Supporting Information

Effect of Thiadiazole Out-Backbone Displacement in Indacenodithiophene Semiconductor Polymers

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Supporting Information Figures



Figure S1. DSC traces of C₁₆IDT-fDTBT (*black line*) and C₁₆IDT-fDTBPT (*red line*).



Figure S2. Cyclic voltammetry traces of drop casted $C_{16}IDT$ -fDTBT (*black line*) and $C_{16}IDT$ -fDTBPT (*red line*) into ITO substrate and measured in 0.1 M [TBA][PF₆] acetonitrile solution. Oxidation potentials were determined by the electrochemical onset and reference to ferrocene $E_{1/2}$ (*green line*).



Figure S3. Absorption spectrum of **3** in chlorobenzene (*solid line*) and simulated by TD-DFT means (*dashed line*). *Inset*: Molecular orbital distribution and energies of **3**.



Figure S4. Photoluminescence emission (*red line*) and excitation (*black line*) spectra of $C_{16}IDT$ -fDTBT in chloroform solution (*left*) and spin-coated film (*right*).



Figure S5. Photoluminescence emission (*red line*) and excitation (*black line*) spectra of $C_{16}IDT$ -fDTBPT in chloroform solution. Note that no photoluminescence was observed for spin-coated film.



Figure S6. ¹H-NMR of C₁₆IDT-fDTBT polymer acquired in TCE-*d*₂ at 120 °C.



Figure S7. ¹H-NMR of C₁₆IDT-fDTBPT polymer acquired in TCE-*d*₂ at 120 °C.

Supporting Information Tables

Polymer	Composition ^a		$\Delta E (nm)^b$	f^c	Character
C ₁₆ IDT-fDTBT	HOMO → LUMO	86 %	661.07	3.915	π-π*
	$HOMO \rightarrow LUMO+3$	54%	488.76	1.0844	π-π*
	$HOMO \rightarrow LUMO+4$	29%			
C ₁₆ IDT-fDTBPT	HOMO → LUMO	46 %	1028.39	0.4584	ICT ^d
	$HOMO \rightarrow LUMO'$	19 %			
	$HOMO \rightarrow LUMO''$	19 %			
	HOMO $-3 \rightarrow$ LUMO"	69%	588.54	0.9399	π-π*
	$HOMO \rightarrow LUMO+1$	19%			
	$HOMO-3 \rightarrow LUMO''$	20%	586.98	3.5027	π-π*
	$HOMO \rightarrow LUMO+1$	66%			
	$HOMO-1 \rightarrow LUMO+2$	90%	475.19	0.8468	π-π*

Table S1. Selected TD-DFT calculated energies and compositions (only oscillator strength higher than 0.2 are shown) at B3LYP/6-31G(d) from the first ten singlet electronic transitions.

^{*a*}Compositions of electronic transitions (only higher than 8% are shown) are expressed in terms of contributing excitations between ground state Kohn–Sham molecular orbitals. ^{*b*}Transition energy from the ground state to the excited state. ^{*c*}Oscillator strength. ^{*d*}Intramolecular charge transfer band.

Complete reference 9

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