.38 nm	f=0.0000	<s**2>=0.000</s**2>
262 -> 265	0.69837				
Excited State 12:	Singlet-A	3.3001 eV	375.70 nm	f=0.0096	<s**2>=0.000</s**2>
263 -> 272	0.62874				
263 -> 274	-0.30847				
Excited State 13:	Singlet-A	3.3237 eV	373.03 nm	f=0.0000	<s**2>=0.000</s**2>





Figure S17. Calculated TD-DFT UV-VIS spectrum 2K₂.

8. Full citation for reference 32

All calculations were carried out using Gaussian 09, Revision A.1, M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Ivengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V.

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9. References

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