Electronic Supplementary Material (ESI) for Journal of Materials Chemistry A. This journal is © The Royal Society of Chemistry 2014

# Forming a p-n Heterojunction on GaP Photocathodes for

## H<sub>2</sub> Production Providing an Open-Circuit Voltage of 710 mV

Mauro Malizia, Brian Seger, Ib Chorkendorff, and Peter C. K. Vesborg\*

CINF, Department of Physics, Technical University of Denmark, DTU

Building 311, Fysikvej, DK-2800 Kgs. Lyngby, Denmark

E-mail: peter.vesborg@fysik.dtu.dk

### **Electronic Supplementary Information**

#### **UV-Vis Spectroscopy**

The absorption spectra for  $TiO_2$  and  $Nb_2O_5$  films are shown in Figure S1. A considerably high absorption in the visible light region is observed for the  $Nb_2O_5$  film. This is most likely due to the presence of Nb in metallic phase, not completely oxidized during the deposition. For  $TiO_2$  and  $Nb_2O_5$  the bandgaps are found to be 3.1 eV and 3.5 eV, respectively, as depicted in the inset of Figure S1.



Figure S1: Absorption spectra for TiO<sub>2</sub> and Nb<sub>2</sub>O<sub>5</sub> sputtered on quartz glass. Inset: Tauc plot for TiO<sub>2</sub> and Nb<sub>2</sub>O<sub>5</sub>.

#### **X-ray Diffraction**

The XRD data of both the TiO<sub>2</sub> and Nb<sub>2</sub>O<sub>5</sub> films is shown in Figure S2. The peaks of the TiO<sub>2</sub> at 25.5°, 38.6°, 56.2°, and 63.8° correspond to anatase. The peak at 32° and 52°, and 67° are undefined peaks, but may be related to the underlying Si support. The extremely broad peaks of the Nb<sub>2</sub>O<sub>5</sub> sample give little useful information. However without clear sharp peaks, it is evident that the Nb<sub>2</sub>O<sub>5</sub> has, at most, a minimal degree of crystallinity.



Figure S2: Cu K $\alpha$  XRD peaks of TiO<sub>2</sub> and Nb<sub>2</sub>O<sub>5</sub> sputtered on Si substrate showing the anatase structure of TiO<sub>2</sub> and the amorphous phase of Nb<sub>2</sub>O<sub>5</sub>.

#### Electrochemical Impedance Spectroscopy

The positioning of flat band potential of GaP,  $TiO_2$  and  $Nb_2O_5$  was determined by electrochemical impedance spectroscopy (EIS) using the Mott-Schottky relationship<sup>1</sup> expressed in the following Equation

$$\frac{1}{C^{-2}} = \frac{2}{e\varepsilon_r \varepsilon_0 A^2 N} \left( E - E_{FB} - \frac{k_B T}{e} \right)$$
(1)

where C is the space-charge region capacitance, e is the electronic charge,  $\varepsilon_r$  is the relative permittivity,  $\varepsilon_0$  is the vacuum permittivity, A is the electrode surface area, N is the number of donors or acceptors, E is the applied voltage,  $E_{FB}$  is the flat band potential,  $k_B$  is Boltzmann's constant, and T is the temperature. Some assumption are made in the derivation of the Mott-Schottky equation, thus a standard nonlinear relationship between C<sup>-2</sup> and E may not be observed if such assumptions are not fulfilled (*i.e.* space-charge region capacitance << Helmholtz layer capacitance, absence of surface states, frequency-independence of the dielectric constant  $\varepsilon_r$ , homogeneous spatial distribution of donors/acceptors, flat surface).<sup>1,2</sup>



Figure S3: Mott-Schottky plot for (a) p-GaP, (b) n-TiO<sub>2</sub> and (c) n-Nb<sub>2</sub>O<sub>5</sub> on Si substrate, measured using a modulation frequency of 1 kHz in 1 M HClO<sub>4</sub> in dark conditions.

The Mott-Schottky plot for p-GaP is shown in Figure S3a. By using Equation 1 and  $\varepsilon_r$ =11.1,<sup>3</sup> the flat band potential was found to be  $E_{FB}$ =0.92  $V_{RHE}$  and the acceptors concentration was found to be  $N_A$ =5.6·10<sup>17</sup> cm<sup>-3</sup>. Mott-Schottky plots of the sputtered TiO<sub>2</sub> and Nb<sub>2</sub>O<sub>5</sub> films deposited on Si are shown in Figures S3b and S3c. For TiO<sub>2</sub>,  $\varepsilon_r$  is 75,<sup>4</sup> while  $E_{FB}$  was calculated to be -0.02  $V_{RHE}$ . N<sub>D</sub> is dependent upon the surface area, and SEM images have shown that the sputtered TiO<sub>2</sub> is quite rough.<sup>5</sup> If only the geometric area is taken into consideration, N<sub>D</sub> would be 5·10<sup>20</sup> cm<sup>-3</sup>. However, it should be noted that using a surface roughness of 3.2 or 10 would correspond to an N<sub>D</sub>

of  $5 \cdot 10^{19}$  cm<sup>-3</sup> or  $5 \cdot 10^{18}$  cm<sup>-3</sup>, respectively. Determining the exact surface roughness and the effect of surface roughness on the Mott-Schottky impedance data is beyond the scope of this work. However, we can reasonably estimate that the actual donor density will be somewhere between  $5 \cdot 10^{18}$  cm<sup>-3</sup> and  $5 \cdot 10^{20}$ .

For Nb<sub>2</sub>O<sub>5</sub> ( $\varepsilon_r$ =42),<sup>5</sup> E<sub>FB</sub> was calculated to be 0.12 V<sub>RHE</sub>. Again, the rough surface of the sputtered Nb<sub>2</sub>O<sub>5</sub> made determining an effective surface area difficult. Using the same approach of TiO<sub>2</sub>, we could estimate the N<sub>D</sub> of Nb<sub>2</sub>O<sub>5</sub> to be somewhere between 4·10<sup>20</sup> cm<sup>-3</sup> and 4·10<sup>18</sup> cm<sup>-3</sup>. Both metal oxides are n-type and characterized by a high donor density. The high donor density often introduces error for the determination of E<sub>FB</sub> and carrier concentration due to the magnitude of the capacitance of the Helmholtz layer. Therefore, this can result in nonlinearity of the Mott-Schottky plot.<sup>6</sup>

We found an apparent  $E_{FB}$  for Nb<sub>2</sub>O<sub>5</sub> that is more positive than the one of TiO<sub>2</sub>, although the majority of studies on Nb<sub>2</sub>O<sub>5</sub> place the value more negative than the one of TiO<sub>2</sub>.<sup>7</sup>

The carrier density found in our  $Nb_2O_5$  film is comparable to the carrier density reported for similar amorphous thin films, which indicates a highly disordered structure and probably a considerable presence of defects.<sup>8</sup>

#### References

- (1) Cardon, F., Gomes, W. P., J. Phys. D: Appl. Phys., 1978, 11, L63-L67
- (2) De Gryse, R.; Gomes, W. P.; Cardon, F., Vennik, J., J. Electrochem. Soc., 1975, 122, 711-712
- (3) Samara, G. A. Phys. Rev. B, 1983, 27, 3494-3505
- (4) King D. M., Du X., Cavanagh A. S., Weimer A., Nanotechnology, 2008, 19, 445401-445406
- (5) Seger, B.; Pedersen, T.; Laursen, A. B.; Vesborg, P. C. K.; Hansen, O., Chorkendorff, I., J. Am. Chem. Soc., **2013**, *135*, 1057-1064
- (6) Di Quarto, F.; La Mantia, F., Santamaria, M., Electrochim. Acta, 2005, 50, 5090-5102
- (7) Alibabaei, L.; Luo, H.; House, R. L.; Hoertz, P. G.; Lopez, R., Meyer, T. J., J. Mater. Chem. A, **2013**, *1*, 4133-4145
- (8) Wang, W., Alfantazi, A., J. Electrochem. Soc., 2013, 160, C1-C11