

Supporting Information

Self-assembly of colloidal sulfur particles influenced by sodium oxalate salt on glass surface from evaporating drops

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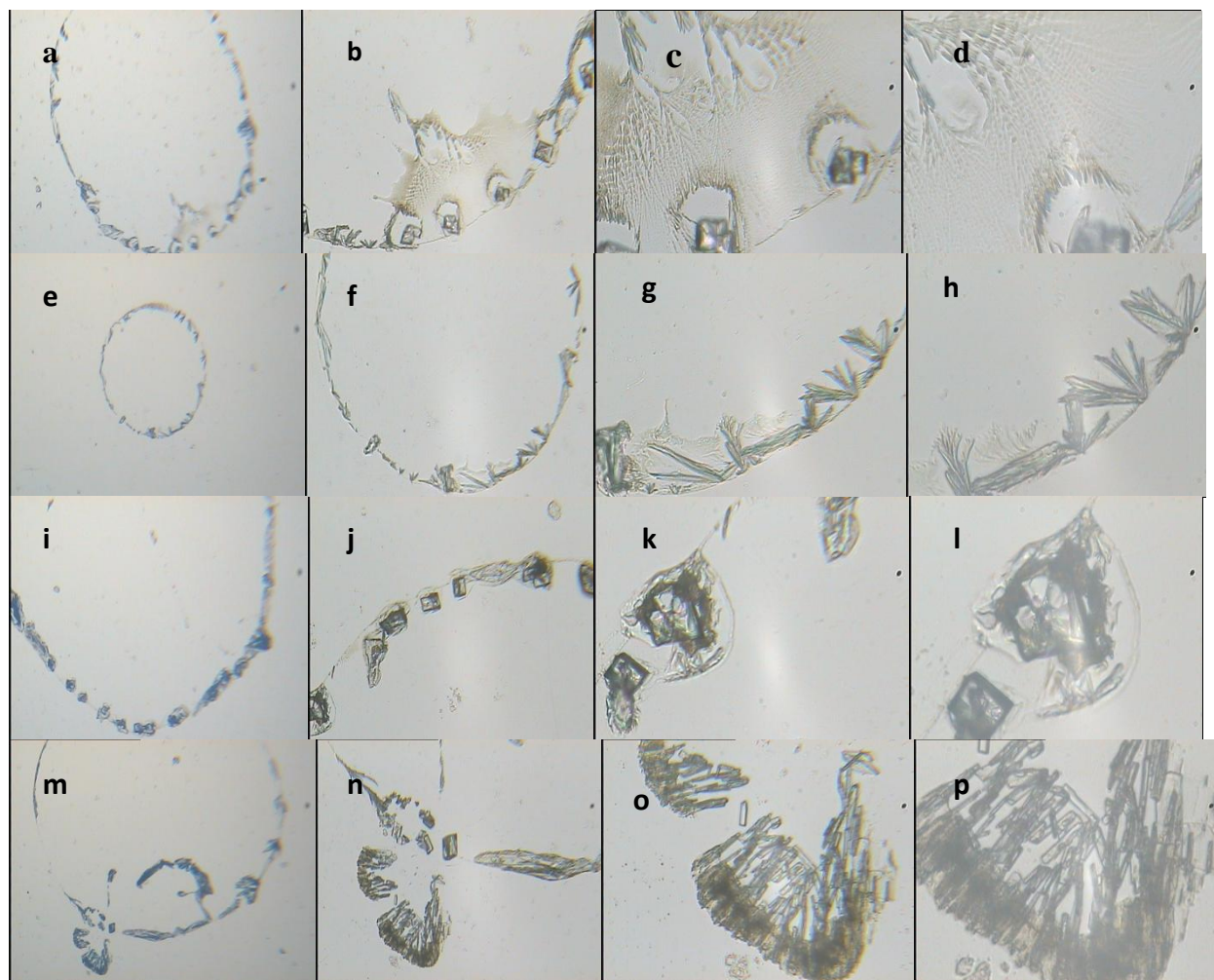
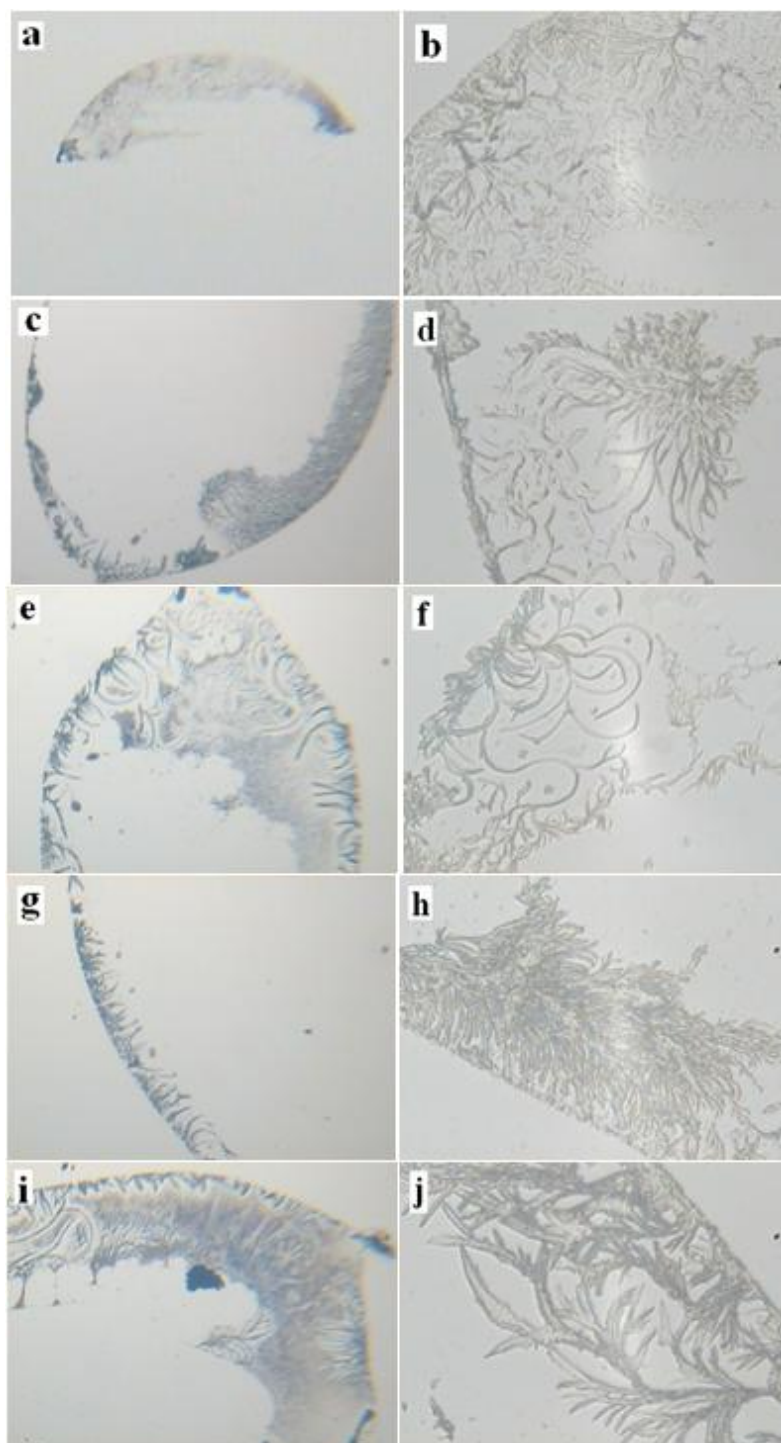


Figure S1. Drying patterns for pure sodium oxalate solution of 10 mM concentration: (a,e,i) using 5X objective; (b,f,e) using 10X objective; (c,g,i) using 25X objective; and (d,h,l,m,n,o,p) using 45X objective. Figures (a,b,c,d) depict the cases where the tree-like crystal forming nature can be faintly seen. The rest of the figures represent the more common crystal formations of

sodium oxalate. As is visible from the images in Figure S1, the sodium oxalate salt solution has more of a tendency to form bulk crystals near the boundary of the drop after drying on a glass surface.



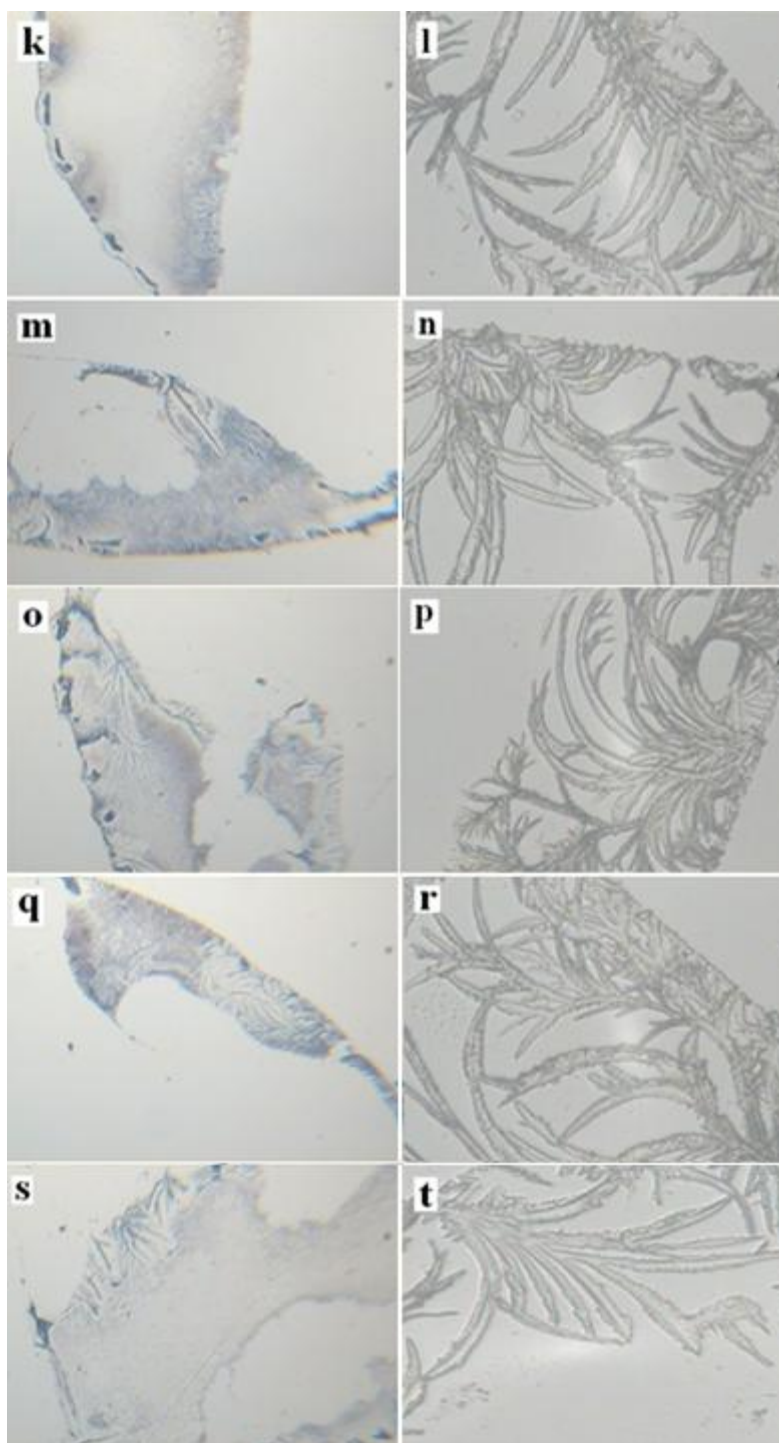


Figure S2. Self-assembled structure of sulfur particle for 1mM particle concentration of (a,b) 1 μl , (c,d) 2 μl , (e,f) 3 μl , (g,h) 4 μl , (i,j) 5 μl , (k,l) 6 μl , (m,n) 7 μl , (o,p) 8 μl , (q,r) 9 μl and (s,t) 10 μl , drop volumes, where the images on the right are at 50X magnifications (5X objective) and the adjacent images are their respective magnifications at 450X (5X objective).

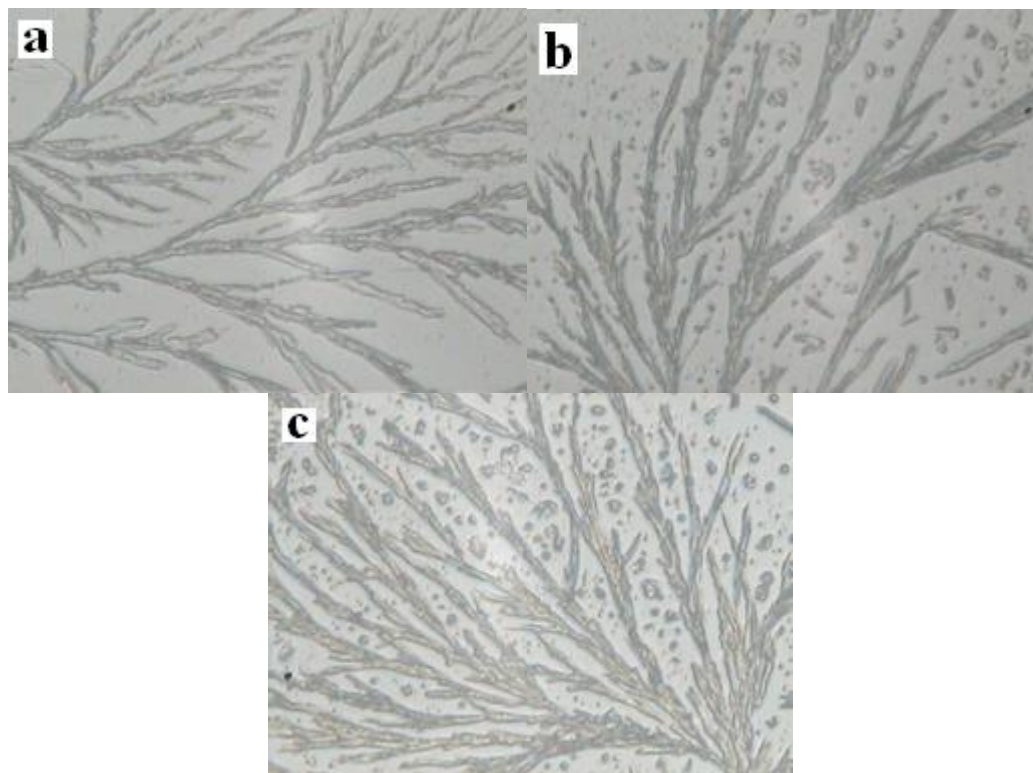


Figure S3. Self-assembled structure of sulfur particle in oxalic acid media using 10 μl drop volume for (a) 1 mM, (b) 5 mM and (c) 10 mM particle concentrations, 1 mM and 5 mM diluted from original 10 mM after reaction.